

Article

Extra-Framework Content in Sodalite-Group Minerals: Complexity and New Aspects of Its Study Using Infrared and Raman Spectroscopy

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Received: 23 March 2020; Accepted: 14 April 2020; Published: 17 April 2020



Abstract: Nine samples of carbonate-free sodalite-group minerals, including those with abnormally high contents of polysulfide groups, fluoride anion and carbon dioxide molecules as well as synthetic fluoraluminate sodalite-type compound $\text{Na}_8(\text{Si}_7\text{Al}_5\text{O}_{24})(\text{AlF}_6)^{3-} \cdot 5\text{H}_2\text{O}$, have been studied by means of electron microprobe analyses, infrared and Raman spectroscopy; the CO_2 content was determined using the selective sorption of gaseous ignition products. This article describes a semi-quantitative method for estimating the content of carbon dioxide molecules in these minerals, based on IR spectroscopy data. The data obtained demonstrate the existence of a sulfide sodalite-group mineral with the idealized formula $\text{Na}_7(\text{Si}_6\text{Al}_6\text{O}_{24})(\text{S}_3^-) \cdot \text{H}_2\text{O}$, which differs significantly from the formula $\text{Na}_6\text{Ca}_2(\text{Si}_6\text{Al}_6\text{O}_{24})\text{S}^{2-}_2$ accepted for lazurite. According to single-crystal X-ray structural analysis, in the F-rich sodalite-group mineral from the Eifel paleovolcanic region, Germany with the idealized formula $\text{Na}_7(\text{Si}_6\text{Al}_6\text{O}_{24})\text{F}^- \cdot n\text{H}_2\text{O}$ fluorine occurs as an isolated F^- anion, unlike synthetic F-rich sodalite-type compounds.

Keywords: mineral; sodalite group; extra-framework species; infrared spectroscopy; Raman spectroscopy; crystal structure; polysulfide radical anion

1. Introduction

Natural aluminosilicates belonging to the sodalite group are feldspathoids whose crystal structures are based on the tetrahedral framework $(\text{Al}_x\text{Si}_{1-x}\text{O}_2)^{x-}$ with large cavities (so-called β -cages) hosting large cations (Na^+ , K^+ , Ca^{2+}), additional anions (Cl^- , F^- , SO_4^{2-} , S^{2-} , S_3^- , etc.) and neutral molecules (H_2O , CO_2) [1–11]. Sodalite-type compounds with different additional anions (CO_3^- , HCOO^- , AlF_6^{3-} , OH^-) have been synthesized [12–16].

Vibrational (infrared and Raman) spectroscopy has proven to be an effective tool with which various anionic and neutral extra-framework groups in minerals of the sodalite group can be identified. However, a number of unresolved problems remain related to the quantitative determination of the relative contents of various groups containing C, S, and F, the limits of carbon dioxide contents, the

writing of chemical formulae for members of the sodalite group with high sulfur and fluorine contents, as well as the present-day problem with the status of lazurite in the nomenclature of minerals. In this work, a combination of infrared and Raman spectroscopy was applied to the study of extra-framework anionic and neutral groups in nine samples of sodalite-group minerals with contrastingly different chemical compositions, including samples with abnormally high contents of polysulfide groups, fluoride anion and carbon dioxide molecules. In addition, this paper describes a semi-quantitative method for estimating the content of CO₂ molecules in these minerals, based on IR spectroscopy data.

2. Materials and Methods

The samples investigated in this work are as follows.

Sample 1 is blue sodalite, forming grains up to 3 cm across in a cancrinite–sodalite pegmatite uncovered at the western slope of Mount Kobelikha, in the bench bottom of the open pit situated in the Vishnevogorskiy syenite-miaskite complex, South Urals, Russia. Associated minerals are potassium feldspar, cancrinite, annite, muscovite, different zeolites, as well as accessory zircon, thorite, ilmenite, columbite-(Fe), Nb-bearing rutile, a pyrochlore-group mineral, various sulfides etc. [17].

Sample 2 is gray nosean forming crystals up to 0.5 cm across and twins on (111) up to 1 cm long in cavities of nosean sanidinite in the In den Dellen (Zieglowski) pumice quarry, 1.5 km NE of Mendig, Laach Lake (Laacher See) paleovolcano, Eifel region, Rhineland-Palatinate, Germany. Metasomatic nosean sanidinites of the Laach Lake area are derivatives of foyaite magmas which are comagmatic with nosean-cancrinite-nepheline syenites [18–22]. Associated minerals are sanidine, augite, annite, and zircon.

Sample 3 is blue rock-forming SO₄-bearing triclinic lazurite (intermediate member of the haüyne–lazurite solid-solution series) from the Malobystrinskoe lazurite deposit, Baikal Lake area, Siberia, Russia [23]. Associated minerals are calcite, diopside, pyrite, fluorapatite, and phlogopite. The sample was kindly provided by A.N. Sapozhnikov.

Sample 4 is lilac haüyne forming imperfect crystals up to 0.4 cm in cavities of a metasomatic rock composed of haüyne, calcite and phlogopite from the Tultuy lazurite deposit, Baikal Lake area, Siberia, Russia.

Sample 5 is light blue (with a greenish tint) rock-forming haüyne from the Malobystrinskoe lazurite deposit. Associated minerals are dolomite and diopside.

Sample 6 (Figure 1) comes from the collection of Anatoly V. Kasatkin (catalogue number 569L). It originates from the Malobystrinskoe lazurite deposit and is a fragment of lazurite-rich metasomatic rock. CO₂-rich, SO₄-bearing lazurite (an intermediate member of the haüyne–lazurite solid-solution series) occurs there as bright blue massive aggregates up to 3.5 cm × 2 cm. Other associated minerals are diopside, calcite, dolomite, anhydrite, and phlogopite.

Sample 7 is a large (2 cm × 3 cm × 4 cm) deep blue twin on (111) of sulfide-rich lazurite from calcite marble (Figure 2). The sample originates from the Sar-e Sang lazurite deposit, Badakhshan, Afghanistan, which is the type locality of lazurite.

Sample 8 is lazurite forming a deep blue outer zone of a rhombic dodecahedral crystal 2 cm across from calcite marble (Figure 3). The inner zone of the crystal is composed of gray haüyne. The sample originates from the Sar-e Sang lazurite deposit.

Sample 9 is an F-rich sodalite-group mineral forming yellowish twins on (111) up to 1 mm long in the cavities of nepheline- and biotite-bearing trachitoid sanidinite from the In den Dellen (Zieglowski) pumice quarry (Figure 4).

Sample 10 (used for comparison) is a synthetic F-rich sodalite-type compound Na_{7.38}(Si_{6.74}Al_{5.26}O₂₄)(AlF₆)_{0.70}·4.88H₂O synthesized hydrothermally (at $T = 650$ °C, $P = 2$ kBar) in a Si-Al-Na-F-H₂O system [14]. The sample was kindly provided by A.R. Kotel'nikov.



Figure 1. Bright blue CO₂-rich, SO₄-bearing lazurite with orange diopside in metasomatic rock. Sample 6. Field width: 9 cm. Photographer: Anastasia D. Kasatkina.



Figure 2. Twin of a deep blue sulfide-rich lazurite extracted from calcite marble (white). Sample 7. Field width: 4.5 cm. Photographer: Nikita V. Chukanov.

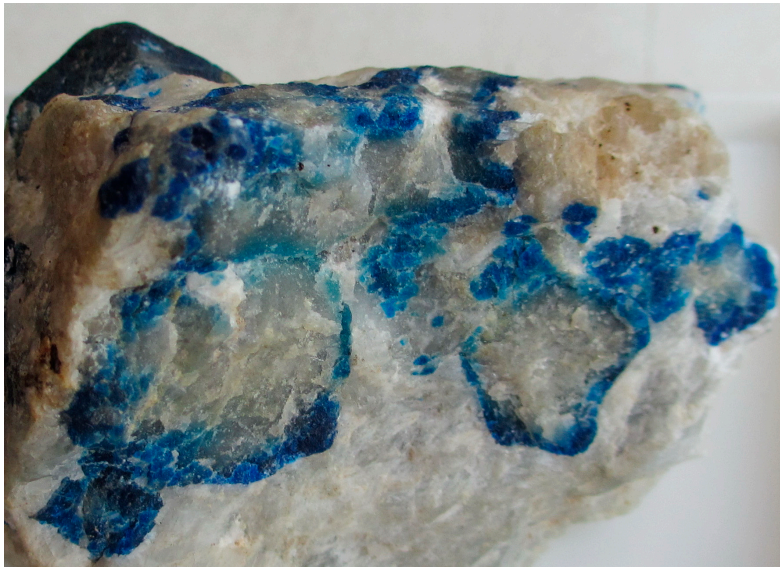


Figure 3. Zoned haiyune–lazurite crystals in calcite marble. Sample 8. Field width: 6 cm. Photographer: Nikita V. Chukanov.



Figure 4. Twin of a F-rich sodalite-group mineral on sanidine. Sample 9. Photographer: Marko Burkhardt. Field width: 1 mm.

Table 1 lists the analytical methods which were used for the characterization of different samples (“+” means that the method was applied). Samples 3, 6, 7, and 9 with unusual compositional characteristics were investigated in the most detail. Samples 1, 2, 4, 5, and 8 played an auxiliary role and were used to obtain calibration dependence for the determination of the content of CO₂ molecules from the IR spectra. Some data for Samples 1 and 10 were taken from literature sources, indicated in Table 1.

Table 1. Techniques applied for the characterization of Samples 1 to 10.

Sample No.	EDS	CO ₂ Sorption	IR Spectroscopy	Raman Spectroscopy	Single-Crystal XRD
1	[17]	+	+	–	–
2	+	+	+	–	–
3	+	+	+	+	–
4	+	+	+	–	–
5	+	+	+	–	–
6	+	+	+	+	–
7	+	+	+	+	–
8	+	+	+	–	–
9	+	–	+	+	+
10	[14]	–	+	+	–

Electron microprobe analyses (three spot analyses for each sample) were carried out using a Tescan VEGA-II XMU electronic microscope (EDS mode, 20 kV, 400 pA). Data reduction was carried out by means of a modified INCA Energy 450 software package. The size of the electron beam was 157–180 nm. The beam was rastered on an area 20 $\mu\text{m} \times 20 \mu\text{m}$ in order to minimize sample damage. The time of data acquisition was 50 s. The sample-to-detector distance was 25 mm. The standards used are: albite for Na and Si, sanidine for K, wollastonite for Ca, Al₂O₃ for Al, BaSO₄ for S, CaF₂ for F, and NaCl for Cl.

The content of fluorine in the F-rich Sample 9 was determined as the mean of 11 spot analyses using a Jeol JSM-6480LV scanning electron microscope equipped with an INCA-Wave 500 wavelength-dispersive spectrometer (Laboratory of Analytical Techniques of High Spatial Resolution, Department of Petrology, Moscow State University), with an acceleration voltage of 20 kV, a beam current of 20 nA, and a 3 μm beam diameter. MgF₂ was used as a reference.

For all samples except Sample 7, CO₂ was determined by the selective sorption of CO₂ on an *askarite* sorbent (an asbestiform matter saturated by NaOH) from gaseous products obtained by heating of the mineral at 1080 °C in oxygen at 1 atm. For Sample 7, CO₂ and H₂O were determined by gas chromatography of products of ignition in oxygen at 1200 °C with a Vario Micro cube analyzer (Elementar GmbH, Frankfurt, Germany).

In order to obtain infrared (IR) absorption spectra, powdered samples were mixed with dried KBr, pelletized, and analyzed using an ALPHA FTIR spectrometer (Bruker Optics) in the range 360–4000 cm^{-1} with a resolution of 4 cm^{-1} and 16 scans. An IR spectrum of an analogous pellet prepared from pure KBr was used as a reference. The optical densities of the bands at 2341–2346 cm^{-1} were measured relative to the optical density at 2500 cm^{-1} . The optical densities of the band used as an internal standard was measured at the absorption maximum in the range 656–668 cm^{-1} ; the common tangents to the two deepest minima in the range 500–830 cm^{-1} were used as baselines.

Raman spectra of randomly oriented samples were obtained using an EnSpectr R532 spectrometer based on an OLYMPUS CX 41 microscope coupled with a diode laser ($\lambda = 532 \text{ nm}$) at room temperature. The spectra were recorded in the range from 100 to 4000 cm^{-1} with a diffraction grating (1800 g mm^{-1}) and spectral resolution of about 6 cm^{-1} . The output power of the laser beam was about 5 mW. The diameter of the focal spot on the sample was less than 5 μm . The backscattered Raman signal was collected with a 40 \times objective; signal acquisition time for a single scan of the spectral range was 1 s, and the signal was averaged over 50 scans. Crystalline silicon was used as a standard.

The single-crystal X-ray diffraction experiment was carried out for Sample 9 using Xcalibur S CCD diffractometer. The measured intensities were corrected for Lorentz, background, polarization and absorption effects. Data reduction was performed using CrysAlisPro Version 1.171.39.46 [24]. The studied sample is cubic, space group *I-43m*, $a = 9.05887(10) \text{ \AA}$, $V = 743.40(2) \text{ \AA}^3$. The crystal structure of Sample 9 was obtained by direct methods and refined using the SHELX software package [25]. The structure was refined for the two-component twin with refined twin ratio 0.52:0.48. Experimental details are given in Table 2.

Table 2. Crystal data, data collection information and structure refinement details for sample 9.

Crystal sizes, mm	0.07 × 0.11 × 0.13
Temperature, K	293
Radiation and wavelength, Å	MoK α ; 0.71073
F ₀₀₀ *	461
Diffractometer	Xcalibur S CCD
θ range for data collection, °	5.514–30.359
h, k, l ranges	−12 ≤ h ≤ 12, −12 ≤ k ≤ 12, −12 ≤ l ≤ 12
Reflections used in the refinement: total/with $I > 2\sigma(I)$	728/725
Data reduction	CrysAlisPro Version 1.171.39.46
Absorption correction	multi-scan
Refinement method	Full-matrix least-squares on F^2
Number of refined parameters	21
Final R indices (with $I > 2\sigma(I)$) $R1/wR2$	0.0430/0.1050
R indices (with all data)	0.0433/0.1052
GoF	1.161
Largest diffraction peak and hole, e/Å ³	0.71 and −0.33

* Calculated based on structural data.

3. Results

3.1. Chemical Composition

Chemical data for the studied sodalite-group minerals are given in Table 3. The CO₂ content in Sample 9 was determined by means of IR spectroscopy (see below).

Table 3. Chemical composition of sodalite-group minerals.

Compo-nent	Sample No.									
	1 ^a	2	3	4	5	6	7	8	9	10 ^b
Na ₂ O	24.97	16.96	18.60	18.33	19.02	20.66	17.88	18.08	19.00	21.5
K ₂ O	bdl	2.90	bdl	0.13	bdl	0.66	1.06	0.53	4.04	bdl
CaO	bdl	1.77	4.80	6.14	6.54	2.17	0.77	6.25	0.59	bdl
Al ₂ O ₃	31.86	26.85	25.75	26.86	28.00	26.28	24.44	27.33	30.21	32.2
Fe ₂ O ₃	bdl	0.47	0.31	0.52	0.23	0.58	0.15	0.31	bdl	bdl
SiO ₂	35.20	36.97	30.83	33.41	33.80	33.28	33.52	33.52	39.83	42.7
CO ₂	bdl	0.48	1.35	0.78	0.24	1.81	0.82	0.65	1.66	-
SO ₃	bdl	9.27	14.97	10.41	11.56	11.88	19.01	13.07	1.42	bdl
F	bdl	0.17	bdl	0.21	0.70	bdl	bdl	0.27	1.65	6.3
Cl	7.97	0.37	bdl	0.27	0.45	0.20	0.56	0.40	0.62	bdl
−O=(F,Cl)	−1.80	−0.15	0	−0.15	−0.40	−0.05	−0.13	−0.20	−0.83	−2.7
Total	98.20	96.06	96.61	96.91	100.14	97.47	98.96^c	100.21	98.19	100.0
Formula coefficients calculated on Al + Si + Fe = 12 atoms per formula unit										
Na	7.98	5.72	7.05	6.52	6.61	7.43	6.66	6.38	5.86	6.21
K	0	0.64	0	0.03	0	0.16	0.26	0.12	0.82	0
Ca	0	0.33	1.00	1.21	1.26	0.43	0.16	1.22	0.10	0
Al	6.20	5.51	5.93	5.80	5.91	5.75	5.54	5.86	5.66	5.65
Fe	0	0.06	0.05	0.07	0.03	0.08	0.02	0.04	0	0
Si	5.80	6.43	6.02	6.13	6.06	6.17	6.44	6.10	6.34	6.35
C	0	0.11	0.36	0.20	0.06	0.46	0.22	0.16	0.36	0
S	0	1.21	2.20	1.43	1.55	1.66	2.75	1.78	0.17	0
F	0	0.09	0	0.12	0.05	0	0	0.16	0.83	2.99
Cl	2.22	0.11	0	0.08	0.32	0.06	0.18	0.12	0.17	0

^a Data from [17]. ^b Data from [14]. ^c The total includes 0.88 wt.% of H₂O which corresponds to 1.13 H₂O molecules. bdl: below detection limit.

It is to be noted that the CO₂ contents in Samples 2–9 correspond to the CO₂ molecules in the sodalite cavities because IR spectra of these samples do not contain characteristic bands of the CO₃²⁻ anion. Samples 6, 7 and 9 are characterized by abnormally high contents of carbon, sulfur and fluorine, respectively.

3.2. Infrared Spectroscopy

The IR spectra of the studied Samples 1–10 are given in Figures 5–8. All of them contain bands of O–H stretching and H–O–H bending vibrations (in the ranges 3400–3700 and 1640–1660 cm⁻¹, respectively) which are due to the presence of H₂O molecules (zeolitic water). Taking into account that in the sodalite-type framework, each tetrahedral atom is coordinated by two O atoms, the presence of OH groups connected to the framework should be excluded. Additional bands in the ranges 3200–3400 and 1687–1709 cm⁻¹ in the IR spectra of Samples 1, 2, 9 and 10 may correspond to H₃O⁺ or H₅O₂⁺ cations [26]. This assumption is confirmed by the presence of a band in the range 1339–1358 cm⁻¹ in the IR spectra of these four samples whereas IR spectra of other samples do not contain this band. The IR band in the range 1320–1390 cm⁻¹ is a characteristic feature of isolated H⁺ cation which may be formed as a result of dissociation of acid groups: H₃O⁺ → H⁺ + H₂O; H₅O₂⁺ → H⁺ + 2H₂O [26].

Characteristic IR bands of S-bearing groups in the sodalite-group minerals are observed in the ranges 1124–1142 cm⁻¹ (asymmetric stretching vibrations of the SO₄²⁻ group, the F₂(ν₃) mode), 614–637 cm⁻¹ (bending vibrations of the SO₄²⁻ group, the F₂(ν₄) mode) and 575–585 cm⁻¹ (antisymmetric stretching vibrations of the S₃⁻ radical anion, the ν₃ mode—A weak band) [26–30]. Taking into account the overlapping of the bands of asymmetric stretching vibrations of the SO₄²⁻ group and antisymmetric T–O–T stretching (T = Al, Si) vibrations of the tetrahedral framework in the range 1020–1150 cm⁻¹, the band in the range 614–627 cm⁻¹ should be considered as the most reliable indication of the presence of the sulfate anion in the sodalite-group minerals. This anion was reliably detected in Samples 2–9. The characteristic band of the S₃⁻ radical anion is observed in the IR spectra of Samples 5 and 7. It is to be noted that the S²⁻ anion does not have inner degrees of freedom and, consequently, its absorption bands could not be observed in the wavenumber range 360–3800 cm⁻¹.

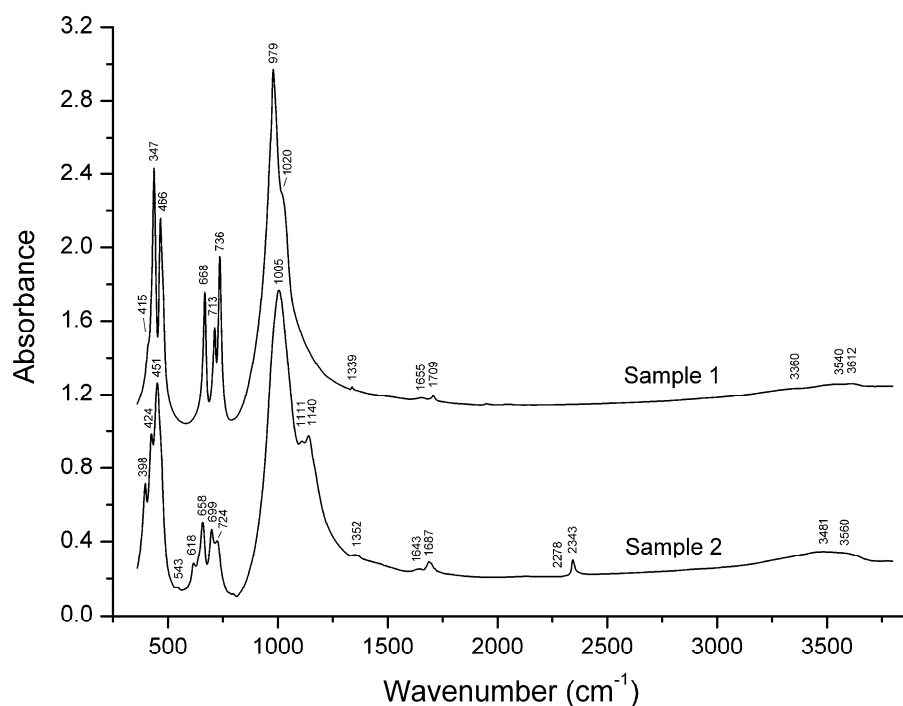


Figure 5. Infrared absorption spectra of sodalite and nosean (Samples 1 and 2).

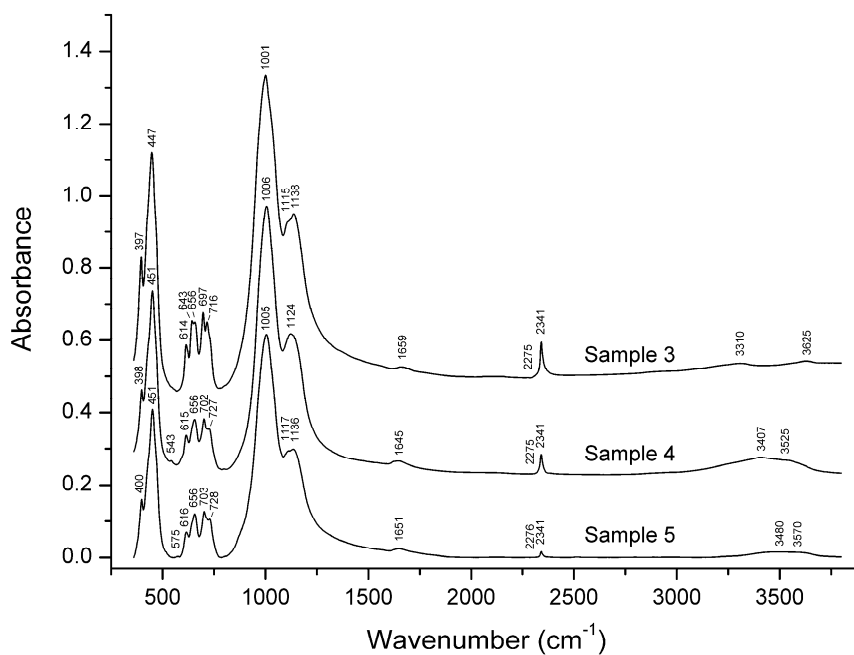


Figure 6. Infrared absorption spectra of the sulfate-rich minerals belonging to the haiyue–lazurite solid-solution series (Samples 3 to 5).

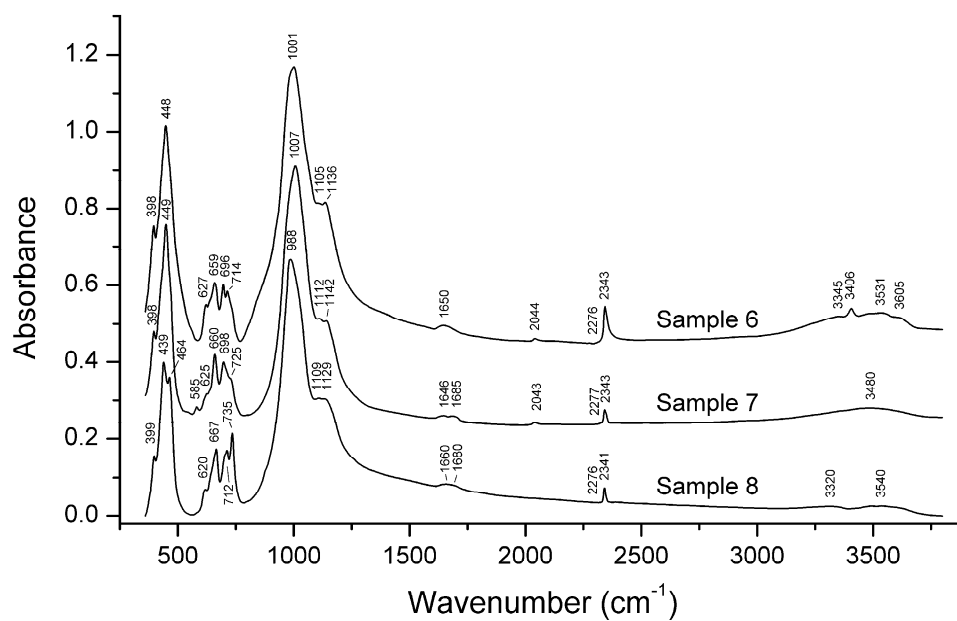


Figure 7. Infrared absorption spectra of the sulfate-poor minerals belonging to the haiyue–lazurite solid-solution series (Samples 6 to 8).

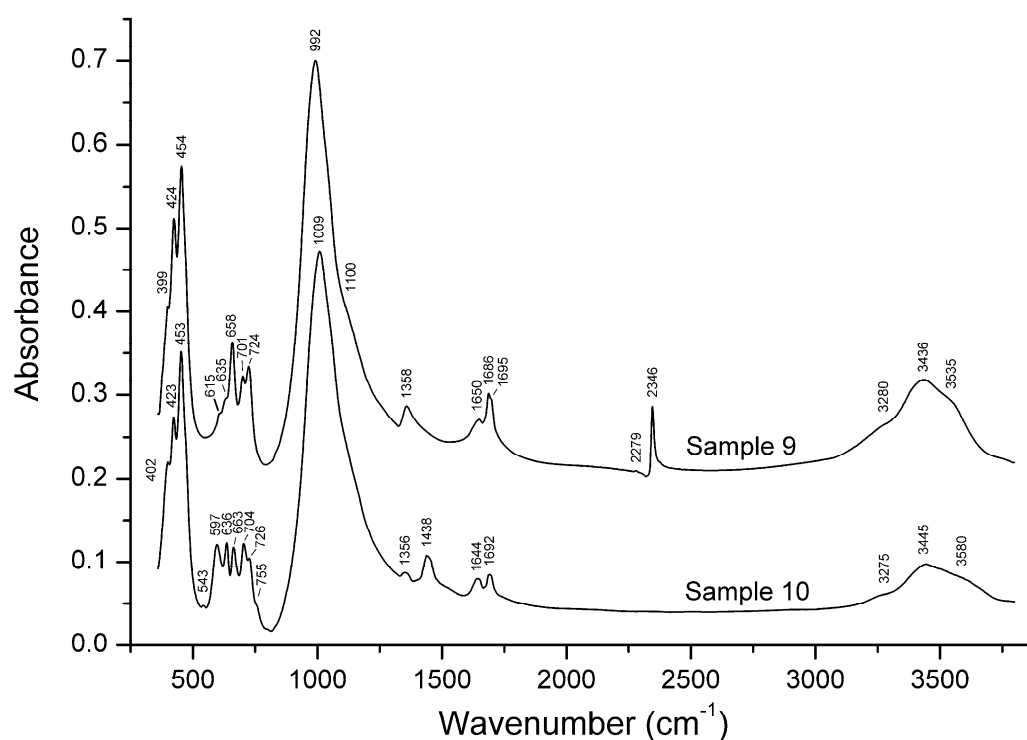


Figure 8. Infrared absorption spectra of the F-rich sodalite-group mineral (Sample 9) and synthetic fluoraluminate sodalite-type compound (Sample 10).

The three bands at about 670, 710 and 740 cm^{-1} are proof that the sodalite-type structure exists. In some spectra, including those of Samples 3, 4 and 7, the arrangements of these three characteristic bands are disturbed. Most probably, this phenomenon is due to structure modulation, which is a characteristic feature of sodalite-type minerals from metasomatic lazurite deposits [23].

The strongest band of the CO_3^{2-} group in sodalite cavities is observed in the range 1436–1450 cm^{-1} [30,31]. In the IR spectra of natural samples investigated in this work, this band is absent, but a band at 1438 cm^{-1} is observed in the IR spectrum of the synthetic sodalite-type compound, with the presumed crystal-chemical formula $\text{Na}_{7.38}(\text{Si}_{6.74}\text{Al}_{5.26}\text{O}_{24})(\text{AlF}_6)_{0.70}\cdot 4.88\text{H}_2\text{O}$ (Sample 10 [14]).

The IR spectrum of Sample 10 was discussed elsewhere [15]. In accordance with structural data and by analogy with IR spectra of natural fluoraluminates [30], the band at 597 cm^{-1} corresponds to stretching vibrations of the AlF_6^{3-} anion. This band is absent in the IR spectra of all other samples investigated in this work, including natural F-rich sodalite-type mineral (Sample 9).

A narrow band at 2341–2346 cm^{-1} present in the IR spectra of Samples 2–9 is due to the antisymmetric stretching vibrations of CO_2 molecules [11,32]. This band is accompanied by a very weak isotope satellite at 2275–2279 cm^{-1} corresponding to $^{13}\text{C}^{16}\text{O}_2$. Samples 3, 6 and 9 are the most CO_2 -rich.

The optical density of the band at 2341–2346 cm^{-1} (D) shows positive correlation with the content of CO_2 determined by selective sorption (Figure 9). Consequently, IR spectroscopy can be used for the estimation of the content of CO_2 molecules in sodalite-group minerals.

We failed to assign the weak band at 2043–2044 cm^{-1} observed in the IR spectra of Samples 6 and 7. Hypothetically, this band may be associated with a minor admixture of an extra-framework component with the S–H covalent bond.

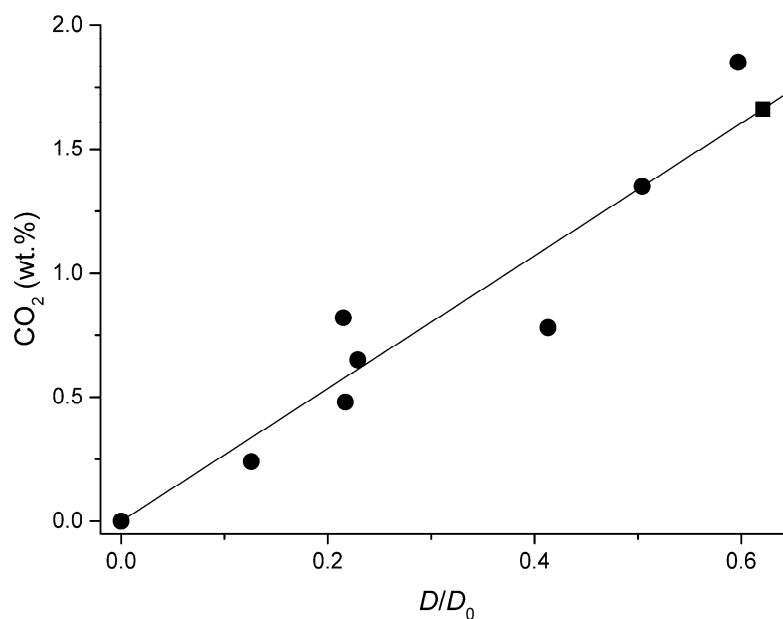


Figure 9. The correlation between the content of CO₂ determined by the selective sorption and relative optical density of the band at 2341–2346 cm^{−1} (D_0 is the optical density at the maximum of the band at 656–668 cm^{−1} used as an internal standard). The circles correspond to Samples 1–8 and the square corresponds to Sample 9 with the CO₂ content determined using this correlation.

3.3. Raman Spectroscopy

Among the nine minerals of the sodalite group studied in this work, there are several samples characterized by anomalous features of the chemical composition (Table 3). These are Samples 3, 6 and 9 with high contents of carbon dioxide molecules (≥ 0.36 CO₂ groups *pfu*), Sample 7 with a high content of sulfur (2.74 atoms *pfu*), and Sample 9 characterized by a very high content of fluorine (1.81 wt.%, 0.83 atoms *pfu*), which has never been noted in natural sodalites before. Among minerals, a higher content of CO₂ molecules (2.2 wt.%) was detected only in cordierite [33]. Figures 10 and 11 show the Raman spectra of these samples as well as the Raman spectrum of the synthetic F-rich compound with the idealized formula Na₈(Si₇Al₅O₂₄)(AlF₆)·5H₂O (Sample 10).

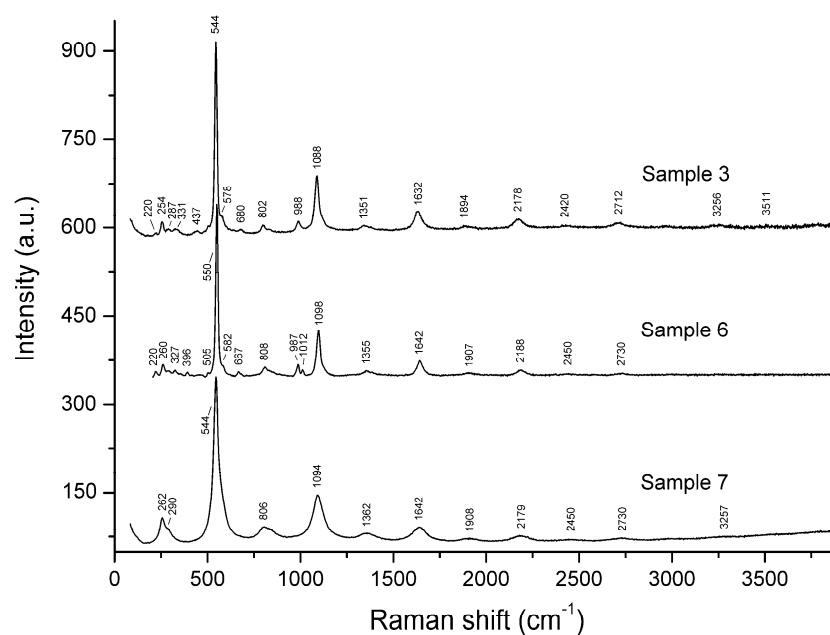


Figure 10. Raman spectra of the members of the h aüyne–lazurite solid-solution series.

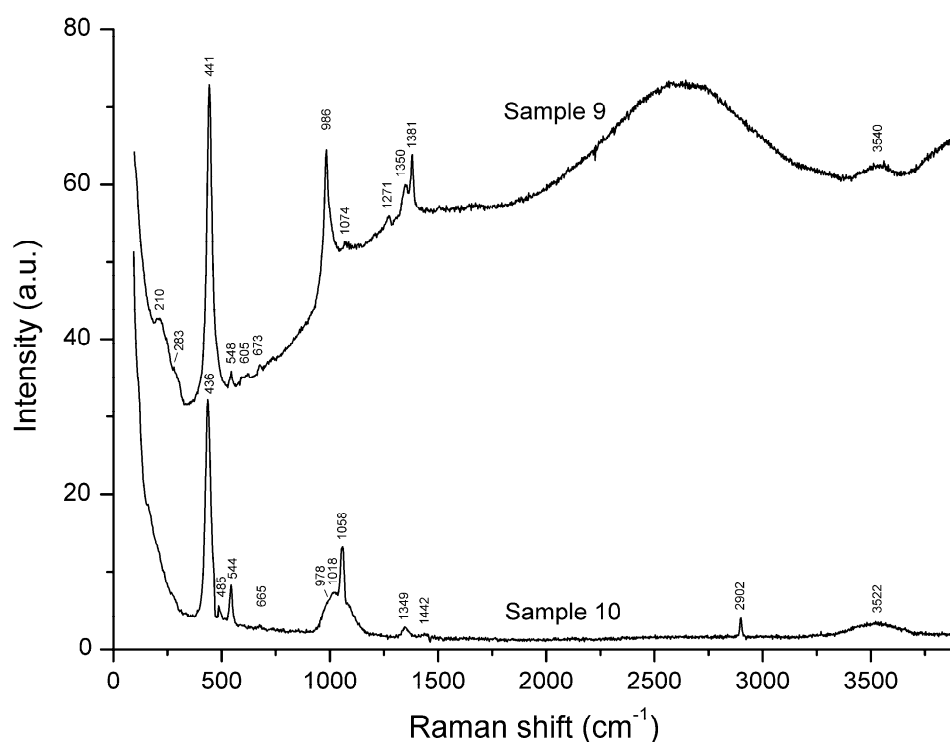


Figure 11. Raman spectra of the F-rich sodalite-group mineral (Sample 9) and the synthetic fluoraluminate sodalite-type compound (Sample 10).

The F-rich Sample 9 shows a strong luminescence under a laser beam (the broad feature in the range 1900–3200 cm^{-1} ; Figure 11). The cause of this phenomenon is unknown. Another specific feature of the Raman spectrum of Sample 9 is the presence of the weak bands at 605 and 1074 cm^{-1} which are absent in the Raman spectra of other samples and may correspond to translational and librational modes of HF [34].

The assignment of Raman bands is given in Table 4. For the assignment of the bands at 327–331 and 667–680 cm^{-1} to the vibrations of the S_4^- radical anion, see the Discussion section.

Table 4. Raman bands of sodalite-group minerals and Sample 10, and their assignment.

Sample 3	Sample 6	Sample 7	Sample 9	Sample 10	Assignment
Raman Shift (cm^{-1})					
220w	220w		210		Combination of low-frequency lattice modes
254	265	262			S_3^- bending (ν_2)
287w		290sh	283sh		Combination of low-frequency lattice modes
331w	327w				S_4^- symmetric stretching (ν_1-A_1)
	396				Anhydrite admixture
437w			441s	436s	SO_4 bending and/or $\delta[\text{O}-\text{Si}(\text{Al})-\text{O}]$
				485	AlF_6 stretching.
	505w				Anhydrite admixture
544s	550s	544s	548w	544	S_3^- symmetric stretching (ν_1) and/or AlF_6 stretching

Table 4. Cont.

Sample 3	Sample 6	Sample 7	Sample 9	Sample 10	Assignment
Raman Shift (cm ⁻¹)					
578sh	582sh				S ₃ ⁻ antisymmetric stretching (ν_3)
			605w		HF translational mode
			673w	665w	HF translational mode?
680w	667w				S ₄ ⁻ stretching (ν_3 -E)
802	808	806			S ₃ ⁻ combination mode ($\nu_1 + \nu_2$)
988	987		986s	978sh	SO ₄ symmetric stretching and/or stretching vibrations of the framework
	1012w				Anhydrite admixture
				1018 br	Admixture of aluminosilicate glass
				1058s	CO ₃ symmetric stretching
			1074w		HF librational mode
1088s	1098s	1094s			S ₃ ⁻ overtone ($2 \times \nu_1$)
			1271w		CO ₂ Fermi resonance
			1350	1349	H ⁺
1351	1355	1362			S ₃ ⁻ combination mode ($2 \nu_1 + \nu_2$)
			1381		CO ₂ Fermi resonance
				1442w	CO ₃ asymmetric stretching
1632	1642	1642			S ₃ ⁻ overtone ($3 \times \nu_1$)
1894	1907w	1908w			S ₃ ⁻ combination mode ($3 \times \nu_2 + \nu_1$)
2178	2188	2179			S ₃ ⁻ overtone ($4 \times \nu_1$)
2420w	2450w	2450w			S ₃ ⁻ combination mode ($4 \times \nu_2 + \nu_1$)
2712	2730w	2730w			S ₃ ⁻ overtone ($5 \times \nu_1$)
				2902	CH ₄
3256		3257w			S ₃ ⁻ overtone ($6 \times \nu_1$)
3511w			3540	3522	H ₂ O stretching vibrations

Note: w—weak band, s—strong band, sh—shoulder, br—broad band.

3.4. X-ray Diffraction Data and Crystal Structure

The studied sample of the F-rich sodalite-group mineral (Sample 9) retains the sodalite-type framework built by corner-linked TO₄ tetrahedra ($T = \text{Si, Al}$). Extra-framework cations (Na⁺ and subordinate K⁺ and Ca²⁺ statistically replacing each other), anions (F⁻, Cl⁻, SO₄²⁻) and neutral molecules (H₂O, CO₂) occur in β -cages. The refined F content of 1.54 *apfu* is significantly higher than the value of 0.83 F *apfu* in the empirical formula. On the other hand, no separate sites were found for H₂O and CO₂ molecules. For this reason, we assume that these components may replace F anions. Additional water molecules may occur with a very low occupancy around O atoms of SO₄ tetrahedron (OS). The Na site is split onto three subsites with partial occupancies. The site occupancy factors for all Na subsites were refined using Na scattering curve and U_{anis} restrained to be equal; then, K (K = K + Ca) admixture was added according to chemical data and the refined number of electrons.

Atom coordinates and displacement parameters for Sample 9 are given in Table 5 and interatomic distances in Table 6. The crystal structure of Sample 9 is shown in Figure 12.

Table 5. Atom coordinates, equivalent displacement parameters (in Å²), site occupancy factors (s.o.f.) and site multiplicities (Q) for the Sample 9.

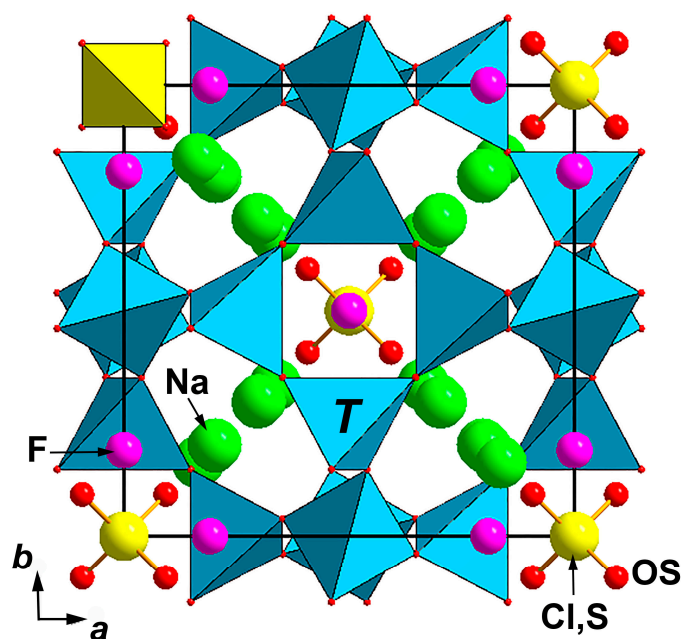
Site	<i>x</i>	<i>y</i>	<i>z</i>	<i>U</i> _{eq}	s.o.f *	Q
<i>T</i>	0.25	0.5	0.0	0.0102(5)	Si _{0.528} Al _{0.472}	12
O	0.1480(4)	0.1480(4)	0.4605(5)	0.0275(12)	O _{1.00}	24
Na1	0.297(4)	0.297(4)	0.297(4)	0.060(8)	Na _{0.159} K _{0.024}	8
Na2	0.164(3)	0.164(3)	0.164(3)	0.060(8)	Na _{0.288} K _{0.045}	8
Na3	0.198(3)	0.198(3)	0.198(3)	0.060(8)	Na _{0.288} K _{0.045}	8
Cl, S	0.0	0.0	0.0	0.024(4)	Cl _{0.096} S _{0.078}	2
OS	0.4074(16)	0.4074(16)	0.4074(16)	0.024(4)	O _{0.078}	8
F	0.0	0.0	0.191(4)	0.034(10)	F _{0.128}	12

* Fixed on the last stages of the refinement, taking into account both structural and chemical data, except F site which showed the s.o.f. significantly higher and is assumed to be partially replaced by H₂O and/or CO₂.

Table 6. Selected interatomic distances (Å) in the structure of Sample 9.

Na1–O	2.415(19) × 3	Na3–O	2.462(16) × 3
–F	2.61(5) × 3	–F	2.54(3) × 3
–O	3.13(2) × 3	–O	2.919(7) × 3
–Cl	3.19(6)	–OS	2.96(4) × 3
		–Cl	3.11(4)
Na2*–OS	2.50(4) × 3	<i>T</i> –O	1.6671(14) × 4
–O	2.70(2) × 3	S–OS	1.45(2) × 4
–O	3.035(13) × 3		

* The Na2 subsite is connected with F and Cl atoms with short distances of 2.11 Å (Na2–F) and 2.57 Å (Na2–Cl); for this reason, these bonds were excluded from the coordination sphere of Na2.

**Figure 12.** The crystal structure of the F-rich sodalite-group mineral (Sample 9).

4. Discussion

The empirical formulae of the studied sodalite-group minerals were calculated based on chemical data (Table 3) and taking into account the above data on the infrared and Raman spectra. According to the spectroscopic data, the anomalously high-sulfur Sample 7 from Sar-e Sang (the type locality of lazurite) has a high content of the radical anions S₃[−] and a low content of the SO₄^{2−} groups; other polysulfide anions are absent, and the presence of the

monosulfide anion S^{2-} cannot be excluded. Consequently, the charge-balanced formula of Sample 7 is $(Na_{6.66}K_{0.26}Ca_{0.16})(Si_{6.44}Al_{5.54}Fe^{3+}_{0.02}O_{24})(S_3^-)_{0.80}(S^{2-},SO_4^{2-})_{0.35}Cl_{0.18}(CO_2)_{0.22}\cdot 1.13H_2O$. The idealized formula of this sample is $Na_7(Si_6Al_6O_{24})(S_3^-)\cdot H_2O$. This formula differs significantly from the idealized formula $Na_3Ca(Si_3Al_3)O_{12}S$ (another spelling being $Na_6Ca_2(Si_6Al_6O_{24})S^{2-}_2$) of lazurite given in the current IMA list of mineral species [35] together with the reference [9]. Moreover, the latter formula contradicts the results of X-ray structural analysis of “lazurite” from [9], which correspond to a SO_4 -dominant mineral.

The charge-balanced empirical formula of another sample from Sar-e Sang (Sample 8) is $(Na_{6.38}Ca_{1.22}K_{0.12})(Si_{6.10}Al_{5.86}Fe^{3+}_{0.04}O_{24})(S^{2-},SO_4^{2-})_{1.30}(S_3^-)_{0.16}(CO_2)_{0.16}\cdot nH_2O$, which is close to the lazurite formula from the IMA list of mineral species. The charge-balanced formula of one more S^{2-} -dominant lazurite sample (Sample 3) is $(Na_{7.05}Ca_{1.00})(Si_{6.02}Al_{5.93}Fe^{3+}_{0.05}O_{24})(S^{2-},SO_4^{2-})_{1.40}(S_3^-)_{0.27}(CO_2)_{0.36}\cdot nH_2O$. These examples show that several different sulfide members of the sodalite group (with different contents of Ca, H_2O and CO_2 and with different dominant sulfide groups) may exist, and the status of lazurite as a mineral species needs to be specified.

The charge-balanced empirical formula of the most CO_2 -rich lazurite (Sample 6) is $(Na_{7.43}K_{0.16}Ca_{0.43})(Si_{6.17}Al_{5.75}Fe^{3+}_{0.08}O_{24})(S^{2-},SO_4^{2-})_{1.21}(S_3^-)_{0.15}Cl_{0.06}(CO_2)_{0.46}\cdot nH_2O$ (here, we neglected trace amounts of the S_4^- anions showing a weak Raman band at 327 cm^{-1}). Taking into account the low intensities of the bands of sulfate groups in the IR spectrum of Sample 6, one can conclude that in this sample, sulfate anion plays a subordinate role. Consequently, the idealized formula of Sample 6 is $Na_8(Si_6Al_6O_{24})S^{2-}_2\cdot nH_2O$, which also differs significantly from the formula $Na_6Ca_2(Si_6Al_6O_{24})S^{2-}_2$ accepted currently for lazurite. Moreover, one cannot exclude the existence of a sulfide sodalite-group mineral with species-defining CO_2 molecules. A higher CO_2 content (2.70 wt.%) was detected in nosean from Bassano Romano, Viterbo, Lazio, Italy [11], but this sample is characterized by a relatively low intensity of the stretching band of carbon dioxide and, therefore, most of the carbon in it belongs to carbonate anions.

Structural carbon dioxide is a typical admixed component of different silicate minerals with microporous and channel structures. The well-known examples are different members of the cancrinite group [36–40], cordierite [41–44], beryl [45,46], pollucite [47], melanophlogite [48] etc. Sodalite-group minerals are not an exclusion from this trend.

The radical ion $(S_3)^-$ is considered by most researchers as the main cause of blue color of lazurite [49–56]. Our data confirm this conclusion.

In the review [57], it is stated that “The radical anion $[S-S-S]^*$ has not yet been observed convincingly by Raman spectroscopy”. However, it is to be noted that the structure and Raman spectrum of the tetrasulfide radical anion is discussed in detail in [54]. The authors of this paper concluded that $(S_4)^-$ has the flat star shape, D_{3h} . The calculated wavenumbers of Raman active bands $\nu_1(A_1) = 329$, $\nu_3(E) = 680$, and $\nu_4(E) = 345\text{ cm}^{-1}$ are in excellent agreement with experimental data. The intensities of the bands ν_1 and ν_3 are nearly equal, whereas ν_4 is observed as a weak shoulder. In our Raman spectra of Samples 3 and 6, the weak bands ν_1 and ν_3 also have nearly equal intensities, and the band ν_4 is not observed. In the Raman spectra of F-rich and S-depleted Samples 9 and 10, the bands at 673 and 665, which are not accompanied by the bands in the range $320\text{--}340\text{ cm}^{-1}$, are tentatively assigned to the vibrations of the HF molecule as a whole, in accordance with [34].

The IR and Raman spectra indicate that F-rich Sample 9 contains sulfate groups, H_2O and CO_2 molecules. As noted above, the band at $1350\text{--}1358\text{ cm}^{-1}$ observed in both IR and Raman spectra may correspond to vibrations of the H^+ cation in the β -cage. Unlike the synthetic F-rich sodalite-type compound (Sample 10), Sample 9 does not show IR and Raman bands of the AlF_6^{3-} anion. Consequently, the empirical formula of the latter sample is $H^{+}_{0.12}(Na_{5.86}K_{0.82}Ca_{0.10})(Si_{6.34}Al_{5.66}O_{24})F_{0.83}Cl_{0.17}(SO_4)_{0.17}(CO_2)_{0.36}\cdot nH_2O$. The addition of H^+ in the empirical formula is formally required for the charge balance, and is in agreement with the above assumption that Sample 10 contains additional H^+ cation. The most interesting feature of

Sample 9 is a significant predominance of F^- over all other extra-framework anions, which makes it possible to consider this mineral as a potentially new mineral species with the idealized formula $Na_7(Si_6Al_6O_{24})F^- \cdot nH_2O$. Most probably, this mineral crystallized at high fugacities of HF and CO_2 .

Hypothetically, anionic and neutral extra-framework components (Cl^- , OH^- , F^- , AlF_6^{3-} , SO_4^{2-} , S^{2-} , S_3^- , CO_3^{2-} , $C_2O_4^{2-}$, CO_2 , H_2O , CH_4 etc.) occurring in β -cages of minerals belonging to the sodalite and cancrinite groups could be used as important geochemical markers, provided that the relationships between their contents and conditions of mineral formation will be investigated in detail experimentally. In particular, one can suppose that the ratios SO_4^{2-} : ($S^{2-} + S_3^-$) and $(CO_3^{2-} + CO_2)$: $C_2O_4^{2-}$: CH_4 may reflect oxygen fugacity in the mineral-forming medium. The geochemical role of different S-bearing species, including the trisulfur radical anion, is discussed in [29,58].

Author Contributions: N.V.C. and I.V.P. wrote the paper. N.V.C. obtained and interpreted the IR spectra. M.F.V. obtained and interpreted the Raman spectra. N.V.Z. solved the crystal structure. D.Y.P. carried out the crystal-chemical analysis. I.V.P. obtained the crystal data. C.S. and A.V.K. collected the samples 6 and 9, respectively. V.O.Y. obtained electron-microprobe data. All authors have read and agreed to the published version of the manuscript.

Funding: This work was supported by the Russian Foundation for Basic Research, grants nos. 18-29-12007-mk (Raman spectroscopy) and by the Russian Science Foundation, grant no. 19-17-00050 (crystal structure analysis). A part of this work (IR spectroscopy and a part of chemical analyses) has been carried out in accordance with the state task, state registration number AAAA-A19-119092390076-7. The authors are grateful to A.N. Saposhnikov and A.R. Kotelnikov for samples provided for the investigations.

Conflicts of Interest: The authors declare no conflict of interest.

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