

VERY SIMILAR TEST 2

Subject – Chemistry

Maximum Mark - 40

1. Atomic size of neon is than the atomic size of fluorine.

- (a) More (b) Less
(c) Equal (d) None of these

2. The acid salt is

- (a) NaCl (b) NH_4Cl
(c) NaHSO_4 (d) K_2SO_4

3. The element which can't displace H from acid is

- (a) Zn (b) Na
(c) Fe (d) Cu

4. The molecular mass of NaCl is:

(A of Na = 23, Cl = 35.5)

- (a) 58.5 (b) 56.5
(c) 12.5 (d) 22.5

14. The number of electrons mutually shared between atoms to form compound are called

- (a) Electrovalence (b) Donation
(c) Covalency (d) None of these

15. The compound having molecular formula is $C_6H_{12}O_6$ is given. Its empirical formula is CH_2O . The value of n is _____

- (a) 6 (b) 3
(c) 2 (d) 12

16. The electrode where the current leaves the electrolyte is called the _____

- (a) Anode (b) Cathode
(c) Insect (d) None of these

17. The number of electrons present in valence shells of halogen is _____

- (a) 1 (b) 3
(c) 5 (d) 7

- 18.** An amphoteric oxide is
- (a) ZnO (b) Al_2O_3
(c) PbO (d) All of above
- 19.** The reaction occurring at cathode during electrolysis is
- (a) Oxidation (b) Reduction
(c) Redox (d) None of these
- 20.** Anhydrous iron (III) chloride is prepared by
- (a) Direct combination (b) Displacement
(c) Decomposition (d) Neutralization
- 21.** The valency of element having atomic number 19 is
- (a) + 1 (b) - 1
(c) + 2 (d) + 3

22. In periodic table, going down in fluorine group

- (a) Reactivity will increase
- (b) Electronegativity will increase
- (c) Ionic Radii increase
- (d) None of these

23. A compound which liberates reddish brown gas around anode during electrolysis in molten state is

- (a) NaCl
- (b) CuO
- (c) CuSO_4
- (d) P_6Br_2

24. On adding methyl orange to a base, the colour change

- (a) Red to blue
- (b) Orange to yellow
- (c) Colourless to pink
- (d) No change

- 25.** Electrochemical cell is also known as
- (a) Voltmeter (b) Ammeter
(c) Galvanometer (d) None of these
- 26.** The salt of elements are colourless
- (a) Representative (b) Transition
(c) Lanthanides (d) Actinoids
- 27.** In electroplating the anode is made of strips of metal
- (a) Which is to be electroplated
(b) Electrolyte
(c) Both (a) and (b)
(d) None of these

- 28.** A compound have a formula MCl . The formula of its oxide is
- (a) MO (b) M_2O
(c) M_2O_3 (d) M_5O_2
- 29.** The compound which contain all three types of bond *i.e.* ionic covalent and co-ordinate is
- (a) $NaCl$ (b) NH_3
(c) $CaCl_2$ (d) NH_4Cl
- 30.** The selective discharge of ions depends on
- (a) Position of ions in ECS
(b) Concentration of ions
(c) Nature of electrodes
(d) All of above

31. Which is not an oxy acid

- (a) HNO_3 (Nitric acid)
- (b) H_2SO_4 (sulphuric acid)
- (c) Phosphoic acid (H_3PO_4)
- (d) Hydrochloric acid (HCl)

32. During electroplating of an article with silver the electrolyte used is

- (a) Silver nitrate
- (b) Silver cyanide
- (c) Sodium argentocyanide
- (d) Nickel sulphate

33. The 12th element in periodic table resembles with Group

(a) I

(b) II

(c) IV

(d) IX

34. The acid having 4 hydrogen is

(a) H_2SO_4 (sulphuric acid)

(b) H_3PO_4 (Phosphoric acid)

(c) HNO_3 (Nitric acid)

(d) Acetic acid | (CH_3COOH)

35. The empirical formula of compound is CH_2O and molecular formula is same. Value of n is

(a) 2

(b) 5

(c) 1

(d) 6

36. A empirical formula of Hexane is

- (a) C_2H_7 (b) C_5H_8
 (c) C_3H_7 (d) C_4H_{17}

37.

A	3
B	4
C	5
D	6
E	7
F	8
G	9
H	10
I	11
J	12

K	13
L	14
M	15
N	16
O	17
P	18

Study the table and answer the following:

- (A) Which of these is a noble gas
 (a) P (b) H
 (c) K (d) Both a and b
- (B) Which element is a halogen?
 (a) G and O (b) G and P
 (c) P and M (d) D and G
- (C) Which element is an alkali metal
- (a) A (b) C
 (c) L (d) P
- (D) Which element have volency 4?
 (a) D and G (b) D and L
 (c) Only D (d) Only L