
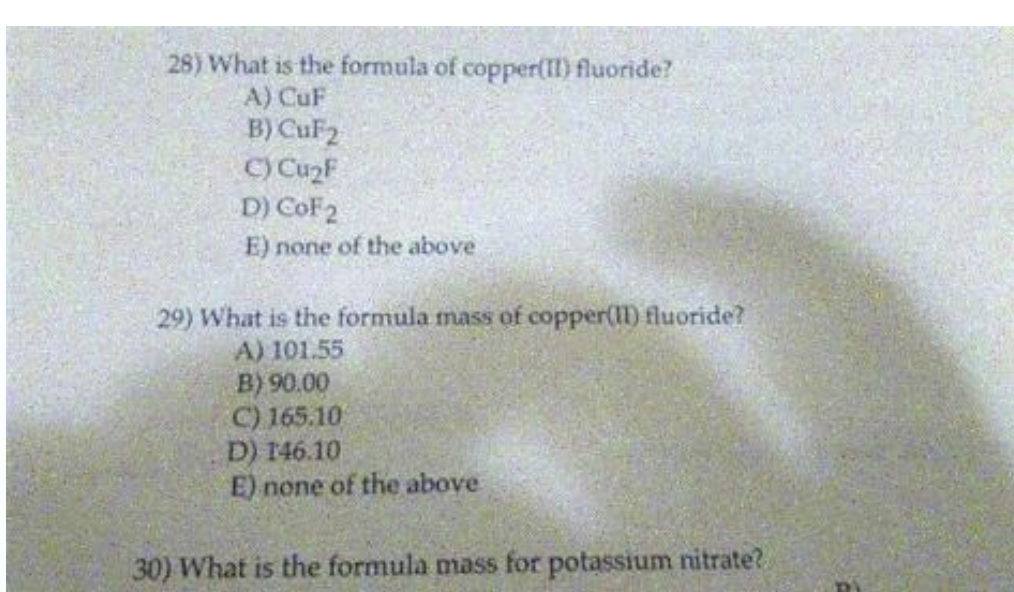


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C [77 °F], 100 kPa). N Mark (What is Yn?) Information box Chemical compound Copper (ii) Fluoride is an inorganic compound with the chemical formula CUF2. The anhydrous form is white, ion, crystalline, hygroscopicEmpirical formula of copper of b/loride (hill designation): CUF2 This article contains general connections, but there are no quotes built in it corresponding. Help us improve this article by adding more accurate quotes. (June 2021) (Find out how and when to delete this model message) Copper fluoride (II) is an anhydrous element of a ballistic model of a crystalline packaging form of dihydrate dihydrate of dihydrate. copper fluoride; Copper Diploma identifiers (2+) CAS number 7789-19-7 Y13454-88-1 (dihydracted) 3D 3D model (JSMOL) Interactive display of Chempidder 74214 Yes, information map ECHA 100.029.225 EC number 232-19- 7 UNIML1257-1407 C1 Y3A38PC42E9 (DIYDRATE) Y Control Panel Comptox (EPA) DTXSID80894782 INCHI INCHE = 1S/CU.2FH/H; 2*1h/q+2;/P-2 YKEY: GWFAVIVIGOHFAOHFAOYSA = F1/F1/F1/F1/F1/F1/F1/F1/F1/F1 UH. H; 2*1h/q+2 ;/P-2knopka: Gwfaviimqduca-Nuqvwonbaf smiles [Cu+2]. [F-] / Mol (dihydrate) white crystalline dust after hydration: blue density 4.23 g/cm3 (cm3 (without water) 2,934 g/cm3 (dihydrate) [1] Fusion temperature 836 \xc2 \xb0c (1537 \xc2 \xb0f; 1109 a) (without water))) 130 \XC2 \XB0C (dihydrate, decomposed) boiling point 1676 \xc2 \XB0C (3049 \XC2 \XB0F; 1949 K) (Anhydro) magnetic susceptibility (\XCF \X87) \ 7 \ 087 \x87 \XCX20. x926 cm3/dangers of Niosh (the limits of the health of the United States): peel (admissible) twissable 1 mg/m3 (in the form of Cu) [2] Rel (recommended) twin 1 mg/m3 (in the form of Cu?) Compounds of copper fluoride (i) unless otherwise indicated, the data are indicated for typical materials (at 25 \xc2 \xb0c [77 \xc2 \xb0f], 100 kPa).

Exercise 3.69

Copper (II) fluoride contains 37.42% F by mass.

H; 2*1h/q+2 ;/P-2knopka: Gwfaviimqduca-Nuqvwonbaf smiles [Cu+2]. [F-] / Mol (dihydrate) white crystalline dust after hydration: blue density 4.23 g/cm3 (cm3 (without water) 2,934 g/cm3 (dihydrate) [1] Fusion temperature 836 \xc2 \xb0c (1537 \xc2 \xb0f; 1109 a) (without water))) 130 \XC2 \XB0C (dihydrate, decomposed) boiling point 1676 \xc2 \XB0C (3049 \XC2 \XB0F; 1949 K) (Anhydro) magnetic susceptibility (\XCF \X87) \ 7 \ 087 \x87 \XCX20. x926 cm3/dangers of Niosh (the limits of the health of the United States): peel (admissible) twissable 1 mg/m3 (in the form of Cu) [2] Rel (recommended) twin 1 mg/m3 (in the form of Cu?) Compounds of copper fluoride (i) unless otherwise indicated, the data are indicated for typical materials (at 25 \xc2 \xb0c [77 \xc2 \xb0f], 100 kPa).

Mark N (what is YN?) Connect the Block Chemical Compositis information Fteride of copper (II) - an inorganic compound with the chemical formula C.Jahn D9 D9(II)[6] and leads to a distorted rutile structure similar to the structure of chromium fluorine, CRF2, which is a D4 compound. [3] Copper fluoride coordination [3] [4] Copper fluoride consumption coordination Copper fluoride consumption can be used to effect the reaction of fluorinated aromatic hydrocarbons with aromatic hydrocarbons in an oxygen-containing atmosphere at temperatures above 450 °C (842 °F). This reaction is simpler than the Sandmeyer reaction, but is only effective in producing compounds that can survive the temperature used. A coupled reaction using oxygen and 2 HF regenerates cupric fluorine(II) to form water. [7] This method was designed as a "green" method for producing fluoromatic substances because it avoids the production of toxic waste such as ammonium fluoride. The chemistry of cupric fluorine(II) can be summarized as copper and fluorine at 400 °C (752 °F). It occurs as a direct reaction. Cu + F2 à CUF2 at temperatures above 950 ° C (1,742 ° F) loses fluorine in the molten state. 2CUF2 à 2CUF + F2 2CUF à CUF2 + CU CUF3à, CUF42' and CUF64 complex are formed when CUF2 is exposed to substances containing fluoride ions f. F. and Cu (OH) ions. [QUOTE Required]

Toxicity There is little detailed information on the toxicity of copper(II) fluoride. Copper and fluoride, however, can be individually toxic to consume. Copper toxicity can affect the skin, eyes and respiratory tract. Serious conditions include metal fume fever and red blood cell embolism. Copper can also damage the liver and other important organs. Metal fluorides are generally safe at low levels, and in many countries they are added to water to protect against tooth decay. At higher levels, they can cause toxic effects from nausea and vomiting to shaking and breathingNiosh Pocket points to chemical dangers. "#0150", National Institute of Occupational Safety and Health (NIOSH). ^ A b c d Grivnads, Normans n.; Ermshev, Alan (1997). Element chemistry (edition 2). Batervort-Heineman. Mr. 1184 1185. ISBN 978-08-037941-8. ^ A b fisher, p.; Chalks, v.; Schwarcenbach, d.; Gamsj Betts, H. (1974). "Copper (II) fluoride magnetic and crystalline structure". J. Physics Chemistry. A strong, 35 (12): 1683 1689. Doi: 10.1016/S0022-3697 (74) 80182-4. ^ C. Billie, H. M. Hendler (1957). "Copper (II) fluoride crystalline structure". Journal of the American Chemistry Society. 79 (5): 1049. 51. Doi: 10.1021/ja01562A011. ^ Greenwood, Normans n.; Ermshev, Alan (1997). Element chemistry (edition 2). Batervort-Heineman. 1190-1191 p. ISBN 978-08-037941-8. ^ M.a. Subramanians; L. E. Manzers (2002).

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