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DICTIONARY

OF THE

ACTIVE PRINCIPLES OF PLANTS:

ALKALOIDS; BITTER PRINCIPLES; GLUCOSIDES:

THEIR SOURCES, NATURE, AND CHEMICAL CHARACTERISTICS, WITH TABULAR SUMMARY, CLASSIFICATION OF REACTIONS, AND FULL BOTANICAL AND GENERAL INDEXES.

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PREFACE.

THE ceaseless flow of new facts into every branch of science is well illustrated in the particular section of organic chemistry embraced under the heading *Alkaloids, Bitter Principles and Glucosides.* The streamlet that started at the commencement of this century, when Derosne and Sertürner made the discovery of Morphine, has swollen with ever-increasing rapidity, until at the present time the flood of matter accumulated almost defies management.

It becomes every day more desirable that something be done towards classifying the substances so constantly being augmented in number, and that the details bearing upon them, now more or less scattered throughout chemical literature, should be tabulated in a convenient form—one that will permit not only that *a given attribute of any substance shall be readily found*, but also will indicate *wherein such substance differs from, or resembles, another of its class.*

LONDON, *October*, 1893.

The present work, which treats of nearly 600 of these bodies, has been prepared with the hope that it may contribute to the attainment of the above objects, and be found of service to those who have occasion to deal with these compounds or study them—particularly the *analyst, professor, research chemist, student, and manufacturer.*

No efforts have been spared to render as full as possible the information supplied (particularly as regards tests and chemical reactions); wherever practicable, it has been drawn from the original sources, and embodies the results of the latest researches.

Easy access to any fact contained in the book has been aimed at, and it is earnestly requested that the reader will peruse the introductory and explanatory part, that he may the better acquire this facility.

INTRODUCTORY AND EXPLANATORY.

In addition to the Active Principles that are members of one or other of the three classes specially dealt with, I have included some substances that are liable to be met with in connection with these, or have bearing on the subjects under treatment.

The term *Alkaloid* (or the letter A used for that word) has been employed to denote any naturally occurring vegetable base; this covers a wider area than would the most modern application of the word, which restricts it to Pyridine and Quinoline derivatives, to the exclusion of such a body as Caffeine. Of course, Salts of Ammonia and albuminoid decomposition products cannot be comprised in this class (one or two of the latter receive notice, however).

The *Glucosides* embrace those substances (other than tannins) that yield sugar (with some other compound) when decomposed by the action of dilute acids or natural ferments. They are generally free from Nitrogen, exceptions being Amygdalin, Solanine, Indican, Linamarin, Myronic Acid, and Sinalbin, the last two of which contain Sulphur also.

The *Bitter Principles* are a very diverse group, as various in composition as in their chemical reactions. They contain no Nitrogen, and yield no sugar on treatment with dilute acids or otherwise. The name *Amaroid* is suggested for such bodies as these, in order to avoid confusion with glucosides and alkaloids, so many of which are also bitter. Being very numerous and frequently of considerable importance, whilst generally neglected in works upon organic chemistry, it has been thought desirable to describe them fully.

In conformity with the system adopted by the Chemical Society in their journal, all alkaloids are spelt with the termination *ine*, whilst glucosides and amaroids end in *in*, without the final *e*.

The book consists of three parts (besides two indexes and sub-indexes).

PART I. in Dictionary form, arranged upon a system which groups together the constituents either of one plant or of a number of botanically or chemically allied plants, whilst retaining as far as practicable an alphabetical order. Supplementary indexes being provided, the manifest inconveniences which a purely alphabetical method would occasion are by these means avoided.

☞ In order that no time may be lost in finding any statement concerning a given substance, the following rotation has been adhered to in the enumeration of the properties and tests:

Botanical Details.—Plants furnishing the compound or compounds, with technical as well as popular names and botanical order.

Parts of plant in which found.

Bibliographical reference (names of investigators and details of published researches).

Each substance (when more than one) is then treated individually thus:

General, Physical and Chemical Details.—Name and synonyms, class (A = Alkaloid, B = Bitter Principle, G = Glucoside).

Chemical formula.

Crystalline form, etc.

Melting point and effects of heat generally.

Rotatory power (action on polarized light).

Reaction (alkaline, neutral, or acid).

Taste and odour (if any).

Various: Character of salts, products of hydrolysis (action of boiling dilute acids), physiological effect.

Solubility in water, alcohol, ether, chloroform, benzene, petroleum ether, amyl alcohol, carbon bisulphide, and sundry other solvents.

Behaviour towards immiscible solvents.

Reactions with precipitants, etc. (in this order):

1. Alkaline hydrates and carbonates.
2. Ammonia.
3. Lead acetate, neutral and basic.
4. Tannic acid.
5. Picric acid, $C_6H_2(NO_2)_3OH$ (Hager's reagent).
6. Ferric chloride.

7. Platinum chloride.
8. Palladious chloride.
9. Gold chloride.
10. Silver nitrate.
11. Copper sulphate.
12. Fehling's solution (alkaline copper tartrate).

Cyanides, etc.:

13. Potassium ferrocyanide.
14. ,, ferricyanide.
15. ,, sulphocyanide.
16. ,, cyanide.
17. ,, nitro-prusside.
18. Silver potassium cyanide.

Chromates, etc.:

19. Potassium chromate.
20. ,, bichromate.
21. Chromic acid.

Phospho-compounds:

22. Phospho-molybdic acid (Sonnenschein's reagent).
23. Phospho-tungstic acid (Scheibler's reagent).
24. Phospho-antimonic acid.

Iodides, etc.:

25. Iodo-potassic iodide (Wagner's reagent).
26. Bismuth-potassic iodide (Dragendorff's reagent).
27. Cadmium-potassic iodide (Marmé's reagent).
28. Zinc-potassic iodide.
29. Mercuric-potassic iodide (Mayer's solution).
30. Mercuric chloride.
31. Chlorine water.
32. Bromine water.
33. Iodine tincture.

Colour tests (in this order), time taken for development of colours and their durability:

Concentrated sulphuric acid alone and with various substances (sugar, potassium bichromate, manganese dioxide, nitric acid, etc.), Erdmann's reagent, sulphuric acid with a trace of nitric acid.

Concentrated nitric acid alone.

Acids in general.

Perchlorate of potash or perchloric acid.

Per-iodic acid.

Fröhde's solution (molybdic and sulphuric acids).

PART II.—A TABULAR SUMMARY designed for ready reference and as a means of contrasting one compound with another for analytical and other purposes; it gives the chief properties and tests of the substances that are more fully described in Part I.

PART III.—A CLASSIFICATION OF REACTIONS for the special use of analysts, showing what compounds are known to respond to a given test.

BOTANICAL INDEX.—Each plant here appears in its alphabetical positions by popular and scientific names. Sub-lists of plants are also inserted for those botanical orders which include members containing alkaloidal, bitter, or glucosidal principles. This facilitates comparison of the constituents of allied plants.

GENERAL INDEX.—Special endeavours have been made to render this comprehensive.

In the TABULAR SUMMARY the order of insertion in Part I. has been preserved, because it permits a survey of all the active principles of a given plant at a glance—a manifest advantage over a *purely* alphabetical arrangement (which might upon a preliminary inspection have appeared desirable), for in that case the reader, dealing, for instance, with Opium, would have had nearly forty different positions to find, whilst here he has but one.

Synonyms and questions of identity cannot, for want of space, be discussed in Part II. (the Tabular Summary), but receive full attention in Part I.; for the same reason, many columns of precipitants have had to be removed from Part II.; but this curtailment is amply compensated for by the classified lists of reactions in Part III. and the numerous reference notes.

In the last section—Part III.—are classed, in a series of alphabetical lists, substances having a common property, or giving a similar reaction.

GENERAL RULES.

☞ 1. To facilitate reference, it is important to remember that the series of details upon any substance is invariably given in a particular rotation (see above).

2. COLOUR.—Since the great majority of substances here dealt with are colourless, colour is only mentioned when coloured compounds are spoken of; *all others are to be understood to be white or colourless.*

3. ODOUR.—For analogous reasons, *all bodies not described as having odour are odourless.*

4. ALKALINITY.—When not otherwise stated, the reaction (alkaline, neutral, or acid) has reference to litmus as indicator.

5. TEMPERATURE is always indicated in degrees Centigrade.

6. SOLUBILITY.—Figures in the solubility columns denote number of parts of solvent required to dissolve one part of substance.

7. DELICACY OF TEST.—Numbers in the columns of precipitants (or colour tests) signify degree of dilution at which the reaction is observable.

8. USE OF REAGENTS.—Except where other indications are given, *precipitants are to be applied to the substance dissolved in water*, and in the case of alkaloids *to the salt*, and not to the free base.

9. COLOUR TESTS.—The dry substance is to be employed when testing with concentrated acids and colour tests generally.

ABBREVIATIONS have as far as possible been avoided, but in the Tabular Summary limited space has necessitated the adoption of a few, thus :

A (after name of substance)= Alkaloid.
acet.= Acetic acid.
ac. eth.= Acetic ether.
alcl.= Alcohol.
alk.= Alkali.
alkd.= Alkaloid.
AmHo=Ammonia hydrate.
amorphs.= Amorphons.
B (after name of substance)= Bitter principle, *i.e.*, amar'oid.
bz. or benz.=benzene (coal-tar benzene, C_6H_6, or benzol).
bc. (in lead acetate column) = Basic.
bn.= Brown.
bl.= Blue.
CS_2=carbon bisulphide.
Cfm. or $CHCl_3$= Chloroform.
crys.=Crystals or crystalline.
df. or diff.= Difficultly.
eth.=Ether.
G (after name of substance)= Glucoside.
G.-d. (after name of substance)= Glucoside derivative.
gn.= Green.
insol.= Insoluble.
M.P.= Melting point.
ndls.= Needles.
nl.= Neutral.
p. (in colour test column, etc.)= Purple.
pp.= Precipitate.
petr. e.= Petroleum ether.
phos.-molyb. = Phospho-molybdic acid.
r. or rd.= Red.
sol.= Soluble.
soln.= Solution.
turp.= Turpentine.
v. or vlt.= Violet.
volat.= Volatile.
w. or wh.= White.
y., yl., or yel.= Yellow.
∞ =Miscible in all proportions.
$\sim$ =Changing or changes to.

DICTIONARY OF THE ACTIVE PRINCIPLES OF PLANTS.

PART I.

§ 1. **ACHILLEA** millefolium (Milfoil or Yarrow); *Compositæ;* Europe and N. America—substance (*a*). A. moschatus ('Iva'), (c) and (*d*). Obtained from the whole plant.

(*a*) *ACHILLEIN* G., $C_2H_{39}N_2O_{15}$ (von Planta); amorphous, reddish-brown, bitter, alkaline.

With boiling dilute acids is converted into sugar and Achilletin (see below).

Soluble in water easily (giving yellow solution), in alcohol with difficulty. **Insoluble** in ether.

Not precipitated by

Alkalies.	Tannic acid.
Lead acetate, neutral or basic.	Ferrous sulphate.

(*b*) *ACHILLETIN* (from above), $C_{11}H_{17}NO_4$, amorphous, dark brown powder, not bitter.

Insoluble in water, and with difficulty in alcohol.

(*c*) *IVAIN* B, $C_8H_{14}O$ or $C_{24}H_{42}O_3$ (von Planta); yellow, amorphous, soft resinous ('Terebinthinate'), bitter.

Soluble in alcohol (yellow solution), not in water.

Not precipitated by neutral lead acetate.

(*d*) *MOSCHATINE* A, $C_{21}H_{27}NO_7$ (von Planta); amorphous, reddish-brown, bitter; melts under water (on water bath).

Soluble with difficulty in alcohol, scarcely in water.

§ 2. **ACHRAS** sapota ('Sapodella Plum'); *Sapotaceæ.* The kernels. Investigators: Michaud, who describes (*a*); and Bernon (c). From the Argan tree (*Sapotaceæ*) of Morocco, S. Cotton (*J. Pharm.* [5] 18, 298) has separated 'ARGANIN,' which would seem to be identical with (*a*).

(*a*) *SAPOTIN* G. (Arganin? see above), $C_{29}H_{52}O_{20}$ (Michaud); microscopic crystals; burning taste; lævo-rotatory ([*a*] D = −32·1 in alcoholic solution); M.P. 240° with decomposition. Dilute sulphuric acid, on boiling, yields Saporetin (see below) and sugar.

Soluble in water easily; in cold alcohol sparingly, easier hot. **Insoluble** in benzene, ether, chloroform.

Precipitants, etc.:

Alkaline hydrates, insol.	Fehling solution, *not* reduced.
Lead acetate basic, pp. sol. in excess.	Concentd. sulph. acid, garnet red.

(*b*) *SAPORETIN*, $C_{17}H_{32}O_{10}$ (from above).

Soluble in alcohol and chloroform. **Insoluble** in water and ether.

(*c*) *SAPOTINE* A. (of Bernon).

Soluble in alcohol, ether, chloroform. **Insoluble** in water, alkalies. Hydrochloride, bitter.

Precipitants:

Mercuric-potassic iodide, brown.
Mercuric chloride, white.
Platinum chloride, yellow.

§ 3. **ACONITUM** Napellus (Monk's Hood, or Wolf's Bane); the alkaloids (*a*), (*b*), (*n*), with traces or, at times, none of (*c*) and (*e*); A. variegatum (*a*); A. Stoerkianum (*a*); A. paniculatum, an alkaloid regarded as

(*a*) by Hübschmann, the investigator of last-mentioned three varieties, but now ascertained to be non-poisonous; A. anthora (c?); A. lycoctonum (*b*?), (*d*?), (*f*) to (*j*), not (*a*). The descriptions of the alkaloids (*f*), (*g*), and (*j*) are due to Dragendorff and Spohn (*Pharm. J., Trans.*, 1884); and alkaloids (*h*), (*i*), to Hübschmann. A. ferox (Indian or Nepaul aconite; Himalaya root), (*c*), trace or no (*a*); A. heterophyllum (Atis root) (*k*); A. sinense ('Shiraka-wauzuware') (*a*). A Japanese plant, 'Kusauzu' (*a*); *Ranunculaceæ*. The plant or root. Investigators very numerous; in addition to chemists already referred to, mention should be made of Wright and Luff, Dunstan and Ince, Jürgens, Duquesnel, Groves, Paul, and Kingzett, etc.

(*a*) *ACONITINE* A. [Napelline; not the Napelline of Dunstan and Ince now being investigated, see (*n*); Japaconitine (Lubbe), see (*j*)]. $C_{33}H_{43}NO_{12}$ or $C_{26}H_{35}NO_7(OH)_3O.(CO.C_6H_5)$, Wright and Luff. (Dunstan and Ince prefer $C_{33}H_{45}NO_{12}$). Crystallizes in rhombic prisms (or amorphous if not pure), scarcely bitter unless impure, causes tingling of the mouth, extremely poisonous ($\frac{1}{10}$-grain may be fatal), very readily decomposed, with formation of Benzoic Acid and Aconine—this less liable to occur when an organic acid is used for extraction; hence advantage of Duquesnel's method, in which tartaric acid is employed. M.P. 188·5° (Dunstan and Ince; other observers have found 60°, 80°, 'above 100°,' 140°—discrepancies due to the difficulty of obtaining a pure specimen). Dextro-rotatory; salts, lævo-rotatory; alkaline reaction.

Solubility: In cold water 1 in 4,431 at 22° C. (Dunstan), or 1 in 726 (Jürgens), easier warm; 1 in 20 boiling alcohol, 1 in 230 chloroform, 1 in 100 boiling ether, and in benzene. **Insoluble** in petroleum ether or carbon bisulphide.

Precipitants ('Reactions not reliable'—Hüsemann):

Alkaline hydrates.	Platinum chloride (Duquesnel); not if dilute.
„ carbonates.	Gold chloride, yellow amorphous (crys. from alcohol, M.P. 135°).
[*Not* bicarbonates unless hot (von Planta).]	[Not by ferro- or ferri-cyanides of potassium.]
Ammonia hydrate, soluble in excess.	Potassium bichromate; not at once, but gradually 1 in 3,000.
Tannic acid.	[Not by Argentic-potassic cyanide.]
Picric acid (Duquesnel); not if dilute.	
Ferric chloride, yellow.	
Phospho-molybdic acid, light yellow flocculent.	Mercuric-potassic bromide, amorphous pp.
Phospho-tungstic acid.	Mercuric-potassic iodide, white amorphous.
Iodo-potassic iodide, yellow-red.	Hydriodic acid, microscopic crystals.
Bismuth-potassic iodide, orange—limit 1 in 40,000.	Iodine tincture, reddish.
Cadmium-potas. iodide; white, 1 in 1,000; cloud at 1 in 2,500.	

The two best tests.—Lubbe.

Colour tests: Mem.—When perfectly pure no colours are given.

Concentrated sulphuric acid, gradual violet.
„ „ „ with sugar, red.
„ „ „ with nitric acid, gradual violet.
Nitric acid, reddish-brown.
Concentrated hydrochloric acid, colourless.
Fröhde's reagent, yellowish-brown.

(*b*) *ACONINE* A. [=Acolyctine (Wright and Luff), this identity questioned by Dragendorff, see (*i*)], $C_{26}H_{39}NO_{11}=C_{26}H_{35}NO_7(OH)_4$ or $C_{26}H_{41}NO_{11}$. Occurs naturally, and formed artificially, together with benzoic acid, by saponification of aconitine by alcoholic potash; amorphous powder forming amorphous salts; M.P. about 130°, alkaline, bitter, does not produce tingling of the tongue. Physiological action feeble.

Soluble in alcohol, chloroform, and water. **Insoluble**, or difficultly soluble, in ether, benzene, petroleum ether.

Precipitants, etc.:

Ammonia; gelatinous.	Gold chloride, with reduction of gold.
[No pp. carbonates or bicarbonates.]	Silver nitrate, reduced.
Lead acetate, soluble in excess.	Fehling's solution, reduced.
Tannic acid, pp.	Phospho-molybdic acid, bluish-gray

No colour with concentrated sulphuric acid.

(*c*) *PSEUDACONITINE* A., $C_{36}H_{49}NO_{11}$, or $C_{27}H_{37}NO_5(OH)_3.O.CO.C_6H_3(OCH_3)_2$;

may be obtained in crystalline needles, but usually separated as varnish; M.P. 104°-105°; very poisonous, resembles aconitine; salts difficultly crystallizable: yields on saponification (by merely heating) pseudaconine and veratric acid (dimethyl-proto-catechuic acid).

Soluble in alcohol, in ether with difficulty—though crystals are obtainable therefrom—scarcely in water.

Precipitants:

Tannic acid.	Mercuric-potassic iodide.
Gold chloride.	Mercuric chloride.

[Platinum chloride, only if concentrated.]

Colour test: Nitric acid fuming (a few drops), yields on evaporation a residue which gives a purple red with a drop of alcoholic potash.

(*d*) *PSEUDACONINE* A. [Lycoctonine, Wright and Luff; identity questioned by Dragendorff, see (*h*)], $C_{27}H_{41}NO_8$. Formed from Pseudaconitine by saponification with alcoholic potash; veratric acid ($C_9H_{10}O_4$ or $C_6H_3(OCH_3)_2COOH$ = Dimethyl-proto-catechuic acid) being formed at same time. Amorphous, but changing gradually under certain conditions to a semi-crystalline mass; bitter (without causing tingling); alkaline; feebly poisonous; salts amorphous.

Soluble in water, alcohol, ether.

Precipitate with silver nitrate.

Ammoniacal silver nitrate, reduced.
[Fehling's solution, not reduced.]

(*e*) *PICRACONITINE* A., $C_{31}H_{45}NO_{10}$ (Wright). [Obtained by Groves from a root sold commercially as Aconitum Napellus, but found only in traces, or not at all in the genuine A. Napellus.] Amorphous, resinous, bitter; feebly, if at all, poisonous; salts crystalline.

Soluble in ether and chloroform. **Insoluble** in water.

Precipitants:

Alkaline hydrates, „ carbonates, Ammonia } a resinous pp. on heating.
Tannic acid.
[Not Platinum chloride.]
Gold chloride, amorphous, pale yellow.
Mercuric-potassic iodide.

No colour reactions with sulphuric acid or other principal colour tests.

(*f*) *LYCACONITINE* A., $C_{27}H_{34}N_2O_6 + 2H_2O$ (Dragendorff and Spohn). Pale yellow, amorphous. Anhydrous at 110° C.; M.P. 114·8° with decomposition; lævo-rotatory: $[a]_D = +31{\cdot}5$ (10 per cent. solution in alcohol); salts amorphous. Readily changed by alkalies, with production of (*g*), lycoctonic acid, and a third substance resembling (*i*).

Soluble very readily in alcohol, chloroform, benzene and carbon bisulphide; less easily in ether, with difficulty in water, and scarcely in petroleum ether.

Removed by ether or chloroform, from a solution previously treated with sodium bicarbonate.

Reactions:

Alkaline hydrates, „ bicarbonates, Ammonia } pp.	Bromine water, yellow amorphous pp. (a tribromide).

Colour tests: Concentrated sulphuric acid, reddish-brown.

Sulpho-selenic acid (8 cc. water, 6 cc. sulphuric acid, and 0·3 gm. sodium selenate), rose or pale red violet coloration. This reaction not given by aconitine, Hübschmann's nepaline, or commercial lycoctonine.

Concentrated phosphoric acid, violet on warming.

(*g*) *LYCACONINE* A. [Lycoctonine? compare (*h*)], $C_{27}H_{47}N_2O_7 + 1\frac{1}{2}H_2O$; crystalline; M.P. 90°-92°, alkaline, fluorescent (blue); dextro-rotatory ($[a]_D = +46{\cdot}4$).

Soluble in 4 parts absolute alcohol, 3 chloroform, 58·4 ether, 64·5 benzene, 247 water.

(*h*) *LYCOCTONINE* A. [Pseudaconine? see (*d*) above; Lycaconine? (*g*)]. Crystallizes in needles and prisms; M.P. about 100°; alkaline reaction, very bitter.

Soluble in water, alcohol, ether, chloroform, petroleum ether, carbon bisulphide.

Precipitants:

Tannic acid.	20,000 after twenty-four hours, 1 in 8,000 after fifteen minutes. *Pp. becomes crystalline* (compare Aconitine).
[Platinic chloride, no pp.]	
[Phospho-molybdic acid, no pp.]	
Iodo-potassic iodide.	
[Potassic iodide, no pp.]	[Mercuric-potassic bromide, no pp., compare Aconitine.]
Bismuth-potassic iodide—1 in 40,000.	[Mercuric chloride, no pp.]
Cadmium-potassic iodide.	Bromine water, pp. crys.
Mercuric-potassic iodide—1 in	

Colour tests:

Concentrated sulphuric acid, Nitric acid, Phosphoric acid } colourless.

(*i*) *ACOLYCTINE* A. (Aconine? see (*b*) above). Bitter, alkaline powder. **Soluble** readily in water, alcohol, and chloroform; *not* in ether.

Precipitants:

Alkaline carbonates.	Tannic acid.
Ammonia, gradually gelatinous.	Gold chloride, yellow.
Lead acetate.	Phospho-molybdic acid, white.

(*j*) *MYOCTONINE*, A. $C_{27}H_{30}N_2O_8{\cdot}5H_2O$ (Dragendorff and Spohn); amorphons, bitter; dextro-rotatory; M.P. 143°-144°. **Salts** amorphous; very poisonous.

Soluble in water with difficulty, easily in alcohol, amyl alcohol, chloroform, benzene, carbon bisulphide. **Insoluble** (or scarcely sol.) in ether or petroleum ether. Heated with alkali, myoctonine gives lycoctonine, lycoctonic acid, an alkaloid resembling acolyctine, and a fourth substance not further examined.

Precipitated by most alkaloid reagents.

Vitali's test (evaporation with nitric acid, then touching with alcoholic potash) gives reddish-brown colour.

(*k*) *ATISINE* A. [from A. heterophyllum (Atis root)], $C_{22}H_{31}NO_2$ (Alder Wright), or $C_{46}H_{74}N_2O_5$ (Broughton); amorphous, oxidizable on exposure (becoming coloured and resinous); M.P. 85°; bitter, non-poisonous. **Haloid** salts are crystallizable and difficultly soluble; nitrate, sulphate, acetate, amorphous.

Soluble in alcohol, ether, and benzene; scarcely in water.

Precipitated by:

Ammonia, white flocculent.	Mercuric-potassic iodide, white.
Tannic acid, yellow to brown.	

Colour reactions:

Concentrated sulphuric acid, yellow~magnificent purple, lasting several days; momentary violet on addition of a drop of water (**Shimoyama**); or faint violet~red~brown (Wasowicz).

Concentrated sulphuric acid and sugar, yellow~reddish~carmine.

Nitric acid, colourless.

Hydrochloric acid, colourless.

Phosphoric acid, colourless, but yellowish-violet on warming.

(*l*) *JAPACONITINE* A. (from Japanese Aconite root), $C_{66}H_{88}N_2O_{21}$ (Wright and Luff), or $C_{33}H_{44}NO_{12}$=Aconitine (Lubbe). The former regard it as a sesqui-anhydro base formed by condensation of $2(C_{33}H_{47}NO_{12})$ with elimination of 3 molecules of water. It bears the closest possible resemblance to aconitine, and yields on saponification benzoic acid (as does the latter), and *JAPACONINE*, $C_{26}H_{41}NO_{10}$, almost indistinguishable from aconine, the analogous derivative of aconitine.

[Lubbe believes that Japaconitine and Japaconine are identical respectively with Aconitine and Aconine.]

(*m*) *APO-* or *ANHYDRO-ALKALOIDS* of the Aconite group are formed from the respective bases by the elimination of a molecule of water; this may be effected by heating to 100° for some hours in a concentrated tartaric acid solution thus:

Aconitine, $C_{26}H_{35}NO_7(OH)_3(CO{\cdot}O{\cdot}C_6H_5) - H_2O = C_{26}H_{35}NO_7{\cdot}O{\cdot}OH(CO{\cdot}O{\cdot}C_6H_5)$ Anhydro or Apo-aconitine.

Aconine, $C_{26}H_{35}NO_7(OH)_4 - H_2O = C_{26}H_{35}NO_7{\cdot}O(OH)_2$ Apo-aconine.

Pseudaconitine and Pseudaconine yield parallel compounds, but Japaconitine being already an anhydro-base (see (*j*) above) does not undergo change during the above treatment.

There is great similarity in the physiological action of these bases and the respective parent alkaloids.

(*n*) *ISACONITINE* A. (of Dunstan and Harrison, see Chem. Soc. Proc., 119, 1893), $C_{33}H_{45}NO_{12}$, colourless, varnish-like, intensely bitter, without the tingling sensation characteristic of Aconitine, much less poisonous than the latter; lævo-rotatory [a] D = −28·74°.

Soluble in alcohol and chloroform readily, less easily in ether (by which means it may be separated from Aconitine), slightly in water.

Forms a substituted compound with gold chloride = $C_{33}H_{44}(AuCl_2)NO_{12}$ analogous to the Caffeine compound described by Dunstan and Shepheard (Chem. Soc. Trans., Feb., 1893).

§ 4. **ACORUS CALAMUS** (**S**weet Flag); *Acoraceæ-aroideæ*; Europe Asia, N. America. Used as tonic and stimulant. The root; (22 grammes were obtained from 12 kilos). Investigators, H. Thoms and others.

ACORIN G., $C_{36}H_{60}O_6$ (Thoms), amorphous, soft resinous, neutral; bitter and aromatic taste. Yields sugar and calamus oil (in a current of hydrogen, otherwise resinification occurs). **S**ee also Acoretin, below.

Soluble in ether, chloroform, benzene, methyl alcohol, acetone; with difficulty in alcohol. **Insoluble** in water. Hydrochloride difficultly soluble in water.

Precipitants:

Tannic acid.	Phospho-molybdic acid, pp.~blue from reduction.
[Platinum chloride, reduced.]	Mercuric-potassic iodide.
[Gold Chloride, reduced.]	Iodine tincture.
Fehling's solution, reduced.	

CALAMINE A., strongly alkaline.

Soluble in alcohol, chloroform, acetone, and dilute acids; not in water or ether.

Precipitants:

Tannic acid.	Phospho-molybdic acid.
[Platinum chloride, reduced.]	Iodo-potassic iodide.
[Fehling's solution, *not* reduced.]	Mercuric-potassic iodide.

§ 5. **ADANSONIA** digitata, A. Gregorii (Baobab tree); *Malvaceæ;* Africa, India, etc. The bark. Investigators, Walz, *Jahrb. f. Pharm.* 24; Wittstein, *Viertelj. schr. f. Pharm.* 4.41.

ADANSONINE A. (?), crystallises in needles, odour like aloes or gentian; bitter.

Soluble in 6 parts cold or 3 parts hot ether, and in alcohol; slightly in water. Forms crystalline compounds with acids (Dupuy).

Reactions:

Alkaline hydrates, Ammonia } Sol. yellow.

No precipitates with metallic salts (Wittstein).

§ 6. **ADONIS** amurensis; *Ranunculaceæ.* From the root, Tahara has isolated Adonin; from A. vernalis, V. Cervello has separated a substance named Adonidin, which gives the same reactions, and is presumably identical with the other glucoside (Y. Inoko).

ADONIN G. (Adonidin ?), $C_{24}H_{40}O_9$, neutral, intensely bitter; converted by boiling dilute acids into glucose, and a resinous matter soluble in ether.

Soluble in water, *cloud on warming and partial separation*, also in alcohol, chloroform, and acetic acid. **Insoluble** in ether.

Precipitants:

[Dilute alkalies, do not decompose.]	Gold chloride.
Gallic acid.	Mercuric chloride.
Picric acid.	

Colour tests: Concentrated sulphuric acid, deep red; nitric acid, indigo blue; hydrochloric acid, rose red.

ADONIDIN (Adonin ?-see above) is stated to have an action resembling that of digitalin, but weaker.

Reactions, see above.

§ 7. **ÆSCULUS** hippocastaneum (Horse Chestnut); *Sapindaceæ;* all the substances below; Æsculetin (*b*) has been found in the free state in the seeds of Euphorbia lathyris (Caper Spurge, or Semen Cataputiæ minoris); Æsculin also in Hymenodictyon excelsum (which see), and in Gelsemium nitidum (Jasmine, see Strychnos group); for Fraxin (existing in horse chestnut), see Fraxinus.

(*a*) *ÆSCULIN* G. (Æsculinic acid, Bicolorin, 'Schillerstoff,' Polychrome), $C_{15}H_{16}O_9 \cdot 1\frac{1}{2}H_2O$ (Liebermann), crys. needles and prisms. M.P. 160°; fluorescent (visible 1 in 1,500,000); acid reaction, slightly bitter; by action of heat or dilute acid æsculetin and sugar are formed.

Soluble in 672 parts water at 10° or 12½ boiling, in 90 of cold alcohol (Sp. G. 0·798) or 24 boiling; in chloroform (which removes it from acid solution), but not in absolute ether, and scarcely in ordinary ether.

Reactions: Alkaline hydrates dissolve more readily than water (solutions fluorescent). Precipitated by basic lead acetate, but not by other metallic salts.

Colour tests: Dilute nitric acid, on shaking, gives a yellow solution, becoming blood-red with ammonia.

Ammonium bisulphite, and then ammonia, give a blood-red, becoming blue on shaking.

(*b*) *ÆSCULETIN* (from above, and in free state; see statement concerning Euphorbia lathyris above), $C_9H_6O_4 + H_2O$, crys. silky needles and plates like benzoic acid. M.P. 270°, part volatilizing, neutral reaction, bitter, fluorescent—feebly blue.

Soluble with difficulty in cold water or cold alcohol, more easily in either when hot. **Insoluble** in ether, chloroform, glacial acetic acid, or carbon bisulphide.

Reactions:

Alkalies dissolve to yellow solution (acids reprecipitate).

Lead acetate, neutral or basic, gives pp.

Copper sulphate, Silver nitrate } reduced.

(*c*) *ÆSCULETIN-HYDRATE*, $4C_9H_6O_4 + H_2O$; crystalline; M.P. 250°, part subliming; isomeric with Daphnetin. Less soluble than Æsculetin. Alkalies and lead acetate, as Æsculin.

(*d*) *ARGYRÆSCIN* G., $C_{27}H_{42}O_{12}$; silvery crystals.

Soluble in alcohol, with difficulty in water (solution frothy). **Insoluble** in ether.

Potassic hydrate (strong solution) converts gradually upon heating to propæscinic acid, which is further changed (see below).

Precipitated by basic lead acetate.

Concentrated sulphuric acid, yellow ~ red with water, gray flocks being precipitated.

(*e*) *PROPÆSCINIC ACID* G. (occurs naturally and producible from above).

Soluble in alcohol and water.

Precipitated by basic lead acetate and by acids.

Potassic hydrate (strong solution)—compare above—converts into æscinic acid and propionic acid.

(*f*) *ÆSCINIC ACID* G. (occurs naturally, and producible artificially from Argyræscin and also from Aphrodæscin), $C_{24}H_{40}O_{12}$; crystalline.

Soluble in hot water (difficultly cold), in alcohol if freshly precipitated. **Insoluble** in ether.

Precipitated by acids and lead acetate (neutral).

Yields on hydrolysis, sugar and Telæscin, which is also a glucoside (see below).

(*g*) *TELÆSCIN* G., $C_{18}H_{30}O_7$, occurs naturally, and producible from above. Yields sugar and Æscigenin (below). In properties it resembles Quinova bitter, which see.

(*h*) *ÆSCIGENIN* (from Telæscin), indistinctly crystalline.

Insoluble in water. **Soluble** in alcohol.

Concentrated sulphuric acid and sugar, blood-red.

(*i*) *CAPSULÆSCINIC ACID* (from the shell of the nut), $C_{13}H_{12}O_8$; crystalline, sublimable.

Reactions resemble gallic acid.

(*j*) *APHRODÆSCIN* G., $C_{52}H_{82}O_{23}$; amorphous substance, the dust of which produces violent sneezing.

Soluble in water (frothy solution), and in alcohol.

Dilute alkaline solutions convert on heating to Butyric and Æscinic acids.

Precipitated by basic lead acetate.

(*k*) *QUERÆSCITRIN* G., $C_{41}H_{46}O_{25}$ (in leaves and nuts of the horse chestnut); yellow crystalline colouring matter resembling Quercitrin (see Quercus tinctoria).

Dilute acid, on boiling, decomposes into sugar and Quercetin.

§ 8. **ÆTHUSA** cynapium (Fool's Parsley); *Umbelliferæ*. Investigator: Ficinus, *Kastn. Arch.* 11, 144.

CYNAPINE A., crys. rhombic prisms, alkaline. **Sulphate** crystalline.

Soluble in water and alcohol, not in ether.

A coniine-like alkaloid was also found by Walz.

§ 9. **AGARICUS** muscarius; *Fungi*.

MUSCARINE A., deliquescent crystals; M.P. 100°; tobacco-like odour; no sublimate; tasteless, alkaline. Contracts the pupil.

Soluble in alcohol or water in all proportions; scarcely in chloroform. **Insoluble** in ether.

Precipitants:

[Not by lead acetate, neutral or basic.]	Bromine water, yellow pp. redissolving.
Ferric chloride.	Pp. with most other alkaloid reagents.
Copper sulphate.	

AMANITINE, $C_5H_{15}NO_2$ (isomer of Choline), yields Muscarine with nitric acid, and evolves trimethylamine on heating.

§ 10. **ALOE** (*Liliaceæ*), various species as A. Barbadensis (Barbadoes Aloes), A. Socotrina (Socotrine Aloes), A. Lucida (Cape Aloes), A. ferox, spicata, plicatilis, vulgaris, etc.

ALOIN B., $C_{16}H_{18}O_7$, Tilden (Barbaloin, Socaloin? see below). Crys. sulphur-yellow needles or granules; soften at 100° and then resinify; neutral, bitter, purgative.

Soluble in 600 parts cold or 10 parts boiling water, also in amyl alcohol, ethyl alcohol, and acetic ether; very difficultly in ether, and *not* in chloroform, benzene, or petroleum ether.

Reactions:

Alkaline hydrates, Ammonia } produce orange yellow solutions.
Lead acetate, neutral, no pp.
" " basic, dark yellow pp. *if concentrated.*
*Tannic acid, pp.
Ferric chloride, green black colouration.
Platinum chloride, red to violet with Barbadoes and Curaçao aloes.
" " greenish-brown with Socotra and Cape aloes.
" " yellowish-brown with Natal aloes.
Gold chloride, raspberry-red coloration~violet.

* Tests so marked (*i.e.*, *) are applicable to the residue obtained on evaporating the amyl alcohol employed to extract aloin from a solution previously treated with neutral lead acetate.

[Silver nitrate, no change.]	*Mercurous nitrate, pp.
[Mercuric chloride, no pp.]	Bromine water, pp.

Special colour tests: Potassic Cyanide to alcoholic extract from aloin (½ mgm.) evaporated with nitric acid—intense rose colour.

*Trace of copper sulphate to a solution previously diluted till colourless—intense yellow; sodium chloride or potassium bromide then added and some alcohol—intense red (with all aloes).

By these tests upon the residue from amyl alcohol (see foot-note), it is possible to detect very small quantities in mixtures, *e.g.*, beer. Similar reactions are not given with Cortex Frangula, Folia Sennæ, Radix Rhei, Baccæ Spinæ.

N.B.—The amyl alcohol residue should have characteristic aloe taste.

BARBALOIN, *NATALOIN*, *SOCALOIN* are regarded by Kranzfeld as differing only from Aloin in degree of hydration. Tilden gives to Barbaloin and Socaloin the same formula $C_{16}H_{18}O_7$. Groenewold ascribes to the active principle of Barbadoes and Curaçao aloes the formula $C_{16}H_{16}O_7$, and to Nataloin $C_{24}H_{26}O_{10}+H_2O$ with M.P. 210°.

§ 11. **ALSTONIA** scholaris (*Echises scholaris;* Dita bark). *Apocynaceæ;* Philippine Isles. The alkaloids, (*b*), (c), (*d*), etc.; **A.** constricta, (*a*), (*e*), etc.; **A.** spectablis, (*b*), (c), (*d*). In A. villosa (Blaberopus) Greshoff has found alkaloids 1·1 per cent. in the bark and 0·4 per cent. in the leaves—no doubt identical with the following, but not further examined. Investigators of first-mentioned varieties: Oberlin and Schlagdenhauffen, O. Hesse (compare Cinchona group).

(*a*) *ALSTONINE* **A.**, (Chlorogenine; *not* the Alstonine of O. and **S.**, see Alstonidine), $C_{21}H_{20}N_2O_4{\cdot}3\frac{1}{2}H_2O$. Amorphous brown (salts also amorphons); **M.P.** 100°, or 195° if anhydrous; alkaline reaction.

Soluble in alcohol, amyl alcohol, chloroform (with green fluorescence), with difficulty in water, and *not* removed by petroleum ether from a bicarbonate solution (compare Porphyrine).

Precipitants:

Alkaline hydrates.	Potassium bichromate, blood-red coloration with HCl solution.
Ammonia, no pp. if dilute, otherwise a pp. sol. in excess.	Acids in excess precipitate salts.

Concentrated acetic solution, crystals on addition of a few drops of HCl.

Charcoal carries down in part.

(*b*) *DITAMINE* **A.**, $C_{16}H_{19}NO_3$ (Hesse); amorphous, somewhat bitter, alkaline: M.P. 75° (becoming red at 130°). Salts mostly amorphous.

Soluble in alcohol, chloroform, ether, and benzene.

Precipitants:

Ammonia.	Gold chloride, pp. dissolves on boiling, crys. on cooling.
Platinum chloride, crystalline pp.	Mercuric chloride, same action.

Concentrated acetic solution, *no* crystals on addition of a few drops of hydrochloric acid (compare Alstonine).

Colour tests:

Concentrated sulphuric acid, red ~ reddish-violet on warming (uncertain).

Concentrated nitric acid, yellow ~ dark-green ~ orange-red.

(*c*) *ECHITAMINE* (Ditaine) **A.**, $C_{22}H_{28}N_2O_4+4H_2O$ (Hesse). Crystallizes in prisms from alcohol. Lævo-rotatory $[a]_D = -28{\cdot}8°$; loses $3H_2O$ at 80° C., 1 molecule at 130°, and then melts anhydrous at 206°; powerfully alkaline, bitter.

Soluble in water, alcohol, ether, chloroform, acetone, with difficulty in benzene, and **insoluble** in petroleum ether.

Reactions: [Ammonia, no pp., but the free alkaloid decomposes ammonium salts.]

Fehling's solution reduced after previously treating the alkaloid with boiling dilute acid.

The free alkaloid *decomposes sodium chloride, liberating NaHO!*

Colour tests: Concentrated sulphuric acid, purple.

Concentrated nitric acid, purple changing to green.

Hydrochloric acid produces an insoluble salt.

(*d*) *ECHITENINE* **A.**, $C_{20}H_{27}NO_4$? Formed from Echitamine by caustic potash. Amorphous brown (salts also amorphous); M.P. 120°; alkaline, bitter.

Soluble in alcohol, with difficulty in water; **insoluble** in petroleum ether. If freshly precipitated, soluble in ether and chloroform.

Precipitants:

Alkaline hydrates	pp. if concentd.	Platinum chloride.
Ammonia		Mercuric chloride, pale yellow.

Colour tests: Concentrated sulphuric acid, reddish-violet.

Concentrated nitric acid, red ~ purple ~ green ~ yellow.

(*e*) *PORPHYRINE* **A.** (Oberlin and Schlagdenhauffen's Alstonicine, presumably), $C_{21}H_{25}N_3O_3$; amorphous (Hesse) or crys. prisms from alcohol (O. and **S.**); M.P. 97°; alkaline; bitter; salts have blue fluorescence.

Soluble in alcohol, ether, chloroform, petroleum ether (removed by the latter from a bicarbonate solution—compare Alstonine).

Reactions: Ammonia, pp. and magnificent blue fluorescence.

Platinum chloride, Gold chloride } pp.

Potassium bichromate, yellow pp. and red coloration (not permanent).

Mercuric chloride, pp.

Colour tests:

Concentrated sulphuric acid
„ „ „ +molybdic (Fröhde's soln.) } purple.
„ nitric acid
„ sulphuric acid and potassium bichromate, greenish-blue.

(*f*) *ALSTONIDINE* A. (Hesse's), closely resembles O. and S.'s Alstonine, but not identical. Crystallizes in needles; M.P. 181°; bitter.

Soluble in acetone. Salts have blue fluorescence.

Precipitated by alkaline hydrates and ammonia.

No colours with concentrated sulphuric acid or Fröhde's solution.

(*g*) *ECHITIN* (Cynanchin? see Asclepias group), $C_{32}H_{52}O_2$; non-alkaloidal, non-glucosidal; crystallises in needles; dextro-rotatory.

Soluble in 1,430 parts of alcohol, and in ether and chloroform.

(*h*) *ECHICERIN* (Cynanchocerin?), $C_{30}H_{48}O_2$; wax-like, crystalline, dextro-rotatory.

Soluble in ether, chloroform, benzene, petroleum ether, acetic ether, and with difficulty in acetone.

Saponifiable, yielding a substance crystallizing in needles.

(*i*) *ECHIRETIN* (=Lactucerin=Cubeb Camphor, Hesse). Aromatic.

Soluble in ether, chloroform, benzene, acetic ether, and hot alcohol.

§ 12. **AMARYLLIS** formosissima (*Amaryllidaceæ*); alkaloid (*a*). Investigators: Arata and Gieger. A. Belladonna, alkaloid (*b*).

(*a*) *AMARYLLINE* A. Crystallises in needles from alcohol; M.P. 196° (previously becoming brown).

Soluble in alcohol, ether, chloroform, and with difficulty in water.

Precipitants:

[Not tannic acid.]	Bismuth potassic iodide }
Picric acid, yellow.	Cadmium potassic iodide } yellow.
[Not platinum chloride.]	Mercuric potassic iodide, yellowish-
[Not potassium bichromate.]	green.

'Pps. also with other alkaloid precipitants.'

Colour tests: Concentrated sulphuric acid, red-brown~brown~green with water.

Fröhde's solution, brownish-green.

(*b*) *BELLAMARINE* A. Crystallizes in needles; M.P. 181° (previously browning).

Soluble in alcohol, ether, chloroform.

Precipitants:

Alkaline carbonates.	Potass. bichromate, crystalline pp
Tannic acid, white.	Bismuth potassic iodide, white.
Picric acid, yellow.	Cadmium potass. iod., yellow-white.
Platinum chloride, yellow.	Merc. potassic iodide, yellow-green.

Colour tests: Concentrated sulphuric acid, gray~red on warming.

Concentrated sulphuric acid+nitre, yellow-green.
„ „ „ +potass. bichromate, yellowish-green~brown.

Fröhde's solution, brown.

§ 13. **AMMI WISNAGA**—the seeds ('Kell'); *Umbelliferæ*. Egypt. Investigator: Mustapha, Compt. R., Aug., 1879.

KELLIN G. Crys. silky needles, neutral, very bitter. Action emetic and narcotic.

Soluble in *hot* water, amyl alcohol, methyl alcohol, and chloroform, but with difficulty in these fluids when cold; readily soluble in ether.

Reactions:

Fehling's solution, reduced.
Nessler's „ white pp., sol. in alcohol.
Loew's reagent „ „ excess.

§ 14. **ANAGYRIS** fœtida; *Leguminosæ*. Investigators: Gallois and Hardy.

ANAGYRINE A., $C_{14}H_{18}N_2O_2$; amorphhous, alkaline, bitter. Salts crystalline.

Soluble in water, alcohol, ether, benzene.

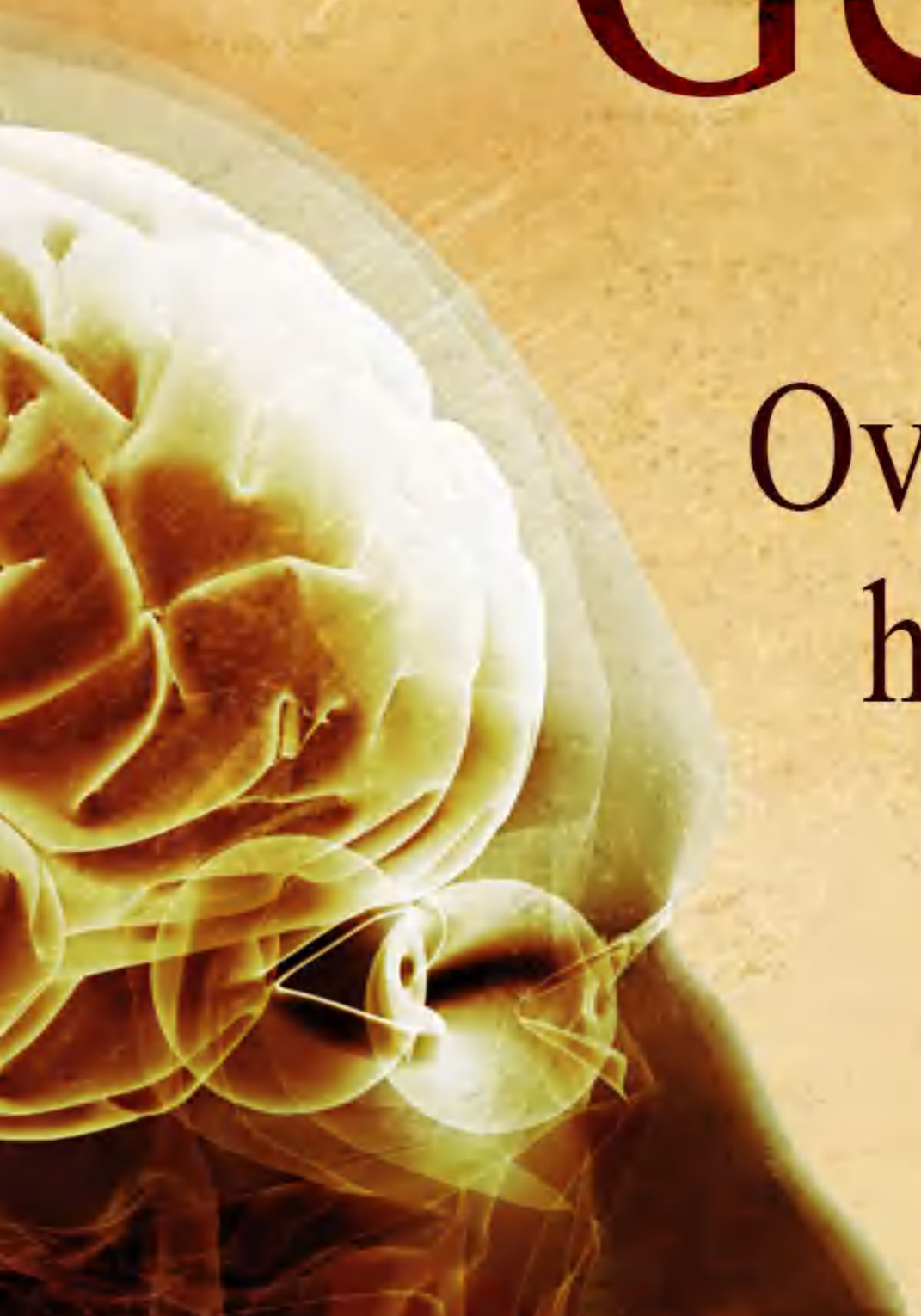
Precipitates with most alkaloid reagents.

Concentrated sulphuric acid, unchanged cold; musk-like odour on warming.

§ 15. **ANCHIETA** salutaris; *Violaceæ*. Investigator: Peckolt, *Archiv. Pharm.* [2].97.271.

ANCHIETINE A. Crys. light yellow needles, feebly alkaline, pungent taste. Salts crystalline. Action emetic.

Soluble in alcohol; *not* in ether, nor cold water, and scarcely when hot.



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Solubility: in water, 1 in 254 at 22°, 1 in 27 boiling; in alcohol, 1 in 70 at 22° C.; in ether, 1 in 2,792 at 22°.

Reactions:

Alkaline hydrates, dissolve.
Ammoniacal silver nitrate, reduced.
[No pp. tannic acid nor metallic salts.]
Concentrated sulphuric acid, yellowish-brown.

(*b*) *ANTIARETIN* G., deriv. (from above). Feathery crystals.
Soluble in alcohol, ether, benzene, and petroleum ether.

§ 23. **APOCYNUM** cannabinum, *Apocynaceæ*. (See also § 23 *a*.)

APOCYNIN G. Poisonous. **Soluble** in alcohol and ether, scarcely in water.

APOCYNEIN G. **Soluble** in water.

§ 23*a*. **APOCYNEÆ** of the Dutch Indies. The following series of plants have been investigated by Greshoff, with results as below:

1. Rauwolfia (Ophioxylon) serpentina and trifoliata, (*a*), (*b*), (*c*).
2. „ canescens, Rauwolfia (Cyrtosiphonia) spectabilis and madurensis, (*b*), (c).
3. Hunteria corymbosa, the cortex, (*c*), (*d*)—0·3 per cent. of the latter.
4. Ochrosia (Lactaria) acuminata, ackeringæ, coccinea, (*e*), (*f*).
5. Melodinus lævigatus, Bl., leaves, cortex, and seed (*g*).
6. Leuconotis eugenifolia, Dec., (*h*).
7. Pseudochrosia glomerata, Bl., (*c*), and a poisonous alkaloid, possibly (*d*).
8. Kopsia flavida, Bl., seeds, 1·85 per cent. alkaloid, sol. ether.
9. „ arborea, alkaloid, sol. ether, and substance possibly (*c*).
10. „ (Calpicarpum) Roxburghii, a tetanus-producing alkaloid, differing from above.
11. Kopsia (Calpicarpum) albiflorum, contains an alkaloid.
12. Vinca rosea, L., contains an alkaloid.
13. Alstonea (Blaberopus) villosa, contains an alkaloid (see Alstonia).
14. Voacanga Orchipeda, (*c*), and a bitter alkaloid, soluble in ether.
15. Tabernæmontana sphæro-carpa, Bl., contains an alkaloid.
16. Rhincodia (Cercocoma) macrantha „ „
17. Chonemorpha macrophylla „ „

(*a*) *OPHIOXYLIN* (Dulong's Plumbagin), see No. 1 above. $C_{16}H_{13}O_6$ or $C_{48}H_{39}O_{18}$; orange crystals of tetragonal system; M.P. 71·8°; burning taste; resembles Juglone.

Soluble in alcohol, in water sparingly, very soluble in chloroform, petroleum ether and carbon bisulphide.

(*b*) *AN ALKALOID*, giving blood-red with nitric acid (from 1 and 2 above).

(*c*) *A BROWN SUBSTANCE*, soluble in ether with blue fluorescence (from 1, 2, 3, 7, 9, and 14, above).

(*d*) *AN ALKALOID*, crystalline and poisonous, with sharp burning taste. Salts crystalline (from 3, and perhaps 7, above).
Frühde's solution produces violet coloration.

(*e*) *AN ALKALOID*, crystalline, colourless, moderately poisonous, soluble in ether (from 4 above).

(*f*) *AN ALKALOID*, insoluble in ether, dissolved by amyl alcohol (from 4 above).

(*g*) *AN ALKALOID*, giving reactions in very dilute solutions (from 5 above). Concentrated sulphuric acid with feeble oxidizers, green~~deep blue~~orange.

(*h*) *AN ALKALOID*, crystalline, poisonous, soluble in ether (from 6 above).
Precipitates with most alkaloid reagents.
No colour reactions.

§ 24. **ARALIA** spinosa, *Araliaceæ* (False Prickly Ash; the true Prickly Ash = Xanthoxylum). The bark. Investigator: J. Lilly, *Pharm. J., T.*, 1882. [For A. quinquefolia, see Panax.]

ARALIIN G. Yellowish powder, neutral reaction, acrid taste.
Soluble easily in water (saponaceous solution) and dilute acetic acid, scarcely in absolute alcohol (sol. in dilute alcohol). **Insoluble** in ether, chloroform, or benzene.

Reactions:

Alkaline hydrates, no effect—amber colour on boiling.
(Ammonia, no effect.)
[Lead acetate, neutral, no pp.]
Lead acetate basic, pp. from which Araliin may be removed by alcohol.
Tannic acid, no pp. cold, but pp. hot (the substance is, however, liable to be carried down in the cold by the pp. formed with other matters in the bark).
[Platinum chloride and other alkaloid reagents, no pp.]
Concentrated sulphuric acid, decolourizes and produces the odour characteristic of the plant.
Nitric acid, no action.
Hydrochloric acid, as sulphuric acid.

MEM.—A tannin is contained in the root, giving red colour with potash and green with ferric chloride.

§ 25. **ARARIBA** rubra (Pickneya rufescens), *Rubiaceæ*.

ARIBINE A., $C_{23}H_{20}N_4$ and with $5H_2O$ (Rieth and Wöhler, 1861). Crys. rhombic octahedra or prisms; optically inactive; neutral; bitter; M.P. 229° with sublimate; salts crystalline.

Soluble in alcohol, with difficulty in amyl alcohol or ether, and requiring 7,762 parts of cold water (easier soluble on heating, with deposition of crystals as solution cools).

Precipitants:

Alkaline hydrates.
" carbonates.
Ammonia.
[Lead acetate, no pp.]
[Tannic acid, no pp.]
Chlorine water, pp. sol. warm.
Bromine water, yel. pp. sol. warm.
Iodine tincture, pp. sol. warm.
Concentrated acids, pp. from solutions.

§ 26. **ARBUTUS** group; *Ericaceæ*, *Pyrolaceæ*, etc. A. uva ursi (Arctostaphylos, or Bearberry), (*a*), (*f*), not (*d*). Arctostaphylos glauca (Manzanita), the leaves (*a*).

Chimaphylla umbellata (Pyrola umbellata = Winter green), (*a*) not (*d*).

Ledum palustre (Labrador Tea, or 'James' Tea'), (*a*) not (*d*).

Calluna vulgaris (*a*).

Erica Rhododendron (c); Pyrola (*c*); Vaccinium (Cranberry; V. vitis idæa = Red Whortleberry), (*c*), (*e*); Azalea Indica, (c), (*d*); Gaultheria (Winter Green, see also above), (*c*) [no (*d*) in G. procumbens]; Clethra (*c*) [Clethra arborea and alnifolia do not contain (*d*)]; Eriodictyon, (c); Epigæa, (c); Rhododendron ponticum, (*c*), (*d*); R. chrysanthemum, Hybridum and Maximum, (*d*) [no (*d*) in R. Hirsutum].

(*d*) Also in all the following: Andromeda japonica [also (*g*), (*h*)], A. Catesbæi, Calyculata, Polyfolia, Angustifolia; Kalmia Latifolia.

(*d*) Is *not* contained in Erica vulgaris, nor in Oxydendron arboreum.

(*a*) *ARBUTIN* G., $C_{25}H_{34}O_{14}$ or $C_{12}H_{16}O_7$(?). Crys. silky needles with $2H_2O$; lævo-rotatory; M.P. 144°-146°, if previously dried at 100°; very hygroscopic; neutral; bitter.

Acids change to sugar and hydroquinone.

Soluble in water and alcohol when hot, with difficulty in these fluids when cold, and scarcely in ether.

Reactions:

Not precipitated by neutral lead acetate nor by other metallic salts.
Ferric chloride, blue coloration.
[Fehling's solution, not reduced.]

(*b*) *METHYLARBUTIN* G., $C_{13}H_{18}O_7$(?). Similar to Arbutin; M.P. 142°-143°. Yields sugar and methyl-hydroquinone.

(c) *ERICOLIN* G., $C_{34}H_{56}O_{21}$; amorphous brown; semi-fusion at 100°. Removed by benzene from an acid solution.

Reactions:

[Lead acetate, neutral or basic, no pp.]
[Tannic acid, no pp., or but slight.]
Gold chloride, no pp., but reduction on warming.
[Silver nitrate, no pp.]
Concentrated sulphuric acid, yellowish-brown.
" " " with sugar, red.
Fröhde's solution, brown.
Warmed with dilute acid, odour of ericinol.

(*a*) *ANDROMEDOTOXIN* (Asebotoxin) B., $C_{31}H_{51}O_{10}$ (de Zaeyer); crystallizes in needles; M.P. 229°; lævo-rotatory) but dextro-rotatory in chloroform solution); neutral reaction, bitter. Extremely poisonous, more so even than aconitine, and more emetic than emetine or apomorphine.

Solubility: 100 *parts of the respective solvents take up the following quantities*—water, 2·81 at 12° C., 0·87 at 100°; alcohol, 11·7 at 12° C.; amyl alcohol, 1·14; chloroform, 0·26; ether, 0·07; benzene, 0·004. **Insoluble** in petroleum ether.

Reactions: Alkaline hydrates remove the colours produced by acids (on re-acidifying the solution, the colours reappear).

No precipitates by lead acetate, neutral or basic, nor by usual alkaloid reagents.
Concentrated sulphuric acid, dark brown, becoming red on warming.
Dilute sulphuric acid solution slowly evaporated, magnificent red—resembling narceine reaction (see Opium).
Solution in 25 per cent. phosphoric acid evaporated with a little hydrochloric acid, red.

(*e*) *VACCININ* B. (nitrogen free). Crys. in prisms, fusible.

Soluble in alcohol and water (crystals from a cooling solution), scarcely in ether.

(*f*) *URSON*, $C_{20}H_{17}O_2$ (from Arbutus uva ursi); crystalline, tasteless; M.P. 198°-200°, sublimable.

Soluble with difficulty in alcohol or ether. **Insoluble** in water, acids, or alkalies.

Concentrated sulphuric acid gives yellow solution.

(*g*) *ASEBOFUSCIN* G., $C_{18}H_{18}O_8$; reddish-brown, yielding the violet substance *Asebopurpurin* on boiling with alcoholic hydrochloric acid.

Soluble in alcohol, acetic acid, and alkalies ; scarcely in water. **Insoluble** in ether, chloroform and benzene.

(*h*) *ASEBOTIN* G., $C_{24}H_{28}O_{12}$. Crys. needles ; M.P. 145·5° ; specific gravity, 1·356 ; bitter, non-poisonous. Yields sugar and *Asebogenin* on treatment with acids.

Soluble in alcohol, hot water, and glacial acetic acid ; with difficulty in cold water, or in ether, chloroform, benzene, petroleum ether.

Reactions :

Alkaline hydrates, dissolve (acids reprecipitate).

Lead acetate, basic, precipitates (no pp. with the neutral salt).

(*i*) *ASEBOGENIN*, $C_{18}H_{18}O_7$? (from Asebotin) ; crys. needles.

Soluble in alcohol, ether, and dilute alkalies ; with difficulty in water or chloroform.

Precipitated by basic lead acetate.

§ 27. **ARECA** Catechu, *Palmaceæ* ('Guvaca,' Betel-nut) ; the nut. Contains, also, Choline. For Catechin, see Catechu.

(*a*) *ARECOLINE* A., $C_8H_{13}NO_2$ or $(C_5NH_7)CH_3{\cdot}CO_2{\cdot}CH_3$ (Methyltetrahydro-nicotinate), E. Jahns. An oil, volatile in vapour of water, boils at 209° ; strongly alkaline, very poisonous. **Salts** crystalline.

Solubility : Miscible in all proportions with water, alcohol, ether, or chloroform.

Reactions (of the salts) :

Alkaline hydrates give no pp., but liberate (*a*) from its salts.

Tannic acid, no pp. with the salts, but free (*a*) gives pp.

Picric acid, a tarry pp., becoming crystalline.

[Platinum chloride, no pp.]

Gold chloride forms a double salt ($A{\cdot}HCl{\cdot}AuCl_3$), difficultly soluble in water.

Phospho-molybdic acid, white pp.

Bismuth-potassic iodide, pp. red, becoming crystalline.

Mercuric-potassic iodide, pp. yellow oily, becoming crystalline.

[Mercuric chloride, no pp.]

Iodine, pp. and brown solution.

(*b*) *ARECAINE* A., $C_7H_{11}NO_2+H_2O$ (E. Jahns) ; crystalline, neutral, non-poisonous ; M.P. 213° (becoming frothy).

Soluble in water ; scarcely in cold alcohol or in chloroform. **Insoluble** in ether or benzene. (Ether does not remove from alkaline aqueous solution—*distinction from Arecoline.*)

Reactions :

Tannic acid, turbidity.

[Picric acid, no pp.]

Phospho-molybdic acid, turbidity.

Potassic iodide, dark crystalline pp.

Bismuth-potassic iodide, pp. red becoming crystalline.

Mercuric-potassic iodide, pp. yellow crystalline.

(*c*) *ARECAIDINE* A., $C_7H_{11}NO_2$ (occurs naturally in the nut, and producible artificially from Arecoline by heating in sealed tube with alkali or hydriodic acid, methyl group being eliminated). Isomer of Arecaine. Crystallizes from 70 per cent. alcohol with 1 molecule H_2O. M.P. of anhydrous alkaloid, 223°. Non-poisonous.

Soluble in water and dilute alcohol ; scarcely in absolute alcohol, ether, chloroform, or benzene.

Ferric chloride gives red coloration.

(*d*) *GUVACINE* A., $C_6H_9NO_2$ or $NH{\cdot}CH_2{\cdot}CO{\cdot}CO{\cdot}CH(CH_3)CH_2$. Crystalline, neutral ; M.P. 271°-272° ; non-poisonous.

Soluble in water and dilute alcohol. **Insoluble** in ether, chloroform, benzene, or strong alcohol.

Salts have acid reaction, and their solubilities are similar to those of the free base.

Ferric chloride gives a deep red.

§ 28. **ARISTOLOCHIACEÆ**, Aristolochia serpentaria, the root (Serpentary root, or Virginian Snake root), (*a*) ; A. Argentina, (*b*), (c). [For A. cava, containing Corydaline, see Corydalis.]

(*a*) *ARISTOLOCHIN* B., $C_{32}H_{22}N_2O_{13}$(?). Bitter, poisonous, resembling Aloin physiologically.

Soluble in water, alcohol, and ether.

Reactions :

Alkaline hydrates, pp. (purple-red on fusion with potash).

Lead acetate neutral, pp.

[Fehling's solution, not reduced.]

Concentrated sulphuric acid, dark green.

Dilute acids, precipitate from solution.

(*b*) *ARISTOLOCHINE* A. (not identical with above—O. Hesse, 1892) ; amorphous, resinous.

Soluble in alchol, ether, chloroform, and benzene.
Precipitated by alkaline hydrates.

(*c*) *ARISTINE* A. Crystallizes in gold laminæ from hot glacial acetic acid ; neutral reaction ; decomposed at 260°.
Soluble in ether, chloroform, and benzene ; sparingly in hot alcohol.
Alkaline hydrates and ammonia convert to red compound.
Concentrated sulphuric acid, blue~greenish.

§ 29. **ARNICA** montana (Leopard's bane), *Compositæ*. Investigator : Walz, *N. Jahrsb. Pharm.*, vols. 13 and 15.

ARNICIN B., $C_{20}H_{30}O_4$(?). Amorphous terebinthinate ('possibly crystallizable.' Dupuy) ; bitter, non-volatile.
Soluble in alcohol and ether ; difficultly or insoluble in water.
Alkaline hydrates } dissolve.
Ammonia }
Precipitated by the following from alcoholic solution :
Lead acetate, basic.
Tannic acid.
Platinum chloride.
Silver nitrate.
Mercuric chloride.

§ 30. **ARTEMISIA** absynthium (Wormwood), *Compositæ* (*a*). A. abrotanum (Southernwood), (*b*) ; A. maritima (Semen santonici, Semen contra, or Wormseed), (c). [A. herba alba, Asso. (A. ramosa, Smith ; or, A. silberi, Besser), contains no (*c*).]

(*a*) *ABSYNTHIIN* B. or G., $C_{40}H_{28}O_8$ or $C_{15}H_{20}O_4$ (Senger) ; pale yellow, amorphous ; M.P. 120°-125°? (65° Senger) ; neutral, bitter ; convertible, according to Senger, into dextrose, a volatile substance, and a resin.
Soluble in alcohol, ether, chloroform, benzene ; with difficulty in water or petroleum ether.
Benzene removes from acid solution ; petroleum spirit, only traces.
Reactions :
Alkaline hydrates, dissolve.
[Lead, acetate neutral or basic, no pp.]
Tannic acid, pp.
Gold chloride, pp. ; reduction on warming.
Colour tests :
Concentrated sulphuric acid, brown~green~blue ; no characteristic odour on warming.
Nitric acid, oxidizes to oxalic and picric acids, etc.
Fröhde's reagent, brown~violet-blue.

(*b*) *ABROTINE* A., $C_{21}H_{22}N_2O_2$ (Giacosa, 1883). Crystalline powder or needles ; fluorescent blue. Sulphate = $Ab_2{\cdot}H_2SO_4{\cdot}6H_2O$, crys.
Soluble with difficulty in hot water.
Platinum chloride forms a compound difficultly soluble.

(c) *SANTONIN* B. (Santonic acid), $C_{15}H_{18}O_3$. Crystals four-sided orthorhombic ; M.P. 169°-170°, yielding sublimate (if rapidly heated becomes brown, alcohol then giving a red solution) ; specific gravity 1·257 ; feebly acid reaction, bitter in alcoholic solution ; lævo-rotatory ; anthelmintic.
Solubility : 1 in 4,000 cold, 1 in 250 hot water ; 1 in 72 cold alcohol, easier warm ; 1 in 4·35 chloroform ; 1 in 42 hot ether ; soluble also in benzene, acetic acid, and ethereal oils.
Chloroform and benzene remove from *acid* solutions on shaking, but not from *alkaline* aqueous mixtures.
Precipitants :
Alkaline hydrates, dissolve (alcoholic sodic hydrate gives a momentary red).
Lead acetate, neutral or basic (the pp. is a lead santonate, soluble in warm alcohol).
Tannic acid.
Ferrous sulphate.
Copper sulphate.
Mercurous nitrate.
Chlorine water, slow deposition of crystals.
Concentrated sulphuric acid, dissolves at first colourless ; on heating to 150°, then adding 1 drop ferric chloride, red~gradually violet.
Dilute acids do not dissolve, but precipitate from solution after some days. (May be 'shaken out' by chloroform.)

(*d*) *HYPOSANTONIN*, $C_{15}H_{18}O_2$ (Grassi-Cristaldi), formed by reduction of santonin-phenylhydrazone with sodium amalgam. Crystallizes in shining plates ; M.P. 152° ; dextro-rotatory : $[a]_D = +30$.
Soluble in benzene, warm alcohol, ether, acetic acid. **Insoluble** in water and cold alkaline solutions, but latter dissolve on warming.
Concentrated sulphuric acid and ferric chloride, violet~green.

(*e*) *ISOHYPOSANTONIN*, from (*d*) by dissolving in potash and reprecipitating with acid. Lævo-rotatory : $[a]_D = -70{\cdot}31$.

§ 31. **ASCLEPIADACEÆ**, various. A. syriaca, A. currassavica (Bastard ipecacuanha), A. incarnata (Swamp silk-weed), A. tuberosa (Butterfly-weed), and Vincetoxicum officinale, (*a*). Solenostemma (Cynanchum) Argel, (*b*), (c). [Amygdalin was found by Greshoff in an Indian Asclepiadea : Gymnema latifolium.]

(*a*) *ASCLEPIADIN* G. (C. Gram) ; amorphous yellow, bitter ; emetic, diaphoretic and purgative. Convertible into a substance, Asclepin.

Soluble in ether, alcohol, and hot water ; with difficulty in cold.

Not precipitated by lead acetate, neutral or basic, but by tannic acid.

Concentrated sulphuric acid, green (at first yellowish, then deepening in colour).

(*b*) *CYNANCHIN* (compare Echitin, *Alstonia*). Crys. glittering plates, M.P. 148° 149°.

Soluble in alcohol, ether, chloroform.

(c) *CYNANCHOCERIN* (compare Echicerin, *Alstonia*). Crystallizes in flat needles.

Soluble in alcohol, ether, and chloroform.

§ 32. **ASIMININIA** triloba, *Papaveraceæ*. The seed.

ASIMININE A. ; amorphous, tasteless, odourless, alkaline.

Soluble in alcohol and ether (which removes from alkaline solution) ; with difficulty in benzene or chloroform, and scarcely in water.

Precipitated by ammonia and by usual alkaloid reagents.

Colour tests :

Concentrated sulphuric acid, produces effervescence, then dissolves green⁓reddish-yellow⁓dark-red⁓colourless.

Nitric acid, carmine⁓purple.

Hydrochloric acid, colourless cold, purple on warming (like morphine).

§ 33. **ATHAMANTHA** oreoselinum. Peucedanum oreoselinum, (*a*), (*b*) ; Peucedanum officinale, Imperatoria ostruthium, (*b*), (*c*), (*d*?), (A. Jassoy found no (*d*) but only (*c*) in I. ostruthium).

(*a*) *ATHAMANTHIN* B. (see Laserpitin), $C_{24}H_{30}O_7$; bitter, rancid taste.

Soluble in alcohol and ether, not in water.

(*b*) *OREOSELON*, $C_{14}H_{10}O_3$; not bitter.

Soluble with difficulty in alcohol or ether. **Insoluble** in water.

Alkalies dissolve to yellow solution.

(c) *OSTRUTHIIN* [=Imperatorin, A. Jassoy, compare (*d*)], $C_{14}H_{17}O_3$ (or $C_{18}H_{20}O_3$, A. Jassoy).

Soluble in alcohol and ether ; scarcely in water.

(*d*) *PEUCEDANIN* B. (Imperatorin ? this denied by Jassoy, see above). $C_{15}H_{14}O_4$ or $C_{14}H_{11}O_2{\cdot}OCH_3$ (Jassoy). Crystalline ; M.P. 76°-82° ; burning taste.

Soluble in hot alcohol, in ether, chloroform, benzene, carbon-bisulphide, oils, and hot acetic acid. **Insoluble** in water.

§ 34. **ATHEROSPERMA** moschata, *Atherospermaceæ* (Plume nutmeg, or Australian sassafras).

ATHEROSPERMINE A., $C_{30}H_{40}N_2$; feebly alkaline, bitter powder ; M.P. 128°, methylamine odour at high temperature. Salts amorphous.

Soluble in 6,000 parts cold water, 32 cold and 2 boiling alcohol, in chloroform, carbon bisulphide, turpentine, essential oils, and with difficulty in ether.

Precipitants:

Alkaline hydrates.	Phospho-molybdic acid, dirty yellow.
„ carbonates.	Iodo-potassic iodide, brownish-yellow.
Ammonia.	Mercuric chloride, white.
Tannic acid, yellowish-white.	(Chlorine water, yellow solution unchanged by ammonia.)
Picric acid.	Iodide is liberated from iodate.
Platinum chloride.	
Gold chloride.	
Potassium ferrocyanide.	
„ sulphocyanide.	

Concentrated sulphuric acid, no colour.

„ „ „ with potassic chromate, green.

§ 35. **ATROPA** group (*Atropaceæ*), and various of the *Solanaceæ* (see also **Solanum**). Atropa belladonna (Deadly nightshade), (*a*), (*g*), (*h*), (*k*) ; Scopolia japonica (Japanese belladonna), (*g*) principally, small quantity (*a*), also (*h*), (*n*), (*o*), (*p*) ; Scopolia carniolica. (*g*) only ; S. Hlardnackiana (*g*), possibly (*p*), Schmidt ; Anisodus luridus. *Solanaceæ*, the plant in flower, (*g*). no (*g*) in seeds but trace of (*a*) ; Hyoscyamus (Henbane). (*q*) ; Lactuca sativa (Lettuce, *Compositæ*), trace of (*g*) ;* Hyoscyamus Niger, (*g*), (*h*) ; Datura stramonium, (*a*), (*g*), (*h*), (*l*) ; Duboisia Hopwoodii, (*g*), (*m*) ; D. myoporoides, (*g*), (*g*2), (*h*) ; Mandragora (Mandrake). (*g*) ; Nicotiana tabacum (Tobacco), (*r*) ; Fabiana imbricata, *Solanaceæ-Nicotianeæ*, S. America, (*s*), (t). (For Solanine, etc., see Solanum.)

(*a*) *ATROPINE* A. (Tropine tropate, Daturine : Erhardt and Poebl dispute the identity of Atropine and Daturine. Commercial Daturine is frequently a mixture of Hyoscyamine and Atropine, or former solely). $C_{17}H_{23}NO_3$ or $C_5H_7(CH_2CH_2O{\cdot}CO{\cdot}CH[C_6H_5]CH_2OH)NCH_3$ (May, 1891), Ladenburg. Crys. pillars and needles ; optically inactive (or feebly lævo-rotatory ?). M.P. variously given as 65°, 90°, 114°, part sublimable. Alkaline reaction (reddens Phenolphthalein, *an exception to nearly all alkaloids* ; for other bases reacting similarly, see 'Reaction,' Part III.)

* For other constituents of Lettuce, see Lactuca.

Disagreeable metallic bitter taste. Produces dilatation of the pupil. Decomposed readily by strong acids and alkalies (except ammonia), forming tropine and tropic acid, $C_9H_{10}O_3$.

Soluble in 300 parts cold, or 58 of boiling water, 3 chloroform, 30 cold and 6 boiling ether, 40 benzene; miscible in nearly all proportions with alcohol; soluble also in glycerine, but scarcely in petroleum ether or carbon bisulphide. It is removed by benzene from alkaline solution, not easily by ether.

Precipitants:

Alkaline hydrates (not if dilute), pp. sol. in excess with gradual decomposition.
Alkaline carbonates.
[*Not* bicarbonates.]
[*Not* sodium acetate, compare Opium.]
Ammonia hydrate.
[*Not* am. carbonate.]
Tannic acid (1 in 3,000 ? sol. in HCl. not 1 in 200.—Dragendorff.* 'No pp. unless acid.'—Dupuy.)
? Picric acid, 1 in 200, not 1 in 500; pp. sol. excess, or pp. with Daturine, not with Atropine ?
[Not ferric chloride.]
? Platinum chloride, *no* pp. Atropine, *pp.* Daturine — (Erhardt).
Gold chloride, lemon-coloured pp. 1 in 3,000 ? (1 in 100, Dragendorff).* No reduction.
Mercuric-potassic iodide, up to 1 in 7,000.
Mercuric chloride, cloud at 1 in 3,000 then pp.; partly sol. HCl.
*Mercuric chloride in alcohol, to the free base, red pp. (HgO).
[*Not* by potassium ferrocyanide.
" " ferricyanide.
" " sulphocyanide.
Not by potassium chromate.]
Potassium bichromate, 1 in 3,000 gradually.
[*Not* chromic acid, 5 per cent. solution.]
[*Not* silver-potassic cyanide.]
Phospho-molybdic acid, yellow flocks~blue with ammonia; cloud 1 in 4,000.
Phospho-tungstic acid.
Phospho-antimonic acid, up to 1 in 5,000, sol. warm.
Iodo-potassic iodide, reddish-brown, 1 in 3,000.
[*Not* potassic iodide.]
Bismuth-potassic iodide, orange; flocculent pp. 1 in 10,000; faint at 1 in 16,000.
Cadmium-potassic iodide (1 in 500—Dragendorff).†
[Not zinc-potassic iodide, or but slightly.]
Mercurous nitrate, dilute solution free from excess of acid, gives,
when added to a dilute solution of free Atropine, a black pp.
Bromine water, yellow, becoming crystalline.
Bromine in hydrobromic acid or in alcohol, gives crystalline compound with Atropine or Hyoscyamine, seen as radiating leaflets under the microscope ('Characteristic'—T. G. Wormly).
Charcoal, partially absorbs.

* 'Ermittelung der Gifte.'

* Gerrard test (see Part III.).

Colour tests (nearly all are negative; see, however, Vitali test):

Concentrated sulphuric acid, colourless. If atropine sulphate be heated first alone till white fumes appear, then sulphuric acid (1·5 gramme) be added, and farther heated till browned, an agreeable odour is produced on subsequent addition of 2 cc. water. On then adding potassium permanganate (a fragment), an odour of bitter almond oil will be perceived.
Concentrated sulphuric acid and potassium bichromate, discoloured merely.
Concentrated nitric acid, crystals become brown; solution colourless.
Vitali test: Evaporated with nitric acid at 100°, then touched with a drop of freshly-prepared alcoholic potash, magnificent violet changing to cherry red gradually.
Concentrated hydrochloric acid, no effect.
Per-iodic acid, not reduced.
Fröhde's solution, colourless.

(*b*) *DEXTRO-ATROPINE*
(c) *LÆVO-ATROPINE*
These have been formed artificially from the respective optically active tropic acids with Tropine; they form crystals from alcohol, the former melting at 110°·111°, and the latter at 111°. Rotatory power of former $=+10°$.

(*d*) *TROPINE* A., $C_8H_{15}NO$ or $C_5H_7(CH_2CH_2OH)NCH_3$ (compare Ecgonine, under *Coca*). Formed from Atropine or Hyoscyamine by acids or alkalies; rhombic leaflets; M.P. 61°, boils at 229°; hygroscopic.

Soluble in water and alcohol readily. Platinum salt soluble.

Concentrated sulphuric acid yields (*e*).

(*e*) *TROPIDINE* A. (from above), $C_8H_{13}N$. An oily liquid with coniine-like odour; boils at 162°.

Soluble in cold water; less easily hot.

Picric acid gives yellow pp.

(*f*) *HOMATROPINE* A., $C_{16}H_{21}NO_3$ (from Tropine and oxytoluic acid).

Crystallizes in prisms; M.P. 95°-98°. Resembles Atropine in physiological action.

Soluble only with difficulty in water, although hygroscopic; readily in ether and chloroform.

Precipitants:

Picric acid, yellow becoming crystalline.
Mercuric potassic iodide, white flocculent.
Mercuric chloride in alcohol, to the free base, red pp. (HgO).

Colour test: Evaporated with nitric acid, then touched with drop of alcoholic potash when cold, yellow coloration.

(*g*) *HYOSCYAMINE* A. (Duboisine; 'Commercial samples with latter name are frequently only scopolamine.'—Schmidt), $C_{17}H_{23}NO_3$; isomeric with Atropine, from which it differs only (according to Ladenburg's suggestion) as tartaric differs from racemic acid. Crys. in needles or plates if pure, otherwise amorphous; M.F. 108° and slightly volatile; feebly lævo-rotatory. No odour if pure, otherwise tobacco-like. Dilates the pupil. Reaction alkaline, both to phenolphthalein as well as litmus (compare Atropine). Sharp and disagreeable taste.

Soluble in hot water; with difficulty in the cold (easier sol. when impure), in alcohol, amyl alcohol, ether, chloroform benzene.

It is removed from *alkaline* solution by benzene and ether (not easily by latter).

It may be obtained in the *crystalline* form from benzene (in needles), and from chloroform (in plates), but only *amorphous* from amyl alcohol or ether.

Precipitants:

(Alkaline hydrates, only in part even with concentrated solutions).
(„ carbonates]
(Ammonia hydrate] „ „ „ „)
Tannic acid, yellowish-white.
Picric acid, yellow.
[*Not* ferric chloride.]
? *Not* Platinum chloride, unless concentrated (pp. sol. in excess—Dragendorff, 'Ermittelung der Gifte').
Gold chloride, yellowish white.
[*Not* potassium ferrocyanide.
„ ferricyanide.
sulphocyanide.
chromate.]
Phospho-molybdic acid, yellowish-white.
Phospho-tungstic acid.
Iodo-potassic iodide, orange.
[Not potassic iodide alone.]
Bismuth-potassic iodide, red amorphous.
Cadmium-potassic iodide, 1 in 10,000.
Mercuric-potassic iodide, white.
Mercuric chloride.
Iodine tincture.

Colour tests (negative):

Concentrated sulphuric acid, colourless.
„ „ „ with sugar, no effect.
„ „ „ with potassium bichromate, discoloured.
„ nitric acid, colourless.
„ hydrochloric acid, colourless.
Fröhde's solution.

(*g*2) *PSEUDOHYOSCYAMINE* A., $C_{17}H_{23}NO_3$. Obtained by E. Merck from Duboisia myoporoïdes. Crys. needles; M.P. 133°-134° (with decomposition); lævo-rotatory.

Readily soluble in alcohol and chloroform; with difficulty in water and ether.

Yields, on hydrolysis, tropic acid and a base isomeric with Tropine.

Gold chloride salt of Pseudohyoscyamine melts at 176° C.

(*h*) *HYOSCINE* A. (Sikeranine; Scopolamine), $C_{17}H_{23}NO_3$ (Ladenburg), or $C_{17}H_{21}NO_4$ (Hesse and Schmidt); amorphous; semi-fluid, giving crystalline salts; M.P. about 55° (Hesse), or 59° (Schmidt); lævo-rotatory: $[a]_D = -13{\cdot}7$]. Dilates the pupil.

Soluble in alcohol, ether, chloroform, benzene, and with difficulty in water. Removed from alkaline solution as atropine (see above).

Precipitants:

Gold chloride; may be separated from commercial Hyoscyamine by fractional precipitation.
Potassium ferrocyanide, white amorphous.
[*Not* by potass. ferricyanide.]
[*Not* by potass. sulphocyanide.]
„ „ chromate].
Phospho-tungstic acid.
Iodo-potassic iodide, pp. oily.
[*Not* potassium iodide.]
Mercuric-potassic iodide, yellow.
Mercuric chloride, amorphous.

No coloration with concentrated sulphuric acid.

(*i*) *ATROPAMINE* A., $C_{17}H_{21}NO_2$; stands to Hyoscine as does Apoatropine

times toxic power of Coniine. Salts crystallize with difficulty, and lose Nicotine on evaporation.

Soluble in all proportions, in water, alcohol, and ether; also soluble in benzene, chloroform, petroleum ether, oils, and in 40 parts of turpentine.

It is removed from alkaline solution by petroleum ether as well as by ether, benzene, and chloroform.

Sulphur is dissolved by it at 100°, and crystallizes out on cooling.

Precipitants:

(It is only displaced from its salts by alkalies and alkaline earths.)

Solid potash causes Nicotine to separate from solution.

Lead acetate, neutral, pp. of *lead hydrate* with the free base.

Tannic acid, cloud at 1 in 500, sol. in warm hydrochloric acid, but reappearing as solution cools.

[Not gallic acid.]

Picric acid *in excess* gives pp., otherwise not in dilute solution. The pp. is amorphous, but changes to crystals.

Ferric chloride, pp. of ferric hydrate by free base.

Platinum chloride, pp. sol. in hot water and in hydrochloric acid; cloud at 1 in 5,000 using HCl salt.

Gold chloride, gradual cloud at 1 in 10,000.

[Not potassium ferrocyanide.
" " ferricyanide.
" " sulphocyanide.
" " chromate.
" silver-potassic cyanide.]

Phospho-molybdic acid, cloud at 1 in 40,000.

Phospho-antimonic acid, cloud at 1 in 250.

Iodo-potassic iodide to acidified solution, reddish-brown pp., up to 1 in 250,000.

Bismuth-potassic iodide, pp.; cloud at 1 in 40,000.

Cadmium-potassic iodide, pp., 1 in 10,000? (1 in 500 Dragendorff, 'Ermittelung der Gifte').

[Not zinc-potassic iodide.]

Mercuric-potassic iodide, white amorphous resinous characteristic; cloud at 1 in 15,000; limit, 1 in 25,000.

Mercuric chloride gives *with free Nicotine*, white crystalline pp., 1 in 3,000.

Mercuric cyanide, not with free Nicotine, but crystalline pp. when neutral Nicotine hydrochloride is added to saturated sol. of mercuric cyanide.

Chlorine gas colours brown to blood-red; substance formed is sol. in alcohol yielding crystals on evaporation.

Iodine dissolved in ether gives with etheral solution of nicotine, after some hours, long needle-shaped crystals.

Colour tests (mostly negative):

Concentrated sulphuric acid, no effect.
" " " with sugar, no effect.
" " " " nitric acid, no effect.
" nitric acid, no effect, or scarcely yellow; but with larger quantity (than ordinarily used for such tests), say half-drop, violet-red~blood-red~colourless.

Concentrated hydrochloric acid, no effect (red amorphous residue on evaporation).

Fröhde's reagent, no effect or yellowish.

(*s*) *FABIANIN* G., giving fluorescent solutions.

Soluble in water, alcohol, chloroform, ether (removed by latter from alkaline solution).

Alkaline hydrates, Ammonia } dissolve to deep yellow fluorescent solutions.

Concentrated sulphuric acid with potassic bichromate, deep green.

Nitric acid, yellow.

Hydrochloric acid, ditto.

(*t*) *FABIIN.* Neutral, crystalline, non-fusible (browns at 270°).

Soluble in absolute alcohol or in boiling 95 per cent. spirit, and in ether or chloroform. **Insoluble** in water.

Unaffected by alkalies.

Concentrated sulphuric acid with potassic bicromate, dark blue.

§ 36. **AURANTIACEÆ.** Citrus aurantium (Sweet orange); C. bigaradia, or C. aurantium, var. Amara (Bitter orange): C. medica ('citron'), the fruits (*a*); Citrus decumana (Shaddock), (*g*), 2 per cent. in dried flowers; C. limonum (Lemon), the seeds (*i*); Murraya exotica (*j*).

(*a*) *HESPERIDIN* G. (see Barosmin), $C_{22}H_{26}O_{12}$ (or $C_{50}H_{60}O_{27}$. C. Tanret, 1888). Forms sphero-crystals; M.P. 250°-251°. With $\frac{1}{2}$ per cent. sulphuric acid in alcoholic solution, it yields Hesperidin + glucose (2 parts) + isodulcitol (1 part).

Soluble in acetic acid and alcohol, in 5,000 parts of hot water (crystals

on cooling); but **insoluble** in ether, chloroform, benzene, carbon, bisulphide, or acetone.

Alkaline hydrates } dissolve; acids reprecipitate.
„ earths (hydrates) }

[No pp. lead acetate, neutral or basic.]

Ferric chloride, brownish-red coloration.

Concentrated sulphuric acid dissolves, and on warming gives intense red.

Treated with sodium amalgam, and then hydrochloric acid, a pp. is obtained that dissolves in alcohol to a reddish-violet solution.

(*b*) *HESPERETIN* (from Hesperidin), $C_{10}H_{16}O_4$. Crystalline plates; M.P. 224°-226°; sweetish taste.

Soluble in alcohol, with difficulty in ether or benzene, and scarcely in water.

Ferric chloride, brown coloration.

(c) *HESPERETINIC ACID* (Iso-ferulic Acid), $C_{10}H_{10}O_4$, or C_6H_3 (CH : CH.COOH)(OH)(OCH_3), is formed, together with phloroglucin, by the action of hot potash solution on Hesperetin; M.P. 228°. Yields protocatechuic acid on fusion with potash.

(*d*) *ISOHESPERIDIN*, $C_{22}H_{26}O_{12}$, or $C_{50}H_{60}O_{27}+5H_2O$; same reactions as Hesperidin, but distinct therefrom (C. Tanret).

(*e*) *AURANTIAMARIC ACID*, $C_{10}H_{12}O_4$; resinous, very bitter, lævo-rotatory ([a] D = −28°).

Soluble in hot water.

(*f*) *AURANTIAMARIN* G.; bitter; probably same formula as Isohesperidin (Tanret).

(*g*) *NARINGIN* G. (De Vrij's Hesperidin; Aurantiin), $C_{23}H_{26}O_{12}+4H_2O$ (E. Hoffmann). Yellow, crystalline, bitter, lævo-rotatory (−84·5° in water, or −87·6° in alcohol); M.P. 171°. Yields Naringenin and isodulcitol on hydrolysis.

Soluble in 300 parts water, in alcohol and acetic acid. **Insoluble** in ether, chloroform, benzene, or essential oils.

Ferric chloride, reddish-brown coloration.

(*h*) *NARINGENIN* (from Naringin) = Phloroglucinol-paracoumarate, or $OH{\cdot}C_6H_4{\cdot}CH : CH{\cdot}COO{\cdot}C_6H_3(OH)_2$.

(*i*) *LIMONIN* G., $C_{22}H_{26}O_7$? (Bernays); microscopic crystals; M.P. 244° (or 275° Paterno); very bitter, neutral.

Soluble in alcohol and acetic acid; scarcely in water or ether.

Reactions:

Alkaline hydrates dissolve. | Tannic acid } pp. with the alcoholic
Acids reprecipitate. | Picric acid } solution.

Concentrated sulphuric acid dissolves blood-red; water reprecipitates.

(*j*) *MURRAYIN* G., $C_{18}H_{22}O_{10}+1\frac{1}{2}H_2O$ (air-dried); crys. fine needles; M.P. 170°, with crystalline sublimate; feebly bitter. Dilute sulphuric acid converts to Murrayetin and glucose.

Soluble in alcohol easily, and in hot water; difficultly in cold water, and scarcely in ether.

Alkaline hydrates } give solutions that are fluorescent on account of Murrayetin formed.
„ carbonates }
„ earths }

(*k*) *MURRAYETIN*, $C_{12}H_{12}O_5$; formed together with glucose from Murrayin. It is a fluorescent, tasteless substance.

Soluble in hot water and alcohol; difficultly in cold water or ether.

Basic lead acetate gives pp.

Ferric chloride, bluish-green coloration.

§ 37. **BACHARIS** cordifolia (*Compositæ*), **S.** America. Investigator: P. Arata, *J. Pharm.*, 1879.

BACHARINE A.; crystallizes in needles.

Solubility, slight in cold water, better hot, and in alcohol and ether; best in amyl alcohol.

Precipitated by most alkaloid precipitants (Dupuy).

§ 38. **BAPHIA** nitida (*Leguminosæ—Cæsalpineæ*), **Sierra Leone**, etc. From the wood ('Barwood' or 'Camwood') Anderson (*J. Chem. Soc.*, 1876) has extracted:

BAPHIIN B. $(C_{12}H_{10}O_4)_n$; crystallizing in lustrous needles or plates which are colourless if pure, otherwise red. M.P. below 200° C. (part only melts, remainder decomposing); odour of orris root on heating. Oxidizes on exposure, becoming yellow, then red. Yields *Baphic* acid on boiling (see over). Baphiin is neutral, insoluble in water, difficultly soluble in benzene and carbon bisulphide, soluble in alcohol and ether.

With alkaline solution decomposition occurs, *Baphinitin* being formed.

A white precipitate is obtained on addition of neutral lead acetate to the alcoholic solution. Dry hydrochloric acid gas gives a red, changing to violet, then green.

BAPHIC ACID, from Baphiin (see p. 19), is a yellowish-white powder, soluble in alcohol and ether, but not in water.

§ 39. **BAROSMA** (Diosma) crenulata, B. serratifolia, and B. betulina (*Rutaceæ*). The leaves (Buchu leaves). Landerer, *Repert. Pharm.*, **34**, 63, and others.

BAROSMIN G. (Diosmin ; possibly identical with Hesperidin) ; crystallizes in microscopic needles ; M.P. 243° C. ; agreeable odour on burning. Dilute acids convert to glucose, and a substance melting at 126° to 130° C. Barosmin is soluble in alcohol (scarcely so if cold) and in ether, volatile oils, dilute acids and alkalies, the latter with yellow colour.

It does not reduce Fehling's solution.

Concentrated sulphuric acid dissolves to a yellow solution.

§ 40. **BEBEERU** bark (Nectandra rodiæi—*Lauraceæ*), Greenheart-tree, British Guiana ; Buxus sempervirens (Box—*Euphorbiaceæ*), the bark and leaves ; Botryopsis platyphylla ('Pareira root') and Cissampelos Pareira (*Menispermaceæ*). [Geshoff has found an alkaloid resembling Bebeerine in Hernandia sonora and H. ovigera—*Lauraceæ*]. Investigators : Fauré, Rodie, Maclaglan, etc.

BEBEERINE A. (Beberine, Buxine, Pelosine), $C_{18}H_{21}NO_3$ (Flückiger) ; crystalline or amorphous, salts amorphous. Dextro-rotatory. M.P. 150° (198° Dupuy). Becomes electric on rubbing. Alkaline reaction ; bitter. Used as tonic—dose, 1 to 10 grains of the Sulphate—is also antiseptic. Very difficultly soluble in water (6,000 parts cold, 1,500 boiling) ; dissolves in 5 parts absolute alcohol, 13 parts ether, also in amyl alcohol, chloroform, acetone, benzene, and carbon bisulphide.

Precipitants :

Alkaline hydrates (sol. in excess).
Alkaline carbonates.
Ammonia (*not* by neutral Lead acetate).
Tannic acid, yellowish-white (sol. in warm hydrochloric acid).
Picric acid, yellow, amorphous at 1 in 3,000.
Platinum chloride, yellow, insoluble in HCl.
Gold chloride, yellowish-white.
Ferro- and ferri-cyanides of potassium, yellow.
Potassium sulphocyanide, white.
Potassium bichromate, an immediate light yellow precipitate, flocculent at 1 in 3,000.
Phospho-molybdic in acid solution, pp. dissolves blue in ammonia.
Iodo-potassic iodide, a Kermes-coloured pp.
Bismutho-potassic iodide, orange-red.
Cadmium-potassic iodide.
Mercuric-potassic iodide (Mayer's reagent), white.
Mercuric chloride, white, dissolved by HCl or by ammonium chloride.
Iridium-sodium chloride, red.

White pp. by sodium phosphate, nitre, nitric acid and platino-potassic cyanide.

Colour tests : Concentrated Sulphuric acid, dirty olive-green, becoming lighter in 15 to 20 hours ; Fröhde's reagent (concentrated sulphuric with molybdic acid), brownish-green, lighter after half an hour, yellowish in 24 hours ; nitric acid, brown.

PARA-BEBEERINE A., $C_{24}H_8N_2O$; the sulphate soluble in hot water but insoluble in alcohol.

PARA-BUXINE A. (Para-bebeerine ?) ; red, amorphous.

Soluble in water and alcohol ; **insoluble** in ether. Possesses acid properties.

Nitric acid produces a permanent green.

SIPIRINE A. ; resinous, brownish-red ; soluble in alcohol (Maclaglan). The name is derived from 'Sipiri,' the Dutch name for the Greenheart-tree. Flückiger doubts the individuality of this alkaloid.

NECTANDRINE A. (regarded by v. Planta as identical with Bebeerine), $C_{20}H_{23}NO_4$ (Maclaglan) ; amorphous powder ; M.P. below 100° (melts under hot water) ; bitter.

Insoluble in water ; dissolves in 250 parts of ether. **Soluble** in chloroform.

Concentrated sulphuric acid, with a little manganese dioxide, gives a magnificent green, becoming violet like the strychnine reaction.

§ 41. **BERBERIS** group. The alkaloid Berberine has been found in members of the *Berberideæ*, *Cassieæ*, *Menispermeæ*, *Papaveraceæ*, *Ranunculaceæ*, *Rutaceæ*, etc. ; notably in the families Coptis, Coccinium, Cocculus, Cœlocline, Geoffroya, and in particular the following plants : Berberis vulgaris (the alkaloids (*a*), (*b*), (*d*) below) ; Coptis Teeta ('Mahmira' —an Indian Ranunculacea—8½ per cent. Berberine, also Coptine) ;

Hydrastis Canadensis, the 'Golden Seal' (the alkaloids (*a*),(*f*),(*g*), including 4 per cent. of Berberine); Cocculus palmatus, 'Columbo root' (see also Calumba); *X*anthoxylum Clava Herculis, Podophyllum peltatum (see also Podophyllum), etc.

[For Artarine resembling Berberine, see Xanthoxylon.]

(*a*) *BERBERINE* **A.** (Xanthopicrit), $C_{20}H_{17}NO_4$, and with $5H_2O$; crystallizes in yellow prisms or needles; M.P. 120° C. (Fleitman), or browns at 110° and blackens at 160° (Perkin); faint quinone odour on warming; neutral to litmus; bitter; feebly toxic to man, poisonous to dogs and other animals. Medicinal dose, 2 to 5 grains.

Solubility: It dissolves in 300 parts of cold water, in alcohol, in benzene with difficulty. **Insoluble** in ether and petroleum ether; chloroform removes traces from *acid* solution. Boiling alkalies convert to a resinous substance, and ammonia dissolves to a reddish-brown solution. The salts (excepting acetate and pyrophosphate) are difficultly soluble.

Precipitants, etc.:

Tannic acid, cloud at 1 in 3,000, increased by hydrochloric acid and dissolved on warming.
Picric acid, pp. amorphous, becoming crystalline at 1 in 3,000.
Platinum chloride.
Gold chloride, immediate orange pp. at 1 in 3,000; part soluble in cold HCl.
Ferro-cyanide potassium, yellow.
Potassio-argentic cyanide, amorphous at 1 in 6,000.
Potassium bichromate, amorphous.
Iodo-potassic iodide, kermes colour (or green plates in warm alcoholic solution.
Bismutho-potassic iodide, orange-red.
Cadmium-potassic iodide, complete precipitation.
Zinc-potassic iodide, amorphous at 1 in 6,000.
Mayer's reagent (mercuric-potassic iodide).
Mercuric chloride, amorphous, insoluble in HCl even warm.
Bromine water, a yellow pp.
Iodine, green crystals if excess of iodine be avoided.
Tincture of iodine, yellowish-brown crystalline pp.

Colour Reactions:

Chlorine water, poured gently on to surface of Beberine solution, gives a red ring distinguishable even at 1 in 250,000.
Concentrated sulphuric acid, yellow (olive-green, Muter).
Oxidizers, such as manganese dioxide, when added to the sulphuric acid solution, give colours resembling those with strychnine.
Potassic nitrate, added after 10 to 15 hours solution in sulphuric acid, changes the olive-green solution to dark brown or orange.
Nitric acid, brown.
Concentrated hydrochloric acid, insoluble.
Fröhde's reagent, brownish-green, becoming gradually lighter.

(*b*) *OXYACANTHINE* **A.** (Vinetine, from *Vinétier*, the French name for Berberis vulg.; *not* the Oxyacanthine from Cratægus oxyacanthus), $C_{18}H_{19}NO_3$, or $C_{19}H_{21}NO_3$ (Rüdel); crystalline; dextro-rotatory $[a]_D = +131{\cdot}6$ for chloroform solution; M.P. of the dried ammonia pp. = 138°-150°, or the crystals from alcohol 208°-214°; alkaline reaction; bitter.

Solubility: Water dissolves it with difficulty; alcohol, 30 parts cold, 4 boiling; soluble in ether (which removes it from alkaline solution), slightly in petroleum ether; miscible with chloroform in all proportions.

Precipitants:

Alkaline hydrates and ammonia, soluble in excess.
[Not by neutral or basic lead acetate.]
Tannic acid, white.
Picric acid, yellow.
[No pp. ferric chloride.]
Platinum chloride, yellow, soluble in hydrochloric acid.
Gold chloride.
Silver nitrate, white.
[Not by copper sulphate.]
Phospho-molybdic acid, yellowish-white.
Mercuric chloride.
Bromine solution.
Iodine, reddish-brown.
Tartar emetic.
Stannous chloride.
[No pp. mercurous nitrate.]

Colour Reactions:

Concentrated sulphuric acid, reddish-brown (colourless to yellow—O. Hesse).
Nitric acid, brownish-yellow, becoming resinous with formation of oxalic acid.
Per-iodic acid, reduced.

Fröhde's reagent (molybdic and sulphuric acids), violet, becoming brownish-green at edges.

(*c*) *HYDROBEBERINE* A., $C_{21}H_{21}NO_4$; formed from Berberine by reduction; crystallized in needles or prisms.

Soluble in alcohol, chloroform, carbon bisulphide. **Insoluble** in water.

Concentrated sulphuric acid gives a yellowish-green.

Nitric acid oxidizes to Berberine. See Corydalis cava (Chelidonium group).

(*d*) *BERBAMINE* A., isomeric with Berberine (or lower homologue by CH_2, Rüdel).

Soluble in alcohol and ether; M.P. 156°.

Platinum chloride compound, white crystalline.

(*e*) *BETA-OXYACANTHINE* A., by heating Oxyacanthine in alkali.

Insoluble in ether. Hydrochloric acid gives a precipitate soluble in excess. **Soluble** in alkalies.

(*f*) *HYDRASTINE* A., $C_{21}H_{21}NO_6$ (Eyckmann). In constitution this alkaloid closely resembles Narcotine, differing only by a methoxy group, CH_3O. It crystallizes colourless, if pure; is lævo-rotatory, $[a]_D = -67{\cdot}8$ (Freund); M.P. 135°; has alkaline reaction, and is bitter when in solution (the solid alkaloid is tasteless).

Soluble in alcohol, ether (about 80 parts), chloroform (very easily), and in benzene; scarcely in water, and not in petroleum ether.

Precipitants:

Alkaline hydrates, white.
Ammonia, slightly soluble in excess.
Tannic and picric acids.
Platinum chloride, yellowish-red.
Gold chloride, similar pp.
Ferrocyanide of potassium, white.
Potassic iodide.
Iodo-potassic iodide, brown.
Mayer's reagent (mercuric-potassic iodide.
Mercuric chloride.
[Not by chlorine water.]

Colour Reactions, etc.:

Potassium permanganate, if not in excess, produces a fluorescent solution.
Concentrated sulphuric acid, a faint yellow with pure acid, orange to red when containing nitric acid.
Nitric acid alone, orange, the solution on dilution being fluorescent.
Fröhde's solution (molybdic in sulphuric acid), green, changing to brown.

(*g*) *XANTHOPUCCINE* A., orange-yellow crystals.

Soluble in hot alcohol. **Insoluble** in ether or chloroform.

Precipitants:

Ammonia.
Iodo-potassic iodide, brown.
Concentrated sulphuric acid, reddish-brown.
Nitric acid, brown.

(*h*) *CANADINE* A., $C_{21}H_{21}NO_4$ (Methyl berberine plus two atoms of hydrogen).

Soluble in ethyl acetate.

Precipitated by ammonia.

(*i*) *COPTINE* A., accompanies Berberine in Coptis trifoliata; colourless.

Dissolves in concentrated sulphuric acid without colour, but becoming purple on warming.

§ 42. **BOLDOA** (Peumus Boldoa = Boldo); *Monimiaceæ*; Chili.

BOLDINE A., alkaline, bitter powder.

Scarcely soluble in water, difficultly in benzene; dissolved by alcohol, ether and chloroform.

Precipitants:

Alkaline hydrates, soluble in excess.
Ammonia.
Iodo-potassic iodine.
Mercuric-potassic iodide.
Iodine tincture, brown.

Colour reactions:

Concentrated sulphuric acid, red.
Nitric acid, red.

§ 43. **BRAYERA** anthelmintica (Hagenia Abyssinica—'Cusso' or 'Kousso'); *Rosaceæ*; Abyssinia. Used as anthelmintic. The flowers.

KOSIN B., $C_{31}H_{38}O_{10}$; yellow rhombic crystals: M.P. 142° (194°?); acid reaction; bitter and sharp taste.

Scarcely soluble in water, but dissolved by alcohol, ether, chloroform, benzene, carbon bisulphide, glacial acetic acid, alkaline hydrates and carbonates.

Precipitated by neutral lead acetate.

Ferric chloride gives red coloration with the alcoholic solution.

Concentrated sulphuric acid produces a red amorphous substance, $C_{22}H_{26}O_{10}$.

§ 44. **BRYONIA** alba (*Cucurbitaceæ*). The root. Investigators: Walz and others.

BRYONIN G., $C_{48}H_{80}O_{19}$? (Walz); amorphous; very bitter; yields sugar and *Bryonetin*.

Soluble in water, alcohol (2 to 3 parts). **Insoluble** in ether.

Precipitants:

Tannic acid.

Platinum chloride.

[Not precipitated by basic acetate of lead.]

§ 45. **CAILCEDRA** bark (Swietenia, or Khaya Senegalensis); *Cedrelaceæ;* Africa. E. Caventou, *J. de Pharm.* (3) **16**, 355 and **33**, 123.

CAILCEDRIN B.; amorphous, brittle, resinous, bitter, neutral, fluorescent; M.P. 70°-80°.

Soluble in alcohol, ether and chloroform; with difficulty in water, to which, however, it imparts its fluorescence.

Tannic acid gives a precipitate soluble in excess.

Not precipitated by lead acetate (neutral or basic), ferric chloride, nor platinum chloride.

§ 46. **CALABAR** bean (Physostigma faba; P. venenosum); *Leguminosæ;* W. Africa.

PHYSOSTIGMINE A. (Eserine), $C_{15}H_{21}N_3O_2$; crystallizable only with great difficulty; lævo-rotatory; tasteless or slightly bitter; alkaline; M.P. 45°, decomposed in part at 100°, and even at lower temperatures, becoming red; extremely poisonous, producing death by asphyxia or cardiac paralysis; contracts pupil when locally applied.

Solubility: It is dissolved easily by alcohol, ether, chloroform, benzene, and carbon bisulphide; with difficulty by water, and not by petroleum ether.

Precipitants:

Alkaline hydrates and carbonates, oily pp., but only with conoentrated solutions; a colour effect is produced, however, namely, an intense red, changing successively to yellow. green and blue.

[No pp. lead acetate, neutral, nor by picric acid or platinum chloride in a solution of 1 in 250.]

Tannic acid, pp. reddish, 1 in 3,000, dissolved by HCl.

The free base precipitates iron as hydrate from a solution of ferric chloride.

Gold chloride, blue coloration, 1 in 2,000, and reduction of gold.

Phospho-molybdic and iodo-potassic iodide give precipitates—the latter kermes-coloured—1 in 25,000.

Cadmium-potassic iodide, yellowish-white, 1 in 1,000.

Mercuric-potassic iodide, 1 in 5,000; the pp. is soluble in a mixture of ether and alcohol, and has a melting point, 70°.

Mercuric chloride, reddish-white, soluble in HCl. [No pp. at 1 in 500.]

Chloride of lime (bleaching powder) and bromine solution both give red coloration, the latter distinguishable at 1 in 5,000.

Concentrated sulphuric acid, yellow, becoming orange-green (gradual red, Muter).

Sulphuric and nitric acids mixed, or nitric acid alone, yellow to red.

Concentrated hydrochloric acid also gives a reddish coloration.

The sulphate, neutralized with ammonia and warmed, gives the series of colours mentioned after alkaline hydrates above.

Barium hydrate produces the red substance *Rubreserine*, soluble in chloroform.

CALABARINE A. (obtained by Poehl, *Pharm. J., Russland*, **17**, 385, from the solution from which Physostigmine has been removed by shaking with ether, which does not dissolve Calabarine); known only in solution; it is more powerfully lævo-rotatory than Physostigmine, differs from the latter in producing tetanus, in being insoluble in ether (it is soluble in water and alcohol), and in the fact that the precipitate produced with mercuric-potassic iodide is insoluble in alcohol.

It is precipitated by phospho-tungstic acid.

[Not by lead acetate, neutral or basic.]

§ 47. **CALENDULA** officinalis (Marigold); *Compositæ*; the leaves and flowers. Investigator: Geiger, *Dissertn.*, Heidelberg, 1818.

CALENDULIN. Neutral principle, amorphous, transparent, yellow, tasteless.

Soluble in alcohol and glacial acetic acid, not in ether; swells up in water. It is dissolved by ammonia and alkaline hydrates.

The alcoholic solution is precipitated by:

Lead acetate, both neutral and basic, and by

Mercuric chloride.

[Not by tannin.]

§ 48. **CALOTROPIS** gigantea and C. procera (C. Mudarii)=Mudar

bark (*Asclepiadaceæ*); the root; used in India against leprosy: an emetic. Investigator: Duncan, *Phil. Mag.*, **10**, 465.

MUDARIN B.: amorphous, yellow, bitter.

Soluble in alcohol and cold water, gelatinizing on warming. **Insoluble** in ether and turpentine.

§ 49. **CALUMBA ROOT** (Cocculus palmatus, or Jateorhiza Calumba); *Menispermaceæ*; E. Africa. *For Berberine, which is contained in the drug, see Berberis group.*

CALUMBIN B. (Columbin), $C_{21}H_{22}O_7$ (discovered by Wittstock, see Pogg., *Annal.* **19**, 248; further investigated by Paterno, *Gazz. Chim.*); neutral, bitter; crystallizes from acetic acid in prisms.

Soluble with difficulty in water, alcohol, and ether (more easily in the latter two fluids when warm); also in chloroform and glacial acetic acid. It is removed from an *acid* solution when shaken with ether or chloroform, but not entirely by the last-named solvent. It is dissolved by alkaline hydrate solution; acids reprecipitate.

It is *not* precipitated by metallic salts, nor by tannin.

Concentrated sulphuric acid dissolves with yellow to red coloration.

CALUMBIC ACID B. (Columbic Acid), $C_{21}H_{22}O_6$ (Boedecker); pale yellow, amorphous, bitter; acid reaction.

Soluble in alcohol, acetic acid, dilute alkalies, and lime-water; soluble with difficulty in ether, and scarcely in water.

From its solution in alkalies it is precipitated by hydrochloric acid.

From alcoholic solution it is thrown down as a yellow pp. by lead acetate.

§ 50. **CALYCANTHUS** glaucus (*Calycanthaceæ*); the kernels of the seed; 3 per cent.

CALYCANTHINE A.; crystalline.

Insoluble in water or alcohol, but dissolved by ether. In consequence of the presence of fat in the seeds, it is liable to be taken up by petroleum ether.

Concentrated sulphuric acid gives a pale yellow, becoming brick-red with bichromate of potash, or purple to blue with sugar.

Nitric acid produces a bright green.

§ 51. [**CANTHARIDES** (Spanish Fly). Though not of vegetable origin, this is included on account of the active principle contained in it, viz.:

CANTHARIDIN, $C_{10}H_{12}O_4$ (Homolka); of acid nature; powerfully vesicant; sublimes at 180°, and volatile at lower temperatures in the vapours of alcohol or water.

Solubility in water very slight (common salt assists); soluble in 800 parts of alcohol (in which its salts are soluble with difficulty even warm), in 80 parts of chloroform (the best solvent), 900 of ether, 500 of benzene, 2,000 of carbon bisulphide, and in amyl alcohol. It may be extracted from an *acid* solution by shaking with amyl alcohol, chloroform, ether, or benzene.

Alkaline hydrates dissolve.

Neutral lead acetate precipitates, white crystalline.

Palladium chloride, yellow hairs.

Silver nitrate, white crystalline.

Copper and nickel sulphates, green crystalline precipitates isomorphous with the palladium salt.

Mercuric chloride, white crystalline.

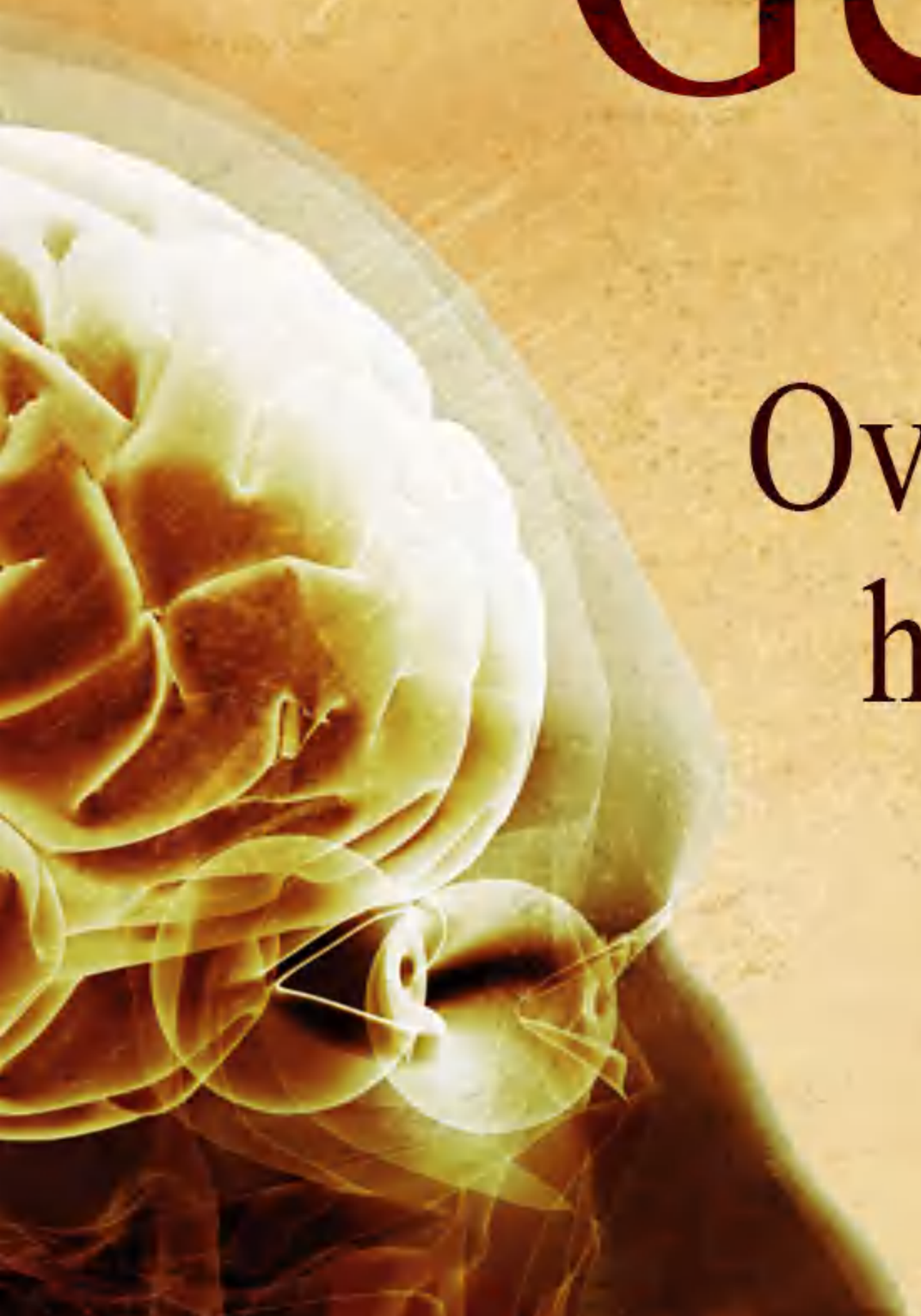
Alkaline permanganate, hydriodic acid, and sodium amalgam appear to be without action.

§ 52. **CAPSICUM** annuum (C. fastigiatum); *Solanaceæ*; the fruits = chillies or cayenne pepper when ground. Investigators: Buchheim, *Arch. Pathol.*, 1872; Thresh, *Ph. J. Trans.*, 1876; Braconnot, *Ann. Chim. Phys.* (2), **6**, 1, 124; and Flückiger.

CAPSICOL A. (of Buchheim)
CAPSAICIN (of Thresh)
CAPSICIN (of Braconnot)

These may be conveniently taken together, as their properties in many respects coincide; in the following statement, B. refers to Buchheim, T. to Thresh, Bc. to Braconnot, and F. to Flückiger. Formula, $C_9H_{14}NO_2$ (T., F.); strong odour, burning taste, (Bc.); M.P. 76° (T.), 58° (Bc.); volatile at 115° C., decomposed at 120° (T.); crystalline (T., F.); [reddish-brown liquid (B.), soft yellowish-brown, amorphous (Bc.)].

Soluble in alcohol and ether (B., Bc., T.), benzene (T., Bc.), petroleum ether (B., Bc.), with difficulty (T.); amyl alcohol (T., Bc.); with difficulty in water, easily in acetic ether and fixed oils; not readily in turpentine. It is dissolved by concentrated acetic acid, being precipitated therefrom by water.



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It reduces platinum chloride, copper sulphate, and Fehling's solutions.
No precipitates are given by tartar emetic, alkaloids, or gelatin (distinction from tannin).
Concentrated sulphuric acid, warm, dissolves with purple-red colour.
Nitric acid decomposes; red fumes are evolved, and oxalic acid is produced.

§ 57. **CEANOTHUS** Americanus L.; *Rhamnaceæ;* N. America. The dried leaves constitute 'New Jersey tea'; the root bark contains ½ per cent. of the following alkaloid. Investigator: F. Gerlach, *Amer. J. Pharm.*, 1891.

CEANOTHINE A. Formula has not been ascertained; it contains nitrogen and closely resembles caffeine; is crystalline, neutral, bitter, and melts at 190°. Forms salts only in acid solution.
Soluble in chloroform and ether.
Fehling's solution is reduced after boiling.

§ 58. **CEPHALANTHUS** occidentalis (Swamp Dogwood); *Rubiaceæ;* N. America. Investigators: Mohrberg and Claasen.

CEPHALANTHIN G., $C_{22}H_{34}O_6$; amorphous; dextro-rotatory; very bitter (1 in 25,000). On boiling with acids, *Cephalanthein*, $C_{16}H_{28}O_3$ (a crystalline substance), and sugar are produced.
Soluble in alcohol, amyl alcohol, and acetic ether; with difficulty in ether, chloroform and water (frothy solution).
Fehling's solution reduced after boiling.

CEPHALIN G.; crystalline, fluorescent (1 in 2,000,000). Yields sugar and *Cephaletin*, a crystalline substance giving fluorescent solutions with alkalies.
Cephalin is **soluble** in alcohol, ether and chloroform; very difficultly so in water.
Dissolved also by alkalies and ammonia.

§ 59. **CERBERA** Thevetia L. (Thevetia neriifolia, Juss); *Apocynaceæ;* the substances (*a*) and (*b*); investigators, De Vrij and Blas, *N. Jahresb. Ph.*, **31**, 1. Thevetia Ycotli, *Apocynaceæ*, the substance (*c*); Herrera, *Pharm. J. Trans.*, 1877. Cerbera Odollam, *Apocynaceæ*, the substances (*d*) and (*e*); Greshoff, 1891.

(*a*) *THEVETIN* G. (Cerberin? see below), $C_{54}H_{84}O_{24}$ (?); crystallizes in microscopic plates; lævo-rotatory: M.P. 170°; bitter, poisonous; decomposes with formation of Theveresin (see below) and sugar (?).
Soluble in cold water (122 parts), in alcohol and glacial acetic acid; not in ether.
Gives *no* precipitates with metallic salts.
Concentrated sulphuric acid, a reddish-brown, changing to cherry and then violet in a few hours; water destroys the colour.

(*b*) *THEVERESIN*, formed from above, is neutral, amorphous, resinous; M.P. 140°.
Soluble in alcohol, with difficulty in water, scarcely in ether, not in chloroform or benzene.
Alkalies give yellow solutions.
Concentrated sulphuric acid, the same colours as Thevetin.

(*c*) *THEVETOSIN* G.; crystallizes in four-sided prisms; poisonous; yields sugar and a resin.
Insoluble in water, slightly in ether.

(*d*) *ODOLLIN*, poisonous.
Soluble in water, alcohol, amyl alcohol. **Insoluble** in chloroform.
Not precipitated by neutral lead acetate.
Violet colour with concentrated sulphuric acid.

(*e*) *CERBERIN* B. (compare Thevetin above); crystalline; nitrogen free; neutral; bitter and burning taste, poisonous; non-glucosidal, but acids decompose; M.P. 165°.
Soluble in alcohol, chloroform, 80 per cent. ether, glacial acetic acid; not in water.
Concentrated sulphuric acid, violet colour (as with Odollin).

§ 60. **CHÆROPHYLLUM** bulbosum; *Umbelliferæ.* Investigator: Polstorff, *Arch. Pharm.* [2], **18**, 176.

CHÆROPHYLLINE A.; alkaline, volatile.
The sulphate is **soluble** in water, alcohol and ether.
Precipitated by

Tannic acid.	Picric acid.

§ 61. **CHAMÆLIRIUM** luteum; *Melanthaceæ* (*Colchicaceæ*); 10 per cent. of rhizome. Investigator: Greene, *Amer. J. Ph.*, **50**, 250. Chamælirin was also found by Nevinny in a specimen of supposed East Indian Ipecacuanha, but which he believes to have been derived from Helouins dioica (*Liliaceæ*).

CHAMÆLIRIN G.; amorphous, light yellowish-red; bitter.
Soluble in water (producing frothy solution), in alcohol and acetic acid;

with difficulty in ether. **Insoluble** in chloroform, benzene, petroleum ether, carbon bisulphide.

§ 62. **CHARA** fœtida, palmella, oscillaria, nostoc; *Characeæ*. Investigator: Phipson, *Pharm. J. T.*, **162**, 479.

CHARACIN. Non-crystalline, fat-like, volatile substance; 'marshy' odour; floats on water.

Soluble in ether and alcohol.

§ 63. **CHELIDONIUM** majus (Celandine), (*a*), (*b*), and Protopine, see Opium; Sanguinaria Canadensis (Puccoon, or Blood-root), (*a*). (*b*), and Protopine; Glaucium luteum, (*a*), (*e*), (*f*); Macleya cordata (Bocconia cordata, native of Japan), (*a*), and Protopine (Macleyine); Eschscholzia Californica, Morphine, (*a*) probably, and two others (Walz); Stylophoron diphyllum, (*b*). *Papaveraceæ*.

(*a*) *CHELERYTHRINE* A. (one of the two Macleya alkaloids; one of the Eschscholzia alkaloids; Sanguinarine? Schmidt gives $C_{21}H_{15}NO_4$ for Sanguinarine, and $C_{21}H_{17}NO_4$ for Chelerythrine, and states that the former gives red salts and the latter yellow; otherwise the two alkaloids closely resemble each other. This is confirmed by G. König. Commercial Sanguinarine was found by latter chemist to contain [in addition to Chelerythrine, which predominated] Protopine, Sanguinarine, and Homochelidonine); $C_{21}H_{17}NO_4$ (E. Schmidt), or $C_{19}H_{17}NO_4$ (Limpricht); crys. needles; alkaline reaction; liberates ammonia from its salts; bitter in alcoholic solution; fluorescent; optically inactive; on heating, it becomes resinous at 65°. Salts are orange-coloured, and mostly soluble in water, possessing burning taste, and giving red coloration with acid vapours.

Soluble in alcohol, ether, chloroform, amyl alcohol, benzene, petroleum ether, fatty and volatile oils; not in water.

Precipitants:

Alkaline hydrates, gray caseous.	Potassium chromate.
Ammonia, gray caseous.	Mercuric chloride, yellowish-white.
Tannic acid, yellowish-red	Iodine tincture.
Gold chloride, dark red (M.P. 233°).	Magnesia.

Concentrated sulphuric acid, yellowish-red.

(*b*) *CHELIDONINE* A. (Stylophorine, F. Selle), $C_{19}H_{17}NO_3$ or $C_{20}H_{19}NO_3$ (F. Selle); crystallizes in plates with $2H_2O$; M.P. 130°; volatile in steam, not poisonous in small doses, reaction alkaline, taste bitter.

Soluble in amyl alcohol, chloroform, volatile, and fatty oils; scarcely in alcohol or ether except after long boiling. **Insoluble** in water.

It is removed by chloroform from *alkaline* solution.

(*c*) *BETA-HOMO-CHELIDONINE*, $C_{21}H_{21}NO_5 = C_{19}H_{15}(OCH_3)_2NO_3$ (E. Schmidt); monoclinic crystals; M.P. 159°.

Precipitated in solutions of 1 in 100 by:

Phospho-tungstic acid.	Cadmium-potassic iodide.
Phospho-molybdic acid.	Bromine water.

Colour tests:

Concentrated sulphuric acid, violet.

Fröhde's reagent, yellow~violet~green.

(*d*) *ALPHA-HOMO-CHELIDONINE* A. (isomeric with preceding); $C_{19}H_5(OCH_3)_2NO_3$; crystalline; melting at 182°.

Precipitated, as preceding base.

Tannic acid also gives pp. soluble in excess.

Colour tests:

Concentrated sulphuric acid, yellow.

Fröhde's solution, dirty-brown green.

(*e*) *GLAUCOPICRINE* A.; amorphous or crystalline, giving crystalline salts (the hydrochloride, rhombic plates or prisms); alkaline, bitter.

Soluble in water and alcohol; with difficulty in ether.

Precipitants:

Alkaline hydrates.	[*Not* neutral lead acetate.]
Ammonia.	Tannic acid.

Colour test: Concentrated sulphuric acid, dark green on warming.

(*f*) *GLAUCINE* A., $C_{18}H_{19}NO_4$; pearly crystals; melting below 100° C.; reaction alkaline, taste sharp and bitter. Darkens on exposure.

Soluble in alcohol, ether, and petroleum ether; also in hot water, but scarcely in cold. The hydrochloride is insoluble in alcohol or ether; the sulphate is soluble in alcohol.

Precipitants:

Alkaline hydrates.	[*Not* neutral lead acetate.]
Ammonia.	Tannic acid.

Colour tests:

Concentrated sulphuric acid, blue~violet~red; after addition of water, ammonia gives a blue pp. (Battandier states that no violet colour is produced with pure unoxidized glaucine unless warmed.

Concentrated sulphuric acid with mercuric nitrate gives rise to intense green striæ on adding a crystal of glaucine.

§ 64. **CHENOPODIUM** album (White Goose-foot); *Chenopodiaceæ*. The juice. Investigator: Reinsch, *N. Jahrb. Pharm.*, 20, 268 and 27, 193.

CHENOPODINE A. (Leucine ? Dragendorff; v. Gorup), $C_6H_{13}NO_4$ ($C_6H_{13}NO_2$?); crystallizes in microscopic needles; M.P. 180°, with crystalline sublimate and noxious odour; neutral reaction; tasteless.

Solubility: 1 in 11 cold, and 1 in 3 to 4 boiling water; 1 in 202 cold, and 1 in 77 boiling alcohol of 90 per cent.

Precipitated by platinum chloride.

§ 65. **CHIOCOCCA** racemosa (Cainça root); *Rubiaceæ*.

(*a*) *CAINCIN* G. (Cainçaic acid, Chiococcin), $C_{40}H_{64}O_{18}$ (Rochleder); crystalline needles, acid reaction, taste gradually bitter. Action emetic and purgative. Salts amorphous.

Soluble in hot alcohol (crystals separate on cooling), difficultly soluble in water or ether.

Reactions:

Alkaline hydrates, Ammonia, Barium hydrate, Calcium hydrate	dissolve; excess of lime precipitates a basic salt.	[Lead acetate, neutral, no pp., or slight.]
		Lead acetate, basic, pp.
		[Ferric chloride, no pp.]

(*b*) *CHIOCOCCAIC* acid is formed together with sugar from the above by momentary heating with acid; and

(c) *CAINCETIN*, $C_{22}H_{24}O_3$, is produced when the heating is continued.

§ 66. **CHIRETTA** (Ophelia Chiretta, or Agathotes Chirayta); *Gentianeæ*. Investigator: Höhn, *Arch. Ph.*, 139, 213.

(*a*) *CHIRETTIN* B. (Chiratin), $C_{26}H_{48}O_{15}$; amorphous, resinous, neutral, very bitter.

Soluble in ether, which removes it from acid solutions. Chloroform does not extract it from either acid or alkaline solutions, or at most but traces.

Reactions:

[No pp. lead acetate, neutral or ammoniacal.]
Tanic acid, pp.
[Fehling's solution, not reduced.]

(*b*) *CHIRETTOGENIN*. From the preceding by long boiling with acids.

Not **precipitated** by tannic acid (compare above).

(*c*) *OPHELIC ACID* B., $C_{13}H_{20}O_{10}$; yellow, syrupy, acid reaction, taste at first sour, becoming bitter; gentian-like odour.

Soluble in water and alcohol, also in mixture of ether and alcohol.

Precipitants, etc.:

Fixed alkalies, Ammonia: dissolve; yellow.	Ammoniacal silver nitrate, reduced.
Lead acetate, neutral, pp.	Fehling's solution, reduced.
	Metallic salts, pp. (Höhn).

§ 67. **CHRYSANTHEMUM** tanacetum, (*a*), (*b*); C. Cinerariæfolium—the flowers, (*c*), (*d*); *Compositæ*. Investigators: Leppig, *Pharm. Z. Russland*, 1882; Marino Zuco, and others.

(*a*) *TANACETIN* B., C 61·46 per cent., H 7·7 per cent.; amorphous, bitter.

Soluble in water and alcohol, not in ether.

(*b*) *TANACETIC ACID*. The individuality of this substance is denied by Leppig.

(c) *CHRYSANTHEMINE* A., $C_{14}H_{28}N_2O_3$ (Marino Zuco); syrupy, but dried *in vacuo* appears in form of silky needles; M.P. above 100°: alkaline reaction, no rotatory power, physiologically inactive. Salts mostly soluble in water, some being deliquescent.

Soluble in water, alcohol, and methyl alcohol. **Insoluble** in ether, chloroform, or benzene.

Precipitants:

[*Not* alkalies in dilute solution.	[Not phospho-molybdic acid.]
„ lead acetate, neutral.	Bismuth-potassic iodide, orange flocculent, becoming crystalline.
„ „ basic.	
„ tannic acid.	
„ picric acid.	Mercuric-potassic iodide, yellowish-white.
„ platinum chloride.]	
Gold chloride, yellow crystalline, sol. hot.	[Not mercuric chloride.]
	Platinic-sodic iodide, brown.

Colour tests (negative):

Concentrated sulphuric acid, no action (slight resinification on heating.

Fuming hydrochloric acid, no action.

(*d*) *PYRETHROSIN* B., $C_{34}H_{44}O_{10}$; crystalline, bitter.

Soluble in hot alcohol and chloroform; less easily in ether and petroleum ether. **Insoluble** in water.

Concentrated hydrochloric acid, red or violet coloration.

§ 68. **CICHORIUM** intybus (Chicory); *Compositæ*. The flowers. Investigator: R. Nietzki, *Arch. Pharm.* [3] 8, 327.

CHICORIN G. (suggested name; R. Nietzki's 'Cichorium glucoside').

$C_{32}H_{34}O_{19}+4\frac{1}{2}H_2O$; crystalline, bitter ; M.P. 215°-220°. Converted by acids into glucose and a substance, $C_{20}H_{14}O_9$, occurring in glittering needles.

Soluble in alcohol and hot water (scarcely cold). **Insoluble** in ether.

Reactions :

Alkaline hydrates, carbonates } dissolve to yellow solution.
Lead acetate, gives pp. in absence of free acetic acid.
Ferric chloride, green coloration.
Silver nitrate, reduced.
Fehling's solution, reduced.
Concentrated nitric acid, red solution.

§ 69. **CINCHONA** group. *Rubiaceæ—Cinchoneæ.* A great number of varieties of bark, derived from some three dozen species of Cinchona, have from time to time been employed for the antifebrile properties which they owe to the alkaloids contained in them ; of these, Quinine is the most valuable, and though occasionally absent or only present in traces, usually forms a large proportion of the total alkaloid of the bark.

Next in importance are Cinchonine, Cinchonidine and Quinidine ; the latter, however, is found in but small quantity (generally in traces, and rarely to the extent of 1 per cent., except in C. pitayensis, etc., see below). Cinchonine and Cinchonidine, on the other hand, frequently exceed the Quinine in amount, or even replace it entirely.

Besides the four alkaloids already referred to, there is a long list of basic and other compounds of less general occurrence, or present in minute proportions ; sixty-six substances are here mentioned.

The following table must not be regarded as more than a rough guide ; records are to a great extent wanting as to the distribution of the rarer bases, hence a complete enumeration of the alkaloidal constituents of each species of Cinchona cannot be given, and as regards the more important alkaloids, much variation is met with in the quantities obtainable from a particular variety of bark.

(Qn. = *Quinine* ; Qnd. = *Quinidine* ; Cn. = *Cinchonine* ; Cnd. = *Cinchonidine.*)

Cinchona affinus, Wedd, see C. micrantha (Ruiz and Pavon).
„ Anglica, cross between C. calisaya and C. succirubra ; Qn. 0·81, Cnd. 1·49, Cn. 0·88, Qnd. 0·29, amorphous 0·44. Total, 3·91 per cent.*
„ Angustifolia, R. and P., see C. lancifolia.

* Analysis of Madras Barks, by D. Hooper.

Cinchona Australis, see C. nitida.
„ *CALISAYA* (Yellow bark, Quinquina jaune royal), very rich in quinine, as much as 11 per cent. having been found (Moens) ; also Cn., Cnd., Quinamine, etc.
„ cordifolia (Hard Columbian bark), Qn., Cn. ; poor in alkaloids (D. Hooper, 1888).
„ excelsa, see Hymenodictyon.
„ glandulifera, Qn. Unimportant.
„ Josephiana. Low percentage of Qn.
„ *LANCIFOLIA*, Mutis (Soft Columbian, Caqueta, Carthagène ligneux) ; composition very various ; Qn. sometimes absent, but frequently abundant.
„ *LEDGERIANA* (Ledger bark), the richest probably of all ; Qn. 5·48, Cn. 1·33, Cnd. 0·82, amorph. 0·88—total, 8·52.* Larger quantities than these have been obtained—also Quinamine.
„ micrantha (Huanoco gray) ; no Qn., Cnd. 1·92, amorph. 0·4—total, 2·32.*
„ morada. See Pogonopus, below.
„ negra. Qn. 5·48, Cn. 0·1, Cnd. 0·00, Qnd. trace, amorphous 0·78—total, 6·26.†
„ nitida, R. and P. (C. scrobiculata, H. Baillon ; C. Australis, Wedd ; Cascarilla Colorada, red Cusco bark, C. Peruviana, Calisaya fibrosa : see Flückiger) ; poor in alkaloid. Quinamine has been isolated amongst other bases.
„ *OFFICINALIS* (Pale or Crown bark ; the chief varieties are Condaminea, Bonplandiana, and Crispa ; also Chahuarguera and Uritusinga ; Quinquina gris de Loxa is obtained from C. officinalis var. Crispa : see Baillon). Rich in Qn. ; Qn. 4·74, Cnd. 1·23, Cn. 0·10, Qnd. 0·07, amorphous 0·42—total, 6·56.†
„ pelletierana ; aricine, cusconine, etc. See No. 8 and onwards.
„ Peruviana. See C. Nitida.
„ pitayensis, Wedd (Pitayo bark) ; rich in Qn. and Qnd.
„ pombiana ; Qn. 4·41, Cnd. 0·34, Cn. 0·02, Qnd. trace, amorph. 0·26—total, 5·03.†
„ pubescens ; poor in alkaloids (Flückiger) or even devoid of them (Hesse, 1871).

* Analysis of Madras Barks, by D. Hooper.
† Analysis of New Grenada Barks, by D. Howard.

Cinchona robusta ; a hybrid.
„ rosolenta ; Cnd., Quinamine, Bicinchonine, Homo-cinchonidine, etc.
„ rubra ; red bark. See also C. succirubra.
„ scrobiculata. See Nitida.
„ *SUCCIRUBRA* (Red bark). More Cnd. than Qn., usually about 3 to 5 per cent. of former to 1½ per cent. of latter (Dr. Paul) ; also Cn. and Quinamine—more of the last named in this variety than in others. Qnd. only trace, or absent.
„ tuna. Qn. 6·78, Cnd. 0·40, Cn. 0·38, Qnd. 0·18, amorphous 0·42—total, 8·16 per cent.*
Pogonopus febrifuga, or Howardia (Bolivian Cascarilla, C. morada). Qnd. (no Cnd.), Moradeine, etc. See (32), (33).
Remijia pedunculata (Cuprea bark). Qn. (1), Cupreine (21), Homoquinine 22. Quinamine, etc. (18), (20), Concusconine, etc. (10), (11), Hydrocinchonine.
„ purdieana (Remijia bark). No Quinine nor Cinchonidine ; traces Cn., also Cinchonamine, Concusconine, Chairamine and analogues (14 to 17), Cinchotine (4 per cent.).

Croton eluteria (Cascarilla bark), *Euphorbiaceæ*. Cascarillin (28). For Bolivian cascarilla, see Pogonopus.
„ pseudochina. Copalchin (29).
Cascarilla hexandra. Paricine (25).
„ riedeliana (China Californica). Californin (31).

The following is a list of the alkaloids and other substances here described, and the order in which they appear.

(1) *QUININE.*
(2*a*) Isoquinine.
(2*b*) Quinicine.
(2*c*) Apoquinine.
(2*d*) Quitenine.
(2*e*) Hydroquinine.
(3) *CINCHONINE.*
(4*a*) Isocinchonine.
(4*b*) Cinchonetine.
(4*c*) Apocinchonine.
(4*d*) Apocinchonicine.
(4*e*) Diapocinchonine.
(4*f*) Hydrocinchonine.
(4*g*) Cincholeuponic acid.
(4*h*) Cinchonibine.
(4*i*) Cinchonifine.
(4*j*) Cinchonigine.
(4*k*) Cinchoniline.
(4*l*) Alpha-oxy-cinchonine.
(4*m*) Beta-oxy-cinchonine.
(4*n*) Cinchonicine.
(4*o*) Carthagine.
(4*p*) Cinchotine.
(4*q*) Dicinchonine.
(5) *QUINIDINE.*
(6*a*) Apoquinidine.
(6*b*) Hydroquinidine.
(6*c*) Quitenidine.
(6*d*) Isoquinidine.
(7) *CINCHONIDINE.*
(7*a*) Isocinchonidine.
(7*b*) Apocinchonidine.
(7*c*) Cinchamidine.
(7*d*) Homocinchonidine.
(8) Aricine.
(9) Cusconine.
(9*a*) Cusconidine.
(9*b*) Cuscamidine.
(10) Concusconine.
(11) Concusconidine.
(12) Cuscamine.
(13) Chairamine.
(14) Conchairamine.
(15) Chairamidine.
(16) Conchairamidine.
(17) Quinamine.
(18*a*) Apoquinamine.
(18*b*) Quinamicine.
(18*c*) Quinamidine.
(19) Conquinamine.
(20) Cupreine.
(21) Homoquinine.
(22) Cinchonamine.
(23) Quinoidine.
(24) Quinetum.
(25) Paricine.
(26) Quinovin.
(26*a*) Beta-quinovin.
(27) Quinovic acid.
(28) Cascarillin.
(29) Copalchin.
(30) Lignoin.
(31) Californin.
(32) Moradeine
(33) Moradin.
(34) Javanine.

* Analysis of New Grenada Barks, by D. Howard.

(1) *QUININE* A., $C_{20}H_{24}N_2O_2+1$ or $3H_2O$; usually amorphous, but crystalline needles, containing $3H_2O$, are gradually formed from the precipitate by ammonia, and crystals are also obtainable from solutions in petroleum ether and benzene, but not from ether. M.P. 171°-172° when anhydrous, 120° hydrate, or 57° the trihydrate (O. Hesse) ; lævo-rotatory, $[a]_R = -141·3$ in alcohol (De Vrij) ; very bitter ; fluorescent in acid solution ; alkaline reaction.

Solubility : 1 in 1,667 cold water and 1 in 902 boiling (Sestini), or 1 in 364 cold and 1 in 267 boiling (Duplos) ; 1 in 2 alcohol of sp. gr. 0·82, and even easier soluble at boiling heat ; 1 in 23 cold ether (Vandenburg), or 1 in 60 (Merck), but much more readily when freshly precipitated ; very soluble in chloroform, about 1 in 1·8 (Pettenkofer) ; also dissolved by benzene, amyl alcohol, petroleum ether, carbon bisulphide, and oils. It is removed from *alkaline* solutions by petroleum ether, benzene, ether, chloroform, and amyl alcohol.

The salts are generally less soluble than those of Cinchonine.

Precipitants:
Alkaline hydrates.
„ carbonates.
Ammonia (somewhat soluble, and also in amm. chloride).
„ carbonate after some hours.
Alkaline bicarbonate if not dilute (no pp. at 1 in 200).
Mem.: Tartaric acid prevents the precipitation by bicarbonates, for a time.
[No pp. sodium acetate, compare Opium.]
[Not lead acetate, neutral or basic.]
Tannic acid, yellowish-white; soluble in warm hydrochloric acid.
Picric acid, amorphous.
[Not ferric chloride.]
Platinum chloride, light yellow, nearly white; insol. in hydrochloric acid.
Gold chloride, yellow amorphous.
Potassium ferrocyanide; sol. warm, and in excess.
„ sulphocyanide, white changing to needles.
(„ cyanide, red coloration.)
Silver-potassic cyanide, white amorphous.
Potassium bichromate, yellow amorphous (cloud at 1 in 3,000).
Chromic acid (5 per cent.), a pp. with neutral salts of Q.
Phospho-molybdic acid, white amorphous (caustic potash dissolves the pp. to light yellow solution, see Morphine).
Phospho-tungstic acid, 1 in 100,000.
Phospho-antimonic acid, pp. (cloud at 1 in 5,000).
Iodo-potassic iodide, reddish-brown.
Potassic iodide.
Bismuth-potassic iodide, orange-red (cloud at 1 in 50,000).
Zinc-potassic iodide.
Mercuric-potassic iodide, amorphous: limit 1 in 125,000.
Mercuric chloride, white amorphous; sol. in amm. chloride.
Chlorine water in not too great excess, then ammonia; green flocks soluble in excess to an *emerald solution*, (thalleioquin reaction); by careful neutralization: blue⁓violet⁓red; becoming again green with excess of ammonia. (Potash to the chlorine solution, yellow; lime-water, red). Test applicable 1 in 20,000.
Bromine water, as with chlorine.

Other **Colour reactions**, negative.
Concentrated sulphuric acid
„ „ „ with sugar } colourless.
Concentrated sulphuric acid with potas. bichromate, merely light yellow.
Per-iodic acid, iodine liberated.
Fröhde's reagent, greenish.
Nitric acid alone
„ „ with sulphuric } colourless.

(2*a*) *ISOQUININE.* From Quinine by solution in strong sulphuric acid.
Not precipitated by sodium tartrate.

(2*b*) *QUINICINE*, isomer of Quinine; formed by heating acid sulphate of quinine to 135°, and occurring in 'Quinoidine,' see (24). Yellowish amorphous, bitter, alkaline, dextro-rotatory; M.P. about 60°, non-fluorescent.
Soluble in alcohol, chloroform, ether; but with difficulty in water.
Reactions:
Absorbs carbonic acid.
Alkaline hydrates } oily pp. incomplete; only cloud in presence of
Ammonia } ammonium chloride.
[No pp. tartrate.]
Potassium sulpho-cyanide, oily pp.
Chlorine water then ammonia, as with quinine, but less intense.

(2*c*) *APOQUININE* A., $C_{19}H_{22}N_2O_2+2H_2O$; from Quinine by heating with hydrochloric acid in sealed tube to 140°, whereby CH_3 is eliminated (escaping as methyl chloride). Crystalline, lævo-rotatory, alkaline, bitter, non-fluorescent; M.P. 160°.
Soluble in alcohol, ether, chloroform, and hot water; difficultly in the latter when cold.
Precipitated by alkaline hydrates and ammonia, but soluble in excess.
Chlorine water then ammonia (thalleioquin reaction), green coloration, but only faintly.

(2*d*) *QUITENINE*=Dihydroxiquinine. From Quinine by oxidation with potassium permanganate.

(2*e*) *HYDROQUININE* A., $C_{20}H_{26}N_2O_2$, and with $2H_2O$. Accompanies Quinine (in the Cinchona barks, and up to 4 per cent. in commercial quinine sulphate; separable from Quinine by fractional precipitation of the acid sulphates). Crystalline, lævo-rotatory (−142·2°), alkaline, bitter; M.P. 168°, fluorescent in sulphuric acid solution.
Soluble in alcohol, ether, chloroform, benzene, carbon bisulphide, aqueous acetone and ammonia; slightly in water.

Precipitated by sodium hydrate, amorphous becoming crystalline.
Chlorine water, etc. (thalleioquin reaction), as quinine.
Potassium permanganate, only slowly decolorized.

(3) *CINCHONINE* A., $C_{19}H_{22}N_2O$ (Laurent, etc.). Crystallizes in needles or prisms (no crystalline hydrate). Dextro-rotatory ; $[a]_R = 237{\cdot}5$ in alcoholic solution. Melting-point has been variously given as 140°, 165°, 250° by different observers ; purple fumes are obtained on heating ; reaction alkaline, taste bitter, non-fluorescent.

Solubility : 1 in 3,800 cold, 1 in 2,500 boiling water ; 1 in 140 alcohol ; 1 in 40 chloroform ; 1 in 371 *hot* ether ; soluble also in hot benzene, in amyl alcohol, and in olive oil (1 in 100) ; but scarcely in cold ether, cold benzene, or petroleum ether, and with difficulty in essential oils.

It is removed from both alkaline and acid solutions by chloroform ; in traces only by ether or benzene from alkaline solution ; not at all by petroleum ether.

Precipitants :

Alkaline hydrates } soluble in excess (Dragendorff) ; very slightly (Hüsemann).
Ammonia } the amm. pp. is amorphous, but becomes crystalline.
Alkaline carbonates.
„ bicarbonates.
Tartaric acid prevents the precipitation by bicarbonate until boiled.
[Not by lead acetate, neutral or basic.]
Potassium ferrocyanide, neutral solution, yellowish-white, insol. on warming or in excess (Dupuy) ; no pp. 1 in 500.
Potassium sulphocyanide, white flocculent becoming crystalline ; sol. alcohol ; no pp. 1 in 500.
Silver-potassic cyanide.
Potassium bichromate, pp. (gradual at 1 in 3,000, Dragendorff).
Chromic acid (5 per cent. solution to the neutral salt), pp.
Phospho-molybdic acid, pp ; cloud at 1 in 200,000, slight opalescence at 1 in 500,000.
Phospho-antimonic acid, pp. ; bluish-white at 1 in 1,000 ; cloud at 1 in 5,000.
Iodo-potassic iodide, kermes pp. up to 1 in 500,000.
Bismuth-potassic iodide, orange-red—limit as phospho-molybdic acid.
Tannic acid, yellowish-white ; sol. in hydrochloric acid on warming ; cloud at 1 in 40,000.
Picric acid, yellow becoming crystalline (sol. in excess, Dragendorff) ; reaction distinct at 1 in 100,000 ; feeble cloud at 1 in 200,000.
Platinic chloride, lemon pp. ; insol. hydrochloric acid, 1 in 500.
Gold chloride, lemon pp. ; distinct cloud 1 in 100,000, feeble at 1 in 200,000.
Cadmium potassic iodide, hair-like crystals up to 1 in 50,000 ; cloud at 1 in 100,000.
[Zinc potassic iodide, no pp., or very slight.]
Mercuric-potassic iodide, pp. ; cloud even at 1 in 600,000 (? limit 1 in 75,000, Dragendorff).
Mercuric chloride, white amorphous ; feeble at 1 in 10,000. Red compound formed on heating concentrated solutions.
Chlorine water, no coloration ; on addition of ammonia, white pp.

Sundry tests :

Potassium persulphide, gradual white pp. on heating.
Cadmium chloride, white crystalline pp.
Heated with tartaric acid, vapours yellow~violet.
Many oxidizing agents give red coloration.

Colour tests (negative) :

Concentrated sulphuric acid } with sugar
„ „ „ } with nitric acid
„ „ „ nitric acid
Acids in general
Fröhde's reagent.
} colourless.
Per-iodic acid, iodine liberated.

(4*a*) *ISOCINCHONINE* A. (see (4*j*), Cinchonigine), $C_{19}H_{22}N_2O$. Formed like the corresponding Isoquinine. Alkaline, lævo-rotatory ; M.P. 125°, volatile in steam.

Soluble easily in alcohol, ether, acetone, benzene, and chloroform. **Insoluble** in water or alkalies.

(4*b*) *CINCHONETINE* A. From Cinchonine by oxidation with potassium permanganate, or lead peroxide and dilute sulphuric acid. A dark violet substance.

(4*c*) *APOCINCHONINE* A., $C_{19}H_{22}N_2O$. Prepared like the corresponding Apoquinine, but no methyl group is eliminated as in the formation of the latter. Crystallizes in anhydrous prisms ; M.P. 209° ; dextro-rotatory (+160° in alcohol) ; alkaline, bitter, non-fluorescent.

Soluble in alcohol, difficultly in ether or chloroform, and not in water.

(hydrate), or 168° (anhydrous). The acid salts show a blue fluorescence, but the sulphate in chloroform a green.

Solubility in water 1 in 2,000 cold, 1 in 750 boiling; 1 in 26 cold alcohol of 80 per cent.; 1 in 3·7 boiling absolute alcohol; in ether, 1 in 30 at 10°, 1 in 22 at 20° (the hydrochloride 1 in 325); also soluble in amyl alcohol and benzene; very difficultly in chloroform, carbon bisulphide, and petroleum ether. It is removed from alkaline solutions by benzene, chloroform (with difficulty), and amyl alcohol.

Precipitants:

No pp. sodium bicarbonate cold, but pp. on heating.
Tannic acid, pp. sol. warm hydrochloric acid.
Picric acid, amorphous pp. becoming crystalline.
Platinum chloride, pale yellow; insol. hydrochloric acid.
Gold chloride, lemon pp., amorphous.
Silver nitrate.
Potassium ferrocyanide, yellowish white crys.; sparingly soluble.
Silver-potassic cyanide.
Potassium bichromate, yellow amorphous; cloud at 1 in 3,000.
Chromic acid (5 per cent., to neutral salt).
Iodo-potassic iodide, kermes colour.
Potassium iodide, white amorphous; *distinctive from all Cinchona alkaloids.*
The hydriodide is difficultly soluble in water or alcohol.
Cadmium potassic iodide, pp. complete.
Zinc-potassic iodide.
Mercuric-potassic iodide, limit 1 in 50,000.
Mercuric chloride, white amorphous.
Chlorine water, then ammonia, as Quinine.

Colour tests (negative, identical with those of Quinine):

Concentrated sulphuric acid
" " " with sugar
" " " " nitric acid
" nitric acid
" hydrochloric acid
} colourless.

Fröhde's solution, greenish.

(6*a*) *APOQUINIDINE* A. (Apoconquinine), $C_{19}H_{22}N_2O_2$; formed from Quinidine like the corresponding Apoquinine from Quinine, a methyl group being eliminated; amorphous, alkaline, non-fluorescent; M.P. (anhydrous alkaloid) 137°.

Soluble in alcohol and ether.

Precipitated by alkaline hydrates and ammonia, the pp. being difficultly soluble in excess.

Gives the thalleioquin reaction, but not so strikingly as does Quinine.

(6*b*) *HYDROQUINIDINE* A. (Hydroconquinine); $C_{20}H_{26}N_2O_2 + 2\frac{1}{2}H_2O$; found in commercial Quinidine; crystallizes in prismatic needles or plates; M.P. 160°-167°; dextro-rotatory; fluorescent (in sulphuric acid).

Soluble readily in alcohol and chloroform, with difficulty in ether.

Gives the thalleioquin reaction.

(6*c*) *QUITENIDINE* A., $C_{19}H_{22}N_2O_4$; from Quinidine by oxidation with potassium permanganate.

(6*d*) *ISOQUINIDINE* (Isoconquinine); formed like Isoquinine; crystalline needles.

Soluble in ether.

(7) *CINCHONIDINE* A. (Cinchovatine; Winckler and Hesse's Quinidine; the Homocinchonidine of Hesse, described at (7*d*), is impure Cinchonidine (Skraup), or Beta-cinchonine); $C_{19}H_{22}N_2O$ (Pasteur); large glittering prisms from alcohol; no crystalline hydrate; M.P. 175°-206° (observers differ); alkaline, bitter; not fluorescent.

Soluble in 1,680 parts of water at 10°, 19 alcohol of 80 per cent., 76 parts ether (Leers' results differ from these); it is also soluble in chloroform, by which it is removed from alkaline solutions.

Precipitants ('No characteristic test,' Hüsemann):

Alkaline hydrates
Ammonia
} scarcely sol. in excess.
Alkaline carbonates.
" bicarbonates.
Rochelle salt.
[Not neutral sodium acetate (see Opium).]
[Not ferric chloride.]
Platinum chloride, pale orange, scarcely soluble.
Gold chloride, yellow; M.P. about 100°.
Potassium ferrocyanide, reddish-yellow crystals, sparingly soluble.
Silver potassic cyanide.
Chromic acid, 5 per cent. solution to neutral salt.
[Not zinc potassic-iodide, or very slightly.]
Mercuric chloride, crystalline, difficultly soluble.
No thalleioquin reaction.

Colour tests (negative):
Concentrated sulphuric acid
„ „ „ with sugar
„ „ „ „ nitric acid } colourless.
„ nitric acid
„ hydrochloric acid

(7*a*) *ISOCINCHONIDINE* A.; prepared like Isoquinine; crystalline plates; **M.P.** 235°.

Readily **soluble** in ether and chloroform.

(7*b*) *APOCINCHONIDINE* A.; formed like Apoquinine, but no methyl group eliminated; crystalline, very bitter; non-fluorescent.

Precipitated by alkaline hydrates, and by ammonia, which gives an amorphous pp., becoming crystalline.

No pp. Rochelle salt.

(7*c*) *CINCHAMIDINE* A. (Hydrocinchonidine, Forst); accompanies Cinchonidine, from which it may be separated by fractional precipitation; $C_{20}H_{26}N_2O$ or $C_{19}H_{24}N_2O$; crystalline plates or needles; **M.P.** 230°; lævo-rotatory; not fluorescent.

Soluble in alcohol and chloroform; with difficulty in ether.

(7*d*) *HOMOCINCHONIDINE* A. [regarded by Skraup as merely impure Cinchonidine, see (7)]; $C_{19}H_{22}N_2O$ (Hesse); accompanies Cinchonidine; crystalline prisms or plates.

Soluble in alcohol and chloroform, scarcely in water.

(8) *ARICINE* A. (Manzini's Chinovatine or Quinovatine; Hesse's Cinchonidine?); exists to extent of 3 to 3½ per cent. in China de Cusco vera (false Calisaya bark), in which Moissan found no Quinine or Cinchonine; $C_{23}H_{26}N_2O_4+2H_2O$ (Gerhardt, Hesse, Moissan); isomeric with Brucine; crystallizes in long transparent needles; M.P. 188°; lævo-rotatory (−58·18° in alcohol, −92·30° in ether); alkaline, not bitter; salts mostly crystalline (but the sulphate gelatinous from water, though crys. from alcohol); acetate difficultly soluble in water.

Soluble in 100 parts cold or 11 parts boiling alcohol, 33 parts ether; scarcely in water.

Precipitants:
Alkaline hydrates.
„ carbonates.
Ammonia, somewhat soluble in excess.
[Not tartrates of soda or potash.]
Tannic acid.
Picric acid.
Gold chloride, yellow amorphous.
Phospho-tungstic acid } pp., cloud at 1 in 50,000.
Iodo-potassic acid }
Potassic iodide.
Mercuric-potassic iodide, pp., cloud at 1 in 50,000.

Colour tests (Wittstein believes the colours are due to impurity):
Concentrated sulphuric acid, dark green.
„ nitric acid, intense green.
Fröhde's reagent, blue, changing to green.

(9) *CUSCONINE* A. (regarded by Wittstein as identical with preceding), $C_{23}H_{26}N_2O_4+2H_2O$; crystalline plates or prisms, with metallic lustre; M.P. 110° (after loss of water); lævo-rotatory (−54·3 in alcohol); feebly alkaline; not bitter; not fluorescent; salts mostly amorphous; sulphate gelatinous.

Soluble in alcohol, acetone, and ether, very readily in chloroform, but scarcely in water.

Precipitated by alkaline hydrates and ammonia; scarcely soluble in excess.

Other precipitants as under Aricine.

Colour tests:
Concentrated sulphuric acid, yellowish-green.
Concentrated nitric acid colours the alkaloid dark green, the solution becoming yellowish-green.
Fröhde's solution, dark-blue ~ olive green warm ~ blue on cooling.

(9*a*) *CUSCONIDINE* A. [possibly amorphous form of Cusconine (Hesse); compare also (8) and (9)]; amorphous; dark brown.

Reactions:
Alkaline hydrates } dirty yellow flocculent pp.
Ammonia }
Gold chloride } yellow amorphous pp., soluble with difficulty in
Platinum chloride } in water.
To the solution in concentrated acetic acid:
Sulphuric acid gives no pp. (distinction from Paricine).
Nitric acid, cloud, changing to pp. of Cusconidine nitrate.
Hydrochloric acid, cloud, disappearing on addition of water.

(9*b*) *CUSCAMIDINE* A.; closely resembles the above, the only difference (observed by O. Hesse, *Ann. Chem. Pharm.*, 200) being that Cuscamidine

gives a precipitate with nitric acid from a *dilute* solution of acetic acid, whereas Cusconidine is only precipitated by nitric acid from a *concentrated* acetic solution.

(10) *CONCUSCONINE* A., $C_{23}H_{26}N_2O_4 + H_2O$; monoclinic prisms; M.P. 144°, solidifying on further heating, melting again at 206°-8°; dextro-rotatory; neutral reaction; tasteless (salts are bitter).

Soluble readily in ether, chloroform, or benzene; very difficultly in alcohol even boiling; scarcely in petroleum ether, and not in water.

Precipitants:

Alkaline hydrates, resinous.
[Not ammonia, nor lead acetate, neutral.]
Platinum chloride, yellow flocculent.
Gold chloride, dirty yellow, with reduction of gold.

Colour tests:

Concentrated sulphuric acid, bluish-green, olive-green on warming.
„ „ „ with potassium bichromate, dark reddish-brown, changing to intense green.
Nitric acid, dark green (without solution); added to the alcoholic solution, nitric acid gives a magnificent green coloration.

(11) *CONCUSCONIDINE* A., $C_{23}H_{26}N_2O_2$; yellow amorphous; M.P. 124°; slightly dextro-rotatory.

(12) *CUSCAMINE* A. crystalline (from alcohol); M.P. 218°; non-fluorescent; taste astringent and somewhat bitter.

Soluble in ether and chloroform (very readily), and in alcohol.

Precipitated by alkaline hydrates and ammonia.

[No colour with Ferric chloride.]

(13) *CHAIRAMINE* A., $C_{22}H_{26}N_2O_4+H_2O$. Crys. needles or prisms; anhydrous at 140°, then melts at 233°; neutral reaction, dextro-rotatory (about +100°).

Soluble readily in ether and chloroform, in 540 parts of 97 per cent. alcohol at 11° C., with difficulty in water, and **insoluble** in dilute hydrochloric acid.

Precipitants:

Alkaline hydrates. } insoluble in excess.
Ammonia. }
Platinum chloride—yellow needles, insoluble in alcohol; formula $(Ch{\cdot}HCl)_2\,PtCl_4$.

Colour tests: Concentrated sulphuric acid, colourless.
Nitric acid, gradual dark green.
Fröhde's solution, colourless.

(14) *CONCHAIRAMINE* A., $C_{22}H_{26}N_2O_4$, with H_2O, and with C_2H_6O (alcohol), prismatic crystals; M.P. 108°-110° the hydrate, 120° anhydrous, or 82°-86° the alcoholate. Reaction almost neutral.

Soluble in hot alcohol (with difficulty cold), in ether, and readily in chloroform; scarcely soluble in dilute hydrochloric acid.

Precipitants:

Alkaline hydrates, pp. becomes crystalline, and is insoluble in excess.
Ammonia, insol. in excess.
Platinum chloride, dark yellow, flocculent.
Potassium sulphocyanide.

Colour tests:

Concentrated sulphuric acid, brown, changing to intense green.
Nitric acid added to the hydrochloric solution, dark green.
Fröhde's reagent, brownish.

(15) *CHAIRAMIDINE* A., $C_{22}H_{26}N_2O_4$; amorphous; M.P. 127°; dextro-rotatory (= +7·3°) neutral.

Soluble in alcohol, ether, chloroform, benzene; not in water.

Removed by benzene from alkaline solutions.

Absorbed by charcoal.

Precipitants:

Alkaline hydrates. } Insoluble in excess. | Platinum chloride, yellow flocculent, pp. = $(ChHCl)_2\,PtCl_4+5\,H_2O$.
Ammonia. }
Alk. carbonates. }

Colour tests:

Concentrated sulphuric acid, yellow, becoming dark green.
Nitric acid does not dissolve the solid alkaloid.
„ „ to the Hydrochloric solution, dark green.

(16) *CONCHAIRAMIDINE* A., $C_{22}H_{26}N_2O_4$, crys. prisms.; M.P. 114°-115° (anhydrous); neutral reaction.

Soluble in alcohol, ether, chloroform, benzene, acetone, and easily in acetic and mineral acids except nitric. **Insoluble** in water.

Removed from alkaline solutions by benzene, etc.

Precipitants:

Alkaline hydrates } pp. oily, becoming crystalline.
„ carbonates }
Ammonia }
Baryta }

Colour tests: Concentrated sulphuric acid, intense green.
Nitric acid, as Chairamidine (15) above.
Fröhde's reagent, intense green.

(17) *QUINAMINE* A., $C_{19}H_{24}N_2O_2$. Crystallizes in prisms (no crystalline hydrate; M.P. 172°; dextro-rotatory (+104·5°); tasteless, though salts are bitter; non-fluorescent.

Soluble in 1,516 parts water at 16° C., 105 alcohol at 20° C., 48 ether at 16° C., in chloroform, boiling benzene, and petroleum ether.

Precipitated by gold chloride, pale yellow, changing to red with reduction of gold.

(18*a*) *APOQUINAMINE* A. From the above by boiling with acids.

(18*b*) *QUINAMICINE* A., $C_{19}H_{24}N_2O_2$. Formed together with (18*c*) from Quinamine by heating to 130° with hydrochloric acid in sealed tube. Crystalline; M.P. 109°.

Soluble in alcohol, ether, and chloroform.

Precipitants:
Alkaline hydrates, carbonates, bicarbonates } insol. in excess.
Ammonia, amorphous, becoming crystalline; insol. in excess.
Gold chloride, amorphous; yellow if pure, or purple if containing Quinamidine.
Potassium iodide.
Salicylate ammonia.
Oxalate „
Sodium chloride.
Acids (hydrochloric, nitric, or sulphuric) also give precipitates.

(18*c*) *QUINAMIDINE* A. Formed in the preparation of the above, from which it is distinguished by giving
No pp. with bicarbonates of the alkalies.
The gold chloride pp. is purple.

(19) *CONQUINAMINE* A., $C_{19}H_{24}N_2O_2$. Triclinic prisms; M.P. 121°; dextro-rotatory (+204°).

Soluble in alcohol, ether, chloroform, benzene, carbon bisulphide, and with difficulty in water.

Precipitate by gold chloride is yellow, becoming purple.

(20) *CUPREINE* A., $C_{19}H_{22}N_2O_2$ and with $2H_2O$. Crystallizes in prisms from ether; becomes anhydrous at 120°-125°, subsequently melting at 198°; alkaline reaction, lævo rotatory (= −175°).

Soluble in alcohol, but with difficulty in ether or chloroform.

It is not removed from a 10 per cent. soda solution, but from ammoniacal liquids by immiscible solvents.

Precipitants:
Alkaline hydrate, soluble in excess.
Ammonia, slightly soluble in excess.
Ferric chloride, dark brown coloration.
Platinum chloride, pp. difficultly soluble. The dry double salt becomes dark green on heating.
Potassium sulphocyanide, cloud changing gradually to crys. pp.
Chlorine water, etc. (thalleioquin reaction), as Quinine.

(21) *HOMOQUININE* A. (Ultraquinine), $C_{19}H_{22}N_2O_2 \cdot C_{20}H_{24}N_2O_2$. A compound of Quinine and Cupreine in molecular proportions (Paul and Cownley). Pearly prismatic crystals in stellar groups; M.P. 177°; lævo-rotatory $[a]_J = -221°$; fluorescent (the acid salts); strongly alkaline.

Soluble in alcohol and chloroform, but with difficulty in ether unless freshly precipitated.

Reactions:
Ammonia, dissolves. Rochelle salt, pp. with concentrated solutions.
[No pp. iodo-potassic iodide.]
Concentrated sulphuric acid with potass. bichromate, deep green.
Concentrated sulphuric acid alone, nitric „ „ } do not decompose.
Gives the thalleioquin reaction (chlorine water, then ammonia).
The soda hydrate pp. contains Quinine and Cupreine, only the former being removable on shaking with ether.

(22) *CINCHONAMINE* A., $C_{19}H_{24}N_2O$. Hexagonal prisms (no crystalline hydrate); M.P. 195°; dextro-rotatory (+121°); bitter, poisonous.

Soluble in alcohol, ether, chloroform, benzene, and carbon bisulphide; difficultly in water and petroleum ether. Hydrochloride sparingly soluble.

Reactions: No colour with ferric chloride, nor thalleioquin test (chlorine water, etc.).

(23) *QUINOIDINE* and (24) *QUINETUM* are mixtures; the former containing principally Quinicine (see 2*b*) with Cinchonicine (see 4*n*); and the latter the alkaloids of red bark.

(25) *PARICINE* A., $C_{16}H_{18}N_2O$, from China Succirubra and Cascarilla

hexandra. Pale yellow powder; M.P. 116° or 130°; bitter; feebly alkaline.

Soluble in alcohol and ether, slightly in petroleum ether, scarcely in water.

Concentrated sulphuric acid, yellowish green.

(26) *QUINOVIN* G. (Chinovin, Esenbeckin), $C_{30}H_{48}O_8$ (Gilm and Hlasiwetz). Amorphous, resinous powder, neutral reaction, taste gradually bitter; dextro-rotatory (+52·4° in alcoholic solution). Hydrochloric acid gas decomposes with formation of Quinovic acid (27) and a sweet substance, Quinovite.

Soluble in alcohol, less easily in ether, in hot water, but scarcely in cold; also in chloroform and fatty or essential oils.

Reactions: Soluble in alkaline hydrates, ammonia, lime water.
Insol. in alkaline carbonates.
[No pp. lead acetate neutral or basic to alcoholic solution.]
Pp. by other metallic salts and by acids from alkaline solutions.
Fehling's solution, not reduced.

Colour tests:
Concentrated sulphuric acid gradually dissolves, dark red.
" " " with sugar, red.
" nitric acid, dissolves on warming, with decomposition and evolution of nitrous fumes.

(26*a*) *BETA-* and *ALPHA-QUINOVINS* have been prepared and described by Liebermann as distinct substances as follows: α=alpha-quinovin and β the beta compound—*a* from cinchona bark, β from cuprea bark; both give glittering scales on adding water to the alcoholic solution. Isomeric: formula = $C_{38}H_{62}O_{11}$. Both dextro-rotatory; $\alpha = +56{\cdot}6°$, $\beta = +27{\cdot}9°$.

Solubility: *a*, nearly 1·2 in alcohol (absolute), with difficulty in ether, chloroform, or benzene, and scarcely in water, hot or cold. β is **insoluble** n ethyl acetate or absolute ether (distinction from *a*); β forms a crystalline compound with 5 molecules of alcohol; in most other respects β resembles *a*.

Colour tests with (*a*): Concentrated sulphuric acid, orange yellow with evolution of carbon monoxide. Glacial acetic acid, pale blue.

(27) *QUINOVIC ACID.* Obtainable from Quinovin, and occurring naturally. $C_{24}H_{38}O_4$. Crys. scales or needles; dextro-rotatory; tasteless.

Solubility very slight in alcohol, ether, chloroform, or glacial acetic acid. **Insoluble** in water.

Reactions: Salts with alkalies, crystalline.
Alkaline hydrates, „ carbonates, Ammonia } dissolve to frothy solutions.

Precipitated by acids from alkaline solutions.
Copper sulphate pp. gradually with green coloration.
Concentrated sulphuric acid, dissolves unchanged (re-precipitated by water).

(28) *CASCARILLIN* B. From China (Quina) nova, etc., and Cascarilla bark (Croton eluteria—*Euphorbiaceæ*); amorphous, resinous, neutral reaction, bitter.

Soluble in alcohol, ether, and benzene, scarcely in water.

Removed by benzene from *acid* solutions.

[**Not Precipitated** by lead acetate neutral or basic nor by tannic acid]

Colour tests:
Concentrated sulphuric acid, blood-red (? reddish-brown, Dragendorff); nitric acid or nitre added to the solution after ten to fifteen hours changes the reddish solution to bluish violet, then blood-red; compare morphine.
Nitric acid, reddish-violet.
Hydrochloric acid, violet.

(29) *COPALCHIN* B. From Croton pseudoquina (*Euphorbiaceæ*). Amorphous, resinous, bitter.

Soluble in alcohol and chloroform, 'partly' in ether—hence presumably impure—scarcely in water.

Reactions:
[Not precipitated by neutral lead acetate.]
Pp. by tannic acid.
Concentrated sulphuric acid, red.

(30) *LIGNOIN* B. From old Huanocho bark. $C_{20}H_{23}NO_8$ ($C_{20}H_{20}O_8NH_3$?). Amorphous, brown, humoid. Yields ammonia when boiled with alkali.

Soluble in alcohol.

Reactions: Alkaline hydrates dissolve brown with evolution of ammonia; acids re-precipitate.
Lead acetate neutral, „ basic } flesh-coloured pp.

(31) *CALIFORNIN* B. (not the Californine from Lotus bark). From China (Quina) Californica=Cascarilla riedeliana or possibly Buena

obtusifolia. Investigator: Winckler. Golden yellow, amorphous, neutral reaction, bitter.

Soluble in water and alcohol; not in ether.

[**Not precipitated** by Tannic acid.

Platinum chloride.

Mercuric chloride.]

Concentrated **Sulphuric acid**, brownish red.

(32) *MORADEINE* **A.** This, with the next substance, obtained by Arata from Bolivian Cascarilla = Quina morada (Howardia, *Wedd.*, or Pogonopus febrifugus, *Benth.-Hook.*), *Rubiaceæ*. Opaque prismatic crystals; **M.P.** 199°-200°.

Soluble very readily in alcohol, ether, and chloroform; slightly in water.

(33) *MORADIN*, $C_{21}H_{18}O_8$ or $C_{16}H_{14}O_6$. Fluorescent. Acts as a hydroxy-quinone.

(34) *JAVANINE* **A.** Obtained in 1877 by O. **Hesse** from the mixture of 'amorphous bases' derived from Cinchona calisaya var. Javanica.

Crys. rhombic plates from water; **soluble** also in ether.

Concentrated sulphuric acid, intense yellow.

§ 70. **CNICUS** benedictus (Centaurea calcitrapa); *Compositæ*.

CNICIN **B.**, $C_{42}H_{56}O_{15}$. Crys. silky needles; fusible, neutral, bitter; dextro-rotatory (130·68°); emetic and purgative.

Soluble in ethyl and methyl alcohols (in all proportions), in chloroform and benzene, scarcely in ether or cold water, with difficulty in hot water, and insoluble in essential oils.

It is removed from *acid* solutions by benzene or chloroform.

Precipitants: Basic lead acetate, partially (Dragendorff).

[Not gold chloride, and no reduction.]

[Silver nitrate not reduced.]

Colour tests:

Concentrated sulphuric acid, red; odour of benzoic acid on warming; on adding water, violet; then ammonia, yellow.

Fröhde's solution, blood-red.

§ 71. **COCA** (Erythroxylon Coca); *Erythroxylaceæ*, Brazil.

(1) *COCAINE* **A.** (Benzoyl-methyl ecgonine; or, taking into account the composition of Ecgonine, the full constitutional name becomes, Benzoyl-methyl-tetrahydromethyl-pyridine-β-hydroxypropionic acid); $C_{17}H_{21}NO_4$. Crystallizes in 4-6 sided prisms; **M.P.** 98°, re-crystallizing on cooling; at higher temperature, partial sublimation with decomposition; lævo-rotatory (−15·8°); alkaline reaction, slightly bitter, succeeded by temporary destruction of sense of taste and the local anæsthesia characteristic of this alkaloid. Dilates the pupil. Readily decomposes into benzoic acid, methyl alcohol, and Ecgonine. In the human system it is entirely (or nearly entirely) destroyed; Vitali has, however, found traces together with Ecgonine in the urine. The hydrochloride is readily crystallizable.

Soluble in 704 parts cold water (easier hot, but liable to decompose), in alcohol and benzene, very readily in ether; also dissolved by the following when hot: chloroform, wood spirit, acetone, and petroleum ether; crystallization taking place on cooling.

Precipitants of the salts (the hydrochloride was used):

Alkaline hydrates, „ carbonates, Ammonia } pp. crystalline, or amorphous changing to crystals. Insol. in excess. Ammonia less liable to cause decomposition than fixed alkalies.

Alk. bicarbonate, pp. if concentrated.

[No pp., lead acetate.]

Tannic acid in acid solution, white; cloud at 1 in 25,000.

Picric acid, yellow. Gradually at 1 in 1,000.

Platinum chloride, whitish yellow even at 1 in 12,500; cloud at 1 in 25,000. *The crystals of the double salt are of remarkable* T *shape.* M.P. 80° (W. C. Howard).

Gold chloride, light yellow fern-like. Immediate pp. at 1 in 3,000; gradual at 1 in 12,500. Benzoic acid produced on warming.

Potassium ferrocyanide, pp. soluble in excess.

[Potassium ferricyanide reduced.]

Chromic acid (5 per cent.) gives pp. that is *only permanent after acidification with hydrochloric acid* (characteristic).

[Potassium bichromate not unless concentrated; but on addition of HCl a pp. See preceding test.]

Phospho-molybdic acid, yellowish-white at 1 in 12,500; cloud at 1 in 50,000.

Phospho-tungstic acid, gelatinous; sol. in ammonia.

Iodo-potassic iodide, brown with black globules under the microscope (Vitali); pink at 1 in 7,500; cloud at 1 in 200,000.

Mercuric-potassic iodide, pp. even very dilute; cloud at 1 in 1,000,000.

Mercuric chloride, white.

Iodine, brown.

Stannous chloride ($SnCl_2$), white.

Potassium permanganate, violet.

Colour tests:

Concentrated sulphuric acid, no coloration; vapours of benzoic acid on warming.

Concentrated hydrochloric acid, decomposes with formation of Ecgonine.

Per-iodic acid, minute fragment of alkaloid dissolved in ½ to 1 cc. concentrated sulphuric acid, then sodium per-iodate or per-iodic acid in small quantity (two or three times as much as alkaloid), produces on gentle warming, bright green changing to blue and violet, with evolution of violet fumes (Vitali).

NOTE: Atropine, Thebaine, Chelidonine, and Corydaline give colour changes under these circumstances, but not same sequence. Reaction is really due to benzoic acid.

(1*a*) *DEXTRO-COCAINE* A., $C_{17}H_{21}NO_4$. Formed artificially from benzoyl-chloride and dextro-methylecgonine (Einhorn and Marquardt). Oily substance becoming crystalline; M.P. 45°.

Soluble in alcohol (less so than ordinary Cocaine), in ether, benzene, and petroleum spirit.

Resembles ordinary Cocaine (Lævo-cocaine) in chemical properties and physiological action, but its anæsthetic effect is more transient.

(1*b*) *ETHYL-COCAINE* A., $C_{18}H_{23}NO_4$ (Homo-cocaine, Cocethyline, Benzoyl-ethylecgonine). Prepared by ethylation of Benzoyl-ecgonine. Crys. glassy prisms; M.P. 108°-109°.

Soluble in alcohol and ether; not in water.

Weaker physiological action than Cocaine.

(2) *BENZOYL-ECGONINE* A., $C_{16}H_{19}NO_4$. $[C_8H_{13}N(COOH)O\cdot C_7H_5O]$. Occurs naturally, and produced by decomposition of Cocaine, which may be reformed from it by methylation. Benzoyl-ecgonine is also prepared by synthesis from benzoic anhydride and Ecgonine. Crystallizes with $4H_2O$ in transparent prisms—M.P. variable (87°-140°)—or in long needles from chloroform. The hydrate in alcohol has lævo-rotatory power of −44·6°. Scarcely anæsthetic; reaction neutral.

Soluble in alcohol, wood spirit, hot acetone, hot water (with difficulty cold), and hot chloroform (crys. needles on cooling); scarcely in ether.

Reactions:

Alkaline hydrates, Ammonia } dissolve.

Gold chloride, pp. yellow crys., sol. hot alcohol, slightly in water.

(3) *ECGONINE* A., $C_9H_{15}NO_3$ ($C_8H_{13}N(OH)\cdot COOH$; Methyl-tetra-hydro-pyridine-β-oxypropionic acid). Occurs in the amorphous bases obtained after separation of Cocaine, and producible (together with Benzoyl-ecgonine and Cocaine benzoate) from Cocaine by mere heating of the aqueous solution to 80° (Hesse). Rhombic crystals with 1 molecule H_2O; M.P. 198°, or when anhydrous 205°; reaction neutral; taste slightly bitter and sweet.

Soluble very readily in water, with difficulty in absolute alcohol or chloroform, and **insoluble** in ether or carbon bisulphide (Mussi).

Reactions:

Alkaline hydrates, Ammonia } dissolve.

[No pp. platinum chloride, the double salt being very soluble.]

Gold chloride to concentrated solution, yellow amorphous pp.

[No pp. 5 per cent. chromic acid solution.]

Phospho-molybdic acid, pp. yellow.

Iodo-potassic iodide, reddish-brown, becoming crystalline, microscopic tufts, rhombic plates or prisms (Vitali), 1 in 500 for crystals.

Treated with phosphoric pentachloride, anhydro-ecgonine is formed: see (3*b*).

(3*a*) *DEXTRO-ECGONINE* A., $C_9H_{15}NO_3$; from preceding alkaloid by heating with potash; M.P. 257°: dextro-rotatory.

(3*b*) *ANHYDRO-ECGONINE* A., $C_9H_{12}NO_2$; from Ecgonine by heating with phosphorus pentachloride, or from Cocaine by heating to 140° with a solution of hydrochloric acid gas in glacial acetic acid; crys. M.P. 235°.

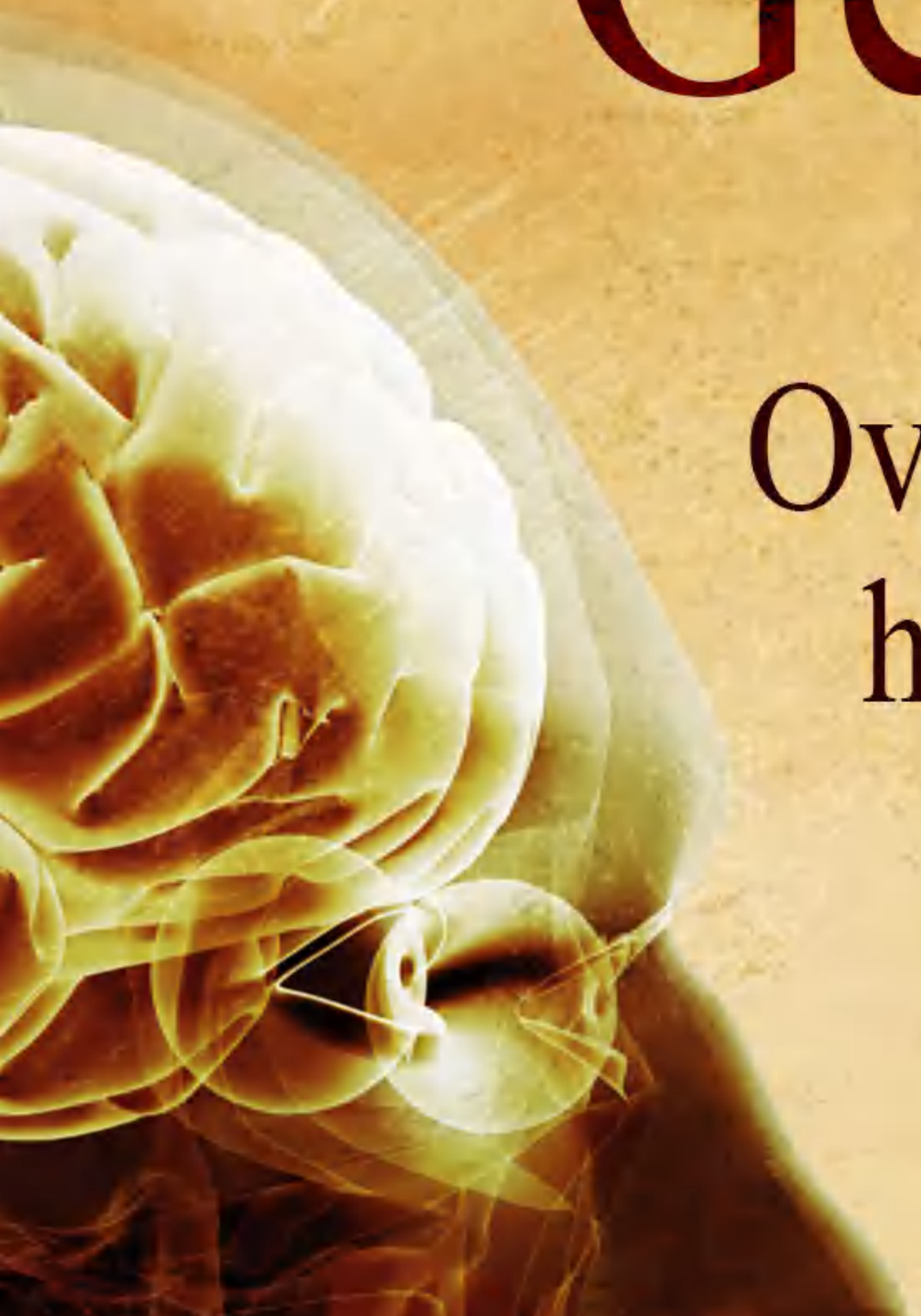
Soluble in water and alcohol, but *not* in ether, chloroform, benzene, or petroleum spirit.

(4) *CINNAMYL-ECGONINE* A., $C_{18}H_{21}NO_4$. May be prepared from Ecgonine and Cinnamic anhydride, and derivable from Cinnamyl-cocaine [see (5)], which occurs naturally.

(5) *CINNAMYL-COCAINE* A., $C_{19}H_{23}NO_4$; the methyl ether (ester) of the preceding, or Cinnamyl-methyl-ecgonine. Occurs naturally (Giesel obtained 5 per cent., together with 15 per cent. of other bases besides Cocaine, from a Java Coca); producible also by methylation of (4). Crystalline; M.P. 121°.

Soluble in ether and alcohol, but not in water.

Odour of bitter almonds on treatment with potassium permanganate in the cold.



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Precipitants: Alkaline hydrates.
Ammonia.

(*f*) *PARAMENISPERMINE* A.; crystalline; M.P. 250°, and volatile.

Soluble in alcohol, almost **insoluble** in ether, and not dissolved by water. Though soluble in acids, no salts are formed.

§ 73. **COLCHICUM** autumnale (C. Alpinum, C. Arenarium, C. Montanum, and C. Neapolitanum); *Colchicaceæ*.

(*a*) *COLCHICINE* A., (Colchiceine-methyl ether) ($C_{22}H_{25}NO_6$, Zeisl); amorphous; yellowish-white, gummy usually, although prismatic crystals are obtainable; M.P. 140°, or 143°-147° (Zeisl); lævo-rotatory); feebly alkaline; bitter. Yields Colchiceine and a yellowish-green resin on boiling with acids.

Soluble in water in all proportions, very readily in alcohol: also in chloroform, benzene, and amyl alcohol. In ether, Dragendorff describes it as soluble, but Zeisl as scarcely so. **Insoluble** in petroleum ether.

It is removed from *acid* solutions by amyl alcohol, also in part by benzene and chloroform.

Precipitants:

[Alkaline hydrate, pp. if acid and warm, soluble in excess, yellow.]
[Not lead acetate, neutral or basic.]
Tannic acid, 1 in 2,500; sol. in alcohol and acetic acid.
[Not picric acid, unless concentrated.]
[No colour ferric chloride?]
Platinum chloride, not at once 1 in 125, but after twenty-four hours 1 in 3,000.
Gold chloride pp., reduction of gold after twenty-four hours.
[Not potass. ferrocyanide, unless concentrated.]
[Silver potassic cyanide, not 1 in 3,000.]
[Potassium bichromate, not 1 in 3,000.]
Phospho-molybdic acid, 1 in 3,000.
Iodo-potassic iodide, 1 in 2,500.
Bismuth-potassic iodide, 1 in 3,000.
[Cadmium-potassic iodide, only if concentrated.]
[Mercuric-potassic iodide, only when acidulated and not dilute.]
[Mercuric chloride, if concentrated.]
Chlorine water, yellow pp.; sol. ammonia, orange.

Colour tests (besides those formed by precipitants above):
Concentrated sulphuric acid, yellow.
„ „ with sugar, yellow.
Concentrated sulphuric acid, with potass. bichromate, green, then brown.
„ „ „ with nitric acid, blue.
Potash added to this after the solution has faded in colour, red.
Nitre or nitric acid after ten hours' solution in sulphuric acid~violet-red~blue~brown.
Nitric acid alone, blue to violet~brown~yellow.
„ „ fuming, violet to indigo.
Acids generally, yellow.
Fröhde's solution, yellow~yellowish-green; in twenty-four hours, yellow.

(*b*) *COLCHICEINE* A., $C_{21}H_{23}NO_6$. Occurs naturally, and prepared from preceding alkaloid by long boiling with acids. Crystallizes in needles; the hydrate melts at 139°-142° in open tube, or 156°-162° in closed tube, and the anhydrous alkaloid at 161°-172° (Zeisl); lævo-rotatory, neutral reaction. Properties feebly acid and basic.

Solubility as Colchicine, except that it is difficultly soluble in cold water.

Removed from acid solutions by benzene.

Reactions:
Alkaline hydrates, Ammonia } dissolve, metallic salts pp.
[No pp. lead acetate, neutral or basic.]
Ferric chloride, green coloration or pp.
Gold chloride, pp.
Phospho-molybdic acid, pp.

Colour tests similar to those of Colchicine.

§ 74. **COLOCYNTHIS** citrullus (Cucumis); *Cucurbitaceæ*.

(*a*) *COLOCYNTHIN* G., $C_{56}H_{84}O_{23}$. Usually amorphous, but capable of crystallization; neutral, bitter, purgative. Yields sugar and Colocynthein on treatment with acids. Not found in body after death.

Soluble in 8 parts water, in alcohol, amyl alcohol, and benzene (not in latter, according to Henke); with difficulty in ether and chloroform, and not in petroleum ether, or carbon bisulphide.

Removed from *acid* solutions by amyl alcohol, chloroform, and benzene.

Reactions: Ammonia, dissolves.
[No pp. lead acetate, neutral or basic. Dragendorff states that pp. is obtained by latter.]
Tannic acid, pp.
Fehling's solution reduced.

[No pp. Cadmium-potassic iodide nor metallic salts generally.]

Colour tests: Concentrated sulphuric acid, deep red.

Nitric acid, light red.

Fröhde's solution, after half-hour, cherry; later, nut colour (Dragendorff).

(*b*) *COLOCYNTHEIN* G.-deriv. Obtained from (*a*) by treatment with acids.

(*c*) *COLOCYNTHITIN.* Microscopic rhombic prisms; not bitter.

Soluble in hot alcohol and ether; not in water or cold alcohol (Walz).

§ 75. **CONIUM** maculatum (spotted hemlock) the alkaloids (*a*) to (*e*); Cicuta virosa (water hemlock, cowbane) the substances (*a*) and (*f*); *Umbelliferæ.*

(*a*) *CONIINE* A. (Conicine, Cicutine, Alpha-propyl piperidine), $C_8H_{17}N$ or $C_5H_{10}(C_3H_7)N$. Colourless liquid, disagreeable odour; B.P. variously given 150°-180°-218° (?). Volatile in vapour of alcohol or water, and somewhat at ordinary temperatures. Alkaline reaction, burning taste; dextro-rotatory $[a]_D = +13{\cdot}8$; specific gravity 0·878 to 0·886. Hydrochloride crystalline; causes dilation of the pupil; very poisonous; salts lose Coniine on evaporation.

Soluble in 100 parts cold water (cloud on warming), in all proportions in alcohol, in 6 parts ether; also in chloroform, benzene, petroleum ether, amyl alcohol, acetone, and ethereal oils; with difficulty in carbon bisulphide.

Removed from alkaline solutions by immiscible solvents, including petroleum ether.

Precipitants: [Not alkaline hydrates.

Not ammonia.

„ sodium acetate.

„ lead acetate neutral.]

Tannic acid, yellowish white.

[Picric acid, not at 1 in 1,000.]

[Not platinum chloride.]

Gold chloride, cloud at 1 in 300.

Copper sulphate.

[Chromic acid produces butyric acid recognisable by odour.]

Phospho-molybdic acid, yellow, 1 in 1,000. Blue with ammonia.

Iodo-potassic iodide, kermes coloured pp.; distinct at 1 in 8,000.

Bismuth-potassic iodide, yellow 1 in 2,000; limit 1 in 6.000.

Cadmium-potassic iodide, 1 in 200 cloud; (? 1 in 10,000, Dragendorff).

[Not zinc-potassic iodide.]

Mercuric-potassic iodide 1 in 800, white amorphous adherent like resin; characteristic, but nicotine similar.

Mercuric chloride, white amorphous (not in very dilute solutions). [The Nicotine pp. is crystalline.]

Free Coniine precipitates mercury, silver, lead, copper, tin, iron, manganese, and zinc oxides from solutions of their salts.

Albumen coagulated, distinction from nicotine and solid alkaloids ('test valueless,' Dragendorff).

Colour tests (mostly negative):

Concentrated sulphuric acid
„ „ „ with sugar
„ „ „ with nitric acid
} no effect except after standing.

Fuming nitric acid, bluish~orange.

Dry hydrochloric acid gas, purple-red~indigo.

Concentrated hydrochloric acid (solution), no effect or pale red.

Fröhde's solution, no effect.

Coniine salts on exposure become gradually red~violet~green~blue.

(*b*) *CONHYDRINE* A., $C_8H_{17}NO$. Separates out from commercial Coniine by cooling to 5° C. Crystalline pearly plates; M.P. 126·6°; B.P. 226°, part sublimes at 100°; volatile without decomposition; dextro-rotatory, powerfully alkaline (displaces ammonia from its salts). Odour faintly Coniine-like; feebly narcotic.

Soluble in water, very readily in alcohol and ether, sparingly in petroleum ether, by which, however, it may be removed from alkaline solutions.

Precipitants, etc.: [Not alkaline hydrates.

Not ammonia.

„ lead acetate neutral.]

Picric acid
Iodo-potassic iodide
Mercuric-potassic iodide
} pp. from more dilute solutions than is the case with Coniine.

Not attacked by cold fuming nitric acid.

(c) *PSEUDO-CONHYDRINE* A., $C_8H_{17}NO$ (Merck). Crys. needles; M.P. about 98° C.; B.P. 229°-231°; volatile.

Soluble in alcohol, ether, and chloroform.

(*d*) *METHYL-CONIINE* A., $C_9H_{17}N$. Accompanies Coniine in hemlock; volatile, oily, lighter than water, odour Coniine-like.

(*e*) *ETHYL-PIPERIDINE* A., $C_7H_{15}N$, occurs with Coniine. Volatile liquid, lighter than water.

(*f*) *CICUTOXIN* B. Amorphous, resinous, poisonous; acid reaction, disagreeable taste. Böhm, *Archiv. exp. Pathol.*, **5**, 281.

Soluble in alcohol, ether, chloroform, and hot water; *not* in petroleum ether.

Alkaline hydrates } dissolve.
Ammonia }

§ 76. **CONVALLARIA** majalis (Lily of the Valley); *Liliaceæ*. Investigator: Walz, *Jahrb. Pharm.*, vols. 7 and 8; also *N. Jahrb. Ph.*, vols. 5 and 10.

(*a*) *CONVALLAMARIN* G., $C_{46}H_{44}O_{24}$. Crystalline, neutral, bitter-sweet taste; physiogical action like Digitalin. Yields sugar and *Convallamaretin*.

Soluble in water, alcohol, chloroform, amyl alcohol (removed by last two solvents from acid aqueous solution. **Insoluble** in ether.

Reactions:

Ammonia dissolves.	Pp., Tannic acid.
[No pp. lead acetate neutral or basic.]	Mercurous nitrate, white pp.
	[No pp., most other reagents.]

Colour tests:

Sulphuric acid added to aqueous solution, violet.
,, ,, with bromine, brown, but violet in presence of water.
Hydrochloric acid warm, red.

(*b*) *CONVALLARIN* G., $C_{31}H_{31}O_{11}$ (Walz); crystallizes in rectangular pillars; neutral reaction; sharp taste.

Soluble in alcohol, scarcely in water (which it renders frothy), and not in ether.

Not **precipitated** by lead acetate, neutral or basic.

§ 77. **CONVOLVULACEÆ** various: Convolvulus purga (Ipomœa Schiedeana, Jalap), (*a*) and (*b*); C. orizabensis and C. scammonia, (*c*), (*d*); Ipomœa turpethi, (*e*), (*f*); I. simulans (Tampico jalap), (*g*).

(*a*) *CONVOLVULIN* G. (Buchner's Jalapin, Kayser's Rhodeoretin), $C_{31}H_{56}O_{16}$ (Mayer); amorphous; M.P. 150°; feebly acid reaction; tasteless. Hydrochloric acid converts first into sugar and Convolvulinol (*b*), then into Convolvulinic acid. It is not traceable in excretions, but may be found in stomach after death.

Soluble in acetic acid in all proportions, also in alcohol and acetic ether, but scarcely in water, amyl alcohol, or chloroform, and not in ether, benzene, or petroleum ether.

Reactions:

Alkaline hydrates } dissolve, with conversion to Convolvulinic acid.
,, carbonates }
Ammonia }
No pp. with metallic salts that are soluble in alcohol, except silver nitrate.
Concentrated sulphuric acid, gradually pure red.

(*b*) *CONVOLVULINOL* [from (*a*)]; crys. flexible needles; M.P. 39°.

Soluble in alcohol and ether, but with difficulty in water.

(*c*) *JALAPIN* G. (Scammonin, Kayser's Para-rhodeoretin), $C_{34}H_{56}O_{16}$ (Mayer; Spirgatis); amorphous, colourless, resinous; M.P. 150°; feebly acid reaction; tasteless.

Soluble in alcohol, ether, chloroform, amyl alcohol, and acetic acid; with difficulty in water, benzene, and carbon bisulphide.

Reactions:

Alkaline hydrates } dissolve.
Ammonia }
Concentrated sulphuric acid, gradual pure red.

(*d*) *JALAPINOL* (from preceding substance), possibly identical with (*b*); crystalline; M.P. 62·5°.

Soluble in alcohol and ether, not in water.

Alkalies convert to Jalapinic acid (Convolvulinic acid ?).

(*e*) *TURPETHIN* G., $C_{34}H_{56}O_{16}$ (Spirgatis); amorphous; brownish-yellow, resinous; M.P. 183°; taste gradually sharp and bitter. Acids produce sugar and (*f*).

Soluble in alcohol; not in water or ether.

Concentrated sulphuric acid, purple.

(*f*) *TURPETHOL* (from above); crys. microscopic needles; M.P. 88°; burning taste; acid reaction.

Soluble in alcohol.

(*g*) *TAMPICIN* G., $C_{34}H_{54}O_{14}$; amorphous, resinous; M.P. 130°. Decomposed by long heating at 100°. Acids convert to Tampicolic acid, or on further treatment Tampicic acid.

Soluble in alcohol and ether.

Concentrated sulphuric acid, purple.

§ 78. **CORIARIA** myrtifolia; *Phytolaceæ*. The leaves and fruit, 6 to 9 parts in 100,000. Riban, *Compt. Rendus*, vols. 57 and 63.

CORIAMYRTIN G., $C_{30}H_{36}O_{10}$; crystallizes in clino-rhombic prisms;

M.P. 220°; dextro-rotatory ([a]J=+24·5°); neutral reaction; tasteless, odourless, poisonous (producing convulsions).

Soluble in 70 parts water, 50 cold alcohol, in chloroform, benzene (and ether ?); scarcely in carbon bisulphide.

Reactions:

[Not precipitated by lead acetate, neutral or basic.
„ tannic acid.
„ picric acid.]
Fehling's solution, reduced.
Phospho-molybdic acid. pp.

Colour tests: Concentrated sulphuric acid, blackens.
Concentrated nitric acid gives a crystalline nitro-derivative.
On reduction with hydriodic acid, pouring off the liberated iodine and then adding potash, fuchsine red coloration.

§ 79. **CORNUS** Florida (Dogwood); *Cornaceæ*. Investigator: Geiger, *Ann. Chem. Ph.*, 14, 206, etc.

(*a*) *CORNIN* B.; crys silky needles; neutral; bitter.

Soluble in water, alcohol, ether.

Precipitants:

[Not lead acetate, neutral.]	Silver nitrate, white crys. pp.
Lead acetate, basic.	[Not mercuric chloride.
[Not tannic acid,	„ iodine solution.
„ ferric chloride.]	„ barium chloride.]

(*b*) *CORNUS RESINOID* (Dogwood Quinine); neutral, tasteless, resinous.

Soluble in ether and hot alcohol; not in water.

§ 80. **CORYDALIS** group. Corydalis *Dicentra* cava *Schwg.*, [C. bulbosa *Pers.*=C. tuberosa *Dec.*=Bulbocapnus cavus *Bernh.*], C. fabacea [=C. intermedia], and C. solida [Bulbocapnus digitatus=C. digitata =C. bulbosa *Dec.*], (*a*); Fumaria officinale (Fumitory), (*b*), (? *a*), etc.; *Fumariaceæ*. Aristolocia cava (Birthwort), (*a*); *Aristolochiæ*. [For A. Argentina, see separate entry.] Compare Berberis group, the alkaloids of which are probably chemically connected with these. See (c) below.

(*a*) *CORYDALINE* A. [Adermann's alkaloid No. 1, see (c); not Adermann's Corydaline, see (*b*) and (*d*)], $C_{18}H_{19}NO_4$ (Wicke) or $C_{22}H_{29}NO_4$ =[$C_{18}H_{17}(CH_3O)_4N$] (Dobbie); crystallizes in prisms or needles; M.P. 134° (Dobbie); alkaline reaction; bitter in alcoholic solution, or in form of salts. Becomes yellow on exposure.

Soluble in ether, chloroform, benzene, amyl alcohol, carbon bisulphide, and turpentine; with difficulty in alcohol; not in water, which precipitates it from alcohol. The best solvent is a mixture of ether and alcohol.

Precipitants:

Alkaline hydrates, sol. in excess (Wicke); insol. (M. Freund).	Potass. sulphocyanide, white crys.
Alkaline carbonates.	Potass. chromate, yellow.
Ammonia.	„ bichromate.
[Not lead acetate, neutral or basic.]	Sodium phospho-tungstate.
Tannic acid (to alcoholic solutn).	Potass. iodide, white.
Picric acid, crystalline.	Iodo-potassic iodide, brown.
Platinum chloride, yellow crys.	Mercuric-potassic iodide, yellowish-white.
Gold chloride, yellow crystalline.	Iodine tincture.

Colour tests:

Concentrated sulphuric acid, colourless ? (Dupuy says yellow to red).
Nitric acid produces a brownish-red resinous substance.

(*b*) *FUMARINE* A. [Adermann's Corydaline (?), Bulbocapnine (?), see (*f*)], $C_{20}H_{19}NO_2$ (Reichwald); six-sided clino-rhombic prisms; optically inactive, alkaline, bitter. Acetate very soluble.

Soluble in alcohol (difficultly, Reichwald), in 11 parts chloroform, 78 benzene, in amyl alcohol and carbon bisulphide; but with difficulty in water, and sparingly in ether or petroleum ether.

Precipitated by

Alkaline hydrates.
[Not by lead acetate neutral.]
Potassium bichromate, (pp. is $F.CrO_4$]
[No coloration or pp. with chlorine.]

Colour tests:

Concentrated sulphuric acid, dark violet, changing gradually to brownish-green.
Concentrated sulphuric acid with potass. bichromate, brown (Reichwald obtained a green with violet streaks).
Nitre added to sulphuric acid solution, green~violet~yellow.
Nitric acid, colourless cold, yellowish-brown on evaporation.
Fröhde's reagent, violet~dark green.
Seleno-sulphuric acid, pure violet.
Vanadyl sulphate, emerald.

(*c*) *ADERMANN'S ALKALOID NO. I.* $C_{20}H_{21}NO_4$; M.P. 138°. Optically aclive; extracted by ether from *acid* solution.

Isomeric with hydro-berberine (see Berberis), and yielding Berberine on oxidation (disputed by Dobbie and Lauder).

(*d*) *ADERMANN'S CORYDALINE* [Fumarine? compare (*b*)], $C_{22}H_{21}NO_4$. An alkaloid resembling Caffeine, and giving following **colour tests:**

Concentrated sulphuric acid, yellow~violet.
" " " with potas. bichromate, colours like those with strychnine.
Fröhde's reagent, violet, or green streaked with violet.
Seleno-sulphuric acid, light violet.

(*e*) Adermann also isolated an amorphous alkaloid.

(*f*) *BULBOCAPNINE,* A., $C_{34}H_{36}N_2O_7$, compare (*b*). Freund and Josephy separated this alkaloid, as well as Corycavine, from commercial 'Corydaline' (which contained (*a*) also). M.P. 198°-199°; dibasic.

Precipitated by alkalies; soluble in excess.

(*g*) *CORYCAVINE,* A., $C_{23}H_{23}NO_5$; M.P. 214°-215°. Less soluble in absolute alcohol than is Corydaline.

Precipitated by alkalies; insoluble in excess.

§ 81. **CORYNOCARPUS** lævigata (Karaka tree); *Primulaceæ;* New Zealand. The nut. Investigator: W. Skey, *Chemical News*, **27**, 190 (1873).

KARAKIN, G. or B.; percentage composition, $C_{64\cdot48}H_{4\cdot48}O_{31\cdot12}$ (free from nitrogen). Crystallizes in needles arranged as stars; M.P. 100°; feebly acid reaction, bitter, poisonous, producing convulsions.

Soluble in alcohol and hot water, with difficulty in cold water, and *not* in ether or chloroform.

It is absorbed by charcoal.

Reactions:

Alkaline hydrates, Ammonia } dissolve.
[No pp. tannic acid.]
Fehling's solution, green pp., with reduction when in certain proportions.
No pp. Zincic-potassic-sulphocyanide.
" Mercuric-potassic iodide.
Concentrated sulphuric acid, dark rose on warming.
Acids generally, dissolve.

§ 82. **CRATÆGUS** Oxyacantha (Thorn); *Rosaceæ.* The bark. Leroy, J. *Chim. Méd.*, **17**, 3.

CRATÆGIN, B. Grayish white crystals; bitter.

Soluble in water, with difficulty in alcohol, not in ether.

§ 83. **CREPIS** fœtida; *Compositæ.* Investigator: Walz, *N. Jahrb. Pharm.*, **13**, 176.

CREPIN; bitter. **Soluble** in ether. **Not** precipitated by lead acetate, neutral or basic.

§ 84. **CUCUMIS** Melo, *Cucurbitaceæ.* Substance (*a*). Investigator: Torosiewicz, *Repert. Pharm.*, **45** 30. C. Prophetarum and Ecbalium officinale (which see). Substance (*b*), Winckler, *N. Jahrb. Pharm.*, **11**, 31.

(*a*) *MELONEMETIN,* B. Brown amorphous, emetic.

Soluble in water and alcohol, not in ether.

(*b*) *PROPHETIN,* G., $C_{20}H_{36}O_7$? Resinous, white amorphous, bitter.

Acids produce prophetein and sugar.

Soluble in alcohol in almost all proportions, also in ether, with difficulty in water.

Precipitated by tannic acid.

[Not by lead acetate, neutral or basic.]

Concentrated sulphuric acid, reddish-brown.

§ 85. **CYCLAMEN** Europæum; Primula veris (small quantity in root); Anagallis arvensis (very small quantity); Limosella aquatica. *Primulaceæ.* Investigators: Saladin, *Journ. Chim. Méd.*, **6**, 417; Martius, *N. Repert. Pharm.*, **8**, 388; Hilger and Mutschler, *Ann. Chem. Pharm.*, **185**, 214, and others.

(*a*) *CYCLAMIN,* G., (Primulin, Arthanitin), $C_{20}H_{34}O_{10}$ or $C_{36}H_{56}O_{18}$ (G. Michaud), crystalline (Saladin), or amorphous (De Luca); M.P. 236°; neutral reaction; sharp taste; feebly lævo-rotatory. The powder causes sneezing. Acids give Cyclamiretin and sugar. Resembles Saponin.

Soluble in water after exposure to moist air; the solution is frothy; dissolved also by methyl and amyl alcohols, acetic ether, glycerine; sparingly in absolute alcohol (Mutschler gives solubility for 96 per cent. alcohol, 1 in 71). **Insoluble** in ether, chloroform, benzene, petroleum ether, or carbon bisulphide.

Precipitants:

Lead acetate, neutral.
Silver nitrate.
Copper sulphate.
Fehling's solution, white pp., no cuprous oxide formed.
[No pp. Cadmium-potassic iodide.]

Colour tests: Concentrated sulphuric acid gives a red solution from which water precipitates Cyclamiretin.

(*b*) *CYCLAMIRETIN*, G.-deriv. (see above), $C_{14}H_{16}O_6$ or $C_{15}H_{22}O_2$. Amorphous, resinous; M.P. 198°; tasteless, odourless; very poisonous.

Soluble in alcohol, not in water. In ether, Hilger states it to be soluble, but De Luca insoluble.

Precipitants:

Silver nitrate. | Fehling's solution, pp.

§ 86. **CYTISUS** laburnum, C. supinus, C. elongata. The seed; about 3 per cent., Ulex Europæus (Furze), about 0·2 per cent. of seed; also in bark and green buds. *Leguminosæ*. Investigators: A. W. Gerard, A. Partheil, etc.

CYTISINE, A. (=Ulexine, A. Partheil); $C_{11}H_{14}N_2O$; crystalline; M.P. 154° (152°-153°, Partheil). When cautiously heated, gives crystalline sublimate. Lævo-rotatory, $[a]_D = -119{\cdot}37$. Taste rather bitter and caustic. Causes dilatation of the pupil; powerful base, displaces ammonia; permanent in air. Salts usually soluble in water and alcohol; not easily crystallizable except the nitrate.

Soluble in nearly all proportions in water and alcohol, also dissolved by amyl alcohol, but scarcely or not at all in ether, benzene, carbon bisulphide, cold chloroform (?). (Partheil now states that the alkaloid is readily soluble in chloroform.)

Precipitants:

Alkaline hydrates, „ carbonates, Ammonia } dissolve.

[Not pp. by basic lead acetate.]

Tannic acid, pp. up to 1 in 300. Free base or alkaline solution.

Picric acid, 1 in 1,000.

Ferric chloride, see colour tests.

[Platinum chloride, not dilute, but yellow pp. in strong solutions.]

Gold chloride.

[Not potass. chromate.]

Phospho-molybdic acid, 1 in 10,000, acid solutions.

Phospho-tungstic acid, 1 in 30,000.

Iodo-potassic iodide, dark brown becoming crystalline, very dilute.

[Cadmium-potassic iodide, strong solution ?]

Mercuric-potassic iodide, 1 in 5,000 (the nitrate).

[Mercuric chloride, pp. with free base but not salts.]

[No pp. chlorine water.]

Bromine water, orange yellow pp.

Colour tests:

Concentrated sulphuric acid, colourless.

„ „ „ with potass. bichromate, yellow~brown~green.

„ „ „ with nitric acid, yellow.

Ferric chloride, red coloration; hydric peroxide added to this removes colour, but a blue tint appears on warming, recognisable with $\frac{1}{20}$ milligramme.

§ 87. **DAMIANA** (Turnera Aphrodisiaca).

DAMIANIN B. (name suggested); amorphous, light brown, bitter substance; free from nitrogen and non-glucosidal.

Soluble in water and alcohol. **Insoluble** in ether, chloroform, benzene, petroleum ether, or carbon bisulphide.

Not **precipitated** by lead acetate, neutral or basic, nor by other of the usual reagents.

§ 88. **DAPHNE** Mezereum (Mezereon bark), D. alpina; *Thymeleæ*. Investigators: Vauquelin, *Ann. Chim.*, **84**, 174; Zwenger, *Ann. Ch. Ph.*, **115**, 1; Rochleder, *Jnl. Pract. Ch.*, **90**, 442, and others.

(*a*) *DAPHNIN* G., $C_{15}H_{16}O_9 + 2H_2O$ (Zwenger), or $C_{30}H_{34}O_{19}$ (Rochleder); crys. in fine needles or rectangular plates; loses water at 100°; M.P. 200°, recrystallizing on cooling; at higher temperatures a sublimate of Daphnetin. Reaction neutral; taste bitter and astringent.

Soluble in hot water, hot alcohol, warm amyl alcohol, and warm acetic acid, also in chloroform; with difficulty in cold water or alcohol, and not in ether.

Removed by chloroform from acid solutions.

Reactions:

Alkaline hydrates, „ carbonates } dissolve yellow.

? Pp. Lead acetate neutral, Dragendf. (not lead acet. neutral, Rochleder).

Pp. Lead acetate basic.

Ferric chloride, a faint blue with concentrated solutions.

[Silver nitrate, no pp.; slight reduction on warming.]

Concentrated sulphuric acid with nitric acid, red (yields oxalic acid).

Nitric acid, red.

(*b*) *DAPHNETIN* G.-derivative [Dioxycoumarin ? (Stunkel)],

$2(C_9H_6O_4)H_2O$; crys. fine clino-rhombic prisms ; M.P. 253°, with sublimate and odour of Coumarin ; feebly acid reaction, astringent taste.

Soluble in hot alcohol and boiling water (yellow solution), with difficulty cold, or in ether, and not in benzene, chloroform, or carbon bisulphide.

Reactions :

Alkaline hydrates } dissolve red.
„ carbonates }
Lead acetate neutral, pp.
Ferric chloride, green coloration with concentrated solutions.
Silver nitrate, reduced.
Fehling's solution, reduced.
Concentrated sulphuric acid, on gently warming, dissolves yellow ; water reprecipitates unchanged.
Nitric acid, gradually intense red.
Hydrochloric acid, dissolves on warming, unchanged.

(c) *COCCOGNIN* B., $C_{20}H_{22}O_8$; crystalline ; sublimate and odour of Coumarin on heating.

Soluble in alcohol, with difficulty in water, not in ether.

§ 89. **DATISCA** Cannabina ; *Datiscaceæ*. Investigators : Braconnot, *Ann. Chim. Phys.*, [2] 3, 277 ; Stenhouse, *Ann. Chem. Pharm.*, 98, 166.

(*a*) *DATISCIN* G. (formerly mistaken for Inulin), $C_{21}H_{22}O_{11}$; crys. silky needles ; M.P. 186° ; neutral reaction, bitter. Yields Datiscetin and sugar on treatment with acids.

Soluble in alcohol and hot water ; with difficulty in cold water or ether.

Reactions :

Alkaline hydrates } deep yellow solutions ; acids reprecipitate.
„ earths }
Ammonia }
Lead acetate neutral } yellow pp., gelatinous.
„ „ basic }
[Tannic acid, no pp.]
Ferric chloride, brownish-green pp.
Copper sulphate, greenish pp.
Stannic chloride, pp.

(*b*) *DATISCETIN* (from above), $C_{15}H_{10}O_6$; colourless, tasteless, crystalline.

Soluble in alcohol and ether ; scarcely in water.

§ 90. **DAUCUS** Carota (Carrot) ; *Umbelliferæ*.

[*CAROTIN*, not bitter. Soluble with difficulty in water, but taken up by soap solution. Partly precipitated by lead acetate.]

[*HYDROCAROTIN* = Angelica bitter (Brummer) ; not identical with Cholesterin, but of same class (Hüsemann) ; $C_{18}H_{30}O$; crystalline ; M.P. 126·5° or 137° (Reinitzer) ; neutral reaction ; not bitter. Floats on water.

Soluble in hot alchol and in ether, benzene, chloroform, carbon bisulphide, acetone. Not in cold water or cold alcohol, but in soap solution.

Alkaline hydrates, no action.

Concentrated sulphuric acid, ruby-red.]

§ 91. **DELPHINIUM** Staphysagria, L. (Larkspur) ; *Ranunculaceæ*. The seeds (Stavesacre seeds), substances (*a*) to (*d*). D. consolida, alkaloid (*e*).

(*e*) *DELPHININE* A., $C_{22}H_{35}NO_6$ (?) ; rhombic crystals (from ether) ; alkaline reaction ; bitter and sharp taste ; M.P. variously given as 65°, 90°, 120°, 191·8° ; sublimate at 149° ; very poisonous. The nitrate and sulphate are sparingly soluble in water, alcohol, or ether.

Soluble in 1,594 parts cold water, 44 absolute alcohol, 53 ether, 20 benzene, 642 petroleum ether, and readily in chloroform.

Benzene and chloroform remove it from alkaline solution, and in traces from acid solution ; petroleum removes it slowly from alkaline solutions.

Precipitants :

Alkaline hydrates, pp. scarcely sol. in excess.
„ carbonates.
Ammonia, pp. scarcely sol. in excess.
Tannic acid, pp. (cloud 1 in 3,000).
Picric acid, amorphous yellow (1 in 3,000).
Platinum chloride, greenish-yellow (slight at 1 in 3,000).
Gold chloride, lemon-coloured pp.
[Not potassium bichromate.]
Phospho-molybdic acid, grayish-yellow.
Iodo potassic iodide, kermes coloured pp. (1 in 3,000).
Bismuth-potassic iodide, orange-red.
Cadmium-potassic iodide.
Mercuric-potassic iodide, yellowish-white.
Mercuric chloride, at 1 in 3,000, a cloud only at first.
Iodine tincture.

Colour tests:

Concentrated sulphuric acid	no effect, or light brown.
„ „ „	with sugar, no effect.
„ „ „	with bromine, violet.

[Not phospho-molybdic acid.]
[Phospho-antimonic acid, cloud at 1 in 1,000; sol. warm, reappearing cold.]
[Iodo-potassic iodide, cloud not permanent.]
[Bismuth-potassic iodide, cloud; pp. in concentrated solutions.]
[Not cadmium potassic iodide.
[Not mercuric-potassic iodide.
[Not iodine.]

Colour tests (authorities differ):

Concentrated sulphuric acid, green (greenish-brown, Dragendorff).
" " " with bromine, red~green with water.
" " " with nitric acid, pale yellow.
" nitric acid alone, colourless ~ yellow (Nativelle), emerald (Homolle.)
Fröhde's reagent, dark orange~cherry; in half-hour, brownish-black; in twenty-four hours greenish-yellow with black flocks (Dragendorff).

(1*a*) *DIGITOGENIN* G.-deriv., $C_{15}H_{24}O_3$. From preceding and from Digitonin.

Soluble in 100 cold or 35 parts boiling 93 per cent. alcohol, 30 cold or 20 boiling chloroform, and in 30 cold glacial acetic acid (Kiliani).

(2) *DIGITOXIN*. [The most important constituent of Digitalis (Kiliani); this denied by Arnaud, see (1).]

Soluble in hot alcohol and in chloroform; with difficulty in cold alcohol or ether (insol. Dupuy); and not in water or benzene.

(3) *DIGITONIN* G., $C_{31}H_{52}O_{17}$ (Houdas), or $C_{27}H_{46}O_{14}$ (Kiliani). Crystallizes from 85 per cent. alcohol, but amorphous from stronger alcohol; M.P. 235° (Kiliani); lævo-rotatory (−50° for a 2·8 per cent. solution in 75 per cent. acetic acid). Heated with hydrochloric acid on water bath for six hours, yields Digitogenin galactose and dextrose, one molecule of each.

Soluble in all proportions in water when amorphous, but the crystals only dissolve with difficulty: if heated they are more readily soluble, but the substance then gives no crystals on evaporation (Kiliani); sparingly in alcohol, and not in ether, chloroform, or benzene.

Precipitants: Ammonia.
Lead acetate, neutral or basic.
Tannic acid.

Colour tests:
Concentrated sulphuric acid, red; violet on dilution.

(4) *DIGITALEIN* G. (Neriin? see Nerium; Schmiedeberg's Digitalein contains, according to Kiliani, seven or eight substances with as much as 60 per cent. of Digitonin); non-nitrogenous, gummy, fusible, bitter.

Soluble in water (in all proportions, Dragendorff), in ethyl and amyl alcohols, but sparingly only in ether.

Precipitated by tannic acid.

Colour tests:
Concentrated sulphuric acid, no colour (reddish-brown, Dragendorff).
Hydrochloric acid at 20°, dissolves brownish-green.

§ 93. **ECBALIUM** elaterium (E. officinalis; Squirting Cucumber); *Cucurbitaceæ*. The dried juice (Elaterium album and nigrum). Investigator: Walz, *N. Jahrb. Pharm.*, **11**, 178, and others.

(*a*) *ELATERIN* B. (Elatin), $C_{20}H_{28}O_5$; crystalline; M.P. 200°; neutral reaction, bitter, purgative.

Soluble in ethyl and amyl alcohols, chloroform, and carbon bisulphide in 125 parts boiling water, with difficulty cold, in 290 ether, also sparingly in benzene, which removes it, however, from acid solutions.

Reactions:
In alkaline hydrates and ammonia, soluble (acids reprecipitate).
Insol. alk. carbonates.
[No pp. lead acetate neutral.]
Concentrated sulphuric acid, yellow to dark red.
" " " with 1 drop phenol, then more acid, crimson.
Fröhde's solution, yellow.

(*b*) *ECBALIN* B. (Elaterio acid; not to confuse with Ecbolin, see Ergot), $C_{20}H_{34}O_4$? (Walz); 'requires confirmation' (Hüsemann). Amorphous, resinous, bitter.

Soluble in water (20 parts), in alcohol and ether; also in alkaline hydrates. Nitric acid dissolves red with decomposition.

(*c*) *ELATERID* B., $C_{20}H_{32}O_{12}$? (Walz); bitter.

Soluble in dilute alcohol, not in water or ether.

Reactions: Alkaline hydrates, dissolve.
[No pp. lead acetate neutral or basic.]
Pp. tannic acid.

(*d*) *HYDRO-ELATERIN*, $C_{20}H_{30}O_6$? (Walz); yellow amorphous, not bitter.

Soluble in water, alcohol, and ether (which removes it from acid solutions).

Reactions: [Not precipitated by lead acetate neutral or basic.]
Pp. Tannic acid.

§ 94. **ECHUGIN** poison, from Adenium Boehmianum; *Apocynaceæ*; Africa. Investigator: R. Boehm, 1889.

ECHUGIN G. (about 10 per cent. of the poison); crys. silky rhombic plates; bitter, cardiac poison.
Soluble in water and alcohol; not in ether.

§ 95. **ELEMI** resins (from Canarium, Icica, etc.); *Amyridaceæ*. (The two following substances are included here for convenience of reference, but do not belong to the groups under treatment.)

(*a*) *AMYRIN*, $2(C_{10}H_{16})H_2O$ (Flückiger); crystalline; M.P. 177° with sublimate.
Soluble in alcohol (27 parts), and in ether, chloroform, carbon-bisulphide.

(*b*) *BRYOIDIN* (Baups' Breïdin; distinct from Bryonin, see Bryonia). $2(C_{10}H_{16})3H_2O$ (Flückiger); prismatic crystals; M.P. 135°, with sublimate.
Soluble in alcohol, ether, chloroform, acetic acid, and glycerine.
Dry hydrochloric acid gas gives red~violet~green.

§ 96. **EPHEDRA** monostachia and E. vulgaris, *var.* Helvetica; *Gnetaceæ*. Investigators: Nagai, Spehr.

(*a*) *EPHEDRINE* A. (the alkaloids so named, by Nagai and Spehr respectively, do not agree in properties; they are here described together, Ng. referring to Nagai's base, and Sp. to that of Spehr), $C_{10}H_{15}NO$ (Ng.), or $C_{13}H_{19}NO$ (Sp.); taste bitter (Ng.), burning (Sp.); M.P. 210° (Ng.), 112° (Sp.).
Soluble with difficulty in water, Ng. (very soluble, Sp.); readily in alcohol, Ng., Sp.; in chloroform easily, Ng., Sp. (1 in 11, Sp.); in ether easily, Ng. (1 in 98, Sp.); in benzene easily, Ng. (1 in 1,180, Sp.).

(*b*) *PSEUDO-EPHEDRINE* A., $C_{10}H_{15}NO$ (Nagai); M.P. 115°; bitter.
Soluble readily in alcohol, in 15 parts ether, 26 benzene; with difficulty in water.

§ 97. **ERGOT** of rye (Secale cornutum; Claviceps purpurea, Tulasne; Sclerotium clavus, D.C.; Spermœdia clavus, Fries); *Fungi*. Investigators, numerous.

In addition to the alkaloids described below, the following substances have been found in Ergot: Scleromucin (2 to 3 per cent.), Sclerotic Acid (1 to 4 per cent.), Cholesterin, Mycose, Mannite, Leucine, Methyl-, and Trimethyl-amine. Several colouring matters have been isolated by Dragendorff (*Dorpat. Naturf. Gesell.*, 4, 392), viz.: Scler-erythrin, Picro-sclerotin, Fusco-sclerotic Acid, Sclero-xanthin, and Sclero-crystallin. For Vernine, also present, see Vicia sativa.

(*a*) *ERGOTINE* A., $C_{50}H_{52}N_2O_3$ (Wenzell, *Amer. J. Pharm.*, 36, 193); amorphous, alkaline, feebly bitter; salts amorphous.
Soluble in water and alcohol; not in ether or chloroform.

Precipitants:
[Not lead acetate, neutral.]
Tannic acid.
[Platinum chloride, in an alcoholic+ether solution.]
Gold chloride.
[Not potass. cyanide.]
Phospho-molybdic acid, separated by this means from trimethylamine, which is not precipitated.
Mercuric chloride, but not in acid solutions (see Ecboline).

(*b*) *ECBOLINE* A.; amorphous, alkaline, slightly bitter. Salts amorphous.
Soluble in water and alcohol; not in ether or chloroform.

Precipitants:

[Not lead acetate, neutral.]	Potass. cyanide, white.
Tannic acid.	Phospho-molybdic acid.
Platinum chloride, yellow.	Mercuric chloride, in acid solution (see Ergotine above).
Gold chloride.	

Colour tests: Concentrated sulphuric acid, dark red.

(*c*) *ERGOTININE* A. (Tanret's), $C_{35}H_{40}N_4O_6$; fluorescent, crystalline, slightly bitter. Salts readily decompose. Acid solutions gradually redden. Alkalies evolve Methylamine.
Soluble in alcohol (becoming green, then brown), also in chloroform and ether (but with difficulty in the latter after exposure); not in water.
It is removed from *acid* solutions by ether and chloroform.

Precipitants:

Tannic acid.	Iodo-potassic iodide.
Platinum chloride.	Mercuric-potassic iodide.
Gold chloride.	[Not mercuric chloride.]
Phospho-molybdic acid.	Bromine water.

Colour tests:
Concentrated sulphuric acid with $\frac{1}{7}$ water, reddish-violet in presence of ether.
Concentrated sulphuric acid with sugar, red, then brown.
Fröhde's reagent, violet~blue.

(*d*) *CORNUTINE* A.; only stable when combined with acid. The aqueous solution decomposes.

Soluble in alcohol.

§ 98. **ERYTHREA** Centaureum (Common Centaury), E. chilensis (Chironia chilensis), Sabattia vulgaris (Chironia angularis); *Gentianaceæ*. Investigators: Méhu, *Journ. de Pharm.*, [4], **3**, 265, etc.; also *Ph. J., Trans.*, [3], **1**, 990; Lendrich, and others.

(*a*) *ERYTHROCENTAURIN* G., $C_{27}H_{24}O_8$? ($C_9H_{14}O_6$, Lendrich); crystalline needles (amorphous, terebinthinate, Lendrich); M.P. 136°; neutral reaction, bitter? (tasteless, Méhu). Reddened by light, but colourless on re-solution; not fluorescent.

Soluble in 1,630 parts cold water or 35 boiling, in 48 of 86 per cent. alcohol, 13½ chloroform, 245 ether, and in benzene, carbon bisulphide, fatty and essential oils.

Reactions:
Alkaline hydrates, Ammonia } dissolve.
No pp. lead acetate, neutral or basic.
Nor tannic acid, Méhu (pp., Dragendorff).
No pp. metallic salts.
Not affected by bromine.
Potass. permanganate, reduced.

No colours or effect with concentrated sulphuric, nitric, or other acids—even chromic.

Solution in sulphuric acid gives odour of Menyanthol on warming.

§ 99. **ERYTHROPHLŒUM** guinense (Sassy tree, Doom or Ordeal bark); *Leguminosæ*. Investigators: Gallois and Hardy, *Pharm. J. Trans.*, [3], **7**, 77.

ERYTHROPHLŒINE A. (the 'Muawine' from 'Muawi'?); crystalline, alkaline, poisonous (acts on heart). Crystalline salts.

Soluble in water, alcohol, acetic ether, and carbon bisulphide; with difficulty in ether, chloroform, or benzene.

Removed from aqueous solution by acetic ether.

Precipitants:

Alkaline hydrates, white.	Iodo-potassic iodide, yellowish-red.
Ammonia, white crystalline.	Bismuth potassic iodide, yellow.
Gold chloride, white.	Cadmium potassic iodide, Mercuric-potassic iodide, Mercuric chloride } white.
Potass. bichromate, yellow.	
Phospho-molybdic acid, dirty green.	

Colour test: Concentrated sulphuric acid with permanganate, violet.

§ 100. **EUGENIA** pimenta (Myrtus cheken—yielding Cheken-leaves—Allspice, pimento); *Myrtaceæ*. The leaves. Investigator: J. Winters England, *Amer. J. Ph.*, 1883, 246.

(*a*) *CHEKENIN* B. (Cheken bitter); very bitter, unpleasant odour, not poisonous.

(*b*) *CHEKENETIN*, $C_{11}H_7O_6+H_2O$; yellowish-green needles, giving green solutions with alkalies, changing to bluish-violet and red.

§ 101. **EUONYMUS** Europæus (Spindle tree), E. atropurpureus (Wahoo bark). Investigator: A. B. Prescott, *Amer. J. Ph.*, 1883 and 1889.

EUONYMIN G.; amorphous, bitter, odourless.

Soluble in alcohol and petroleum ether, slightly in water and ether · not in carbon bisulphide or benzene.

Precipitants (with alcoholic solution)·
Tannic acid, slight white pp.
Picric acid, pp., but not at once.
Sodium phospho-molybdate, pp. greenish-yellow, becoming blue with ammonia.
Iodo-potassic iodide, reddish-brown.
Mercuric-potassic iodide, white.

Colour tests:
Concentrated sulphuric acid, yellow~reddish-brown (potass. bichromate intensifies the colours).
" nitric acid, " hydrochloric acid } yellow.

§ 102. **EUPATORIUM** cannabinum (*Compositæ—Tubulifloræ*). Investigators: Shamel (*Amer. J. Pharm.*, 1892, **14**, 224), isolated (*a*); G. Latin (*Pharm. J. Trans.*, [3], **11**, 192), substance (*b*).

(*a*) *EUPATORINE*, A. $C_{20}H_{25}O_{26}\cdot HNO_3$? (Shamel). Powder consisting of microscopic needles; alkaline reaction; bitter; not fusible without decomposition. Sulphate crys. silky needles.

Soluble in alcohol, ether, chloroform, not in water.
Removed from alkaline solutions by ether.
Reactions (with the nitrate):
Alkaline hydrates, Ammonia } soluble.
Picric acid, crystalline pp., soluble in alkalies, deep red.
[Gold chloride, slight coloration.]
[Phospho-molybdic acid, green coloration.]
'No reactions with other alkaloid precipitants.'—Shamel.
Concentrated sulphuric acid, insoluble.
Nitric acid, dissolves.
Hydrochloric acid, insoluble.

(*b*) *EUPATORIN*, G. Acid taste. Gives sugar and a red substance on boiling with acids—an odour of raspberries being evolved.
Soluble in hot water, and in alcohol, ether, chloroform.
Not precipitated by neutral lead acetate.
Colour tests:
Concentrated sulphuric acid, dissolves, reddish-brown.
„ nitric acid, „ hydrochloric acid } light yellow solution.

§ 103. **EUPHORBIA** resinifera; *Euphorbiaceæ*. Yielding euphorbium resin. Investigator: Flückiger, *Viertelj. Pract. Pharm.*, **17**, 82, and others.

EUPHORBON, $C_{15}H_{22}O$ (Flückiger), or $C_{15}H_{24}O$ (O. Hesse). Powder or crystals (needles from ether or benzene, and prisms from chloroform); M.P. 106°-116°; neutral reaction; burning taste; no odour.
Soluble in hot alcohol (crystallizes out on cooling), also in ether, chloroform, benzene, amyl alcohol, but requiring 38,000 parts of cold water for solution.
Alkalies, Ammonia } dissolve to only slight extent.
Colour tests:
Concentrated sulphuric acid, yellowish-brown.
„ „ „ with nitric acid, violet.
Acids in general exert very little solvent action.

§ 104. **FRAGARIA** vesca (Strawberry); *Rosaceæ*. The root.

FRAGARIAMARIN, G. Yields sugar and a red amorphous substance, FRAGARIN, on treatment with acids. Fragariamarin is soluble with difficulty in water, alcohol, or ether.

§ 105. **FRAXINUS** excelsior (Ash), *Oleaceæ*; Æsculus hypocastaneum (Horse-chestnut) and Æ. Pavia, *Sapindaceæ*. Mem.: For other constituents of horse-chestnut, see Æsculus; and for Quercitrin (contained in the leaves of the ash), see Quercus tinctoria. Following from the root·

(*a*) *FRAXIN*, G. (Paviin),

$$C_{27}H_{30}O_{17}, \text{ or } C_{16}H_{18}O_{10} = OCH_3 \cdot C_6H(OH)\,(O \cdot C_6H_{11}O_5) < \begin{matrix} CH{:}CH \\ | \\ O - CO \end{matrix}$$

(Koerner and Beginelli). Crystals resembling zinc sulphate (Rochleder); M.P. 320°; neutral, feebly bitter. Acids give Fraxetin and sugar. The ammoniacal solution shows bluish-green fluorescence.
Soluble in 1,000 parts cold water, more readily hot, with difficulty in cold alcohol but more easily warm, slightly in ether to which it communicates fluorescence.
Reactions:
Alkaline hydrates, „ carbonates, Ammonia } dissolve to yellow solutions, fluorescent.
Lead acetate, neutral or basic, to alcoholic solution, yellow pp.
„ „ aqueous solution, no pp.
Concentrated sulphuric acid, yellow.
Ferric chloride, green~yellow pp.

(*b*) *FRAXETIN*, G.-derivative.

$$C_{15}H_{12}O_8 \text{ or } C_{10}H_8O_5 = OCH_3 \cdot C_6H(OH)_2 < \begin{matrix} CH{:}CH \\ | \\ O - CO \end{matrix} \text{ (Koerner).}$$

Yellow crystalline needles and plates with $1\frac{1}{2}H_2O$, anhydrous at 100° C.; M.P. 227° (Koerner), re-crystallizing on cooling. Acid reaction; not bitter.
Soluble in 10,000 parts of cold water, or 33 on boiling, with difficulty in alcohol, scarcely in ether.
Colour tests:
Concentrated sulphuric acid, yellow.
„ nitric acid, dark violet~red~yellow~colourless.

§ 106. **FRITILLARIA** Imperialis (Crown Imperial); *Liliaceæ*. The bulbs, 0·08 to 0·12 per cent.

IMPERIALINE, A., $C_{35}H_{60}NO_4$? (K. Fragner). Crys. needles, lævorotatory ($[a]_D = -35{\cdot}4°$ in chloroform); M.P. 254° (browns at 248°); very bitter. Hydrochloride solution is fluorescent, the sulphate hygroscopic.

Soluble slightly in cold water, readily in chloroform, also in hot alcohol, but sparingly in amyl alcohol, ether, benzene or petroleum ether.

Precipitants:

Alkaline hydrates.
,, carbonates.
Tannic acid, yellowish-gray flocculent.
Picric acid, yellow flocculent.
Potassium bichromate, yellow crystalline.
Iodo-potassic iodide, dark yellow amorphous.
Bismuth-potassic iodide, orange.
Cadmium-potassic iodide, white flocculent.
Mercuric-potassic iodide, reddish-yellow.

Colour tests:

Concentrated sulphuric acid, pale yellow.
,, ,, ,, with sugar, yellowish-green~flesh colour ~cherry red, finally dark violet after long exposure.
Nitre added to solution in sulphuric acid—dark reddish-yellow.
Nitric acid, yellow.
Hydrochloric acid, gradually brownish-green on warming~red on longer heating. Solution is fluorescent.
Potassic perchlorate and sulphuric acid—orange.
Fröhde's reagent, greenish-yellow.

§ 107. **GARCINIA** mangostana (Mangosteen); *Guttiferæ* or *Clusiaceæ*. The shells of the fruit. Investigator: W. Schmid, *Ann. Chem. Pharm.*, 93, 83, and P. B. Liechti.

(*a*) *MANGOSTIN* B., $C_{20}H_{22}O_5$ (Liechti); yellow crystalline plates; M.P. 190° (173° Liechti), with sublimate; neutral reaction, tasteless, odourless. Separable with difficulty from the resin accompanying it.

Soluble in alcohol, ether, chloroform, carbon bisulphide, glacial acetic acid, and acetone; not readily in benzene, and **insoluble** in water or petroleum ether.

Reactions:

Alkaline hydrates dissolve yellow.
[No pp. neutral lead acetate.]
Pp. basic lead acetate.
Ferric chloride, dark green to black coloration.
Platinum chloride, reduced.
Gold chloride, reduced.

Not precipitated by other metallic salts.
Concentrated sulphuric acid, dissolves yellowish-red.
Nitric acid decomposes with production of oxalic acid.

(*b*) *ISOMANGOSTIN* $(C_{20}H_{22}O_5)^n$; prepared from the preceding substance; dark brown amorphous. The alkaline alcoholic solution shows green fluorescence.

§ 108. **GARDENIA** lucida (G. balsamifera, yielding Decamalee Gum); *Cinchonaceæ* or *Rubiaceæ*. Investigators: Stenhouse and Grove, *J. Chem. Soc.*, 1877.

GARDENIN B., $C_{14}H_{12}O$; dark yellow glittering crystals; M.P. 163°-164°.

Soluble in alcohol, easily in ether, scarcely in water, and not in petroleum ether.

Reactions: Alkaline hydrates do not dissolve.
Concentrated nitric acid, momentary crimson.
Hydrochloric acid, dissolves on warming.

§ 109. **GASTROLOBIUM** bilobum; *Leguminosæ*. Investigator: Rummel, *J. Chim. Min.*, 1880, 1,032.

GASTROLOBIN G.; saffron odour, hygroscopic.

Soluble in boiling water and boiling alcohol.

§ 110. **GENTIANA** lutea (Gentian); *Gentianaceæ*. The root. Investigator: Kromayer, *Arch. Pharm.*, [2], 110, 27, and others.

(*a*) *GENTIOPICRIN* G. (not Gentianin), $C_{20}H_{30}O_{12}+H_2O$; crystals only obtainable from fresh root (4 gms. from 6 lbs.); yellow, neutral, bitter; M.P. (when anhydrous) 120°-125°; acids give Gentiogenin and sugar.

Soluble in water, alcohol (with difficulty if absolute), chloroform, difficultly also in benzene, and sparingly in ether.

It is removed from acid solutions by chloroform in part, by benzene with difficulty, and in traces by ether.

Reactions: Absorbed by charcoal.
Alkaline hydrates dissolve yellow.
Ammonia, warm, dissolves yellow.
[Not precipitated by lead acetate, neutral.]
Ammoniacal lead acetate. pp. (Dragendorff); no pp. (Hüsemann)?
[Ferric chloride, no change.]
Silver nitrate ammoniacal, reduction and deposit of silver.
[Fehling's solution, not reduced.]

Concentrated sulphuric acid dissolves yellow on warming; the alkaline solution on treatment with sulphuric acid becomes first colourless, then carmine on warming.

(*b*) *GENTIOGENIN* G.-derivative; from above by boiling with dilute acids; yellowish-brown amorphous powder (permanent); taste bitter; reaction neutral.

Soluble in alcohol and in a mixture of ether with alcohol, but difficultly in water.

(*c*) *GENTIANIN* (gentianic acid; Gentisin), $C_{14}H_{10}O_5 = C_{13}H_5O_2(OCH_3)(OH)_2$; yellow crystals, neutral reaction, not bitter.

Soluble scarcely in water, and with difficulty in alcohol or ether.

Alkalies give red coloration.

(*d*) *GENTISEIN* = Gentisin less 1 methyl group = $C_{13}H_8O_5$; yellow crystals; acids give deep red precipitates; dyes wool yellow (with alumina mordant).

§ 111. **GEUM** urbanum; *Rosaceæ*. Investigator: Buchner, *Repert. Pharm.*, **85**, 184.

GEUM BITTER; amorphous, yellow, neutral.

Soluble in alcohol and ether, with difficulty in water.

Forms compounds with potash, lime and lead oxide that are soluble in alcohol.

Dissolved by alkalies.

§ 112. **GLEDITSCHIA** triandra; *Leguminosæ*.

GLEDITSCHINE A.; resinous amorphous powder.

Soluble in alcohol, not in water. Salts crystalline.

§ 113. **GLOBULARIA** alypum ('Wild Senna' of Germany), *Selaginaceæ*. Investigators: Walz, *N. Jahrb. Pharm.*, vols. 7 and 13; Heckel and Schlagdenhauffen, *Ann. Chim. Phys.*, [5], **28**, 67.

(*a*) *GLOBULARIN* G., $C_{30}H_{44}O_{14}$ (Walz), or $C_{15}H_{20}O_8$ (Heckel and Schlagdenhauffen); amorphous, bitter, neutral powder. Acids give rise to sugar and *Globularetin*, $C_{12}H_{14}O_3$ (Walz), or C_9H_6O (Schlagdenhauffen); a substance yielding cinnamic acid on decomposition.

Globularin is **soluble** in water and alcohol, but not in ether.

Precipitated by tannic acid.

GLOBULARESIN (not the derivative of Globularin) is an agreeable smelling resin.

§ 114. **GLYCYRRHIZA** glabra (Liquorice), G. echinata, Chrysophyllum glycyphlœum; *Leguminosæ—Papilionaceæ;* also in **Juglans** regia (walnut).

GLYCYRRHIZIN; $C_{44}H_{63}NO_{18}$ (an acid ammonium salt of Glycyrrhizic acid); prismatic needles from glacial acetic acid solution; M.P. 200°; optically inactive, acid reaction, bitter-sweet taste.

Soluble in hot water (gelatinous cold), in alcohol (sparingly if absolute), and in warm ether (Dragendorff says almost insoluble).

Reactions:

Alkalies dissolve reddish-yellow, with characteristic odour.

Ammonia, reddish-yellow solution.

Lead acetate, neutral or basic, precipitates (the alcoholic solution).

Salts of heavy metals give pp. (Habermann).

[Cadmium-potassic iodide, no pp.]

Sulphuric acid, precipitates from aqueous solution.

§ 115. **GONOLOBUS** condurango; *Asclepiadaceæ*. The bark. Investigators: G. Carrara, see J. *Ch. Soc.*, **91**, 1387; also Vulpius. (A wax $C_{30}H_{50}O_2$, M.P. 52°, and Cinnamic acid are also present.)

GONOLOBIN G. (Carrara's Glucoside), $C_{40}H_{74}O_8$; yellow powder; M.P. 112°.

Soluble slightly in water and alcohol; not in ether or petroleum ether.

Not **precipitated** by iodo-potassic iodide or mercuric-potassic iodide.

Another glucoside apparently also present, but details as yet discordant.

§ 116. **GRATIOLA** officinalis (Hedge Hyssop); *Scrophulariaceæ*. Investigator: Walz, *Jahrb. Pharm.*, vols. 14, 21, 24; and *N. Jahrb. Pharm.*, vol. 10.

(*a*) *GRATIOLIN* G., $C_{42}H_{36}O_{14}$ or $C_{40}H_{34}O_{14}$; crys. needles from water; M.P. 200° without change; bitter, poisonous (slows heart's action). Acids split into *Gratioletin*, *Gratiolaretin*, and sugar.

Soluble in 893 parts cold water or 476 boiling, in alcohol, chloroform, benzene, and in 1,000 parts cold or 666 boiling ether.

It is removed from acid aqueous solutions by chloroform and benzene, but not by ether.

Reactions:

Ammonia dissolves (water reprecipitates).

[Not precipitated by lead acetate, neutral or basic.]

Precipitated by tannic acid.

Colour tests:

Concentrated sulphuric acid, orange~brown~red at edges (water reprecipitates).

Concentrated nitric acid, yellow (water reprecipitates).

(*b*) *GRATIOSOLIN* G., $C_{18}H_{16}O_{10}$ or $C_{46}H_{42}O_{25}$; red or yellow powder; amorphous or crystalline; M.P. 125°: bitter, with characteristic odour;

poisonous. Acids readily change into sugar and *Gratiosoletin*, which is very bitter and gives precipitate with tannic acid.

Solubility (of Gratiosolin) 1 in 7 parts cold or 5 of boiling water, 3 cold or 2 boiling alcohol, 1,700 cold or 1,100 boiling ether.

Reactions:

[No precipitate by lead acetate, neutral or basic.]
Tannic acid, pp.

Colour tests:

Concentrated sulphuric acid, brownish-red } (water precipitates
Nitric acid, dissolves } yellow).

§ 117. **GUACHAMANIA** toxifera; *Apocynaceæ*.

GUACHAMANINE A.; amorphous, yellow, resinous, alkaline, bitter. Salts amorphous.

Soluble in water, not in alcohol.

Precipitated by alkaloid reagents.

§ 118. **HARMALA** (Peganum Harmala); *Zygophyllaceæ*. The seeds —about $2\frac{2}{3}$ per cent. (*a*), and $1\frac{1}{3}$ per cent. (*b*). Fritsche, *Ann. Chem. Pharm.*, vols. 64, 68, 72, 88, 92.

(*a*) *HARMALINE* A., $C_{13}H_{14}N_2O$; colourless rhombic octahedra, but salts yellow and fluorescent; M.P. 238° (with decomposition); alkaline reaction, slightly bitter (saliva becomes coloured yellow.

Soluble in 1,600 parts water at 0° C., in hot alcohol, with difficulty in cold alcohol and ether, slightly in petroleum ether and turpentine.

Precipitants:

Alkaline hydrates.
„ carbonates.
Ammonia hydrate.
„ carbonate.
Platinum chloride, light yellow microscopic crys.
Potass. ferrocyanide, red.
Potass. ferricyanide, pp. oily, changing to greenish-blue crys.
Potass. sulphocyanide, yellow crystalline.
Potass. bichromate (acid soln.), pp. oily, becoming crystalline.
Mercuric chloride, white.

Acids and salts precipitate from solutions.

(*b*) *HARMINE* A., $C_{13}H_{12}N_2O$; colourless, glittering rhombic prisms; salts either colourless or slightly yellow, showing blue fluorescence; M.P. 256°-257°, with sublimate; reaction alkaline; taste hitter (in alcoholic solution, but not the solid).

Soluble with difficulty in water, alcohol (cold), or ether; more easily in hot alcohol.

Precipitants:

Alkaline and ammonia hydrates and carbonates.
Platinum chloride, yellow amorphous, becoming crystalline.
Potassium sulphocyanide, white crystalline.
„ chromate, pp. decomposing with formation of bichromate.
„ bichromate.
Mercuric chloride, white.

§ 119. **HEDERA** helix (Ivy); *Araliaceæ*. Investigator: Vernet, *Berichte d. Ch. Ges.*, **14**, 685, and others. (For Carotin, contained in the leaves, see Daucus.)

(*a*) *HELIXIN* G. (Hedera Glucoside), $C_{32}H_{54}O_{11}$ (Vernet), or $C_{32}H_{52}O_{10}+2H_2O$ (H. Block); crys. needles; M.P. 233°; lævo-rotatory ($[a]_D=-47{\cdot}5°$); acid reaction, slightly bitter-sweet taste. Acids convert to a non-fermentable sugar, which reduces Fehling's solution, and a substance $C_{26}H_{44}O_6$; see (*b*).

Soluble in alcohol, hot benzene, and hot acetone; with difficulty in ether and cold benzene, and not in water, chloroform, or petroleum ether.

Reactions:

Alkaline hydrates dissolve with green colour when hot.
Ammonia, dissolves yellow.
Lead acetate, neutral pp.
Ferric chloride, green coloration.
[Fehling's solution not reduced till the Helixin has been boiled with acid.]
Potassium sulphocyanide, rose coloration (not permanent).
[Mercurous nitrate, no action.]

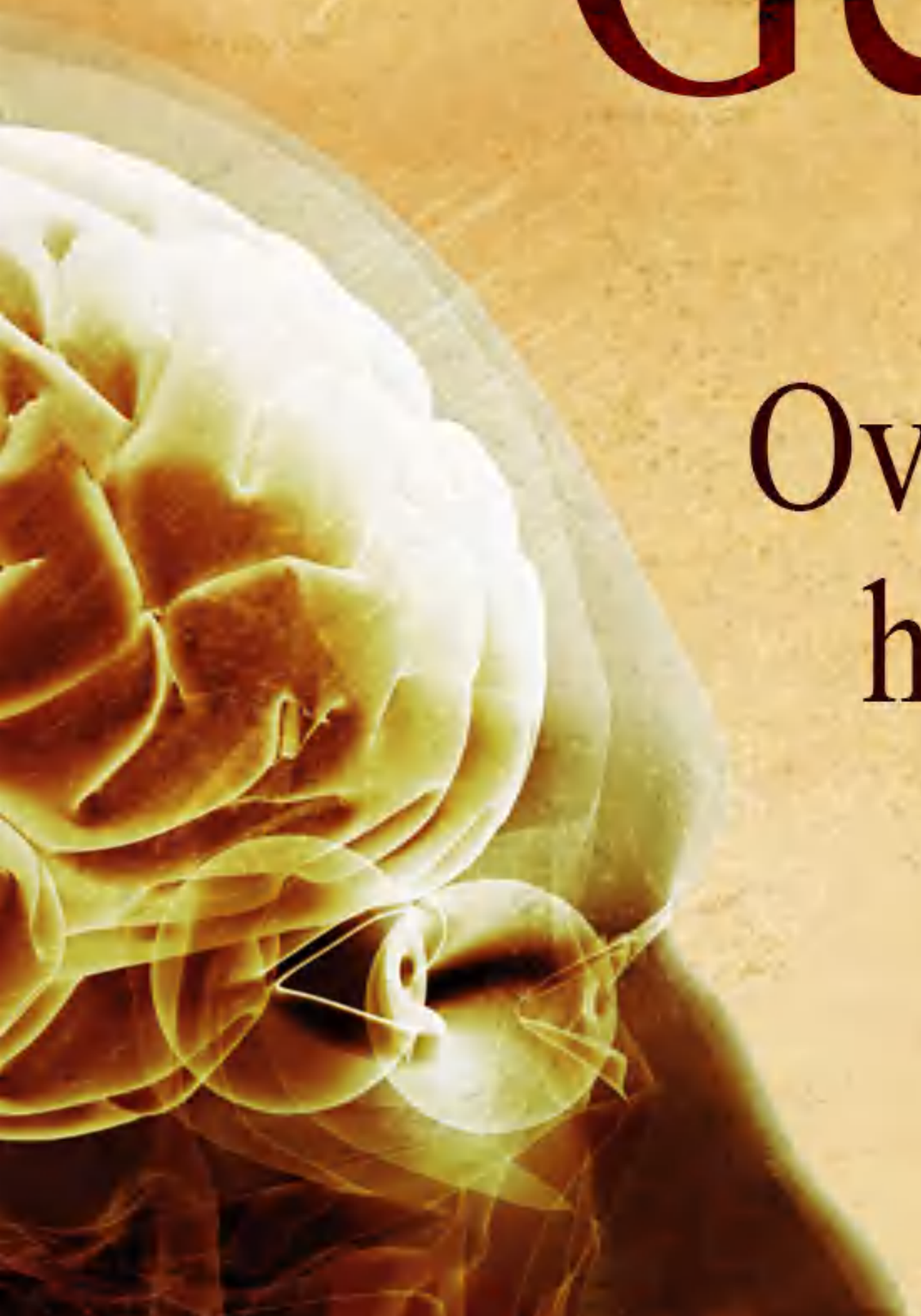
Colour tests:

Concentrated sulphuric acid, bright red.
„ „ „ with potassium bichromate, red, becoming green.
Nitric acid, no colour.
Hydrochloric acid, yellow.

(*b*) *HELIXIGENIN* (suggested name), from preceding substance by action of acids, $C_{26}H_{44}O_6$; M.P. 276°-280°; dextro-rotatory ($[a]_D=+42{\cdot}6°$).

Solubility similar to that of Helixin.

§ 120. **HELIOTROPUM** Europæum. Cynoglossum vulgare, etc.; *Boraginaceæ*. Investigators: Battandier, *Repert. Pharm.*, vol. 32; Schlagdenhauffer, Buchheim, and others.



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softer and has hop odour, presumably due to admixture with hop-oil, for, on adding water to an alkaline alcoholic solution of the resin, a precipitate forms having hop odour, from which the resin separable from the filtrate by acidification is free.

Solubility as Alpha-resin.

Reactions:

Lead acetate, alcoholic, no pp.
„ „ aqueous, pp.
Other reactions as Alpha-resin (with the cupric test, the ether becomes coloured emerald-green).

(4) *GAMMA-RESIN* (Hayduck's); hard, brittle, dark brown; *not bitter*.

Solubility as Alpha and Beta, except that it is not dissolved by petroleum ether.

Reactions:

Lead acetate (alcoholic), no pp.
Does not give the cupric reaction with ether.

§ 123. **HURA** crepitans ('Ajuapar'); *Euphorbiaceæ*. Investigator: Boussingault, *Ann. Chim. Phys.*, [2] 28, 430 (1825).

HURIN B.; oily, becoming crystalline; M.P. above 100°, but volatile at lower temperatures; vapours cause inflammation; no odour, but burning taste.

Soluble in alcohol, ether and oils, but not in water.
Unchanged by alkalies.
Resinified by nitric acid.

§ 124. **HYÆNANCHE** globosa (Toxicodendron capense); *Euphorbiaceæ*. The husk (about 3 per cent.). Investigator: Engelhardt, *Arbeiten d. pharmak. Inst. Dorpat*, 8, 1892.

HYÆNANCHIN B.; crystalline after purification; bitter. Physiological action like Strychnine, but much weaker. Non-glucosidal, but destroyed by acids or alkalies at 100° C.

Soluble in water, alcohol (the best solvent), ether, benzene, fatty oils.
Not precipitated by neutral lead acetate.

§ 125. **HYMENODICTYON** excelsum (Cinchona excelsa, Roxb.) and H. obovatum; *Rubiaceæ*. [See also Cinchona; and for Æsculin, which is said to have been found in this plant, refer to Æsculus. Naylor disputes the presence of Æsculin.] Investigator: Naylor, *Pharm. J. Trans.*, 1883.

(*a*) *HYMENODICTYONINE* A. (Hymenodictine), $C_{24}H_{40}N_2$; 'requires further study' (Hüsemann); amorphous or needles; M.P. 66°; non-fluorescent, reaction alkaline, taste bitter. Salts amorphous.

Soluble in alcohol, ether, chloroform, benzene, petroleum ether; difficultly in water.

Precipitants:

Alkaline hydrates } pp. gelatinous.	Phospho-molybdic acid, yellow.
Ammonia }	Potassium iodide.
Tannic acid, brown.	Iodo-potassium iodide, brown.
Picric acid, yellow.	Bismuth-potassium iodide, red.
Platinum chloride.	Cadmium-potassium iodide.
Potassium ferrocyanide.	Mercuric-potassium iodide, white.
„ ferricyanide.	Mercuric chloride.
„ sulphocyanide in excess, oily drops.	Also by sodium phosphate.
	„ „ chloride.
„ bichromate.	„ „ nitrate.

(*b*) *A bitter principle*, $C_{25}H_{49}O_7$; neutral.

Slightly **soluble** in hot alcohol. **Insoluble** in ether or chloroform.

§ 126. **ILEX** aquifolium (Holly), *Aquifoliaceæ*. The leaves.
[For Ilex Paraguayensis containing Theine see Thea group.]

(*a*) *ILIXANTHIN*, a colouring matter; $C_{17}H_{22}O_{11}$; pale yellow microscopic needles; M.P. 198°; non-sublimable, non-glucosidal.

Soluble in hot water and alcohol, scarcely in cold water, not in ether.

Reactions: Alkaline hydrates, orange pp.
Lead acetate, neutral or basic, yellow pp.
[Fehling's solution not reduced.]

(*b*) *ILICIN* B. Concerning this substance the statements of various investigators are discordant.

§ 127. **ILLICIUM** religiosum ('Sikimi'); *Magnoliaceæ*. The seeds. Investigator: Eyckman, *Pharm. J. Trans.*, 1881, 1,050.

SIKIMIN B. Nitrogen free; crystallizes in stellate prisms; the hydrochloric acid compound has M.P. 175°; non-glucosidal.

Soluble in hot water, alcohol, ether, chloroform, glacial acetic acid; with difficulty in cold water (not in petroleum ether?).

Reactions:

Alkaline hydrates } insoluble.
Ammonia }
[Not precipitated by neutral lead acetate.]

[Fehling's solution not reduced.]
Mercuric-potassic iodide, turbidity.
Hydrochloric acid warm, bluish-violet to green.

§ 128. **IPECACUANHA**, etc.; *Rubiaceæ—Cinchonaceæ, Violaceæ*, etc. Emetine in the following: Cephælis Ipecacuanha (Hooper found following quantities of (*a*) in plants grown in India: root 1·79, leaves 1·45, stalk 1·13 per cent.; none in seed—Flückiger); Ronabia emetica, Richard. (Psychotria emet., L.); Ionidium Ipec., Venten (Viola Ipec., L.); I. indecorum, St. Hil.; Richardsonia scabra, St. Hil. (Richardia scabra, L.); doubtful in Chiococca racemosa—*Rubiaceæ* (=Cainça root, see Chiococca). The substances (*b*) (c) in Viola odorata (violet) the root. [For 'White Ipec. from Isle de France' see Tylophora; for 'East India Ipec.' see Chamælirium.]

(*a*) *EMETINE* A. Numerous formulæ have been suggested by Lefort, Würtz, and others, viz.: $C_{28}H_{40}N_2O_5$, $C_{30}H_{22}NO_8$, $C_{35}H_{25}NO_9$, $C_{36}H_{25}NO_{10}$, $C_{37}H_{27}NO_{10}$, $C_{15}H_{22}NO_2$, $C_{30}H_{40}N_2O_5$, the latter by H. Kunz. Amorphous or crystalline (needles or plates); M.P. 50°-70° (different observers); optically inactive, bitter, alkaline, fluorescent (blue), emetic.

Soluble slightly in cold water, or 1 in 1,000 at 50° C., very readily in alcohol and chloroform; also in hot ether, hot petroleum ether, in benzene, methyl and amyl alcohols, oils, turpentine, acetic acid. Dissolved difficultly by ether and petroleum ether in the cold.

Removed from alkaline solutions by chloroform, amyl alcohol, and in traces by benzene and petroleum ether. Not extracted by solvents from an acid solution.

Precipitants:
Alkaline hydrates.
Ammonia.
Alk. carbonates, „ bicarbonates } insoluble in excess.
Tannic acid, pp. if a little free hydrochloric acid; 1 in 5,000.
Picric acid, yellow amorphous; 1 in 25,000.
Platinum chloride, yellowish-white; 1 in 2,500.
Gold chloride, lemon-yellow amorphous; 1 in 2,500.
Potass. ferrocyanide, 1 in 1,000.
„ sulphocyanide, yellow; 1 in 2,500.
„ chromate (avoid excess), 1 in 5,000.
„ bichromate, gradually yellow; 1 in 3,000.
Phospho-molybdic acid, yellow; 1 in 25,000.
Iodo-potassic iodide, Bismuth-potassic iodide } 1 in 25,000.
Cadmium-potassic iodide, yellow amorphous; 1 in 25,000.
Mercuric-potassic iodide, yellow amorphous; 1 in 25,000.
Mercuric chloride, white; 1 in 1,000; cloud in 3,000.
Potassic nitrate, Sodic „ } white.

Colour tests:
Concentrated sulphuric acid, dissolves slowly, dirty brown.
„ „ „ with potass. bichromate, brown.
„ „ „ with nitric acid, brownish-green.
Nitric acid, yellowish-brown.
Fröhde's reagent, red changing gradually to green; 1 in 100,000 (reliable, Dragendorff); concentrated hydrochloric acid added to this, deep blue.

(*b*) *VIOLINE* A. (possibly Emetine?); pale yellow amorphous, fusible, bitter, emetic.

Soluble in water more readily than Emetine; in alcohol less readily than Emetine; scarcely in ether.

(c) *VIOLA-QUERCITRIN* G., $C_{42}H_{42}O_{24}$; yellow crystalline; convertible by acids into 3 molecules of sugar and 1 of Quercetin (see Quercus tinctoria). Gives fluorescent solutions with alkalies.

§ 129. **ISATIS** indigofera (Indigo), I. tinctoria (Woad), *Cruciferæ;* Nerium tinctorium, Calanthe veratrifolia, Tankewillia cantonensis, Polygonum tinctorium. The leaves of Crotolaria retusa (L.), *Leguminosæ* (see also Leguminosæ Indian); also in the urine of man and ox.

INDICAN G., $C_{26}H_{31}NO_{17}$ (Schunk); amorphous, yellow, syrupy (decomposes on drying); acid reaction, bitter taste; decomposes in aqueous solution. Dilute acids give *Indigo*, C_8H_5NO, and Indiglucin, $C_6H_{10}O_6$, with other products.

Soluble in alcohol and water, difficultly in ether (insoluble, Dragendorff).

Reactions:
Alkalies decompose with production of Indiglucin, etc.
Precipitated by lead acetate neutral and basic.
Fehling's solution, reduced.

§ 130. **ISOPYRUM** thalictroides; *Ranunculaceæ.*

(*a*) *ISOPYRINE* A.; yellowish-white bitter powder; hydrochloride amorphous.

Soluble very readily in ether, not in water.

Precipitated by ammonia (not by ammonium chloride).

(*b*) *PSEUDISOPYRINE* A.; resembles above, except that hydrochloride is precipitated by ammonium chloride.

§ 131. **JUNIPERUS** communis (Juniper); *Coniferæ*. The berries. Investigator: Steer, *Wien. Akad. Ber.*, **21**, 383.

JUNIPERIN B.; light yellow, amorphous.

Soluble in 60 parts water, in hot alcohol, and in ether (which removes from aqueous solution).

Ammonia dissolves with golden yellow.

Concentrated sulphuric acid, dissolves with light yellow.

§ 132. **LACTUCA** sativa (Lettuce, contains also a trace of Hyoscyamine; see Atropa), L. virosa, L. altissima; *Compositæ—Ligulifloræ*. Investigators: Kromayer, Lenoir, Flückiger, Hesse, Franchemont, and others.

(*a*) *LACTUCIN* B., $C_{22}H_{18}O_7$ or $C_{22}H_{14}O_8$ (Kromayer); pearly crystals, fusible, bitter, neutral, non-glucosidal.

Soluble in hot water and in alcohol, with difficulty in cold water, not in ether.

Reactions:

Alkalies dissolve with gradual claret colour.

[Not precipitated by lead acetate, neutral or basic.]

Ammoniacal silver nitrate, reduced.

Fehling's solution, reduced.

Concentrated sulphuric acid, dissolves colourless, gradually becoming cherry red.

(*b*) *LACTUCOPICRIN* B., $C_{44}H_{32}O_{21}$ (Kromayer); brown, amorphous, bitter; reaction acid.

Soluble in water and alcohol.

(*c*) *LACTUCON* (Lenoir's), $C_{15}H_{24}O$ ($C_{16}H_{26}O$, Flückiger); crystallizes in stars; M.P. 185°; non-volatile except in current of carbon dioxide; tasteless, odourless; not physiologically active.

Soluble in alcohol, ether, petroleum ether, essential and fatty oils, scarcely in water.

Not precipitated by reagents soluble in alcohol.

(c) *GALLACTUCON* (Franchemont's 'Lacton,' from L. altissima), crystals; M.P. 296°.

Soluble in alcohol.

(*d*) *LACTUCERIN*, alpha and beta of O. Hesse. Isomeric with sycoceryl alcohol (Sycocerol), Hydrocarotin?, Quebrachol, Cupreol, Cinchol—possibly identical with Lenoir's Lacton. $C_{20}H_{32}O_2 = C_2H_3O \cdot C_{18}H_{29}O$.

Soluble in alcohol and petroleum ether.

§ 133. **LASERPITIUM** latifolium L.; *Umbelliferæ*. Investigators: Feldmann, *Ann. Chem. Ph.*, 135, 236; Kulz, Dissertation, Halle a. S., 1882.

(*a*) *LASERPITIN* B., $C_{15}H_{22}O_4$ or $C_{24}H_{36}O_7$; prismatic crystals; M.P. 114°-118°; bitter in alcoholic solution (not when dry), neutral reaction. Allied to Athamanthin (see Athamantha). Yields, on decomposition, Laserol and Angelic acid.

Soluble in 9 parts absolute alcohol, 3·6 ether, 12 carbon bisulphide, and in chloroform, benzene, petroleum ether, ethereal oils; not in water.

Reactions:

Alkaline hydrates, do not dissolve, or but slightly.

Lead acetate, neutral, precipitates (no pp.? Feldmann).

Concentrated sulphuric acid, dissolves red.

" nitric " " "

Dilute acids do not dissolve.

(*b*) *LASEROL* (aromatic principle), $C_{14}H_{22}O_4$, crystalline, or brown amorphons; taste, hot peppery.

Soluble in alcohol and ether.

Alkaline hydrates, dissolve yellow; acids precipitate from the alkaline solution.

§ 134. **LAURINEÆ** various—*e.g.*, Litsœa chrysocoma, Litsœa Javanica (about 1 per cent.), Tetranthera citrata, T. amara, T. lurida, T. intermedia, Notaphœbe umbellifera, Aperula sp., Actinodaphne procera, Illigira pulchra: Substance (*a*). In Hassia firma and H. squarrosa an alkaloid was found resembling Laurotetanine (M. Greshoff).

LAUROTETANINE A. Crystalline; very poisonous, action like Strychnine.

Soluble with difficulty in chloroform, scarcely in ether.

Precipitants:

Alkaline hydrates, soluble in excess.
Sodium carbonate.
Tannic acid.
Picric "
Platinum chloride.
Gold chloride.
Potassium sulphocyanide.
Phospho-molybdic acid.
" tungstic acid.
Mercuric-potassic iodide.
" chloride.
Iodine solution.

Colour tests :
Concentrated sulphuric acid, pale rose-red.
" nitric acid, dirty brown.
Fröhde's solution, indigo blue, yellow on addition of water.

§ 135. **LEGUMINOSÆ** Indian. Investigated by Greshoff.

Derris (Pongamia) elliptica—Benth. ; the root cortex, substance (*a*). Pachyrhizus angulatus—Rich. ; a substance 'Pachyrhizid,' possibly identical with (*a*). Erythrina (Stenotropis) Broteroï—Hassk. ; E. (Hypophorus) subumbrans (*b*). Crotolaria retusa contains Indican (see Isatis), and an alkaloid found also in C. striata, L. ; a strong poison, probably related to Cytisine, Sparteine, etc. Millettia atropurpurea—Benth., a poisonous glucoside resembling Saponin. Acacia tenerrima—Jungh., the cortex contains a bitter poisonous alkaloid, soluble in ether and chloroform. Pithecolobium bigeminum, P. saman ; substance (*c*), cortex of former gave 0·8 per cent.

(*a*) *DERRID* B. (Pachyrhizid ? see above). Nitrogen free ; slightly acid reaction ; sharp taste. Powerful fish poison at 1 in 5,000,000.

Soluble in alcohol, ether, chloroform, amyl alcohol, slightly in water.

Reactions :
Alkaline hydrates (solution), dissolves slightly.
" " on fusion, give salicylic and protocatechuic acids.

(*b*) *ERYTHRINE* A. Poisonous.

Soluble in ether. Sulphate crystalline.

Precipitants :

Tannic acid.	Potass. bichromate.
Picric acid.	Mercuric-potassic iodide.
Platinum chloride.	Mercuric chloride.
Gold chloride.	Iodine solution.
Platinum sulphocyanide.	

(*c*) *PITHECOLOBINE* A. Amorphous ; burning taste ; acts corrosively on the skin ; fish poison, 1 in 400,000. Salts crystalline. Milky solution with water.

Precipitated by :
Picric acid, 1 in 100,000.
Mercuric-potassic iodide, limit, 1 in 200,000.

§ 136. **LINARIA** vulgaris ; *Scrophulariaceæ*. Investigator : Walz., *Jahrb. Pharm.*, **27**, 16.

(*a*) *LINARIN* B. Crystalline, bitter.

Soluble in water and alcohol.

[Lead acetate, neutral or basic, no pp.]

Precipitated by tannic acid.

(*b*) *LINARACRIN*. Yellowish-brown, amorphous, resinous ; taste pungent.

Soluble in ether, not in water.

(c) *LINARESIN*. Yellow.

Soluble in alcohol, not in ether or water.

§ 137. **LINUM** catharticum (Purging Flax) ; *Linaceæ*. Investigators : Schroeder, *N. Repert. Ph.*, **11**, 11, and others.

LININ B. Percentage composition, C 62·92, H 4·72, O 32·36. Silky needles, chars on heating ; neutral reaction, bitter (in alcoholic solution), purgative.

Soluble in hot water, not cold ; almost in all proportions in absolute alcohol ; also dissolved by ether, chloroform, glacial acetic acid.

Removed from acid solutions by ether.

Reactions :
Alkaline hydrates, Ammonia } dissolve yellow (acids precipitate).
Concentrated sulphuric acid, dissolves dark violet.

§ 138. **LINUM** usitatissimum (Linseed) ; *Linaceæ*. Investigators : A. Jorissen and E. Hairs, *Journ. de Pharm. Anvers.* One kilo. of the germinated seeds yielded 15 gms. of the glucoside.

LINAMARIN G. Percentage composition, C 47·88, H . . . N 5·55, O . . . Silky needles ; **M.P.** 134° ; neutral reaction ; bitter, odourless. Yields hydrocyanic acid with the ferment (zymase) contained in powdered linseed.

Soluble in alcohol and in its own weight of water, but not in ether.
Fehling's solution, not reduced.
Concentrated sulphuric acid, dissolves colourless.

§ 139. **LIRIODENDRON** tulipifera, L. ('Tulip tree') ; *Magnoliaceæ*. The root-bark (yield, about 2 to 3 per cent.). Investigator : Emmet, *Journ. Ph.*, [2], **17**, 400 ; *Repert. Ph.*, **75**, 88.

LIRIODENDRIN B. ; crystallizes in needles and leaflets ; M.P. 82° (part sublimable) ; neutral, bitter ; oxidizes on exposure.

Soluble in alcohol, ether and hot water ; scarcely in cold water.

Precipitated by alkalies, acids, and neutral salts.

Concentrated sulphuric acid, dissolves orange-yellow ; water precipitates therefrom a non-bitter resin.

Nitric acid, dissolves colourless.

Hydrochloric acid hot, gives rise to a green substance.

§ 140. **LOBELIA** inflata L. ; *Campanulaceæ.* Investigators : F. Mayer, *Viertelj. pract. Pharm.*, **15**, 233 ; Lewis, *Pharm. J., Trans.*, [3], **8**, 561.

(*a*) *LOBELIINE* A.-G.* (Lobelina) ; oily yellowish-white, alkaline, taste acrid. Yields sugar on long boiling. Salts crystalline (not the acetate).

Soluble in water (with yellow colour), in alcohol, amyl alcohol, chloroform, ether (the best solvent), benzene, petroleum ether, carbon bisulphide, fatty and essential oils.

Contrary to the behaviour of substances of its class, it is removed from acid solution by petroleum spirit and chloroform.

Precipitants:

[Alkaline hydrates, decompose with production of substance having aromatic odour.]
Ammonia, white pp.
Lead acetate, neutral, white needles (no pp. ?).
„ „ basic (no pp. ?).
Tannic acid, pp. solution in excess, or in ammonia.
[No pp. gallic acid.]
Picric acid.
Ferrous sulphate, brown pp.
Platinum chloride, pp. floats (aqueous solution).
Gold chloride, pp. insoluble in hydrochloric acid (aqueous solution).
Silver nitrate, white pp. sol. ammonia or nitric acid (aqueous solution).
Fehling's solution, reduced after hydrolysis of L.
Phospho-molybdic acid, yellowish-white ; ammonia changes blue, then colourless.
Metatungstic acid.
Iodo-potassic iodide (acid solution), brown.
Mercuric-potassic iodide (acid solution), pale yellow.
[Mercuric chloride, no change.]
[Albumen, not coagulated.]
Charcoal, absorbs.

Colour tests :

Concentrated sulphuric acid, reddish-brown ; intensified by potass. bichromate.
Fröhde's reagent, reddish-brown.

(*b*) *LOBELACRIN* is considered by Lewis to be a salt of Lobeliine, with an acid resembling gallic, and named Lobelie acid.

Alkaloidal glucoside.

§ 141. **LOLIAM** telumentum (poisonous darnel grass) ; *Graminaceæ.* The seed. Investigators: Ludwig and Stahl, *Archiv Pharm.*, [2], **119**, 59 ; P. Antze, and others.

LOLIIN G. ; dirty white, amorphous, bitter. Yields sugar and volatile acids. Forms crystalline salts with sulphuric and hydrochloric acids.

Soluble in water and alcohol, not in ether.

§ 142. **LONICERA** xylosteum, L. (Honeysuckle) ; *Caprifoliaceæ.* The berries. Investigator : Enz, *Viertelj. pract. Pharm.*, **5**, 196, and others.

XYLOSTEIN G. : crystalline needles ; M.P. 100° (fusible to colourless liquid, recrystallizing on cooling) ; neutral, slightly bitter, poisonous. Acids give sugar and a derivative of Xylostein.

Soluble in hot water (scarcely cold), in alcohol and ether (which removes it on shaking with aqueous solutions).

Not precipitated by neutral lead acetate.

Concentrated sulphuric acid, brown.

§ 143. **LOTUS** bark (Symplocos racemosa) ; *Styraceæ.* The bark gave 0·24 per cent. (*a*), and 0·02 per cent. (*b*).

(*a*) *LOTURINE* A. ; crystallizes in efflorescent prisms ; M.P. 234° ; alkaline reaction. Salts with mineral acids show bluish-violet fluorescence.

Soluble in alcohol, ether, chloroform, acetone ; scarcely in water.

The alcoholic solution is precipitated by potass. sulphocyanide.

(*b*) *COLLOTURIN* A. ; crystallizes in non-efflorescent prisms, giving sublimate at 234° ; alkaline reactions. Salts fluorescent as (*a*).

Soluble in alcohol and ether.

Precipitants :

Gold chloride, yellow flocculent.
Potass. sulphocyanide (to the alcoholic solution).

(c) *LOTURIDINE* A. ; amorphous ; bluish-violet fluorescence.

Potass. sulphocyanide gives no pp. (compare above).

(*d*) *CALIFORNINE* (of Winckler) contains the 3 alkaloids above. (For Californin, see Cinchona group.)

§ 144. **LUPINUS** albus, L. luteus, and others (*Leguminosæ*), substances (*a*), (*b*), (*c*), etc. ; L. angustifolius (Blue Lupines), (*d*), 165 gms. from 2 cwt. of the seed.

Arginine (*e*) is a proteid decomposition product found in etiolated Lupine cotyledons, as well as in Soja hispida and Cucurbita pepo, grown in the dark.

After fermentation, the following decomposition compounds from pro-

teid were obtained: Asparagine (the principal product), Phenyl-amido-propionic and Amido-valerianic acids, Leucine, Tyrosine, Xanthine, Hypoxanthine, Lecithin, Arganine.

(*a*) *LUPININ* G. (Lupiniin; not the alkaloid Lupinine, Baumert), $C_{29}H_{32}O_{16}{\cdot}7H_2O$ (Schulze and Barbieri); crystallizes in yellowish-white needles; bitter taste and fruity odour. Acids, and even boiling water alone, convert to sugar and *Lupigenin*, which is partly sublimable.

Soluble with difficulty in water or alcohol.

Reactions:

Alkaline hydrates, Ammonia } dissolve; acids reprecipitate.

Precipitated by basic lead acetate.

(*b*) *LUPINIDINE* A., $C_8H_{15}N$; pale yellow, heavy oily, alkaline; pungent, bitter.

Soluble in water and alcohol; with difficulty in ether.

(c) *LUPININE* A., $C_{21}H_{40}N_2O_2$ (Liebscher, *Ber. landwirths. Inst. Halle*, 1880); rhombic crystals; M.P. 67°-80°; B.P. 257°-8°; volatile in steam, lævo-rotatory, apple-like odour, alkaline reaction, bitter.

Soluble in cold water (*less soluble warm*), in alcohol, ether, chloroform, benzene, petroleum ether; dissolved, but decomposed, by carbon bisulphide.

Precipitated by alkalies (concentrated solution).

(*d*) *LUPANINE* A., $C_{15}H_{24}N_2O$ (Hagen and C. Siebert); light yellow syrupy liquid, even at −16° C.; non-volatile; does not boil even at 290° C.; fluorescent, green; intensely bitter and powerfully alkaline. Liberates ammonia and gives cloud with hydrochloric acid vapour. Unchanged by boiling potash solution, or by hydrochloric acid at 200° C.

Soluble in cold water, nearly all precipitated on warming (Siebert found it to be freely soluble; dissolved also by ether and chloroform, but only difficultly by alcohol.

Removed from alkaline solutions by ether.

Precipitants:

Lead acetate neutral, if concentrated; after precipitation it is difficultly soluble even dilute.
Tannic acid, white.
Gold chloride, yellow; sol. on warming, crystallizing out when cold.
Iodine, brownish-red.

(*e*) *ARGININE*, decomposition product of proteid (see under plant-heading above), $C_6H_{14}N_4O_2$. Urea is obtained on boiling with Baryta water (saturated solution).

[Not precipitated by tannic acid nor lead acetate.]
Picric acid, gold needles; sol. hot water.
Precipitated also by phospho-tungstic acid.

§ 145. **LYCOPODIUM** complanatum; *Lycopodiaceæ*. Investigator: Bœdecker, *Ann. Chem. Ph.*, **208**, 263, substance (*a*), **L.** saururus (Piliganum), (*b*).

(*a*) *LYCOPODINE* A., $C_{32}H_{52}N_2O_3$; crystallizes in clino-rhombic prisms; M.P. 114°-115°; alkaline, bitter.

Soluble in water, more readily in chloroform or benzene; also in alcohol, ether, and amyl alcohol.

Reactions:

Alkaline hydrates, resinous pp. if concentrated.
[Not precipitated by neutral lead acetate.]
Iodine solution, brown cloud.

(*b*) *PILIGANINE* A. (Amyl-nicotine? Arata), $C_{15}H_{24}N_2O$ or $C_{10}H_{15}N_2{\cdot}OC_5H_{11}$; light yellow, amorphous, alkaline reaction, disagreeable odour, poisonous; gives Nicotine (?) on distillation in current of hydrogen. Salts mostly crystalline.

Soluble in water, alcohol, chloroform; with difficulty in ether.

Precipitants: Alkaline hydrates.
[Not lead acetate, neutral or basic.]
Tannic acid, white.
Ferric chloride, reddish pp.
[Not platinum chloride; the Pt. salt is very soluble.]
Gold chloride, yellow; altered by light.
Potassium ferricyanide, greenish.
Potassium bichromate, yellow.
Phospho-molybdic acid, whitish-yellow.
Iodo-potassic iodide, light brown.
Mercuric-potassic iodide, white.
Bromine water, Iodine tincture } yellow.

§ 146. **LYCOPUS** Europæus, **L.**; *Labiatæ*. The leaves. Investigator: Geiger, *Repert. Ph.*, **15**, 11.

LYCOPIN B.; amorphous, bitter, odourless.

Soluble in alcohol, ether, in 500 parts water, and in warm acetic acid.

§ 147. **MARRUBIUM** vulgare (Horehound); *Labiatæ*. The whole plant. Investigators: Kromayer and Harms, *Archiv. Pharm.*, vols. 83, 108, 116. Harms obtained 2 grammes of Marrubiin from 25 lb. of the plant.

MARUBIIN B.; crys. needles (from cold alcohol); M.P. 148° (Harms), or 160° (Kromayer); neutral reaction, bitter.

Soluble in alcohol, ether, chloroform, slightly in water, not in petroleum spirit.

Removed by ether from aqueous solution.

Absorbed by charcoal.

Not precipitated by lead acetate, neutral or basic (to alcoholic solution), nor by alkaloid reagents.

Fehling's solution, not reduced.

No colour reactions with acids (concentrated sulphuric acid, brownish-yellow—Kromayer).

§ 148. **MEGARRHIZA** Californica (Echinocystis); *Cucurbitaceæ*. Investigator: Heaney, *Amer. J. Ph.*, 48, 451.

(*a*) *MEGARRHIZIN* G.; poisonous. Acids give sugar and *Megarrhizonetin*.

Soluble in alcohol, not in ether.

(*b*) *MEGARRHIN*, resembles Saponin; enlarges pupil.

(c) *MEGARRHIZITIN* B.; crystalline, poisonous.

§ 149. **MENYANTHES** trifoliata (Buckbean); *Gentianaceæ*. Investigators: Kromayer, *Archiv Ph.*, 1861; Lendrich, *idem*, **230**, 38.

MENYANTHIN G., $C_{30}H_{46}O_{14}$ (Kromayer) or $C_{33}H_{50}O_{14}$ (Lendrich); amorphous, terebinthinate becoming gradually solid; M.P. 100°-115°; neutral reaction, bitter. Acids convert to Menyanthol, C_8H_8O (Hüsemann), or $C_7H_{11}O_2$ (Lendrich), with characteristic odour, and a lævorotatory sugar.

Soluble in hot water (with difficulty cold), in alcohol, chloroform, benzene; not in ether.

Removed from acid solutions by benzene and chloroform; absorbed by charcoal.

Precipitants:

[Not lead acetate, neutral or basic, or mere cloud.]

Tannic acid, white.

[Not gold chloride, but reduction on warming.]

Ammoniacal silver nitrate, reduced.

Fehling's solution, reduced.

Phospho-molybdic acid, yellow.

Bismuth-potassic iodide, yellow.

Mercuric-potassic iodide, white.

Iodine solution, yellow.

Colour tests, etc.:

Concentrated sulphuric acid, yellowish-brown changing to violet.

Dilute sulphuric acid on warming, odour of Menyanthol.

§ 150. **MIKANIA** Guaco; *Compositæ*. The leaves. Investigators: Faure, *Journ. Pharm.*, [2] **22**, 291; Pettenkofer, *Repert. Ph.*, **86**, 311. Used in America against snake-poisoning, etc.

GUACIN B.; light brown amorphous powder; M.P. 100°; neutral, bitter.

Soluble in hot water (difficultly cold), in alcohol, and in ether.

Colour tests:

Concentrated sulphuric acid, dissolves reddish-brown.

„ nitric acid, dissolves dark yellow.

„ hydrochloric acid, little change.

§ 151. **MORINDA** citrifolia (*Rubiaceæ*) contains MORINDIN, a glucoside yielding sugar, and MORINDON, a colouring matter. Morindin resembles Ruberythric acid (from Rubia tinctoria or Madder), the glucoside of Alizarin (Dihydroxi-anthraquinone).

§ 152. **MORRENIA** brachystephana (tisis); *Palmaceæ*; Argentine Republic. The root.

MORRENINE A.; dark, brown amorphous; M.P. 106°; bitter taste, sharp odour.

Soluble in water, alcohol, amyl alcohol, chloroform.

'Precipitated by most alkaloid reagents.'

§ 153. **NANDINA** domestica; *Berberidaceæ*. The root bark.

NANDININE A., $C_{19}H_{19}NO_4$; amorphous powder; poisonous. Salts amorphous.

Soluble in alcohol, ether, chloroform, benzene; not in water.

Alkaline hydrates, dissolve (acids re-precipitate).

Platinum chloride, whitish-yellow pp.

Concentrated sulphuric acid containing nitric acid, blue solution.

§ 154. **NARTHECIUM** ossifragum; *Juncaceæ*. Investigator: Walz, *N. Jahrb. Ph.*, **14**, 345.

NARTHECIN B.; crystalline; M.P. 35°; acid reaction, biting taste.

Soluble in alcohol, scarcely in water.

benzene, petroleum ether, methyl alcohol, fatty oils, and methyl iodide; not in cold water, but slightly on boiling.

Precipitants:

Alkaline hydrates, oily pp.	Phospho-molybdic acid, white.
„ carbonates.	Iodo-potassic iodide, brownish-purple, becoming crystalline.
Ammonia, oily pp.	Cadmium-potassic iodide, white.
Picric acid.	Mercuric-potassic iodide, white.
Platinum chloride.	Mercuric chloride, white.
Palladious chloride.	Nessler's reagent, grayish-brown pp.
Gold chloride.	
Potassium bichromate, yellow.	

Colour tests:

Concentrated sulphuric acid with potassium bichromate, blood-red to violet.

Concentrated sulphuric acid with nitric acid, violet.

„ nitric acid (alone), gradually reddish-violet.

§ 157. **NUPHAR** luteum (Yellow Water Lily); *Nymphæaceæ*. The rhizome.

NUPHARINE A., $C_{14}H_{24}N_2O_2$; amorphous; M.P. 65°; optically inactive; not bitter (but salts are bitter). Salts amorphous. Odour on acidification.

Soluble in alcohol, ether, chloroform, amyl alcohol, acetone; almost **insoluble** in petroleum ether.

Removed by chloroform from alkaline solutions.

Not precipitated by neutral lead acetate.

Pp. by most alkaloid reagents.

Concentrated sulphuric acid, on warming, brown, then greenish; water then gives yellowish-brown pp.

§ 158. **ONONIS** spinosa, L.; *Leguminosæ*. The root. Investigators: Rheinsch, *Repert. Pharm.*, vols. 76, 78; Hlasiwetz, *J. Pract. Pharm.*, 65.

(*a*) *ONONIN* G., $C_{30}H_{34}O_{13}$ (?); crys. microscopic plates and prisms; M.P. 235° (recrystallizing on cooling); no odour; not bitter. Acids yield sugar and Formonetin, (*b*).

Soluble in alcohol after long boiling; not in ether or in cold water; with difficulty in hot water (crystals appear on cooling).

It is carried down by the lead sulphide formed in freeing the vegetable extract from excess of lead acetate.

No pp. with cadmium-potassic iodide (Dragendorff).

Colour tests:

Concentrated sulphuric acid, yellowish-red changing to cherry-red.

„ „ „ with manganese dioxide, carmine.

„ nitric acid, yellow, with formation of oxalic acid.

Fröhde's reagent, pure red, lasting some time.

(*b*) *FORMONETIN* G.-deriv. from preceding; pale violet colouring matter, yielding Ononetin and formic acid on treatment with Baryta water.

(*c*) *ONONID* B. [Ononis-glycyrrhizin, Glycyrrhizin? (Hlasiwetz)], $C_{18}H_{22}O_8$; dark yellow, amorphous; acid reaction; slightly bitter, then sweet.

Soluble in alcohol and water.

Precipitants:

Lead acetate neutral.	[Not tannic acid.]
„ „ basic.	Copper, mercury and silver salts.

Acids precipitate from aqueous solution.

(*d*) *ONOCERIN*; crystalline.

Soluble in alcohol and warm turpentine, with difficulty in water, not in ether.

No change with acids or alkalies.

§ 159. **OPIUM**, from Papaver somniferum (Poppy); *Papaveraceæ*. The whole of the alkaloids described below, except those which are stated to be artificial derivatives of the naturally occurring bases.

Only the following six are found in appreciable quantities, and only the first three of these are physiologically active (some of the rarer alkaloids of Opium are also poisonous, as will be seen):

Morphine	average say	10 per cent.	active and basic.
Codeine	0·25 to 0·5	„	
Thebaine	0·15 „ 1·0	„	
Narcotine	average	6 „	comparatively inactive and neutral.
Papaverine	„	1 „	
Narceine	„	·05 „	

As would be expected, the relative proportions vary considerably. Morphine may be said roughly to fluctuate from about half the above figure (though rarely to that extent) to half as much again. Narcotine varies through wider limits. (French Opium frequently contains none, and East Indian more Narcotine than Morphine.) Narceine sometimes exceeds Thebaine in quantity.

Excluding Morphine and Narcotine, the total of other alkaloids may amount to 0·5 to, say, 2 per cent. Of Cryptopine, T. and H. Smith

obtained only 1 part from 30,000 of Opium; and Kauder (*Arch. Pharm.*, 228, 419) found the following proportions of the rare alkaloids: Cryptopine 70, Laudanine 20, Protopine 3½, Tritopine 2, to 1 part of Laudanosine.

By the 'British Pharmacopœia' Opium is required to assay 9½ to 10½ per cent. of Morphine, calculated on the dried Opium.

In addition to alkaloidal substances, Opium contains: Acetic, lactic (1¼ per cent.), meconic (4 per cent.) acids, meconin, gum, pectin, sugar, bumoid matters, resin, caoutchouc, albuminoids, water (10 to 30 per cent.), and the usual inorganic salts.

Papaver rhœas (Red Poppy) contains Rheadine. Argemone Mexicana (Prickly Poppy) and Eschscholtzia Californica (*Papaveraceæ*) have been found to contain Morphine; and Chelidonium, as well as Sanguinaria—Protopine. For other alkaloids of the *Papaveraceæ*, see Chelidonium group; for Berberine, found in some members of the order, see Berberis group.

The alkaloids, etc., here described appear in the following order:

(1) *MORPHINE.*
(2*a*) Apomorphine.
(2*b*) Pseudomorphine.
(3) *NARCOTINE.*
(3*a*) Hydrocotarnine.
(3*b*) Cotarnine.
(3*c*) Tarconine.
(3*d*) Nartinic Acid.
(3*e*) Cupronine.
(3*f*) Tarnine.
(3*g*) Cuprine.
(3*h*) Dibromapophylline.
(3*i*) Apophyllic Acid.
(4) *PAPAVERINE.*
(5*a*) Papaveramine.
(6) *CODEINE.*
(7*a*) Codenine.
(8) *THEBAINE.*
(9*a*) Thebaicine.
(9*b*) Thebenine.
(10) *NARCEINE.*
(11) Rheadine.
(12) Oxynarcotine.
(13) Gnoscopine.
(14) Codamine
(15) Meconidine.
(16) Lanthopine.
(17) Cryptopine.
(18) Laudanine.
(19) Protopine.
(20) Tritopine.
(21) Laudanosine.

Non-basic Substances:

(22) Meconin.
(23) Pseudomeconin.
(24) Meconic Acid.
(25) Meconoisin.
(26) Opionin.

(1) *MORPHINE* A. (Morphia), $C_{17}H_{19}NO_3+H_2O$; crystallizes in rhombic prisms; M.P. about 230°, with partial decomposition; lævo-rotatory, alkaline. hitter. Fluorescent in sulphuric acid solution.

Solubility varies greatly with the physical condition; very slightly soluble in cold water (1 in 33,333 at 3°, Chastaing), 1 in 400 to 500 boiling; alcohol, 1 in 40 cold, 1 in 30 boiling; chloroform, 1 in 175 if ordinary chloroform, or 1 in 10,000 if free from alcohol (Prescott, *J. Chem. Soc.*, 22, 405, gives 1 in 4,379, 1,977 and 861 for the crystallized, amorphous and freshly liberated or 'nascent' Morphine respectively); cold amyl alcohol, 1 in 400, Van der Burg (Prescott's figure 91 for both crystallized and 'nascent'), easier warm; acetic ether, 1 in 500, Van der Burg; scarcely in petroleum ether, benzene, or ether (Prescott's figures for the latter, 6,148, 2,112, 1,062 for the three conditions).

In consequence of its limited solubility in the usual solvents, special methods have to be employed to remove it from an aqueous mixture; this may be most conveniently effected by warming the *acidified* solution, adding warm amyl alcohol, and then, after alkalizing with ammonia, the Morphine thus freshly liberated will be taken up by the warm amyl alcohol.

Minute quantities are removable by ether (this affords a means of separation from Urea, which is soluble in amyl alcohol, but not in ether).

Precipitants:

Alkaline hydrates, Ammonia, Lime-water } pp. soluble in excess.
Alkaline carbonates, soluble in excess, but not as readily as above.
,, bicarbonate, not soluble in excess. The precipitates become crystalline.
[Not lead acetate.]
[Not Tannic acid. cloud at 1 in 1,000.]
Picric acid, 1 in 100.
[Ferric chloride neutral solution, blue coloration 1 in 5,000, destroyed by acid or alcohol.]
Platinum chloride, 1 in 100; or after 24 hours 1 in 3,000.
Gold chloride, 1 in 3,000 at once.
Silver nitrate, red coloration; 1 in 1,000 after 15 minutes.
[Potassium ferrocyanide, not 1 in 100.]
[Potassium ferricyanide, reduced; no pp. 1 in 60 except after standing, Plugge.]
Potass. sulphocyanide, white crystalline; not 1 in 100.
,, cyanide, pp.; cloud at 1 in 5,000.
Silver potassic cyanide, a crystalline deposit after some hours.
Potass. chromate, the pp. contains free M. and M. chromate.

[Potass. bichromate, scarcely a cloud at 1 in 100; see also mercuric chloride below.]
[Not 5 per cent. chromic acid.]
Phospho-molybdic acid, pp.; cloud at 1 in 5,000; see colour tests below.
[Phospho-tungstic acid, cloud at 1 in 1,000.]
[„ antimonic acid, not at 1 in 1,000.]
Iodo-potassic iodide, red; 1 in 5,000.
Bromo-potassic bromide, red; 1 in 5,000.
Bismuth-potassic iodide, pp.; slight at 1 in 5,000.
Cadmium-potassic iodide, crys. needles; 1 in 1,000 after 2 hours.
[Not zinc-potassic iodide.]
Mercuric-potassic iodide, amorphous 1 in 1,000; limit 1 in 2,500.
„ chloride, white crystalline 1 in 100; remains white with potass. bichromate (compare **Strychnine**).

Colour tests:

Concentrated sulphuric acid, colourless, becoming gradually red; violet on warming.
Concentrated sulphuric acid, heated with the alkaloid to 150°, converts to **Sulpho-morphine**, with which ammonia gives white pp., and chloroform a red solution. Compare **Codeine** and **Pseudo-morphine**.
Concentrated sulphuric acid with potassium chlorate, bluish-violet.
„ „ „ warmed gently, then potass. perchloride, brown. **Strychnine** does not interfere.
Concentrated sulphuric acid with sugar, red (perceptible with $\frac{1}{100}$ milligramme).
Concentrated sulphuric acid with the phospho-molybdic pp., blue; dark brown on warming.
The phospho-molybdic pp. with potash, brownish-orange.
Concentrated sulphuric acid with nitric acid, violet.
Nitric acid added after 10 hours' solution in sulphuric acid (Dragendorff), reddish⁓bluish-violet⁓blood-red⁓orange.
Nitric acid alone, orange⁓light yellow.
Hydrochloric acid (concentrated), no effect.
Dilute acids, dissolve readily.
Per-iodic acid, iodine liberated.
Fröhde's reagent, magnificent violet ⁓ green ⁓ brownish-green ⁓ yellow; after 24 hours, bluish-violet.

(2*a*) *APOMORPHINE* **A.**, $C_{17}H_{17}NO_2$=Morphine less H_2O; formed by heating the latter alkaloid in a sealed tube to 140° or 150° C. with zinc chloride or large excess of hydrochloric acid; greenish crystals, alkaline reaction, emetic. The solutions become green on keeping.

Soluble in alcohol, ether (purple colour), chloroform (violet), benzene; with difficulty in water.

On shaking an alkaline aqueous solution with petroleum ether, benzene, or chloroform, a red decomposition product is removed.

Precipitants:

Alkaline hydrates, greenish-white or turning green; sol. in excess.
Ammonia (sol. excess), dirty violet.
Lime-water (sol. excess).
Alkaline bicarbonate, white when pure, but turns green.
[Not lead acetate.]
Tannic acid, yellowish-green.
Picric acid, yellow; sol. warm.
Ferric chloride, red, becoming black.
Platinum chloride, yellow.
Gold chloride, purple.
Potassium ferrocyanide, reddish-yellow; green on warming.
„ ferricyanide, white, changing to violet, then black
„ sulphocyanide, white; sol. warm.
„ bichromate, orange.
[Not 5 per cent. chromic acid.]
Iodo-potassic iodide, blood-red; sol. warm.
Potassic iodide alone, white, changing to green.
Bismuth potass. iodide, limit 1 in 10,000.
Mercuric chloride, white.
Zinc chloride } white, soluble on warming.
„ acetate }

Colour tests (besides those formed by reagents above):

Concentrated sulphuric acid, red.
Concentrated nitric acid, violet red.
Per-iodic acid, reduced.
Fröhde's solution, violet.

(2*b*) *PSEUDOMORPHINE* **A.** (of Pelletier; Donath's Dehydro-morphine; Schützenberger's Oxymorphine; Polstorff's Oxydimorphine); $C_{17}H_{18}NO_3$ or $C_{34}H_{36}N_2O_6$ (O. Hesse). Occurs naturally in opium, though rarely, and obtained from morphine by moderate oxidation, as with potassium ferricyanide. Its occasional presence in opium extracts may arise also from oxidation.

Crystallizes in leaflets or needles with $3H_2O$; not fusible without decomposition ; neutral reaction, not poisonous and not bitter.

Soluble in hot amyl alcohol (and in excess of ammonia, by means of which it is best purified) ; **insoluble** in water, alcohol, ether, chloroform, or carbon bisulphide.

Reactions :

Alkaline hydrates, dissolve.
Ammonia, pp. sol. in excess.
[No pp. lead acetate.]
Ferric chloride, blue or brownish-green coloration.
[Picric acid, not dilute.]

Colour tests with acids :

Concentrated sulphuric acid, olive-green ; or colourless if pure H_2SO_4, then yellow ; but trace of iron in the acid causes a blue~violet~brownish-green.
Concentrated sulphuric acid with sugar, blue to green if trace of iron, or green to brown in absence of latter.
Dissolved in 50 per cent. sulphuric acid, gradual bluish-green ; then more water~red.
Nitric acid added to this solution, deep violet (Morphine under these conditions gives rose-raspberry red.
Nitric acid alone, blood-red.

(3) *NARCOTINE* A. (Meconine-hydrocotarmine ; Opianine ? ; Aconelline ; the aldehyde of Oxynarcotine) ; $C_{22}H_{23}NO_7$, Matthieson. Crystallizes in rhombic prisms or needles ; sp. gr. 1·37 to 1·39 ; M.P. 115°, 155°, 176° (different observers) ; lævo-rotatory in neutral solution, or dextro-rotatory in acid ; reaction neutral ; tasteless, but salts are bitter, and are soluble in alcohol, ether and water. Decomposes, on heating, to Meconine and Cotarnine. See (3*a*) and (22).

Solubility (results of different observers are not concordant. As in the case of Morphine, solubility is apparently influenced by physical conditions) : Cold water, 1 in 25,000 ; boiling water, 1 in 7,000 (or 1,500 and 600 respectively for the base freshly precipitated by ammonia) ; alcohol, cold 1 in 80, boiling 1 in 20 ; amyl alcohol, 1 in 300 ; chloroform, 1 in 2·69 ; ether, 1 in 166 cold, 1 in 48 boiling ; benzene, 1 in 22 ; acetic ether, 1 in 60 ; also in fatty oils ; scarcely soluble in petroleum ether.

Removed by benzene or ether from alkaline solutions (also from dried opium).

Precipitants :

Alkaline hydrates, insoluble in excess cold, but dissolved on heating.
„ carbonates } insoluble in excess.
„ bicarbonates }
Ammonia }
Lime water, soluble in excess.
[Not lead acetate.]
[Tannic acid, cloud ; slight pp. with hydrochloric acid.]
[Not gallic acid.]
Picric acid, yellow crystalline ; cloud at 1 in 4,000.
[Ferric chloride, no blue coloration].
[Platinum chloride, in concentrated solutions, yellow pp.]
[Gold chloride, in concentrated solutions, yellow pp. with reduction of gold.]
[Potassium ferrocyanide, concentrated solutions, but not 1 in 200.]
[Not potassium ferricyanide.]
Potassium sulphocyanide, 1 in 200 readily sol. in acids.
Silver potassic cyanide, immediate pp. amorphous.
Potassium chromate, pp. free Narcotine.
„ bichromate, slight at 1 in 400 ; pp. = Narcotine-bichromate.
Phospho-molybdic acid, brownish-yellow, 1 in 3,000 ; limit 1 in 4,000.
„ tungstic acid, limit 1 in 8,000.
„ antimonic acid, yellow flocculent, 1 in 1,000 ; cloud at 1 in 2,500.
Iodo-potassic iodide, red, 1 in 8,000.
Bismuth-potassic iodide, orange-red, limit 1 in 4,000.
Cadmium-potassic iodide, 1 in 3,000 ; limit 1 in 8,000.
[Zinc-potassic iodide, scarcely.]
Mercuric-potassic iodide, white ; slight at 1 in 8,000 ; limit 1 in 50,000.
Mercuric chloride, cloud, then pp. ; cloud 1 in 3,000.
Chlorine water, greenish, changing to yellow with ammonia ($\frac{1}{2}$-gramme in solution).
Bromine water, yellow pp.
Potassium bromide, slight at 1 in 8,000.

Colour tests :

Concentrated sulphuric acid, colourless, gradually yellow (red after several days.
Concentrated sulphuric acid, on gradually warming to 200°, bluish-violet, but not below 90°-100° (see Curarine).

Concentrated sulphuric acid with nitric acid, reddish-violet.
Concentrated sulphuric acid with nitre added after 10 hours, brown ~light yellow~yellowish-red.
Concentrated nitric acid alone, yellow~colourless.
Concentrated hydrochloric acid, no effect.
Per-iodic acid, not reduced.
Fröhde's solution, green~cherry-red with more molybdic acid.

(3*a*) *HYDRO-COTARNINE* A., $C_{12}H_{15}NO_3+1\frac{1}{2}H_2O$ (Beckett and Wright). From Opium and also from Narcotine and Cotarnine by reduction. Monoclinic prisms; M.P. 50°-55°, partly volatile at 100°; optically inactive; alkaline reaction; a strong base; more poisonous than Cotarnine and Morphine.

Soluble in alcohol, ether, chloroform, benzene.

Concentrated sulphuric acid, yellow; red on warming.

(3*b*) *COTARNINE* A., $C_{12}H_{13}NO_3+H_2O$, or

$$CH_2<^{O}_{O}>C_6H(OCH_3)(CHO)\cdot\underbrace{CH_2\cdot CH_2\cdot NH\cdot CH_3}_{5\ :\ 4\ :\ 2\ :\ 3}$$

(W. Roser), see J. *Chem. Soc.*, 1890, p. 531. From Narcotine. Crystallizes in radiating needles, anhydrous near 100°; M.P. 132° (Pictet).

Soluble in alcohol (solution turns brown), readily in ether, difficultly in cold water, easier warm.

Reactions:

Alkaline hydrates, scarcely dissolve.	Ferrous salts, pp.
Ammonia, dissolves easily.	Platinum chloride, yellow crystalline pp.
Tannic acid, pp.	Copper sulphate, pp.

Concentrated nitric acid, dissolves red, yielding oxalic acid.

(3*c*) *TARCONINE* A.,

$$C_{11}H_9NO_3 \text{ or } CH_2<^{O\cdot}_{O\cdot C}\!\!>\!\!\overset{O}{\underset{:C\cdot CH:C\cdot CH:CH}{C\text{———}\cdot C:CH\cdot NCH_3}}$$

(W. Roser), see J. *Chem. Soc.*, 1890. From Narcotine.

(3*d*) *NARTINIC ACID*, $C_{20}H_{16}N_2O_6$. From Narcotine. Orange coloured.

Reactions:

Alkaline hydrates dissolve, with separation of green flocks.
Ferric chloride, reddish-brown coloration.
Silver nitrate, reduced.

(3*e*) *CUPRONINE* A., $C_{20}H_{18}N_2O_6$. From Narcotine. Black powder. The hydrobromide is copper-coloured and difficultly soluble in water.
Alkaline hydrates dissolve, reddish-brown.
Concentrated sulphuric acid, fuchsine-red, bluish-violet with water.

(3*f*) *TARNINE* A., $C_{11}H_9NO_4+1\frac{1}{2}H_2O$. From Narcotine. Orange needles. M.P. 220°.

The hydrobromide soluble in water (compare Cupronine).

Soluble in hot water or hot alcohol, not in ether.

(3*g*) *CUPRINE* A., $C_{11}H_7NO_3$. From Narcotine. (Not to confuse with Cupreine, see Cinchona.) Copper-coloured.

Soluble in water and alcohol with green colour. **Insoluble** in ether.

Acids give deep blue solutions.

(3*h*) *DIBROMAPOPHYLLINE* A., $C_{14}H_{10}N_2Br_2O_4$. From Narcotine. M.P. 229°.

Soluble in water.

Silver nitrate, reduced.

(3*i*) *APOPHYLLIC ACID.*

$$C_8H_7NO_4 \text{ or } COOH\cdot C_6H_3\cdot NCH_3\cdot C<^{O}_{O}$$

(Wohler, etc.). M.P. 241°-242°. Difficultly **soluble** in water. **Insoluble** in alcohol or ether.

(4) *PAPAVERINE* A., $C_{20}H_{21}NO_4$ or $C_6H_3(O\cdot CH_3)_2\cdot CH_2\cdot C_9H_4N(OCH_3)_2$ (Goldschmidt). Crystallizes in prisms; M.P. 98°, sublimate 155° (Guy), or 147° (Pictet); sp. gr. 1·308 to 1·337; very feebly alkaline; feebly lævo-rotatory (or inactive—Goldschmidt); not poisonous.

Soluble scarcely in water, in about 50 parts cold alcohol, readily in hot, giving crystals on cooling; in 36 parts benzene, 258 ether, 77 acetic ether; in chloroform, acetone, and warm petroleum ether (separates from the latter on cooling).

Removed by chloroform from both acid and alkaline solutions, by benzene with difficulty.

Precipitants:

Alkaline hydrates.
„ carbonates.
Ammonia.
[Not lead acetate]
Tannic acid, yellow; 1 in 5,000.
Picric acid, yellow amorphous becoming crys.; slight at 1 in 500.

Platinum chloride, nearly white ; not 1 in 500. Sol. in hydrochloric acid.
Gold chloride, dirty yellow ; distinct at 1 in 5,000.
Potassium ferrocyanide ; the pp, is Pa_4H_4 Fcy.
„ ferricyanide ; slight at 1 in 1,000.
„ sulphocyanide ; soluble on warming.
Silver potassic cyanide, 1 in 3,000 to 4,000 at once.
Potassium chromate ; the pp. contains Papaverine chromate and free Papaverine ; or, if warm, free Papaverine only.
Potassium bichromate, at 1 in 3,000 gradual ; the pp. is Papaverine bichromate.
[Not phospho-molybdic acid if dilute ; *this is characteristic.* At 1 in 1,000 a cloud only.]
Iodo-potassic iodide, up to 1 in 50,000.
Bismuth-potassic iodide, orange red ; slight at 1 in 10,000.
Cadmium-potassic iodide, distinct at 1 in 1,000.
Zinc-potassic iodide, crystalline ; 1 in 10,000.
Mercuric-potassic iodide, white ; distinct at 1 in 1,000.
[Mercuric chloride, slowly ; not 1 in 500.]
Chlorine water, greenish : ammonia changes to brown.
Iodine tincture, crystalline needles.

Colour tests :

Concentrated sulphuric acid, colourless *if pure ;* otherwise momentary bluish-violet, then violet.
Concentrated sulphuric acid on warming, blue.
„ „ „ with nitre, orange becoming yellow.
Nitre added to the solution in sulphuric acid after ten hours, momentary violet, then orange~dirty yellow.
Concentrated sulphuric and nitric acids (together), yellow~orange.
„ nitric acid alone, yellow~dark orange.
„ hydrochloric acid, no effect.
Fröhde's solution, green~blue~violet~cherry-red.

(5*a*) *PAPAVERAMINE* A., $C_{21}H_{21}NO_5$ (Hesse) ; crystalline prisms ; M.P. 142°. Scarcely soluble in alkalies.

(6) *CODEINE* A. (Morphine Methyl ether), $C_{18}H_{21}NO_3$, and with $3H_2O$. Octahedral crystals if anhydrous, prisms when hydrated ; M.P. (anhydrous) 155° ; alkaline reaction, slightly bitter (salts very bitter) ; lævo-rotatory. Physiological action somewhat similar to that of Morphine.

Soluble in 80 parts cold or 17 of boiling water, 7 amyl alcohol, 10 chloroform, 12 benzene, in carbon bisulphide, and readily in alcohol and ether ; scarcely in petroleum ether.

Precipitants :

Alkaline hydrates (slightly soluble in excess).
[In ammonia, Codeine is about as soluble as in water, see above.]
[Not alkaline carbonates.]
[Not lead acetate, but free Codeine gives pp. of PbO, etc.]
Tannic acid, white.
Picric acid, yellow ; distinct at 1 in 250.
[Ferric chloride, no colour ; see sulphuric acid.]
Platinum chloride, yellow ; not 1 in 250.
Palladious chloride, yellow ; not 1 in 250.
Gold chloride, brown ; not 1 in 1,000.
[Copper salts give a pp. of CuO with free Codeine.]
Potassium ferrocyanide, white pp. (alcoholic solution) ; slight at 1 in 1,000.
Potassium ferricyanide (aqueous solution), white crystalline ; not 1 in 70 (Plugge).
Potassium sulphocyanide, pp. gradually becoming crystalline ; sol. on warming.
Silver potassic cyanide, crystalline deposit after some hours.
Potassium chromate, pp. = C. chromate.
„ bichromate, pp. = C. bichromate ; gradually at 1 in 3,000.
Phospho-molybdic acid, limit 1 in 50,000.
„ antimonic acid, at 1 in 1,000, dirty white cloud.
Iodo-potassic iodide, kermes coloured.
Bismuth-potassic iodide, orange ; distinct at 1 in 50,000.
Cadmium-potassic iodide, pp. gradually crystalline, 1 in 500.
Zinc-potassic iodide, long hair-like crystals, very abundant, so that solution when not too dilute becomes almost solid (not at 1 in 1,000).
Mercuric-potassic iodide, abundant pp. at 1 in 5,000 ; limit 1 in 50,000.
Mercuric chloride, not 1 in 500.
Chlorine water, colourless solution ; ammonia gives reddish-brown.
Bromine water, pp. distinct at 1 in 50,000.

Colour tests :

Concentrated sulphuric acid, colourless, but deep blue if containing ferric salt.

Concentrated sulphuric acid, heated to 150°, as morphine.
" " " with sugar, red.
" " " with potass. bichromate, olive-green.
" " " with nitric acid, gradually blue.
Nitre added to the solution in sulphuric acid after ten hours, the blue solution becomes cherry-red~blood-colour~orange.
Nitric acid alone, yellow.
Per-iodic acid, not reduced.
Fröhde's reagent, dirty green; indigo colour after twenty-four hours.

(7*a*) *CODENINE* A., $C_{18}H_{21}NO_3{\cdot}H_2O$; isomer of Codeine, and formed from it by heating with sulphuric acid and water; crystallizes in needles.

(8) *THEBAINE* A. (Paramorphine), $C_{19}H_{21}NO_3$ (Anderson). Crystallizable (but not easily) in needles or prisms; M.P. 193°; lævo-rotatory; alkaline reaction, sharp taste, very poisonous (action tetanic).

Solubility: 1 in 10 cold alcohol, 1 in 60 amyl alcohol, 1 in 18 chloroform, 1 in 140 cold ether—more readily hot, 1 in 20 benzene. **Insoluble** in water and petroleum ether.

Precipitants:
Alkaline hydrates, Ammonia } insoluble in excess.
[No pp. alkaline carbonates.]
[Not lead acetate.]
Tannic acid, yellow; 1 in 3,000; soluble in hydrochloric acid.
Picric acid, yellow amorphous.
Platinum chloride, 1 in 1,000 (not 1 in 5,000).
[Palladious chloride, not 1 in 250.]
Gold chloride, yellowish-brown; 1 in 5,000; limit 1 in 10,000.
Potassium ferrocyanide, pp. 1 in 500; sol. in excess; the pp. = $T_4{\cdot}H_4$ Fcy.
Potassium chromate, the pp. is T. chromate.
" bichromate, gradually at 1 in 3,000; the pp. is T. bichromate.
Phospho-molybdic acid, distinct at 1 in 5,000; extreme limit 1 in 50,000.
Iodo-potassic iodide, distinct at 1 in 5,000; extreme limit 1 in 50,000.
Bismuth-potassic iodide, orange; slight at 1 in 10,000.
Cadmium-potassic iodide, distinct at 1 in 10,000; not at 1 in 50,000.
Zinc-potassic iodide, 1 in 500.
Mercuric-potassic iodide, distinct at 1 in 5,000; extreme limit 1 in 50,000.
Mercuric chloride, cloud, then gradual pp.; not 1 in 250.
Chlorine water, colourless solution; reddish-brown with ammonia.
Bromine water, pp. distinct at 1 in 5,000; extreme limit 1 in 50,000.

Colour tests:
Concentrated sulphuric acid, blood-red.
" " " with nitric acid, red.
Nitre added to solution in sulphuric acid after ten hours, the yellow solution~orange~yellow.
Nitric acid alone, yellow.
Concentrated hydrochloric acid, no effect.
Fröhde's solution, orange; after twenty-four hours, nearly colourless.

(9*a*) *THEBAICINE* A., $C_{19}H_{21}NO_3$; formed together with (9*b*) by action of concentrated hydrochloric acid on Thebaine. Amorphous, and giving amorphous resinous salts.

Soluble with difficulty in hot alcohol; **insoluble** in water, ether, or benzene.

Reactions:
Alkaline hydrates, pp. slightly soluble (brown solution).
Ammonia, insoluble.

Colour tests: Concentrated sulphuric acid, blue.
Concentrated nitric acid, red.

(9*b*) *THEBENINE* A., $C_{19}H_{21}NO_3$. For formation see (9*a*). Amorphous.

Soluble with difficulty in hot alcohol, not in water, ether, or benzene.

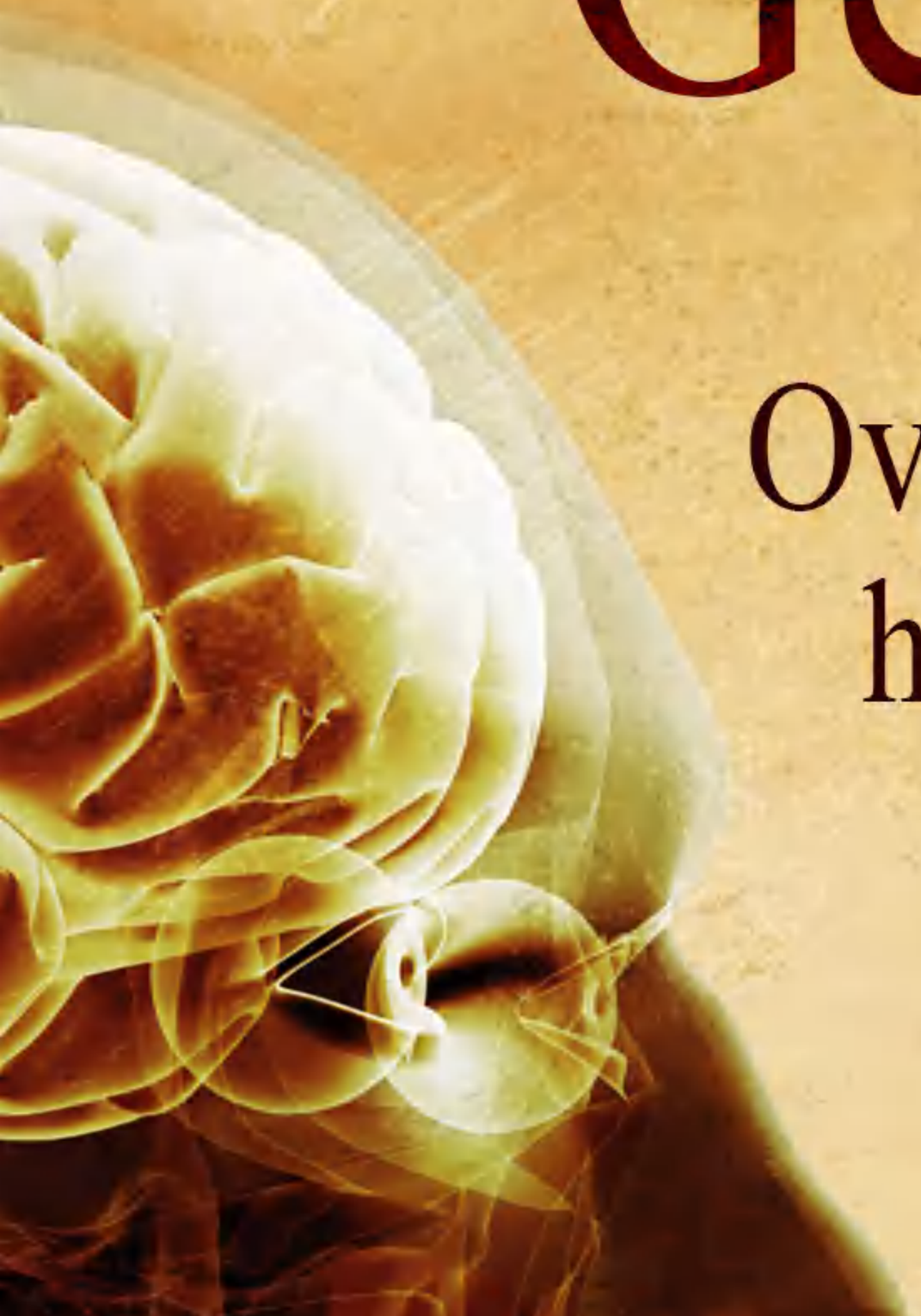
Reactions: Alkaline hydrates, soluble; ammonia, insoluble.

Colour test: Concentrated sulphuric acid, blue, becoming colourless on addition of water, colour reappearing with more acid.

(10) *NARCEINE* A. (Pseudo-narceine=Narceine, W. Roser), $C_{23}H_{29}NO_9$ or $C_{13}H_{20}NO_4{\cdot}CO{\cdot}C_6H_2(OCH_3)_2COOH$: rhombic prismatic needles; M.P. 92°, Pelletier (145°, Hesse; 171°, Merck); neutral reaction, feebly bitter, hypnotic.

Soluble slightly in water, with difficulty in cold alcohol (easier hot), very little in amyl alcohol or chloroform, and **insoluble** in ether, benzene, or petroleum ether.

Precipitants:
Alkaline hydrates; the pp. dissolves in excess more readily than in water, but is again produced with large quantity of alkali.



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Colour tests:
Concentrated sulphuric acid, yellow.
„ „ „ with nitric acid, red.
„ nitric acid, red.

(14) *CODAMINE* A., $C_{20}H_{25}NO_4$ (Hesse); crystallizes in six-sided prisms from ether; M.P. 121°, with decomposition and crystalline sublimate; bitter taste, alkaline reaction. Salts amorphous.

Soluble in alcohol, ether, chloroform, benzene (very readily in last two solvents), and in boiling water (difficultly cold).

Reactions:
Alkaline hydrates, pp., soluble in excess.
„ bicarbonates, pp., becoming resinous.
Concentrated sulphuric acid, green.
„ nitric acid, green cold, violet hot.

(15) *MECONIDINE* A., $C_{21}H_{23}NO_4$ (Hesse); brownish-yellow, amorphous; M.P. 58°; alkaline reaction, tasteless (salts are bitter). Readily decomposed by acids. Salts very soluble.

Soluble in alcohol, ether, chloroform, benzene, acetone; not in water.

Not removed by immiscible solvents from solution in caustic alkali, but from an ammoniacal or calcic hydrate solution.

Precipitants:
Alkaline hydrates, soluble.
Ammonia, Lime-water } soluble with difficulty.
Platinum chloride, yellow amorphous, changing to red.
Gold chloride, dirty yellow amorphous.

Colour tests:
Concentrated sulphuric acid, olive-green.
„ nitric acid, reddish colour.
Acids in general, decompose readily; dilute sulphuric acid gives gradual red, then purple (ammonia gives with this solution a white pp.). Only slight decomposition with acetic acid.

(16) *LANTHOPINE* A., $C_{23}H_{25}NO_4$ (Hesse); microscopic prismatic crystals; browns on heating to 190°-200°; alkaline reaction, tasteless. Salts crystalline, but sometimes gelatinous.

Soluble in chloroform, with difficulty in alcohol, ether, or glacial acetic acid, scarcely in benzene; not in water.

Removed by chloroform from an ammoniacal, but not from a solution containing caustic soda or potash.

Reactions:
Alkaline hydrates dissolve (ammonium chloride precipitates therefrom).
Ammonia, pp., insoluble.
[Gold chloride salt, M.P. 182°.]

Colour tests:
Concentrated sulphuric acid, colourless or pale violet; brown on heating to 150°.
Concentrated nitric acid, resinifies, then dissolves yellow to red.
[No reaction with ferric chloride.]

(17) *CRYPTOPINE* A., $C_{21}H_{23}NO_5$ (T. and H. Smith); crystallizes in six-sided prisms from alcohol; M.P. 217° (176°-177° with sublimate, Guy); optically inactive; specific gravity 1·351; bitter, then hot peppermint-like taste; alkaline, hypnotic (also convulsant action upon dogs), and mydriatic. Salts tend to gelatinize.

Soluble in 1,200 parts cold alcohol (better hot), and in chloroform; slightly in water, giving jelly or crystals; scarcely in benzene; not in ether unless freshly precipitated (and even then it is slowly re-deposited), nor in petroleum ether or turpentine.

Reactions:
Alkaline hydrates, pp. insoluble.
The platinum chloride salt, soluble on warming.
Iodo-potassic iodide, pp.
The mercuric chloride salt, soluble on warming.

Colour tests:
Concentrated sulphuric acid, intense blue.
Nitre added to this solution, green~yellow; on adding more alkaloid, intense green.
Nitric acid alone, gradually yellow.

(18) *LAUDANINE* A., $C_{20}H_{25}NO_4$ (Hesse). Isomeric with Codamine (14); crystallizes in radiating prisms; M.P. 166°, no sublimate; lævo-rotatory; specific gravity 1·255, alkaline, bitter, very poisonous.

Soluble in hot alcohol (with difficulty cold), in 647 parts ether, in benzene, and readily in chloroform.

Removed from alkaline solution by chloroform readily.

Reactions:
Alkaline hydrates, pp. soluble in excess.
Ferric chloride, an emerald-green coloration.

Concentrated sulphuric acid, pink ; dark violet on warming.
" " " containing ferric oxide, pink.
" nitric acid, orange.

(19) *PROTOPINE* A., $C_{20}H_{19}NO_5$ (Hesse), or $C_{20}H_{17}NO_5$ (G. König). From opium, also in Chelidonium and Sanguinaria ; see Chelidonium group. Crystalline ; M.P. 204° (König) ; alkaline.

Soluble in chloroform, with difficulty in alcohol and ether, scarcely in benzene, not in water.

Reactions :
Alkaline hydrates, pp. insoluble in excess.
Ammonia, dissolves.
Iodo-potassic iodide, pp.

(20) *TRITOPINE* A., $C_{42}H_{54}N_2O_7$ (E. Kander) ; prismatic crystals ; M.P. 182°.

Soluble easily in chloroform, slightly in ether.

Reactions :
Alkaline hydrates, Ammonia } pp. insoluble in excess.
Iodo-potassic iodide, pp.

(21) *LAUDANOSINE* A., $C_{21}H_{27}NO_4$ (Hesse) ; prismatic crystals ; M.P. 89° ; dextro-rotatory ; alkaline, bitter, very poisonous.

Soluble very readily in alcohol and chloroform, in 19 parts ether, also in benzene and petroleum ether ; not in water.

Reactions :
Alkaline hydrates, Ammonia } insoluble.
[No coloration with ferric chloride.]
Platinum chloride, pp. yellow.
Concentrated sulphuric acid, reddish-brown.
" " " at 150° C., green~dirty violet.

(22) *MECONIN* B. (Opianyl) ; $C_{10}H_{10}O_4=(CH_3O)_2:C_6H_2<^{CH_2}_{CO}>O$; exists in opium in amounts varying from 0·01 to 0·8 per cent ; also obtainable from Narcotine, and from Opianic acid $\left.\begin{matrix}CH_3O\\CH_3O\end{matrix}\right\}C_6H_2\left\{\begin{matrix}COOH\\COH\end{matrix}\right.$ by reduction. Crystallizes in hexagonal prisms ; M.P. 110°, yielding sublimate, and recrystallizing on cooling, or melts under water at 77° ; optically inactive ; neutral reaction ; bitter.

Soluble in 700 parts cold water or 22 boiling ; also in alcohol, ether, chloroform, benzene, amyl alcohol, ethereal oils, and glacial acetic acid.
Removed from acid solutions by amyl alcohol, chloroform and benzene.

Reactions :
Alkaline hydrates dissolve, forming meconate.
Ferric chloride, blood-red not decolorized by hydrochloric acid.
[No precipitates by metallic salts.]
Concentrated sulphuric acid, colourless ; becoming purple on warming, or gradually green in the cold ; reddish after twenty-four hours.

(23) *PSEUDOMECONIN* B., $C_{10}H_{10}O_4$; resembles preceding ; M.P. 123°-124°.

(24) *MECONIC ACID*, $C_7H_4O_7+3H_2O$; characteristic opium acid, present to the extent of about 4 per cent. on an average ; crystallizes in prisms or scales ; evolves carbonic acid on heating to 120° C.

Soluble in alcohol and amyl alcohol, less readily in ether, with difficulty in water, scarcely in chloroform.
It may be removed from acid solutions by amyl alcohol and ether.

Reactions : Alkaline hydrates dissolve.
Calcic chloride to neutral solution, precipitates.
Pp. lead acetate.
Ferric chloride, blood-red coloration not discharged by hydrochloric acid.

(25) *MECONOISIN* B., $C_8H_{10}O_2$ (T. and H. Smith) ; crystalline ; M.P. 88° ; distils at 280°.

Soluble in alcohol easily, also in ether, but scarcely in water.

(26) *OPIONIN* B. ; nitrogen free (Hesse, 1885) ; crystallizes in needles ; M.P. 227° ; neutral reaction.
Alkaline hydrates, Ammonia } dissolve ; acids re-precipitate.

§ 160. **PALICOUREA** marcgravii ; *Rubiaceæ*. Investigator : Peckholt, *Archiv. Pharm.* [2], **127**, 93.

PALICOURINE A., Crystallizing in silky needles, and forming crystalline salts.

Precipitates with most alkaloid reagents (Dupuy).

§ 161. **PANAX** quinquefolius (Aralia quinquefolia, Decaisne ; American ginseng) ; *Araliaceæ*. The root. Investigator : Garrigues, *Ann. Chem. Pharm.*, **90**, 231.

PANAQUILON B., $C_{24}H_{25}O_{18}$. Yellow, amorphous, fusible; with bitter-sweet taste.

Soluble in water and alcohol, not in ether.

Reactions:

Alkaline hydrates, dissolve brown.
Tannic acid, pp.
[No pp. platinum chloride.]

Concentrated sulphuric acid, purplish-red; water precipitates from the solution a substance *Panacon*.

PAPAVER, see Opium.

§ 162. **PAPAYA** (Carica papaya); *Passiflorinæ* or *Papayaceæ*. The fruit. Investigators: Würtz, Peckholt, Greshoff, and others.

Besides (*a*) below, the two alkaloids CARICINE and PAPAINE have been found.

The principle which gives to Papaya such extraordinary powers of dissolving fibrin, is a soluble nitrogenous ferment (Papaïn or Papayotin), yielding precipitates with the following (which include the usual reagents for albuminoids), viz., hydrochloric, nitric, and meta-phosphoric acids, ferrocyanide of potassium with acetic acid, mercuric chloride (slight), tannic and picric acids, platinum chloride, Millon's solution, and cupric sulphate (the latter on boiling only).

From the young leaves, Greshoff has extracted 0·25 per cent. of a base he calls

(*a*) *CARPAINE* A., crystalline; M.P. 115°, part sublimes without decomposition; extremely bitter (perceptible 1 in 100,000), not very poisonous (acts on heart).

Soluble in alcohol, chloroform and ether; easier soluble in the latter solvent when freshly precipitated than in the crystalline condition; dissolved in traces by water.

Not removed from acid solutions by ether or chloroform.

Precipitants:

Alkaline hydrates, insoluble in excess.
Sodium carbonate.
Tannic acid.
Picric acid, limit 1 in 30,000.
Gold chloride, limit 1 in 25,000.
Potassium sulphocyanide.
Phospho-molybdic acid, limit 1 in 75,000.
Mercuric-potassic iodide } limit 1 in 300,000.
Iodine solution }

Colour tests (negative):

Concentrated sulphuric acid
Concentrated sulphuric acid with nitric acid } no colours.
Concentrated nitric acid
Acids generally

§ 163. **PAREIRA** brava (Botryopsis platyphylla; not Cissampelos pareira as usually stated = Geissospermum); *Menispermaceæ*. Alkaloid (*a*). Both contain also Pelosine = Bebeerine (see Bebeeru).

Pau-Pareira, *Apocynaceæ*. alkaloid (*b*).

(*a*) *GEISSOSPERMINE* A. (Hesse's Geissospermine), $C_{19}H_{21}N_2O_2+H_2O$. Crystallizes in small prisms; M.P. 160°; lævo-rotatory $[a]_D = -93{\cdot}37$. The hydrochloride is amorphous, but crystalline salts are formed with sulphuric and oxalic acids.

Soluble in hot alcohol (difficultly cold), also in benzene and chloroform; not in water or ether.

Removed by benzene or chloroform from *acid* solution, and by amyl alcohol from an *alkaline* infusion.

Reactions:

Alkaline hydrates } crystalline pp.
" carbonates }
Ammonia
Ferric chloride, blue coloration.
Platinum chloride, yellow amorphous pp.
Gold chloride, deep red coloration.

Colour tests:

Concentrated sulphuric acid, colourless, changing to blue, then again colourless.
Concentrated sulphuric acid with potassium bichromate, as Strychnine.
Concentrated nitric acid, purple.
Fröhde's solution, persistent dark blue.

(*b*) *PAREIRINE* A. (Pereirine), $C_{19}H_{24}N_2O$? Grayish amorphous powder; M.P. 124°; salts very soluble.

Soluble in alcohol, ether, chloroform, and with difficulty in water.

Removed from alkaline solutions by chloroform, benzene and petroleum ether.

Colour tests:

Concentrated sulphuric acid, violet.
Concentrated nitric acid, blood-red, but not blue with stannous chloride, as Brucine.

§ 164. **PARIS** quadrifolia L. ; *Liliaceæ*. Investigators : Walz, Delffs, sec *Jahrb. Pharm.*, vols. 4 to 6, *N. Jahrb. Pharm.*, vols. 9 and 13.

(*a*) *PARIDIN* G., $C_{16}H_{28}O_7$ (Delffs). Crystallizes in silky needles containing water ; neutral reaction, pungent taste. Acids convert to *Paridol* and sugar.

Soluble in 50 parts alcohol of 94 per cent. ; scarcely in ether ; difficultly (?) in water. (Though Walz so describes its solubility in water, Delffs separates it from Paristyphnin by its readier solution in that fluid.)

(*b*) *PARISTYPHNIN* G., $C_{33}H_{64}O_{18}$; yellowish-white amorphous ; bitter. Converted by acids into sugar and *Paridin*, the latter again undergoing change, see (*a*), into Paridol and sugar.

Soluble in water and alcohol, not in ether.

Reactions :

Alkaline hydrates, dissolve yellow.
Ammonia dissolves.
Tannic acid, precipitates.

§ 165. **PARMELIA** ceratophylla (P. physodes) ; *Lichenes*. Investigators : O. Hesse, *Ann. Chem. Pharm.*, 119, 365 ; Gerding, *Archiv. Pharm.* [2], 87, 1. Mem. : P. parietina contains chrysophanic acid.

(*a*) *CERATOPHYLLIN*. Prismatic crystals ; M.P. 147°, with sublimate ; neutral reaction, burning taste. Hydrochloric acid precipitates from solution.

Soluble in boiling water (scarcely cold), also in alcohol and ether.

Reactions :

Alkaline hydrates, dissolve when hot, (a) separating on cooling.
Lime water, Ammonia } dissolve.
[No pp. alcoholic lead acetate.]
Ferric chloride, purplish-violet coloration.
[Silver nitrate, no reaction.]
Chloride of lime to the alcoholic solution, blood-red.
Concentrated sulphuric acid, dissolves unchanged.
Concentrated nitric acid, dissolves slightly yellow.

(*b*) *PHYSODIN*. Crystallizes in microscopic pillars ; M.P. 125° ; neutral reactions.

Soluble in ether and hot alcohol.

Reactions :

Alkaline hydrates, dissolve, yellow, changing to red.
Concentrated sulphuric acid, violet solution from which water precipitates bluish-violet flocks.

§ 166. **PASTINACA** sativa (Parsnip) ; *Umbelliferæ* ; substances (*a*) and (*b*), the latter also in Heracleum. Investigators : Wittstein, *Repert. Pharm.*, 68, 15 ; Gutzeit. *Beiträge z. Pflanzenchemie*, 1879.

(*a*) *PASTINACINE* A. Volatile liquid with disagreeable odour ; sharp taste and alkaline reaction. Sulphate crystalline.

(*b*) *HERACLIN* B., $C_{32}H_{22}O_{10}$; M.P. 185°.

Soluble in chloroform but not in water.

§ 167. **PAULLINIA** pinnata ; *Sapindaceæ* (for P. sorbilis containing Caffeine, see Thea group). Investigator : Martin, *Bull. Gen. de Thérap.*

TIMBONINE A. Poisonous. Sulphate crystalline.

§ 168. **PAYTA** bark ('White Payta quinine bark,' from a species of Aspidosperma ; compare Quebracho bark) ; *Apocynaceæ*. Investigator : Hesse, *Ann. Chem. Pharm.*, vols. 154, 166, 178, 211, 249.

(*a*) *PAYTINE* A., $C_{21}H_{24}N_2O + H_2O$; prismatic crystals ; M.P. 156° (loses water at 130°) ; lævo-rotatory [a] D = −49·5. Salts crystalline.

Soluble in alcohol, ether, chloroform, benzene, petroleum ether, with difficulty in water.

Removed from alkaline solution by ether.

Reactions :

Alkaline hydrates, Ammonia } dissolve slowly.
Platinum chloride, dark yellow pp. (dark red on warming).
[Gold chloride, reduced.]
Potassic iodide, pp.

Colour tests :

Concentrated nitric acid gives a colourless solution changing to red.
Perchloric acid, fuchsine-red on heating.
Chloride of lime, dark red, then blue.

(*b*) *PAYTAMINE* A., $C_{21}H_{24}N_2O$ (isomeric with Paytine) ; amorphous, as are the salts.

Soluble in ether.

Reactions :

Platinum chloride, pp.
[Gold chloride, reduced.]
Potassic iodide, pp.
Perchloric acid, fuchsine-red on boiling.

§ 169. **PETROSELINUM** sativum (Apium petroselinum, Parsley); *Umbelliferæ*. In small quantity also in A. Graveolens (Celery).

(*a*) *APIIN* G., $C_{27}H_{32}O_{16}$ (v. Gerichten). Crystallizes in fine needles from alcohol, but usually amorphous; M.P. 228°, leaving glassy mass on cooling; dextro-rotatory in alkaline solution, [a] J = +173°; neutral, tasteless. Warmed with 40 times its weight of hydrochloric acid, sp. gr. 1·04, it gives Apigenin, and 41 to 45 per cent. of glucose.

Soluble in hot water, forming a jelly on cooling (even 1 in 1,500), also in 390 parts cold alcohol, more readily boiling, gelatinizing as solution cools; not in ether.

Reactions:
Alkaline hydrates, „ carbonates, Ammonia } dissolve; acids reprecipitate.
Ferric chloride, dark-red coloration.
Ferrous sulphate, deep blood-red on warming.

(*b*) *APIGENIN* G.-derivative, from preceding; $C_{15}H_{10}O_5$; M.P. 292°-295°. Properties similar to those of Apiin.

§ 170. **PHILLYREA** latifolia, Ph. angustifolia and Ph. media; *Oleaceæ*. The bark. Investigators: Bertagnini and De Luca, *Compt. Rend.*, **51**, 368, also *Ann. Chem. Pharm.*, vols. 24 and 92.

(*a*) *PHILLYRIN* G., $C_{27}H_{34}O_{11}+1\frac{1}{2}H_2O$; silvery crystalline scales; M.P. 160° (when anhydrous); at first tasteless, then a gradually perceptible bitterness; odourless. Acids convert to Phillygenin and glucose.

Soluble in 1,300 parts water at 9° C., or 40 parts alcohol at same temperature; more readily in either solvent boiling, also in warm glacial acetic acid, scarcely in ether; insoluble in oils.

Reactions (negative):
Alkaline hydrates, Ammonia } no action, insoluble.
No pp. metallic salts (Bertagnini).
[Fehling's solution not reduced until after inversion of the Phillyrin.]

Colour Tests:
Concentrated sulphuric acid, reddish-violet.
Nitric acid, decomposes.
Dilute hydrochloric acid, produces yellow crystalline substance and oxalic acid.

(*b*) *PHILLYGENIN* G.-derivative, $C_{21}H_{24}O_6$; pearly crystals (from ether); fusible.

Soluble in alcohol less readily than Phillyrin, also in ether, but scarcely in water.
Alkaline hydrates, Ammonia } dissolve.
Concentrated sulphuric acid, red.
Nitric acid, active decomposition.

§ 171. **PHYLLANTHUS** Niuri, *Euphorbiaceæ*. Java. Investigator: M. Ottow.

PHYLLANTHIN B., $C_{30}H_{37}O_8$. Crystallizes in needles, volatilizes at 200° C.; intensely bitter, poisonous.

Soluble in alcohol, ether, chloroform, benzene, petroleum ether, and glacial acetic acid; slightly in water.

§ 172. **PHYSALIS** Alkekengi (Winter Cherry); *Solanaceæ*. Investigators: Dessaignes and Chautard, *Journ. Pharm.* [3], **21**, 24.

PHYSALIN B., $C_{14}H_{16}O_5$; amorphous; semi-fluid on heating to 190° C.; bitter taste; becomes electrified on rubbing.

Soluble in alcohol and chloroform, with difficulty in water or ether.
Removed from aqueous solution by chloroform.

Reactions:
Ammonia dissolves.
Lead acetate, basic, gives yellow pp. = $C_{14}H_{15}PbO_5 \cdot Pb_2O$.

§ 173. **PILOCARPIUS** pinnatifolius (Jaborandi); *Rutaceæ*. Alkaloids, (*a*), (*b*), (*c*), etc.; Piper reticulatum (false Jaborandi), alkaloids (*c*), (*d*). Investigators: Hardy and Calmels.

(*a*) *PILOCARPINE* A. (Pyridine - methyl - betaine), $C_{11}H_{16}N_2O_2$ or $(C_5H_4N \cdot CH_3)=C=[CO \cdot O \cdot N(CH_3)_3]$; crystallizable with difficulty; dextro-rotatory $[a]_D = +103°$ in alcoholic solution. Gives Jaborine, (*c*), on heating with acid.

Soluble in water (easily), and in alcohol, ether, chloroform, benzene.
Removed from alkaline solution by immiscible solvents.

Precipitated by most alkaloid reagents, including gold chloride [see (*b*)] and phospho-tungstic acid.
[Not by sodium acetate, etc. (compare Opium).]
[Not by potassium ferrocyanide, the salt being readily soluble, nor by 5 per cent. chromic acid.]

Concentrated sulphuric acid, colourless.
" " " with sugar, no effect.
" " " with potass. bichromate, green.
" nitric acid in large excess, converts to Jaborandine, (*e*).

(*b*) *PILOCARPIDINE* A. (from (*a*) by 12 hours' boiling with Baryta, and naturally in Jaborandi leaves), $C_{10}H_{14}N_2O_2$ or $(C_5H_4N\cdot CH_3)C[COOH\cdot N(CH_3)_2]$; deliquescent. Separated from (*a*) by readier solubility of the nitrate of Pilocarpidine.

Soluble in water, alcohol, chloroform (very readily), amyl alcohol and acetic ether; with difficulty in ether or benzene, and not in petroleum ether.

[Not **precipitated** by gold chloride.]
Phospho-tungstic acid, white pp., dissolved by washing with water.

(*c*) *JABORINE* A. (from Jaborandi and Piper reticulatum, and from Pilocarpine), $C_{22}H_{32}N_4O_4$ or

$$C_5H_4N\cdot CH_2):C:[CO\cdot ON(CH_3)_3][N(CH_3)_3O\cdot CO]:C:(CH_3\cdot C_5H_4N);$$

amorphous, non-volatile, alkaline reaction, very poisonous.

Soluble in ether, also in alcohol and water, but with less facility than Pilocarpine.

Precipitated by ammonia.

(*d*) *JABORIDINE* A. (Jaborandine?), $C_{10}H_{12}N_2O_3$; from Piper reticulatum, and from Pilocarpine by oxidation; crystalline.

Soluble in amyl alcohol and benzene; with difficulty in ether or dilute acids.

(*e*) *JABONINE* A., $C_9H_{14}N_2$ or $(C_5H_4NCH_3)CH\cdot N(CH_3)_2$; volatile.

§ 174. **PIMPINELLA** saxifraga; *Umbelliferæ*. Investigator: Buchheim, *Archiv. Pathol.*, 1872.

PIMPINELLIN B., percentage composition, C 63·48, H 4·07; nitrogen free; crystallizes in needles; M.P. 37°.

Soluble in alcohol, with difficulty in ether, scarcely in petroleum ether, not in water.

Concentrated sulphuric acid dissolves red.

§ 175. **PINUS** sylvestris and other of the *Coniferæ*. Glucoside (*a*), which has been found also in asparagus and beetroot; (*b*) has been obtained from P. sylvestris and Thuya Occidentalis (Kawalier, *Wien Akad. Ber.*), and from the leaves of the latter the glucoside (*c*)—only a gramme or so from a cwt.

(*a*) *CONIFERIN* G. (Laricin, Abietin), $C_{16}H_{22}O_8+2H_2O$; silky needles; M.P. 185°; lævo-rotatory. On hydrolysis, sugar and a substance $C_{10}H_{12}O_3$ is obtained, soluble in ether, difficultly in alcohol, not in water. With sodium amalgam, the solution being kept weakly alkaline, Eugenol is formed (coniferyl alcohol being an intermediate product).

Solubility, 1 in 196 cold water, freely warm; with difficulty in alcohol; not in ether.

[Not **precipitated** by lead acetate, neutral or basic, nor ferric chloride.]
Concentrated sulphuric acid, violet-blue.
" " " with potassium bichromate, gives rise to slow formation of Vanillin.
Concentrated sulphuric acid (2 drops), gives fine blue with one drop of a solution prepared as follows: To alcoholic thymol, water is added as long as clear, then potassium chlorate; let stand and afterwards filter. (A test for Coniferin in wood cells, and an indirect test for wood cell).
Hydrochloric acid and phenol, blue.

(*b*) *PINIPICRIN* G., $C_{22}H_{36}O_{11}$; hygroscopic yellow crystals; M.P. 100° (softens at 55°); bitter. On hydrolysis, Ericinol, $C_{10}H_{16}O$, and two molecules of sugar are formed.

Soluble in water, alcohol, and aqueous ether, but not in absolute ether.
Not **precipitated** by lead acetate, neutral or basic.

(*c*) *THUYIN* G. (Thujin), $C_{20}H_{22}O_{12}$; yellow microscopic four-sided plates; astringent taste. Acids give Thuyigenin, (*d*), and glucose.

Soluble in alcohol and hot water, scarcely in the latter cold.

Reactions:
Alkaline hydrates give yellow solutions, turning brownish-red.
Lead acetate, neutral and basic, yellow pp.
Ferric chloride, dark green coloration.

(*d*) *THUYIGENIN* G.-derivative, $C_{14}H_{12}O_7$; crystallizes in microscopic needles. Acids convert to Thuyetin, $C_{14}H_{14}O_8$ (taking up H_2O); see (*e*).

Soluble in alcohol, with difficulty in water.
Ammonia gives bluish-green coloration.
Precipitated by basic lead acetate (not by the neutral salt).

(*e*) *THUYETIN*, $C_{14}H_{14}O_8$ (from Thuyigenin, possibly identical with Quercetin); yellow crystals.

Soluble in alcohol and ether; with difficulty in water.
Ammonia gives bluish-green solution.
Lead acetate, red pp.
(*f*) *PIMARIC ACID,* from resin, has bitter taste.
Soluble in hot alcohol or hot ether; not in water.

§ 176. **PIPER** longum (Pepper), P. nigrum, P. cubebum, P. caudatum; *Piperaceæ.* Schinus mollis, *Terebinthaceæ;* alkaloid (*a*). The fruit and seed of P. longum, 7 to 9 per cent. of (*a*), besides (*b*).

(*a*) *PIPERINE* A., $C_{17}H_{19}NO_3$ or

$$CH_2<^{O}_{O}>C_6H_3\cdot CH:CH\cdot CH:CH\cdot CO(C_5H_{10}N);$$

monoclinic prisms; M.P. 100° (128°-129·5°, Dupuy); neutral reaction, almost tasteless if pure, optically inactive. The salts dissociate with water.

Soluble in 30 parts cold alcohol, 1 part boiling, 60 cold ether; also in chloroform, benzene, and somewhat in petroleum ether; difficultly in water even boiling.

Removed in part from *acid* solutions by benzene and petroleum ether.

Reactions:

Ferric chloride, pp. (hydrochloric solution).
[Platinum chloride, pp. only if concentd.; double salt very soluble.]
Phospho-molybdic acid, brown pp.
[Phospho-antimonic acid, yellow coloration in dilute solutions.]
Cadmium-potassic iodide, pp.
Mercuric-potassic iodide, yellowish-white pp.
[Mercuric chloride, pp. if concentrated.]

Colour tests:

Concentrated sulphuric acid, yellow; green in 20 hours.
" " " with nitric acid, gradually green.
Nitre added to the solution in sulphuric acid after 10 to 15 hours, the greenish-brown solution becomes reddish-brown.
Nitric acid alone, orange changing to yellowish-green (dissolves slowly).
Fröhde's solution, yellow, turning brown to nearly black.

(*b*) *PIPERIDINE* A. (Hexahydro-pyridine), $C_5H_{11}N$; a liquid with peppery and ammoniacal odour, burning taste; boils at 106°; powerfully alkaline, giving crystalline salts.

Soluble in alcohol, and in all proportions with water.

Reactions with metallic salts like ammonium hydrate.
Bismuth-potassic iodide gives a pp. with the hydriodide.

§ 177. **PISCIDIA** erythrina; *Leguminosæ—Papilionaceæ.* The root-bark. Investigator: E. Hart.

PISCIDIN B., $C_{29}H_{21}O_8$; prismatic crystals.

Soluble in hot alcohol, also in benzene and chloroform; with difficulty in ether; not in water. Dilates the pupil.

§ 178. **PLUMERIA** lancifolia (Agoniada bark); *Apocynaceæ.* Investigators: Peckolt, *Archiv Pharm.*, [2], 192, 34; and Geuther. Used as febrifuge.

(*a*) *AGONIADIN* G., $C_{10}H_{14}O_6$ (Genther); crystallizes in radiating silky needles; M.P. 155° (no sublimate); very bitter. Dilute acids produce a brown amorphous substance and sugar.

Soluble with difficulty in boiling water, scarcely in the cold; not readily in cold alcohol, but better in warm; dissolved also by a mixture of ether and alcohol, and by carbon bisulphide; not by chloroform or petroleum ether.

Reactions:

Alkaline hydrates, dissolve brown on warming.
Ammonia, dissolves cold.
[Not precipitated by lead acetate, neutral or basic.]
Tannic acid, pp. gradually.
Ferric chloride, pp.
[Not precipitated by other metallic salts.]

Colour tests:

Concentrated sulphuric acid, dissolves light yellow, changing to green
" " " with nitric acid, gradually yellow.
" nitric acid, golden-yellow.
" hydrochloric acid, dissolves colourless.

(*b*) *AGONIAPICRIN* B.; brownish amorphous, neutral, bitter, not poisonous.

Soluble in water, alcohol and ether.

Reactions:

Tannic acid, brown pp.
[Ferric chloride, brown coloration.]
Platinum chloride, dark brown pp.

§ 179. **PODOPHYLLUM** peltatum (May Apple); *Ranunculaceæ.* Investigator: Podwyssotski, *Archiv Pharmak. u. exper. Pathol.*, 13, 29.

Soluble in 20 parts water; miscible in all proportions with alcohol, ether, or chloroform.

Removed by chloroform from an alkaline, but not from a bicarbonate aqueous solution.

Precipitants:

Alkaline hydrates (but not if dilute).	Iodo-potassic iodide.
[Not bicarbonates.]	Cadmium-potassic iodide.
Tannic acid.	Bromine water.
[Not platinum chloride.]	Gold chloride; reduction of gold on warming.
Phospho-molybdic acid.	

Concentrated sulphuric acid with potassium bichromate, green.

(*b*) *METHYL-PELLETIERINE* A., $C_9H_{17}NO$; liquid; B.P. 215°; hydrochloride dextro-rotatory; salts very hygroscopic.

Soluble in 25 parts water, and in alcohol, ether, chloroform.

Removed by chloroform from alkaline as well as from bicarbonate solutions.

Precipitated by alkaline hydrates and bicarbonates (from strong solutions).

(*c*) *PSEUDO-PELLETIERINE* A., $C_9H_{15}NO+2H_2O$; prismatic crystals; M.P. 46° (or 48°, Ciamician and Silber); B.P. 246°; slightly volatile even cold, optically inactive, reaction alkaline, odorous.

Soluble in water, alcohol, chloroform, petroleum ether, and ether (9 parts of latter).

Removed by chloroform from alkaline and from bicarbonate solutions.

Precipitants:

Alkaline hydrates } (if not too	Iodo-potassic iodide, pale brown needles.
,, bicarbonates } dilute).	Cadmium-potassic iodide.
Picric acid, yellow needles.	The free alkaloid precipitates lime and alumina from their salts.
Gold chloride.	
Phospho-molybdic acid, pale yellow.	

Concentrated sulphuric acid and potass. bichromate, intense green.

(*d*) *ISO-PELLETIERINE* A., $C_8H_{13}NO$; liquid; B.P. 195°; optically inactive.

Solubility as Pelletierine.

Removed by chloroform from alkaline, but not from bicarbonate solution.

Precipitated by alkaline hydrates (from strong solutions), but not by bicarbonates.

§ 183. **PYRUS** malus (Apple); *Rosaceæ*; also in cherry, peach and pear. The root-bark. 'Isophlorrhizin,' from the leaves of the apple, has been shown by Schiff to be identical with (*a*). For Amygdalin, obtained from various species of Pyrus, etc., see Prunus.

(*a*) *PHLORRHIZIN* G. (Phloridzin, Isophlorrhizin), $C_{21}H_{24}O_{10}+2H_2O$ (Strecker); crystallizes in needles; loses water at 100°, melts 106°-109°, rehardens at 130°, then remelts at 160°; lævo-rotatory, $[a]_R=-39{\cdot}9$ (Bouchardat); neutral reaction; slightly bitter, then sweet. Acids convert slowly into (*b*) and sugar.

Soluble in 2,000 parts cold water, readily warm, and in all proportions at 100°; also in alcohol, wood spirit, glacial acetic acid, and in a mixture of ether and alcohol, but scarcely in ether alone.

Reactions:

Alkaline hydrates dissolve.

Ammonia dissolves, solution becoming yellow~red~blue on exposure.

Precipitated by basic acetate lead (not by the neutral salt).

Concentrated sulphuric acid dissolves yellow, changing to red at 30°-60° C.

Fröhde's reagent, pure blue, lasting for a few minutes.

(*b*) *PHLORETIN* G.-derivative (from Phlorrhizin), $C_{15}H_{14}O_5$ (Strecker); M.P. 180°; sweet taste. Boiling potash gives Phloretic Acid and Phloroglucin, $C_6H_6O_3$.

Soluble in alcohol, hot glacial acetic acid, and in wood spirit, but scarcely in water or ether.

Alkaline hydrates dissolve.

Neutral lead acetate precipitates.

Concentrated sulphuric acid } as Phlorrhizin.
Fröhde's solution }

§ 184. **QUASSIA** amara, Picræna excelsa, Simaruba amara; *Simarubaceæ*. From Brucea (quassioides?) F. Eyken has isolated 'Brucamarin,' the melting point and solubilities of which coincide exactly with those of Quassiin, but he finds it to be nitrogenous, and to give a violet coloration with concentrated sulphuric acid; it is soluble in alkalies.

QUASSIIN B. (Picrasmin?—Massube, *Arch. Pharm.*, [3], **23**, 147, considers Picrasmin to be a higher homologue of Quassiin. Brucamarin (?), see above), $C_{10}H_{12}O_3$ (F. Massube gives $C_{35}H_{46}O_{10}$); crystallizes in opaque pillars or rectangular plates; M.P. 215°-217°; dextro-rotatory, neutral reaction, extremely and persistently bitter taste.

Soluble in 222 parts cold water, in alcohol, chloroform and benzene; with difficulty in ether or petroleum ether.

Removed from acid aqueous solution by chloroform and benzene.

Reactions:

[Not precipitated by lead acetate; the basic acetate gives slight cloud.]
Tannic acid, precipitates the alcoholic solution.
Ferric chloride, brown coloration on warming.
[Not precipitated by gold chloride, or silver nitrate.]

Colour tests:

Concentrated sulphuric acid, colourless, then gradually brown.
" " " with sugar, gradual red.
Fröhde's solution, brown.

§ 185. **QUEBRACHO** Colorado (Loxopterigium Lorentzii); *Anacardiaceæ*. Investigator: O. Hesse, *Ann. Chem. Pharm.*, **211**, 249.

LOXOPTERYGINE A., $C_{26}H_{34}N_2O_2$; M.P. 81°; amorphous, alkaline, bitter.

Soluble in alcohol, ether, chloroform, benzene, acetone; with difficulty in water.

Colour tests:

Concentrated sulphuric acid with potass. permanganate, violet.
" " " " chromic acid, violet.
Nitric acid, red.
Perchloric acid, reddish-brown on warming.
Fröhde's reagent, violet~blue.

§ 186. **QUEBRACHO** bark (Aspidosperma Quebracho; compare also Payta bark); *Apocynaceæ*. Investigators: Frande, *Ber. d. Chem. Ges.*, 1878, 1879, 1881; Hesse, *loc. cit.*, 1880, and *Ann. Chem. Pharm.*, **211**, 249.

(*a*) *ASPIDOSPERMINE* A., $C_{22}H_{30}N_2O_2$; crystallizes in needles or prisms; M.P. 205°-206°; feebly alkaline, bitter; lævo-rotatory, $[a]_D = -100{\cdot}2$ for alcoholic solution. **Salts** amorphous.

Soluble in 6,000 parts cold water, 48 alcohol (99 per cent.) at 18° C., 106 absolute ether, and in benzene, chloroform, and amyl alcohol.

Removed from ammoniacal solution by last three solvents.

Precipitants:

Alkaline hydrates.
Ammonia.
Tannic acid.
Picric acid, yellow.
? Not platinum chloride, but blue coloration (yellow pp., Dupuy).
[Fehling's solution, reduced.]
Potassium sulphocyanide, white.
Phospho-molybdic acid, white.
Iodo-potassic iodide, brown.
Mercuric-potassic iodide, yellow flocculent.
Mercuric chloride, white.

Colour tests:

Concentrated sulphuric acid with lead peroxide, red (violet if not pure).
Concentrated nitric acid, red.
" hydrochloric acid warm, red.
Perchloric acid, sp. gr. 1·13 to 1·4, intense red (white pp., Dupuy).

(*b*) *ASPIDOSPERMATINE* A., $C_{22}H_{28}N_2O_2$; crystallizes in needles; **M.P.** 162°; lævo-rotatory; $[a]_D = -72°$; salts amorphous.

Soluble in water if freshly precipitated; also in alcohol, ether, chloroform, and boiling petroleum ether.

No colour reaction with concentrated sulphuric acid and potassium bichromate.

(*c*) *ASPIDOSAMINE* A., $C_{22}H_{28}N_2O_2$; isomeric with (*b*). Crystalline, or amorphous, becoming gradually crystalline; M.P. 100°.

Soluble in alcohol, ether, chloroform, benzene; with difficulty in petroleum spirit; scarcely in water.

Precipitated by:

Sodium platinum chloride.
Gold chloride.
Potassium sulphocyanide.
Mercuric chloride.

Colour tests:

Concentrated sulphuric acid with potassium bichromate, blue.
Perchloric acid, fuchsine red.
Fröhde's reagent, blue.

(*d*) *QUEBRACHINE* A., $C_{21}H_{26}N_2O_2$; crystalline, colourless, but turning yellow on exposure; M.P. 214°-216°; dextro-rotatory; $[a]_D = +62{\cdot}5$, in alcohol, or 18·6 in chloroform; alkaline reaction, bitter. The alcoholic solution gradually acquires a blue tint. **Salts** crystalline.

Soluble in hot alcohol, hot ether, or hot petroleum ether (difficultly in either of these when cold), in amyl alcohol and chloroform; scarcely in water.

Removed by chloroform from *acid* solution (distinction from strychnine) and from ammoniacal solution by amyl alcohol, though with difficulty; not removed from acid solution by benzene or petroleum ether.

Precipitants:
Alkaline hydrates.
,, carbonates.
Gold chloride, yellow amorphous.
Potassium sulphocyanide.
Mercuric chloride.

Colour tests:
Concentrated sulphuric acid with potass. bichromate, blue.
Perchloric acid, yellow.
Fröhde's reagent, blue.

(*e*) *HYPOQUEBRACHINE* A., $C_{21}H_{26}N_2O_2$ (isomeric with Quebrachine); amorphous, yellowish; M.P. 80°; alkaline, bitter, salts amorphous.

Soluble in alcohol, ether, chloroform.

Precipitants:
Alkaline hydrates, pp. resinous.
[Ammonia to concentrated solutions.]
Ferric chloride, red coloration with the hydrochloride.
Sodium platinic chloride, yellow~~red.
Gold chloride, yellow~~violet.

Colour tests:
Concentrated sulphuric acid, violet.
Perchloric acid, yellow.
Fröhde's reagent, violet.

(*f*) *QUEBRACHAMINE* A.; crystallizes in anhydrous needles or leaflets; M.P. 142°; bitter.

Soluble with difficulty in alcohol, ether, chloroform, or benzene, and scarcely in water.

Precipitated by ammonia.
Other reactions similar to those of Quebrachine.

Colour tests:
Concentrated sulphuric acid, bluish.
,, ,, ,, with molybdic acid } intense blue.
,, ,, ,, with potass. bichromate }

§ 187. **QUERCUS** tinctoria (Black Oak); *Amentaceæ*; substance (*a*), which is further found in Chinese tea leaves, vine leaves, Sumach, Carya tomentosa, Rhus coriaria, Fraxinus leaves, etc. Substance (*b*), Quercetin, the derivative of (*a*), is also obtainable from Robinin (see Robinia) and Rutin (see Ruta); a similar compound is present in the rhizome of Podophyllum peltatum (see Podophyllo-quercetin). Compare also Rhamnetin (see Rhamnus infectorius).

(*a*) *QUERCITRIN* G. (Quercimelin; *not* identical with Rutin, Robinin, Rhamnin, Morindin, or Thuyin), $C_{36}H_{38}O_{20}$ (Liebermann), or $C_{21}H_{22}O_{12}$ (Herzig and Smoluchowski, January, 1893). Yellow microscopic plates; loses water at 100°, anhydrous at 165°; M.P. between 160° and 200°, giving some sublimate of Quercetin; neutral reaction, bitter (in alcoholic solution). On hydrolysis, Isoldulcite and Quercetin are formed.

Solubility, 1 in 2,435 cold water or 145 boiling; 23 parts cold alcohol or 3·9 boiling; also in amyl alcohol and warm acetic acid; scarcely in ether; not in benzene, chloroform, petroleum ether, or carbon bisulphide.

Reactions:
Alkaline hydrates } dissolve yellow.
Ammonia }
Lead acetic, neutral or basic, orange pp. solution in acetic acid.
Ferric chloride, dark green coloration.
Gold chloride, reduced in the cold with precipitation of gold.
Silver nitrate, similar reaction (precipitation of silver).
Fehling's solution, reduced after boiling.

(*b*) *QUERCETIN* G.-derivative [=Rhamnetin (Stein); Meletin, Monohydroxi-fisetin (J. Herzig)]; $C_{24}H_{16}O_{11}$ or $C_{15}H_{10}O_7$ (J. Herzig and Smoluchowski, January, 1893) (Zwenger finds $C_{13}H_{10}O_6$). Crystallizes in yellow needles; on heating to 120°, 7 to 10 per cent. of water is evolved; M.P. 250°; recrystallizing on cooling; partially sublimable; neutral reaction, bitter.

Soluble in 229 parts cold alcohol or 18 boiling; not easily in ether; scarcely in water, hot or cold; it dissolves in warm glacial acetic acid, separating on cooling.

Reactions:
Alkaline hydrates, dissolve yellow.
Ammonia gives yellow solution, changing to red.
Lead acetate precipitates the alcoholic solution red.
Ferric chloride, dark green coloration cold or dark red on heating.
Gold chloride
Silver nitrate } as Quercitrin.
Fehling's solution }

§ 188. **RATANHIA** (Krameria triandra, Rhatany): *Krameriaceæ*.

Investigators: Ruge, *Jahresb. Chem.*, 1862, 493; Gintl, *Chem. Cent.*, 1869 and 1870.

Fereira spectabilis, *Rubiaceæ;* the resin (Angelin Pedra resin) gave 86·5 per cent., and the tree 2·45 per cent. Gintl, *Wiener. Akad. Ber.*, **58**, 443.

Geoffroya surinamensis (Andira retusa), *Leguminosæ—Papilionaceæ;* 30 gms. from 1 lb. of the bark. Investigators: Huttenschmidt, *Mag. Pharm.*, **7**, 287; Winckler, *Jahrb. Pharm.*, **2**, 159.

The identity of the substance obtained from the above various sources was suggested by Gintl, and has now been established by Hiller-Bombien (*Archiv Pharm.*, 1892, **230**, 513).

RATANHINE, Amido-acid (Angeline, Geoffroyine, Surinamine, Methyl-tyrosine), $C_{10}H_{13}NO_3$; crystallizes in pearly needles or leaflets; M.P. 40°-45°, with some sublimate; neutral reaction; tasteless and odourless, but evolving disagreeable odour on heating. Forms salts with both acids and alkalies.

Soluble in 9,480 parts cold water or 125 boiling, scarcely in alcohol (1 in 2,300 boiling), and only dissolved in traces by ether, chloroform, benzene.

Reactions:

Alkaline hydrates } pp. soluble in excess.
„ carbonates }
Ammonia }

[No pp. lead acetate.]

Ferric chloride, violet coloration.

[No pp. platinum chloride even concentrated.]

With freshly-precipitated cupric hydrate, a dark blue solution is obtained, which on evaporation deposits violet needles of composition $(C_{10}H_{12}NO_3)_2Cu$.

Phospho-molybdic acid, partial precipitation.

Mercuric-potassic iodide to Rhatanine salt, partial precipitation.

Nessler solution, pp.

[Mercuric chloride } no pp. with free Ratanhine.]
[„ acetate }

Colour tests:

Concentrated sulphuric acid, colourless.

Fuming nitric acid, pink, then successively red, bluish-violet, green (concentrated nitric acid to Ratanhine suspended in water, then warmed).

Pure nitric acid free from nitrous acid, colourless.

§ 189. **RHAMNUS** group, *Rhamnaceæ;* R. cathartica (Buckthorn), substances (*a*), (*b*), (*c*); R. infectorius (yielding Avignon berries), R. oleoides, R. amygdalina and R. saxatilis (furnishing Persian Yellow Berries), (*a*), (*b*); Rhamnus frangula (Alder Buckthorn), substances (*d*) and (*e*); Cascara sagrada (Rhamnus Purshiana), (*d*), (*e*). Emodin, (*e*), is also found in the root of Rheum officinale (Rhubarb), Nephroma lusitanica (*Lichenes*), and as a product of decomposition of Frangulin, (*d*). Regarding Rhamnetin, compare also Quercetin from Quercus tinctoria, etc.

(*a*) *RHAMNIN* G. (Gellatly's Xantho-rhamnin, Rhamnegin; not Fleury's Rhamnin), $C_{22}H_{28}O_{14}$ (Gellatly); crystallizes in microscopic needles or plates, giving yellow solutions; neutral reaction, nearly tasteless. Acids convert to Rhamnetin and Rhamnin-sugar.

Soluble in water, alcohol (if absolute spirit, only when hot), and glacial acetic acid; scarcely in ether, chloroform, benzene, or carbon bisulphide.

Reactions:

Lead acetate precipitates the alcoholic solution, yellow, but no pp. with aqueous solutions.

Ferric chloride, olive-green colour if dilute, brownish-red when concentrated.

Silver nitrate, reduction to silver.

Fehling's solution, no pp.

Copper acetate, brown pp.

(*b*) *RHAMNETIN* G.-derivative. (Fleury's Rhamnin, Quercetin, see Quercus), $C_{11}H_{10}O_5$ (Gellatly), or $C_{23}H_{16}O_{11}$ (Stein); crystallizes in yellow quadratic, prismatic plates; tasteless.

Solubility varies considerably with degree of dryness. Scarcely soluble in water, though colouring it yellow; 1 in 58 boiling alcohol, 1 in 76 ether, also in phenol.

Reactions:

Alkaline hydrates dissolve yellow; Baryta precipitates therefrom.

Aluminic hydrate gives precipitate.

Ferric chloride to alcoholic solution, brownish-green coloration.

Black dyes formed with iron salts.

Silver nitrate } reduced in the cold.
Fehling's solution }

Chloride of lime, dark green, or brownish-red with concentrated solution.

(*c*) *RHAMNO-CATHARTIN;* bitter, purgative, non-crystallizable.

(*d*) *FRANGULIN* G. (Rhamnoxanthin, Kubly's Avornin, Leprince's

Cascarin), $C_{20}H_{20}O_{10}$ (Casselmann), $C_6H_6O_3$ (Hesse), $C_{21}H_{20}O_9$ (Thorpe and Robinson, 1892); crystallizes in yellow silky needles; M.P. 226°, partly sublimable in yellow needles. Acids convert to Emodin and Rhamnose, $C_6H_{12}O_5$.

Soluble with difficulty in cold alcohol, in 160 boiling; also in chloroform, benzene, carbon bisulphide, and very readily in hot glacial acetic acid; scarcely in ether (the impure substance is soluble), not in water.

Reactions:

Alkaline hydrates, Ammonia } give purple-coloured solutions.
[Tannic acid, Metallic salts } no pp.]
Metallic oxides produce coloured lakes.
Concentrated sulphuric acid, emerald changing to purple, then dark red; water reprecipitates.

(*e*) *EMODIN* G.-derivative (Tri-hydroxi-methylanthraquinone), $C_{15}H_{10}O_5$ (Schwabe, Thorpe); from Frangulin by action of acids, and naturally in various plants, see above; microscopic orange-coloured needles; M.P. above 250°, sublimes slowly at 240°.

Soluble in alcohol more readily than Frangulin, also in amyl alcohol, hot benzene, and hot acetic acid; less difficultly in ether than is the case with Frangulin. Insoluble in water.

Reactions:

Alkaline hydrates, cherry red.	Copper sulphate, Mercuric chloride } pp.
Alkaline earths, red to brown pp.	
Lead acetate, neutral or basic, pp.	

§ 190. **RHINACANTHUS** communis; *Acanthaceæ*.

RHINACANTHIN B., $C_{14}H_{18}O_4$; amorphous.

Soluble in alcohol.

§ 191. **RHINANTHUS** buccalis; Alectorolophus hirsutus (the seeds), Melampyrum, Antirrhinum majus (Snapdragon, the leaves).

RHINANTHIN G., $C_{20}H_{52}O_{20}$? ($C_{64}H_{56}O_{40}$? Phipson); microscopic needles, seen to be prisms when magnified 300 diameters; neutral reaction, bitter-sweet taste. Acids convert to Rhinanthogenin, as dark blue or black flocks, and sugar.

Soluble in water and alcohol; not in ether.

Reactions:

Alkaline hydrates, Ammonia } give yellow solutions.
[Not precipitated by neutral or basic lead acetate.]
[No coloration with ferric chloride.]

§ 192. **RHUS** cotinus (Wig tree); *Anacardiaceæ* or Sumach Order. (Rhus coriaria contains Quercetrin; see Quercus.)

(*a*) *FUSTIN* G., $C_{58}H_{46}O_{32}$; crystallizes in needles; M.P. 218°-219°. Acids produce sugar and Fisetin.

Soluble easily in alcohol, also in hot water, but with difficulty in ether.

Reactions:

Soluble in alkaline hydrates.
Precipitated by lead acetate (neutral).

(*b*) *FISETIN* G.-derivative, $C_{15}H_{10}O_6$ (J. Herzig), compare Quercetin; yellow needles from dilute alcohol, or prisms from acetic acid; becomes anhydrous at 110°, browns at 270° C.; fluorescent in alcoholic alkaline solution. Yields on hydrolysis Protocatechuic acid, $C_7H_6O_4$, and Fisetol, $C_8H_8O_4$.

Soluble in alcohol, acetone, and acetic ether; with difficulty in ether, chloroform, benzene, or petroleum ether; scarcely in cold water.

Reactions:

Lead acetate to alcoholic solution, yellow pp., soluble in acetic acid.
Ferric chloride, greenish-black coloration; black pp. on adding ammonia.
Fehling's solution, reduced on warming.

§ 193. **ROBINIA** pseudacacia (North American Locust tree); *Leguminosæ*. The flowers. Investigators: Zwenger and Dronke, *Ann. Chem. Pharm., Suppl.*, 1, 257.

ROBININ G., $C_{25}H_{30}O_{16}+5\frac{1}{2}H_2O$; silky straw-yellow needles; M.P. 195° (anhydrous at 100°); neutral reaction, feebly astringent taste. Acids convert to Quercetin (see Quercus and a non-fermentable sugar).

Soluble in hot water and hot alcohol, with difficulty in cold water. Insoluble in ether.

Reactions:

Alkaline hydrates, „ carbonates } give yellow solutions.
Ammonia, dissolves yellow changing to brown.

Lead acetate, neutral, a slight pp. in strong solutions, disappearing on dilution.
Lead acetate, basic, pp.
Gold chloride, reduced.
Silver nitrate, slowly reduced.
Fehling's solution, reduced.
[Not precipitated by other metallic salts.]
Concentrated nitric acid oxidizes with formation of some picric acid.

§ 194. **ROTTLERIA** tinctoria (Mallotus Philipinensis); *Euphorbiaceæ*. The dye, 'Kamala dye.' Investigator: Anderson, *Edinb. N. Phil. Jnl.*, new series, 1, 300.

ROTTLERIN B., $C_{22}H_{20}O_6$; yellow fusible crystals.

Soluble in water, hot alcohol (difficultly cold), and in ether.
Alkaline hydrates give deep red solutions.
[Not precipitated by lead acetate in alcoholic solution.]

§ 195. **RUBUS** villosus; *Rosaceæ*: N. America. The root. Investigator: G. A. Krause.

(*a*) *VILLOSIN* G.; M.P. 173°-175°; bitter in alcoholic solution. On prolonged boiling in water alone, or more readily with dilute acid, glucose and Villosic acid, (*b*), are produced.

Soluble in methyl, ethyl and amyl alcohols, with difficulty in water, scarcely in ether, not in chloroform.

Reactions:
Alkaline hydrates dissolve with gradual production of lemon coloration.
Lead acetate, basic, pp.
[Not neutral lead acetate.]
[Ferric chloride, no change.]
[Silver nitrate, not reduced.]
[Fehling's solution, not reduced until (*a*) has been inverted.]

Colour tests:
Concentrated sulphuric acid, light brown; with water, deep violet.
„ „ „ with nitric acid, deep orange, destroyed on warming.
Fröhde's solution, dark brown, changing to red with a drop of nitric acid.

(*b*) *VILLOSIC ACID* G.-derivative; bitter, resinous (crystals from ether).

Colour tests:
Concentrated sulphuric acid, 2 drops water added, then warmed, blue.
„ „ „ with 2 drops nitric acid, blood-red.

§ 196. **RUTA** graveolens (Rue); *Rutaceæ*. Capers (from Capparis spinosa), Chinese Yellow Berries (Sophora japonica, *Leguminosæ—Papilionaceæ*), and, according to Stein, Safflower (Carthamus tinctorius, the *yellow* colouring matter), and the straw of Polygonum fagopyrum, etc.; substance (*a*). Foerster (*Berlin Ber.*, **15**, 214) regards the glucoside of Sophora as not identical with (*a*), but a distinct substance, (*b*). (For Sophora speciosa, see Sophora.)

(*a*) *RUTIN* G. (Rutic Acid, Phytomelin, Melin, Sophorin?), $C_{25}H_{28}O_{15}+2\frac{1}{2}H_2O$, Zwenger (if Herzig's formula for Quercetin be correct, see Quercus, this must be altered correspondingly); crystallizes in yellow silky needles; carbonizes on heating, with production of caramel odour; neutral reaction, bitter (when in solution). Acids produce Quercetin and Isodulcite.

Soluble in hot water, scarcely cold (but colouring the water yellow), and in hot alcohol (with difficulty cold), also in hot glacial acetic acid; but not in ether.

Reactions:
Alkaline hydrates } dissolves yellow, solutions turning brown.
Ammonia }
Barium and calcium hydrates also dissolve.
Lead acetate (neutral) yellow pp.
Ferric chloride, dark green coloration.

(*b*) *SOPHORIN* G. (see note above), resembles Rutin, but gives on hydrolysis 57·5 per cent. of Isodulcite besides Sophoretin (Foerster).

§ 197. **SAFFRON** (Crocus sativus); *Iridaceæ*. The stigma and part of style of the flowers. Used for colouring and flavouring. Investigators: Quadrat, Rochleder and Mayer, Weiss, Kayser. Crocin is also found in the fruit of Gardenia grandiflora (*Rubiaceæ*) and in Fabiana Indica.

(*a*) *CROCIN* G. (Polychroite, not Weiss' Crocin); $C_{44}H_{70}O_{28}$, — Kayser (the older formulæ were: $C_{20}H_{13}O_{11}$, Quadrat; $C_{58}H_{42}O$ Rochleder [? $C_{29}H_{42}O_{15}$]; $C_{48}H_{60}O_{18}$, Weiss); red colouring matter; no odour, but sweet taste. Hydrolyzed very readily by boiling dilute acids and alkalies, etc., as follows:

$$2C_{44}H_{70}O_{28}+7H_2O=C_{24}H_{46}O_9\ (\text{Crocetin})+9C_6H_{12}O_6\ (\text{sugar}).$$

Soluble in water (red solution), with difficulty in alcohol, scarcely in ether.

Absorbed by charcoal, removed therefrom by alcohol.

Reactions:

Alkaline hydrates, dissolve with almost immediate production of sugar and Crocetin.
Lime-water, yellow pp. on warming.
Lead acetate, a red pp. on warming only, hydrolysis taking place.
Copper sulphate, green pp.
Concentrated sulphuric acid, deep blue, becoming in turn violet, cherry-red, then brown.
Nitric acid, sp. gr. 1·4, momentary blue, then brown.
" " 1·120, yellow only.

(*b*) *CROCETIN* G.-derivative (Weiss' Crocin), from above, $C_{34}H_{46}O_{19}$ (Kayser); red-coloured.

Soluble scarcely in water, but easily in alcohol and ether.

Reactions:

Alkaline hydrates, dissolve readily, forming yellow solutions; acids pp.
Calcic hydrate, Baric ", Lead acetate, basic } pp.
Concentrated sulphuric acid, as Crocin.

(*c*) *PICROCROCIN* G. (Saffron Bitter), $C_{38}H_{66}O_{17}$; crystallizes in prisms from ether; M.P. 75°; bitter and characteristic persistent taste.

Soluble in water and alcohol easily, less so in chloroform, and with difficulty in ether. By the action of dilute acids or alkalies it undergoes the following decomposition:

$$C_{38}H_{66}O_{17}+H_2O=3C_6H_{12}O_6(\text{sugar})+2C_{10}H_{16}(\text{volatile essential oil}).$$

Other reactions:

Lime or Baryta water, no pp., but intense saffron odour and cloudiness due to the above change.
Fehling's solution, reduced on warming; green flocks also separate.

§ 198. **SALIX** (Willow); *Salicaceæ;* many species. The barks; also Spiræa and Synantheræ various; and in Castoreum Canadense. For Benzoyl Salicin (Populin) see Populus.

(*a*) *SALICIN* G., $C_{13}H_{18}O_7$; crystallizes in plates, needles, or wide pillars; M.P. rather over 100°—Piria (198°—O. Schmidt); lævo-rotatory, neutral, bitter. Acids and ferments (as Emulsin or Ptyalin) decompose into glucose and Saligenin; when acids are used the Saligenin is further resolved into Saliretin and water; alkalies give Saliretin, salicylic acid, and some salicylous acid.

Soluble in 29 parts cold water, and about its own weight of boiling water; also in alcohol, amyl alcohol, and glacial acetic acid; not in ether.

Reactions:

Alkaline hydrates, Ammonia } dissolve.
[Not precipitated by lead acetate or other metallic salts.]
[Basic lead acetate gives pp. in concentrated solutions.]

Colour tests:

Concentrated sulphuric acid, red; water precipitates from the solution a red substance insoluble in water='Rutilin.'
Fröhde's solution, magnificent violet turning to dark cherry (very permanent: compare Morphine).

(*b*) *SALIGENIN*, G.-derivative (Salicyl alcohol), $C_6H_4OH\cdot CH_2OH$; white pearly plates or rhombohedra; M.P. 82°.

Soluble 1 in 15 cold water, and in almost all proportions when boiling; readily also in alcohol and ether.

Decomposes with dilute acids into Saliretin, C_7H_6O, and water.

Reactions:

Ferric chloride (to the aqueous solution), indigo blue coloration.
Concentrated sulphuric acid, intense red.

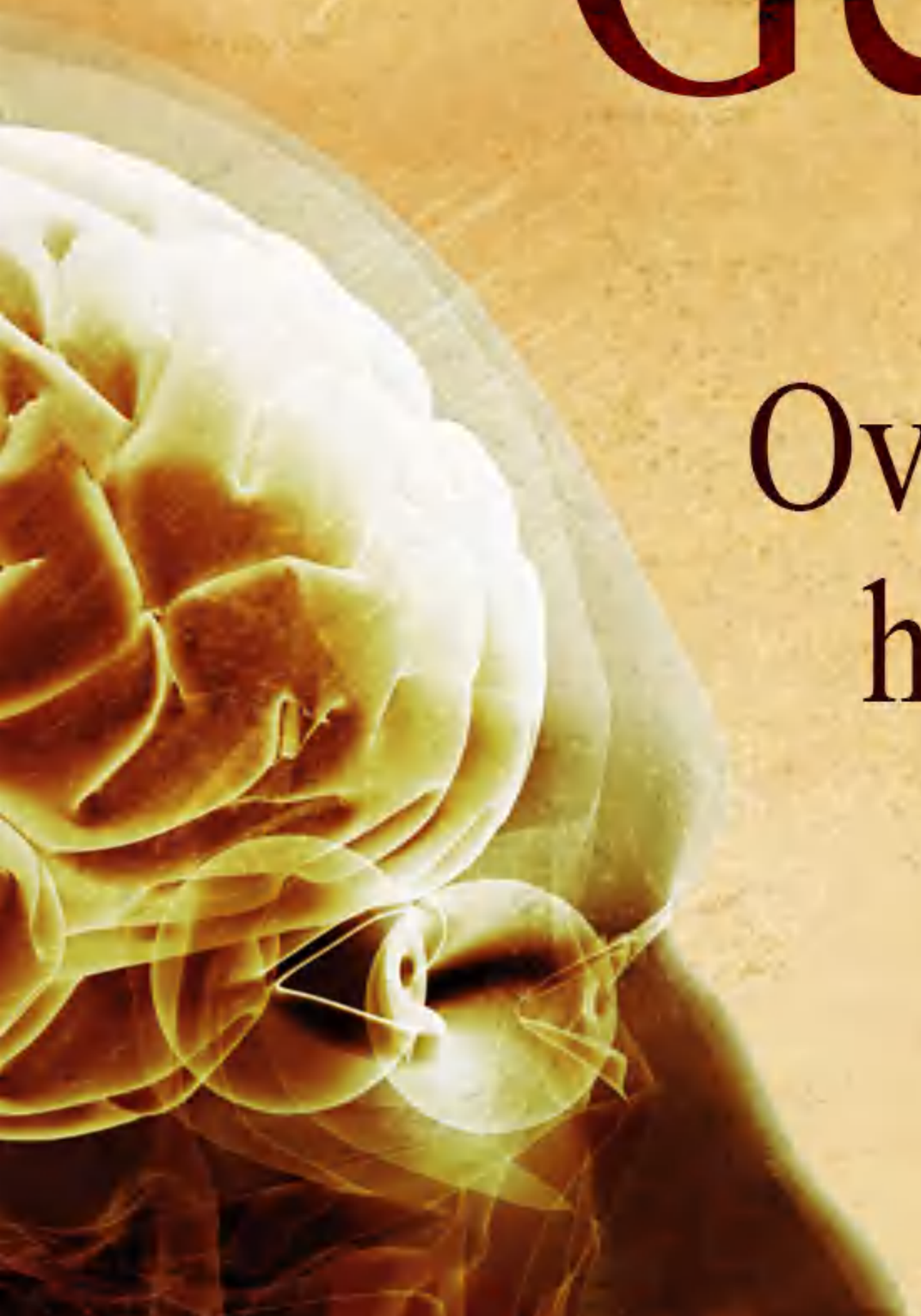
(*c*) *HELICIN* G. From Salicin by action of nitric acid; $C_{13}H_{16}O_7+\frac{3}{4}H_2O$ (the glucoside of salicylous acid). Crystallizes in needles; M.P. 175°, recrystallizing on cooling; neutral, slightly bitter.

Soluble in 64 parts cold water and in alcohol; not in ether.
Concentrated sulphuric acid, yellow.
[No violet coloration with Fröhde's solution.]

§ 199. **SOPHORA** speciosa; *Leguminosæ.* The beans. See also Ruta for Sophora Japonica. Investigator: Wood, *Amer. J. Ph.*, **50**, 203.

SOPHORINE A.; amorphous; alkaline; hydrochloride salt crystalline.
Soluble in water, alcohol, ether, chloroform.
Removed from alkaline solution by latter solvent.
Ferric chloride, blood-red coloration.

§ 200. **SAPONARIA** officinalis, various species of Dianthus, Lychnis, Silene (*Caryophyllaceæ*); Agrostemna Githago (6·5 per cent. of the seed),



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Reactions :
Alkaline hydrates dissolve.
Precipitated by *basic* lead acetate.
[Not by neutral lead acetate.]

§ 203. **SAXIFRAGA** Siberica and **S.** crassifolia ; *Saxifragaceæ.* Investigators : Garreau and Machelart.
BERGENIN G., ? $C_6H_8O_4$ (Garreau and Machelart) ; crystalline, bitter.
Soluble in alcohol and water.
Fehling's solution, reduced.

§ 204. **SCILLA** maritima, or Urginea **S**cilla (**S**quills) ; *Liliaceæ.* **S**ubstance (*a*), Jarmerstedt, *Arch. exp. Pathol.*, **11**, 22 ; (*b*), (*c*), (*d*), Merck, *Pharm. Zt.*, 1879, and others.
(*a*) *SCILLAIN* G. (Jarmerstedt's) ; nitrogen free, amorphous, bitter, poisonous. Resolved by acids into sugar and a resin.
Soluble in alcohol ; with difficulty in water, ether, or chloroform.
[**Not** precipitated by basic lead acetate.]
Tannic acid precipitates.
Colour tests :
Concentrated sulphuric acid, brown, with green fluorescence.
" " " with bromine, bluish-red.
" hydrochloric acid, red, with gradual greenish pp.
(*b*) *SCILLIN* B. (not Riche and Rémoud's **S**cillin, which is a carbohydrate) ; light yellow, crystalline.
Soluble in alcohol and hot ether, with difficulty in water.
Concentrated sulphuric acid, reddish-brown.
" nitric acid, yellow changing to green.
(*c*) *SCILLIPICRIN* B. ; yellow, amorphous, hygroscopic, bitter.
Soluble in water.
(*d*) *SCILLITOXIN* ; brown, amorphous.
Soluble in alcohol. **Insoluble** in water or ether.
Concentrated sulphuric acid, red changing to brown.
" nitric acid, yellow, changing to green on warming.

§ 205. **SCROPHULARIA** aquatica, *L.*, and **S.** nodosa ; *Scrophulariaceæ.* Investigator : Walz, *Jahrb. Pharm.*, **26**, 296.
SCROPHULARIN B. ; crystalline, bitter.
Reactions :
[No pp. lead acetate, neutral or basic.]
Tannic acid, pp.

§ 206. **SIMABA** cedron ; *Simarubaceæ.* The seed, substance (*a*). Investigator : Lewy, *Journ. Chim. Méd.*, 1851, 282. Used against snakebites. **S.** Waldivia (distinct from preceding), substance (*b*), from the fruit (Ch. Tanret, *Compt. Rendu*, **91**, 886).

(*a*) *CEDRINE* (A. ?), (Valdivin ?, Tanret) ; glittering needles, neutral, bitter, emetic. Forms fluorescent compounds.
Soluble in water, alcohol and ether.
Precipitated by tannic acid and by most alkaloid reagents (C. Tanret).

(*b*) *VALDIVIN* G. (?), $C_{36}H_{48}O_{20}+5H_2O$ or $C_{18}H_{24}O_{10}+5H_2O$; crystallizes in hexagonal prisms ; M.P. 230°, non-volatile ; sp. gr. 1·46 ; bitter. Forms compounds showing a magnificent fluorescence. Optically inactive.
Soluble in 600 parts cold water (frothy solution), 30 boiling, 190 absolute alcohol, or 60 spirit of 70 per cent., easily in chloroform ; not in ether.
Removed from aqueous solution by chloroform.
[**Not** precipitated by neutral or basic lead acetate.]
Ammoniacal lead acetate, Tannic acid } pp.
Fehling's solution, reduced after treatment with alkalies (acids ?).
Concentrated sulphuric acid, Nitric acid } dissolve colourless.
Acids in general increase the solubility in water.

§ 207. **SINAPIS** alba (White Mustard) ; *Cruciferæ :* substances (*a*), (*b*). Will and Laubenheimer, *Ann. Chem. Pharm.*, 193, 150. **S.** nigra (Black Mustard), substance (*c*) ; Turritis glabra, (*b*) ; Brassica rapa, (*c*).
(*a*) *SINALBIN* G., $C_{30}H_{44}N_2S_2O_{16}$; crystallizes in needles ; only faint yellow unless impure ; fusible. Decomposed by Myrosin (ferment of mustard) into glucose, **S**inapine sulphate, and Sinalbyl thiocyanide, $C_7H_7O{\cdot}NCS$.
Soluble in water, in 3·3 parts boiling 86 per cent. alcohol ; not in cold absolute alcohol, ether, or carbon bisulphide.
Reactions :
Alkaline hydrates, Ammonia } dissolve yellow.
[No pp. barium chloride until after decomposition.]
Alkaline earths dissolve, the yellow coloration being perceptible even with traces only.

Ferric chloride, no colour till boiled with alkali, then blood-red (sulphocyanide test).
Silver nitrate, white pp., becoming black from production of silver sulphide.
Fehling's solution, reduced, some cupric sulphide being formed.
Mercuric chloride, pp., glucose being left in solution.
Concentrated nitric acid, transient blood-red.

(*b*) *SINAPINE-THIOCYANIDE* A. (Sulphosinapisine, Sulphosinapine; the thiocyanide was formerly mistaken for Sinapine), $C_{16}H_{23}NO_5 \cdot HSCN$ (Gerhardt); glittering prisms; M.P. 130°, recrystallizing on cooling; neutral reaction (free Sinapine is alkaline), bitter. (Free Sinapine rapidly decomposes; not separated pure either by alcohol or ether.

Soluble in water and alcohol (yellow solutions); not in ether, carbon bisulphide, or turpentine.

Alkaline hydrates dissolve yellow.

Ferric chloride, blood-red coloration on warming.

(*c*) *POTASSIUM MYRONATE* G., $C_{10}H_{18}KNS_2O_{10}$; crystallizes in silky prisms, fusible, neutral, bitter. (The free acid decomposes readily, with production of sulphuretted hydrogen; it has acid reaction, with bitter and sour taste). Myrosin converts the glucoside to glucose, Allyl-thiocyanide, and acid potassium sulphate.

Soluble in water, with difficulty in alcohol. **Insoluble** in ether, chloroform, or benzene.

Precipitated by neutral lead acetate, yellowish-white pp., soluble in acetic acid.

(*d*) *SINCALINE*, by action of alkalies on Sinapine; crystalline, non-volatile, powerfully alkaline like potassic hydrate; absorbs carbonic acid.

§ 208. **SOLANUM** various (*Solanaceæ*), *e.g.*, S. nigrum (Nightshade), substance (*a*); S. tuberosum (Potato, the leaves and shoots), (*a*) and (c); S. dulcamara (Bittersweet), (*a*), (*f*), Geissler, *Archiv. d. Pharm.*, [3], **7**, 289, 1875; S. grandiflora var. pulverulentum (Wolf-fruit of Brazil), investigator, H. Freire, *Compt. Rendu*, **105**, 1,074, (*g*); S. paniculatum (the fruit), F. v. Greene, substance (*e*). See also Atropa group for other members of the *Solanaceæ*.

(*a*) *SOLANINE* A.-G., $C_{42}H_{75}NO_{15}$ (Hilger) or $C_{52}H_{92}NO_{18}$ (Firbas); crystallizes in four-sided prisms; M.P. 235°, with some sublimate in needles; feebly alkaline; taste bitter and hot. Dilute acids resolve into sugar and Solanidine. Salts react acid. Sodium amalgam produces butyric acid and nicotine.

Soluble scarcely in water, 1 in 8,000 boiling (the salts are soluble), in 125 boiling alcohol of sp. gr. 0·839 (giving crystals on cooling), in amyl alcohol, 4,000 ether. Benzene dissolves traces.

Removed from aqueous solutions by amyl alcohol.

Reactions:
Alkaline hydrates, Ammonia } amorphous pp.
Tannic acid, pp.; at 1 in 3,000, flocks after 24 hours.
[Picric acid, not dilute.
Platinum chloride, not dilute.]
Gold chloride, Silver nitrate } reduced.
[Fehling's solution, not reduced.]
Silver potassium cyanide, gradually amorphous pp.
[Potassium bichromate, not 1 in 3,000.]
Phospho-molybdic acid, light yellow pp.
[Iodo-potassic iodide, not 1 in 3,000.]
[Bismuth-potassic iodide, pp. only if concentrated.]
[Cadmium-potassic iodide, not dilute.]
[Mercuric-potassic iodide, not 1 in 3,000].
[Mercuric chloride, not 1 in 3,000.]

Colour tests:
Concentrated sulphuric acid, light reddish-yellow.
" " " with potassium bichromate, light blue.
Trace of Solanine + 1 per cent. sulphuric acid evaporated on watch-glass, gives four-sided prisms, which on warming become successively red, purple, brownish-red, and, on cooling, violet, blackish-blue, green (colourless crystals still present).
Concentrated sulphuric acid with nitric acid, light reddish.
Nitric acid added to the solution in sulphuric acid after 10 hours, light yellow.
Nitric acid alone, colourless, then purple, finally colourless.
Concentrated hydrochloric acid, yellow.
Fröhde's solution, cherry-red ~ brownish-red ~ yellow; after 24 hours greenish-yellow with black flocks.
Sulphuric acid and alcohol, red on warming.

(*b*) *SOLANEINE* G., $C_{52}H_{83}NO_{13}$? (Firbas). Amorphous, yellow, horny, M.P. 208°.

Soluble in 85 per cent. alcohol. Acids gives Solanidine and sugar.

(*c*) *SOLANIDINE* A. (G.-derivative). In young shoots of potatoes, and from Solanine. $C_{25}H_{14}NO_2$ (Hilger) or $C_{40}H_{61}NO_2$ (Firbas); crystallizes in glittering needles; M.P. 208° (191°, Firbas); alkaline reaction; bitter and sharp taste. Forms salts difficultly soluble in water.

Soluble in alcohol, ether, chloroform, benzene, scarcely in water.

Removed by chloroform from acid solutions.

Alkaline hydrates, Ammonia } pp. as jelly.

(*d*) *SOLANICINE* A. $C_{50}H_{76}NO_2$? Formed from (*a*) by long-continued action of acids. Light yellow crystals from ether; M.P. 250°, with decomposition; scarcely alkaline, almost tasteless. Salts reddish-yellow, amorphous.

Soluble scarcely in alcohol, in 2,000 parts boiling ether, not in water.

(*e*) *JURUBEBINE* A. Amorphous (hydrochloride crystalline); bitter taste; aromatic odour.

Soluble in alcohol and chloroform, with difficulty in water.

No pp. picric acid or platinum chloride but by most other alkaloid precipitants (F. v. Greene).

(*f*) *DULCAMARIN* B. Bitter.

Soluble in acetic ether; not in ether, chloroform, benzene, or carbon bisulphide.

Reactions:

Alkaline hydrates, Ammonia } dissolve reddish-brown.

Precipitated by basic lead acetate and tannic acid.

Concentrated sulphuric acid, red changing to rose colour.

(*g*) *GRANDIFLORINE* A. Molecular weight 236·4; amorphous, bitter, poisonous.

Soluble in alcohol, not in water.

Reactions:

Alkaline hydrates, evolution of ammonia on heating.

Ammonia, white pp.

Tannic acid, cloud.

Platinum chloride, Iodo-potassic iodide, Mercuric-potassic iodide } yellow pp.

Colour tests:

Concentrated sulphuric acid, yellow turning red.

Concentrated sulphuric acid with manganese dioxide, green changing to violet, then yellow.

Nitric acid, purplish.

§ 209. **SPARTIUM** scoparium (Broom), or Sarothamnus scoparium (Cytisus scops.), *Leguminosæ*. Investigators: Stenhouse, *Philos. Trans.*, 1851; Mills, *Chem. Soc. Qu. J.*, **15**, 1, etc., and others.

(*a*) *SPARTEINE* A., $C_{15}H_{26}N_2$; liquid; sp. gr. greater than water; volatile; B.P. 276°-288° or 311° (different observers); alkaline reaction; bitter taste and aniline-like odour. Lævo-rotatory, −14·6° for D line. Dilates the pupil; becomes brown on exposure; sulphate very soluble.

Soluble in alcohol, ether, chloroform, petroleum ether, scarcely in water, not in benzene.

Removed from *alkaline* aqueous solutions by ether and petroleum ether.

Precipitants:

Alkaline hydrates, Ammonia } pp. soluble in excess.

Lead acetate, neutral, basic; „ „ } white.

Picric acid, yellow.

Platinum chloride, yellow.

[Not silver nitrate.]

[Not potassium ferrocyanide, the compound formed being easily soluble.]

Sodium phospho-molybdate, white.

Cadmium-potassic iodide, white (Bamberger).

Colour tests (negative):

Concentrated sulphuric acid, „ „ „ with nitric acid, Concentrated nitric acid, Acids generally } colourless.

(*b*) *SCOPARIN* G.? $C_{21}H_{22}O_{10}$. Pale yellow crystalline or amorphous. Diuretic.

Soluble in alcohol and hot water (greenish-yellow solution), scarcely cold.

Reactions:

Alkaline hydrates, „ carbonates, Ammonia, Barium and calcium hydrates } soluble, greenish-yellow.

points (anhydrous), but appears in other forms, *e.g.*, scales, octahedra, according to the nature of the solvent, etc. Authorities differ as to melting point, Beckurts finds 265°, but preceding observers have given 221°, 285°, etc. ; læve-rotatory, alkaline reaction, intensely bitter (perceptible at 1 in 600,000), extremely poisonous (tetanic).

Solubility, 1 in 6,667 parts cold water, 2,500 boiling ; 120 cold alcohol sp. gr. 0·863 (requiring more of weaker alcohol), 200 amyl alcohol, 5 to 7 chloroform, about 1,250 commercial ether, 160 of benzene, 12,500 petroleum ether (Wormley), and in oils.

Strychnine nitrate is soluble in 60 parts cold or 2 of boiling water (Wittstein), the sulphate in 50 parts cold water (Regnault).

Removed from alkaline solutions by chloroform ; best to shake with the latter *at once* after alkalizing. A. H. Allen (*Analyst*, 6, 141) finds a convenient menstruum to be a mixture of ether and chloroform in equal measures.

Precipitants:

Alkaline hydrates
„ carbonates
„ bicarbonates, if without free carbonic acid
Ammonia
Calcic hydrate
} pp. insoluble in excess and becoming crystalline.

[Not lead acetate, neutral or basic.]
Tannic acid, difficultly sol. in hydrochloric acid.
Picric acid, yellow amorphous, becoming crystalline ; limit 1 in 20,000.
[Ferric chloride, light brown if concentrated ; not 1 in 100.]
Platinum chloride, yellow amorphous, becoming crystalline ; insol. HCl.
Gold chloride, soluble in hydrochloric acid.
Potassium ferrocyanide, nearly colourless prisms, very difficultly soluble (brucine salt more soluble).
Potassium ferricyanide, greenish-yellow crys.
„ sulphocyanide, white crys. (sensitive).
„ nitroprusside, light brown crys. ; up to 1 in 5,000.
Silver potassium cyanide, hair-like crystals.
Potassium bichromate, yellow crystalline pp. which gives a transient violet with sulphuric acid ; the pp. is very difficultly soluble in cold water (means of separation from Brucine).
Chromic acid (5 per cent. solution).
Phospho-molybdic acid, yellowish-white.
„ tungstic acid, 1 in 200,000.
„ antimonic acid, white flocks at 1 in 5,000 ; cloud at 1 in 25,000.
Iodo-potassic iodide, kermes coloured pp. ; faint cloud at 1 in 50,000.
Bismuth-potassic iodide, orange-red pp. ; limit 1 in 250,000.
Cadmium-potassic iodide, flocculent.
Zinc-potassic iodide, white.
Mercuric-potassic iodide, white—limit 1 in 150,000.
Mercuric chloride, pp. soluble in alcohol (the pp. with free strychnine is soluble). Potass. bichromate gives with this pp. a bright yellow.
Chlorine water, white pp. ; even $\frac{1}{500}$th of a milligramme.

Colour tests:

Concentrated sulphuric acid, colourless.
„ „ „ with fragment of potassium bichromate (avoid excess), violet-blue~red~green. Similar colours with other oxidizers. Compare Geissospermine ; see Pareira.
Nitric acid added to the solution in sulphuric acid after 10 hours, unchanged.
Nitric acid, slightly yellow.
Concentrated hydrochloric acid, no effect.
Potassium perchlorate, white crystalline pp.
Perchloric acid, reddish-brown coloration.
Per-iodic acid, wine-red ; red crystals on evaporation.
Fröhde's solution, colourless.
Strychnine has been found 3 years after death.

(*b*) *BRUCINE* A. (Contramine), $C_{23}H_{26}N_2O_4$: crystallizes in four-sided prisms or plates, also as cauliflower-like aggregates, containing $4H_2O$. M.P. : anhydrous, 170° ; hydrated, rather over 100° ; læve-rotatory, alkaline, bitter. Many oxidizing agents give red coloration.

Soluble in 850 parts cold or 500 boiling water (the hydrated base in 320 and 150 respectively ; crystals only slowly formed from a hot solution after cooling), in $1\frac{1}{2}$ parts boiling alcohol, about 2 of chloroform (Rettendorfer) (7 parts, Schlimpert), 64 benzene ; also in amyl alcohol, with difficulty in oils or petroleum ether ; not in absolute ether.

Removed from alkaline solutions by benzene and chloroform.

Precipitants:

Alkaline hydrates, 1 in 100 to 500.
„ carbonates.
„ bicarbonates, after expulsion of carbonic acid.
Lime water.
Magnesia.
Morphine, strychnine (the free bases).

Ammonia, not at once, at 1 in 100 ; oily pp. becoming crystalline.
[Not lead acetate, neutral or basic.]
Tannic acid, 1 in 2,000 ; limit 1 in 10,000.
Picric acid, pp. gradually crystalline ; limit 1 in 10,000.
Platinum chloride, 1 in 1,000, yellow crys. (insoluble in acetic acid).
Gold chloride, 1 in 25,000, dirty yellow, amorphous ; soluble in hydrochloric acid.
Potassium ferrocyanide, yellow crystalline (not 1 in 500, Dupuy), becoming blue on exposure.
Potassium sulphocyanide (gradually at 1 in 100).
Silver potassium cyanide, pp. becoming crystalline.
Potassium chromate, yellow crystalline, limit about 1 in 500.
Potassium bichromate ; at first no pp., then crys. up to 1 in 3,000.
Chromic acid, 5 per cent.
Phospho-molybdic acid, 1 in 5,000, orange flocks.
„ antimonic acid.
Iodo-potassic iodide, 1 in 50,000, kermes pp.
Potassium iodide alone, crystalline pp.
Bismuth-potassium iodide, 1 in 10,000, orange-red pp.
Cadmium-potassic iodide, crystalline ; precipitation complete.
Mercuric-potassic iodide, whitish yellow, amorphous ; feeble at 1 in 50,000.
Mercuric chloride, amorphous unless concentrated.
Chlorine water, yellow coloration becoming red, then colourless, finally white flocculent pp.
Bromine water, violet coloration.
Iodine tincture, brown crystals.

Colour tests :

Concentrated sulphuric acid, pure, colourless.
„ „ „ with sugar, no effect.
„ „ „ with potass. bichromate, orange ; red with dilute acid.
Concentrated sulphuric acid, with trace of nitric acid, intense red.
Nitre added after 15 hours' solution in sulphuric acid, rose~orange~yellow.
Nitric acid alone, scarlet to blood-red, becoming yellowish-red, then yellow ; on then adding stannous chloride or ammonium sulphide, reddish-violet.
Concentrated hydrochloric acid, no effect.
Fröhde's solution, red~yellow ; after 24 hours, colourless.

(*c*) *IGASURINE* A. Schützenberger and W. A. Shenstone have shown this to be a mixture. Desnoix' description (*Journ. Pharm.* [3], 25, 202) is here given as an illustration of the results one may obtain when dealing with mixed alkaloids from Nux Vomica.

Silky crystals with 10 per cent. of water, lævo-rotatory, alkaline, bitter.

Soluble in 200 parts of boiling water, crystals being rapidly formed on cooling ; also in alcohol, amyl alcohol, and essential oils ; with difficulty in ether.

Reactions :

Alkaline hydrates, Ammonia } pp. soluble in excess.
Alk. bicarbonates precipitate in presence of tartaric acid, 'distinction from Brucine.'
Tannic acid, white pp.
Platinum chloride, yellow pp.
Potassic iodide, pp., not immediately.
Iodo-potassic iodide, brown pp.
Concentrated sulphuric acid with trace of nitric acid, „ nitric acid alone } intense red.

(*d*) *CURARINE* A., $C_{10}H_{15}N$, Preyer ($C_{18}H_{35}N$, Sachs) ; crystals from chloroform, becoming oily on exposure ; feebly alkaline reaction, bitter, extremely poisonous. **Salts** non-crystallizable and very soluble.

Soluble in all proportions in water or alcohol, with difficulty in amyl alcohol or chloroform. **Insoluble** in ether, benzene, petroleum ether, carbon bisulphide, or turpentine.

Not removed from alkaline aqueous solutions by immiscible solvents.

Precipitants :

Tannic acid, pp. soluble in hydrochloric acid.
Platinum chloride, yellow, becoming crystalline.
Potassium ferrocyanide.
„ sulphocyanide.
„ bichromate, amorphous.
Phospho-molybdic acid.
Bismuth-potassic iodide.
Cadmium-potassic iodide, completely.
Mercuric-potassic iodide.
Mercuric chloride.
Potassium platinous cyanide.
„ per-iodate.

Colour tests:

Concentrated sulphuric acid, red; or blue first, then red after some hours, or on heating to 90°-100° (compare Narcotine).
Concentrated sulphuric acid with potassium bichromate, violet.
" " " with sugar, red.
" " " with nitric acid, red.
" nitric acid, purple.

(*e*) *AKAZGINE* A. (Icajanine); amorphous; fusible, becoming yellow on heating; alkaline reaction, bitter, poisonous. Tartrate crystalline.

Soluble in 13,000 parts cold water, 60 of absolute alcohol or 16 of 85 per cent., 120 absolute ether (more easily in commercial ether); further, in chloroform, benzene and carbon bisulphide.

Precipitants (all following precipitates are amorphous):

Alkaline hydrates.
" carbonates.
Ammonia.
Sodium phosphate.
Tannic acid.
Picric acid.
Platinum chloride.
Gold chloride.
Potassium ferrocyanide.
Potassium sulphocyanide.
" bichromate.
Iodo-potassic iodide.
Mercuric-potassic iodide.
Mercuric chloride, soluble on warming.
Iodine tincture.
Stannous chloride.

Concentrated sulphuric acid, as Strychnine.

(*f*) *GELSEMINE* A., $C_{32}H_{38}N_2O_4$ or $C_{24}H_{28}N_2O_4$ (Gerrad); crystallizable with difficulty from alcohol; M.P. 45°; bitter; dilates pupil; very poisonous. Hydrochloride not readily soluble.

Soluble in ether, chloroform, carbon bisulphide, benzene; with difficulty in alcohol or water.

Removed from *alkaline* solutions by benzene, etc.

Precipitants:

Alkaline hydrates, soluble in excess.
[Not lead acetate, neutral or basic.]
Picric acid, yellow crystalline.
Platinum chloride, yellow; soluble on boiling.
Gold chloride, soluble on boiling.
[Potassium ferricyanide, reduced.]

Colour tests:

Concentrated sulphuric acid, colourless, or yellow changing to reddish-brown.
Concentrated sulphuric acid with potassium bichromate, violet-red.
" " " with sugar, red.
" nitric acid, greenish-yellow.

(*g*) *GELSEMININE* A.; dark brown, resinous, alkaline.

Soluble in alcohol, ether and chloroform; sparingly in water and petroleum ether. Hydrochloride easily soluble.

Precipitate by ammonia, reddish-coloured, flocculent.

§ 213. **SYRINGA** vulgaris (Lilac); *Oleaceæ*. Substance (*a*), in the bark during spring (March-April), after which period it is replaced by (*b*). Walz, *Arch. Pharm.*, vols. 105, 108, 109, 113: also G. Koerner and others. Ligustrum vulgare, (*b*), (*c*).

(*a*) *SYRINGIN* G. (Hydroximethylconiferin. Lilacin probably. $C_{17}H_{24}O_9$ or $C_6H_2(OC_6H_{11}O_5)^I(OCH_3)_2^{II,\ VI}(C_3H_4OH)^{IV}$ (Koerner): crystalline hydrate in needles up to half-inch in length; loses water at 100°-115°, melts 185°-190° (191°-192°, Koerner); neutral reaction, tasteless. Acids convert to sugar (fermentable) and Syringogenin, $C_{11}H_{14}O_4$, or $OH{\cdot}C_6H_2(O{\cdot}CH_3)_2{\cdot}C_3H_4{\cdot}OH$ = Hydroximethylconiferyl alcohol, appearing as gray flocks.

Soluble in hot water, hot alcohol, and in chloroform; with difficulty in cold water; not in ether.

[**Not** precipitated by lead acetate or other metallic salts.]

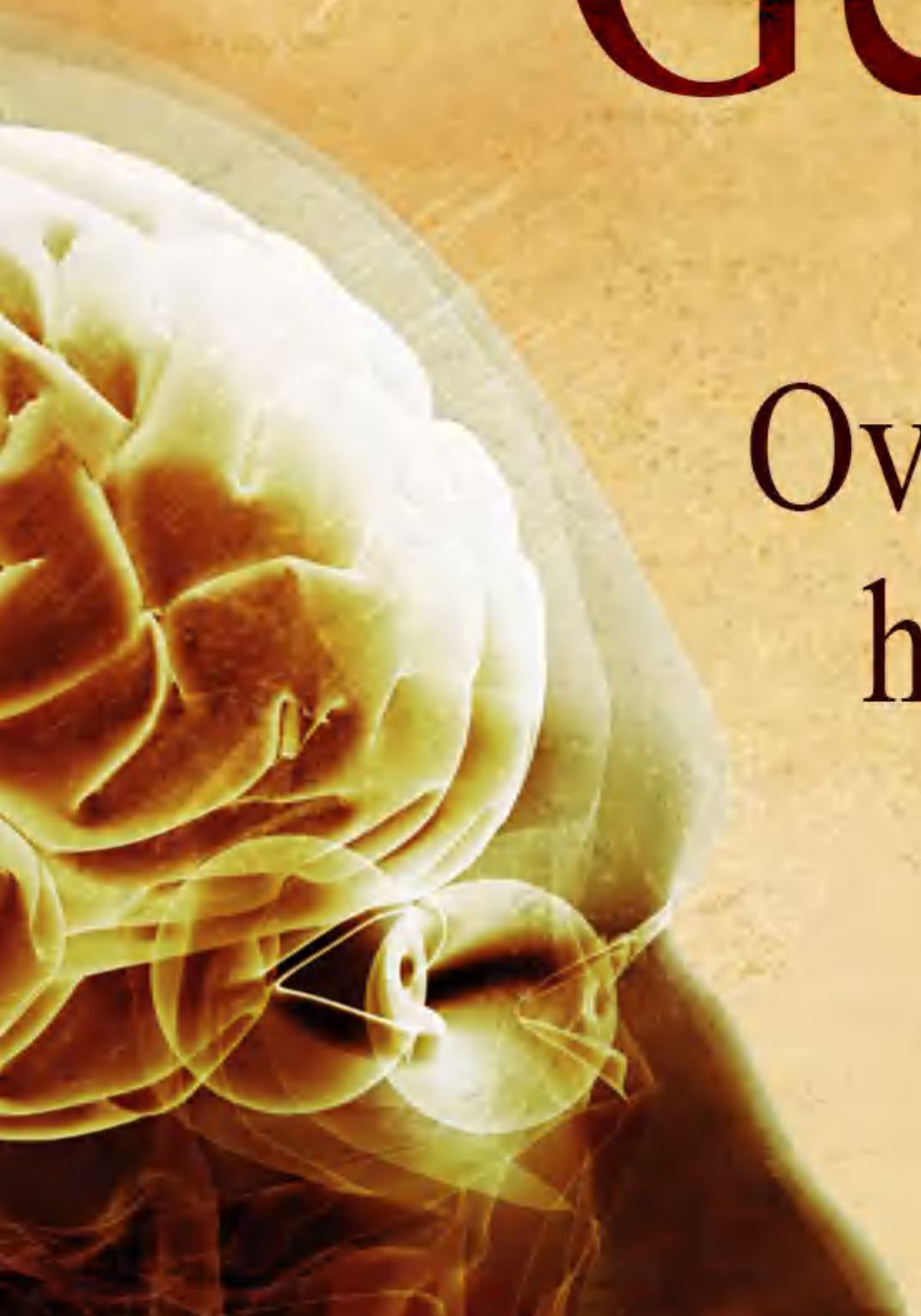
Colour tests:

Concentrated sulphuric acid added to the aqueous solution, dark blue, turning violet with more acid.
Concentrated nitric acid, blood-red.
" hydrochloric acid, colourless; on boiling, blue flocks separate, the fluid becoming violet-red.
Fröhde's solution, blood-red turning to violet-red.

(*b*) *SYRINGOPICRIN* B., $C_{26}H_{24}O_{17}$; amorphous, yellow; M.P. below 100°; slightly acid reaction, bitter.

Soluble in water and alcohol; not in ether.
Absorbed by charcoal.

[**Not** precipitated by lead acetate, neutral or basic.]
Tannic acid, pp.
[No reaction ferric chloride or Fehling's solution.]
Concentrated sulphuric acid, violet (greenish to brown, Kromayer).



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cardiac stimulant. Combines only with strong acids; salts only in acid solution—the compounds with acids not being very stable.

Solubility (authorities differ to some extent) of anhydrous Caffeine: 1 in 98 cold water, 1 in 10 hot, 1 in 97 cold alcohol, 1 in 194 cold ether (less readily in absolute ether); also in amyl alcohol, chloroform and benzene; scarcely in carbon bisulphide or petroleum ether.

Removed by benzene and chloroform from alkaline and acid solutions.

Precipitants (precipitation imperfect):

[Not alkaline hydrates.
„ „ carbonates.
„ ammonia.]
[Tannic acid cloud; insoluble in hydrochloric acid cold, but dissolved on warming.]
[Picric acid, not dilute.]
[Not ferric chloride.]
[Platinum chloride only in very strong solutions, gradual crystalline pp. soluble in about 20 parts cold water.]
[Gold chloride, after some hours, yellow crystalline pp.]
[Fehling's solution, not reduced if pure.]
[Not silver potassium cyanide.
„ potassium bichromate, 1 in 3,000.
„ chromic acid 5 per cent. solution.]
Phospho-molybdic acid, yellow.
[Not phospho-antimonic acid.]
[Iodo-potassic iodide, dirty brown, but not dilute.]
Bismuth-potassic iodide, 1 in 3,000.
[Not cadmium-potassic iodide.
„ zinc-potassic iodide.
„ mercuric-potassic iodide.]
Mercuric chloride, crystalline needles; 1 in 1,000 after a time.

Colour tests (mostly negative):

[Concentrated sulphuric acid, colourless (twenty-four hours).
„ „ „ with sugar, no effect.
„ „ „ with potass. bichromate, no effect.
„ „ „ with nitric acid, colourless.
Nitre added after ten to fifteen hours' solution in sulphuric acid, colourless.]
Concentrated nitric acid, colourless; but on evaporation over water-bath and cautious addition of ammonia, purple (murexide test). This reaction is more sensitive when Chlorine or Bromine water is substituted for the nitric acid, avoiding excess.
[Concentrated hydrochloric acid, no effect.
Fröhde's solution, colourless.]

(*b*) *CAFFEIDINE*, from Caffeine by boiling with Baryta water; $C_7H_{12}N_4O$ (Caffeine + $H_2O - CO_2$); crystals are obtainable; M.P. 94° (Wernecke). The free base readily decomposes into Ammonia, Methylamine, and Cholestrophane. Salts crystalline.

(*c*) *THEOPHYLLINE*,

$$C_7H_8N_4O_2 + H_2O, \text{ or } CO \left\{ \begin{matrix} NCH_3 \cdot CH : C \cdot NH \\ NCH \text{——} C : N - \end{matrix} \right\} CO + H_2O.$$

Crystalline; becomes anhydrous at 110°. Closely resembles Theine, to which it may be converted by methylation.

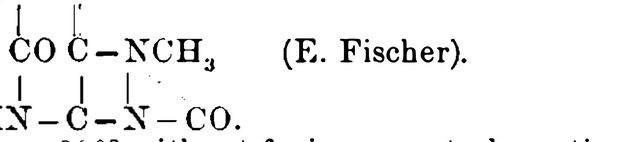
Soluble in water and alcohol.

Ammonia dissolves readily.

Chlorine water gives on evaporation, a bright red, changing to violet with ammonia.

(*d*) *THEOBROMINE*,

$C_7H_8N_4O_2$, or



(E. Fischer).

Microscopic needles, subliming at 290° without fusion; neutral reaction, bitter; similar physiological action to that of Theine; feebly basic; salts separate readily into acid and free alkaloid.

Solubility (authorities differ somewhat): 1 in 1,600 parts water at 0° C., 660 at 20°, 55 boiling; 1,400 cold alcohol, or about 50 boiling: very difficultly soluble in cold ether (1 in 17,000 ?) or in about 600 boiling; with less difficulty in amyl alcohol and chloroform (about 100 parts of the latter boiling; scarcely in benzene; **insoluble** in petroleum ether.

Precipitants:

Alkaline hydrates, Ammonia } pp. soluble in excess, forming salts.
[Not neutral lead acetate, but carried down with other matters from an aqueous extract of cocoa by *basic* lead acetate.]
[Tannic acid, only cloud in dilute solutions.]
[Not picric acid.]
[Not ferric chloride.]
Platinum chloride, brown flocks, slowly 1 in 3,000.
Gold chloride, slowly needles.

Silver nitrate, crystalline pp., gradually ($C_7H_8N_4O_2 \cdot AgNO_3$), *characteristic*.
Potassium bichromate, cloud, then gradual pp.; amorphous at 1 in 3,000.
Phospho-molybdic acid, yellow.
Phospho-tungstic acid in large excess, slimy pp.
Phospho-antimonic acid, cloud 1 in 1,000.
[Iodo-potassic iodide, cloud, pp. if concentrated.]
[Bismuth-potassic iodide, cloud.]
[Not cadmium-potassic iodide.]
[Not mercuric-potassic iodide.]
Mercuric chloride, pp. in strong solutions, white crystals; cloud at 1 in 3,000.

Colour tests (mostly negative):
Concentrated sulphuric acid, colourless (24 hours).
" " " with nitric acid, unchanged.
Concentrated nitric acid, colourless.
Fröhde's solution, colourless.
Evaporated with chlorine water, etc. (murexide test), as Theine.

(*e*) *CAMELLIN* G., $C_{53}H_{84}O_{19}$. From the seeds of Camellia Japonica.
Soluble in alcohol.
Fehling's solution, reduced.
Concentrated sulphuric acid } red.
Concentrated nitric acid }

(*f*) *ASSAMIN* G. From the seeds of Thea Chinensis, var. Assamica. Produces tetanus.

(*g*) *ASSAMIC ACID* G., $C_{18}H_{29}O_{10}$. Yields Sapogenin (see Saponaria), 45·55 per cent., and glucose 42·56 per cent.
Soluble in water, gelatinous pp. by alcohol.
Precipitated by neutral and basic lead acetate.
Concentrated sulphuric acid, yellow~red~blue~violet.
Fuming nitric acid, yellow.

§ 219. **TRIANOSPERMA** filicifolia; *Cucurbitaceæ*. Brazil. The root. Investigator: Peckolt, *Arch. Pharm.* [2], **63**, 104.

(*a*) *TRIANOSPERMINE* A. Crystallizes and sublimes in needles; alkaline; sharp taste, no odour.
Soluble in water and alcohol, not in ether.
Precipitated by platinic chloride.

(*b*) *TRIANOSPERMITINE* A. Crystalline, sublimable, tasteless, odourless.
Soluble in ether, with difficulty in alcohol, not in water.

§ 220. **TRIGONELLA** fœnum Græcum; *Leguminosæ*. The seed.

TRIGONELLINE A. (The methylbetaine of nicotinic acid: isomer of Von Gerichten's pyridine-betaine). Crystallizes in flat prisms from alcohol, loses water at 100° and decomposes on further heating. Neutral reaction, deliquescent.
Very **soluble** in water. **Insoluble** in ether, chloroform or benzene.
Reactions:
[No pp. picric acid.]
Red coloration with ferric chloride.
[No pp. platinum chloride.]
[No pp. mercuric chloride.]

§ 221. **TYLOPHORA** asthmatica; *Asclepiadaceæ*. Ipecacuanha substitute. East Indies. The root. Investigated by D. Hooper.

TYLOPHORINE A. Crystalline, alkaline, emetic. Forms neutral salts.
Soluble in ether (by which it may be removed from alkaline solutions), sparingly in water.
'Precipitated by usual alkaloid reagents.'
[No action ferric chloride.]
Colour tests:
Concentrated sulphuric acid, reddish-brown, changing to green, then blue.
" " " with potass. bichromate, violet-brown.
" " " with potass. permanganate, decoloration.
Concentrated nitric acid, purplish-red; portion soluble, with orange colour.
Concentrated hydrochloric acid, yellow solution.
Fröhde's reagent, bright green.

§ 222. **URECHITIS** suberecta; *Apocynaceæ* (called 'Nightshade' in Jamaica). The leaves. Investigator: Bowrey, *Chem. News*, **37**, 166.

(*a*) *URECHITIN* G. $C_{28}H_{42}O_8 + Aq$. Cystallizes in needles or four-sided prisms; very bitter (perceptible 1 in 40,000); very poisonous. Gives sugar on treatment with acids.
Soluble in hot alcohol, easily in chloroform, also in amyl alcohol, ether, benzene, glacial acetic acid; not in water or dilute alcohol.
Concentrated sulphuric acid, yellow, changing successively to orange, red, mauve, purple; oxidizing agents hasten colour changes.

(*b*) *URECHITOXIN* G., $C_{15}H_{20}O_5$; possibly a derivative of (*a*). Less readily crystallizable than latter. Acids give sugar and Urechitoxetin.

Soluble in chloroform, amyl alcohol, ether, benzene (less readily in last two than is the case with (*a*), also in petroleum ether; with difficulty in water, and not in dilute alcohol.

Removed from alkaline solutions by petroleum ether, contrary to the behaviour of most similar substances.

Precipitated by basic lead acetate.

Concentrated sulphuric acid, as (*a*).

§ 223. **USTILAGO** Maydis ('Smut' of Indian corn). *Fungi.* Investigators: Rademacher and Fischer.

USTILAGINE A. Bitter: physiological action resembles that of Ergot.

Soluble in water, alcohol, and ether.

Reactions:

Ferric chloride, dark yellow coloration.

Precipated by mercuric-potassic iodide.

Concentrated sulphuric acid, brown changing to intense green.

§ 224. **VALLARIA** (Hydrocotyle Asiatica); *Umbelliferæ.* Investigator: Lepine, *Journ. Pharm.* [3], 26, 47.

VELLARIN B.; yellow, oily, neutral, bitter; strong odour; thickens and darkens on exposure.

Soluble in alcohol, ether, volatile and fatty oils; emulsion with water.

Reactions:

Alkaline hydrates, insoluble.

Ammonia, soluble (acids precipitate).

§ 225. **VARIOLARIA** dealbata; *Lichenes.* Substance (*a*). Investigator: Robiquet, *Ann. Chem. Phys.*, **42**, 236. V. amara, substance (*b*), Alms, *Ann. Chem. Pharm.*, **1**, 61; Vogel, *N. Jahrb. Pharm.*, **8**, 201.

(*a*) *VARIOLARIN* B.; crystallizes in needles, subliming with some decomposition.

Soluble in alcohol and ether, not in water.

No coloration with alkalies [see (*b*)].

„ „ „ concentrated sulphuric or other acids.

(*b*) *PICROLICHENIN* B., $C_{12}H_{20}O_6$ (Vogel and Wuth). Transparent rhombohedric crystals; M.P. 111°; not volatile without decomposition; acid to litmus; bitter, odourless. Sp. gr. 1·176.

Soluble in alcohol, ether, carbon bisulphide, volatile and fatty oils, hot glacial acetic acid; difficultly in hot water and not in cold.

Reactions:

Alkaline hydrates dissolve with gradual red coloration; acids precipitate from the solution a *bitterless* substance.

Ammonia dissolves, solution turning gradually red.

Chlorine water, yellow coloration.

No colour with concentrated sulphuric acid.

§ 226. **VERATRUM** group; *Melanthaceæ (Colchicaceæ).* V. album (White Hellebore), alkaloid (*b*) principally; also (*g*), (*j*), (*l*), (*m*), small quantity (*k*), trace (*h*), (*i*). V. viride (Green Hellebore), (*b*), (*j*), (*a*), (*g*); Sabadilla seeds from Asagrea officinalis, Lindl. (Sabadilla officinalis, Brandt; Veratrum officinale, Schecht), (*a*), (*h*), (*e*), (*d*). [For Black Hellebore see Helleborus, *Ranunculaceæ.*] Investigators numerous; mention should be made of Wright and Luff who have recently done so much towards removing the obscurity surrounding these alkaloids.

(*a*) *VERATRINE* A. (Merck's Veratrine; Wright and Luff's Cevadine; *not* Couerbe's Veratrine), $C_{32}H_{49}NO_{11}$ (W. and L.). Needles or compact crystals from alcohol (varnish from ether); M.P. 205° (W. and L.); alkaline reaction, burning taste; no odour, but produces violent sneezing; dilates pupil; poisonous; solutions fluorescent; without action on polarized light. On saponification, methyl crotonic acid, $C_3H_4{\cdot}CH_3{\cdot}COOH$, and a base Cevine, $C_{27}H_{43}NO_9$, are obtained. Salts mostly amorphous (picrate, gold, and platinum double salts have been obtained crystallized—E. Merck).

Soluble in alcohol, ether, chloroform, amyl alcohol, benzene; with difficulty in petroleum ether, not in cold water, 1 in 1,000 boiling (Pelletier and Caventou).

Removed from *acid* solution by chloroform, from alkaline solution also by benzene, and in traces by petroleum ether.

Precipitants:

Alkaline hydrates } precipitation not complete in the cold, Dupuy.
„ carbonates }
„ bicarbonates, if no free carbonic acid.

Ammonia, somewhat soluble in excess.

Tannic acid, gradually (cloud at 1 in 5,000); pp. difficultly soluble in dilute hydrochloric acid.

Picric acid, amorphous pp. 1 in 1,000 (a crystalline picrate is obtainable however).

Ferric chloride, pp. in hydrochloric acid solution.

[Platinum chloride, not in very dilute solution.]

Gold chloride (the double salt melts at 182°).

appearance ; M.P. 238°-240° with decomposition. Green fluorescence in sulphuric acid solution ; sternutatory.

Soluble in alcohol and ether readily, sparingly in petroleum ether.

Reactions :

Alkaline hydrates
„ carbonates } dissolve, but pp. on warming.
Ammonia

Colour tests :

Concentrated sulphuric acid, yellow, changing to blood-red, then violet.
Concentrated nitric acid, no colour.

(*d2*) *SABADININE* A., $C_{27}H_{45}NO_8$ (Merck) ; hair-like crystals from ether ; no definite M.P. ; non-sternutatory.

Soluble in water, alcohol, chloroform, and difficultly in ether.

Reactions :

Alkaline hydrates
„ carbonates } as above.
Ammonia

Concentrated sulphuric acid, persistent blood-red.

(*e*) *SABATRINE* A. ; soluble in 40 parts cold water (Dragendorff), and in ether or chloroform.

Removed from alkaline solutions by chloroform, and in traces by petroleum ether.

Reactions : } as Sabadilline.
Colour tests : }

(*f*) *JERVIC ACID*, from Jervine.

(*g*) *VERATROIDINE* A., $C_{51}H_{78}N_2O_{16}$ (Bullock) ; or $C_{22}H_{53}NO_9$ (Pehkschen). Crystalline ; M.P. about 149° ; salts amorphous ; optically inactive.

Solubility at 22°C ; alcohol dissolves in almost all proportions ; 1 in 5·9 chloroform, 9 absolute ether, 13 benzene; not soluble in petroleum ether.

Precipitated by mercuric-potassic iodide (cloud at 1 in 5,000), and by most alkaloid precipitants.

Colour tests :

Concentrated sulphuric acid, yellow, becoming successively orange and red, with green fluorescence.
Concentrated sulphuric acid and sugar, brown? (violet, Muter).
„ „ „ with nitric acid, gradually cherry-red.
„ nitric acid alone, transient rose colour, then yellow.
Hydrochloric acid 11 per cent., rose colour.
Fröhde's solution, gradually cherry-red.

(*h*) *COUERBE'S VERATRINE* A., $C_{37}H_{53}NO_{11}$; amorphous (varnish) ; M.P. 180° ; not fluorescent. Salts crystalline. Nitrate insoluble. On saponification, a base VERINE, $C_{28}H_{45}NO_8$, and Veratric acid (dimethylprotocatechuic acid) are obtained.

Concentrated sulphuric acid, reaction similar to that with (*a*).

(*i*) *VERATRALBINE* A., $C_{28}H_{43}NO_5$; amorphous ; non-sternutatory. Fluorescent green solutions.

Precipitated by sodium hydrate and phospho-molybdic acid (cloud at 1 in 3,500).

Colours with sulphuric acid as with (*a*).

(*j*) *PSEUDOJERVINE* A., $C_{29}H_{49}NO_{12}$, Pehkschen (Veratrum album gave ·006 per cent. only) ; large rhombic crystals, blackened at 259° C. (300°-307°, Salzberger) ; optically inactive.

Soluble in 184 parts absolute alcohol, 4 chloroform, 372 benzene, 10 petroleum ether, 1,021 absolute ether.

Precipitated by phospho-molybdic acid ; cloud at 1 in 10,000.
Mercuric-potassic iodide ; cloud at 1 in 6,000.

No colours with acids when pure, but if admixed with trace of Jervine, colorations due to latter are obtained.

(*k*) *RUBIJERVINE* A., $C_{26}H_{43}NO_2+H_2O$ (Salzberger) ; M.P. 240°-246°.
Concentrated sulphuric acid, red.

(*l*) *PROTOVERATRINE* A., $C_{32}H_{51}NO_{11}$ (Salzberger) : crystallizes in microscopic four-sided plates ; M.P. 245°-250° ; sternutatory ; very poisonous.

Soluble with difficulty in boiling alcohol or in chloroform, slightly in hot ether ; **insoluble** in cold ether, benzene, petroleum ether.

Precipitants :

Alkaline hydrates.	Phospho-molybdic acid.
Ammonia.	„ tungstic acid.
[Not tannic acid.]*	Cadmium-potassic iodide.*
Picric acid.	Mercuric-potassic iodide.
[Not platinum chloride.]	[Not mercuric chloride.]
Gold chloride.	

Colour tests :

Concentrated sulphuric acid, dissolves slowly with gradual green, then blue, coloration ; violet after some hours.
Concentrated sulphuric acid with sugar, greenish turning to brown.

(*m*) *PROTOVERATRIDINE* A., $C_{26}H_{45}NO_8$ (Salzberger) ; crystallizes in

* See Protoveratridine.

four-sided microscopic plates; M.P. 265°; non-sternutatory; non-poisonous; bitter.

Soluble scarcely in alcohol, chloroform, methyl alcohol, or acetone, and not in ether, benzene, or petroleum ether.

Precipitants:

Ammonia.	Phospho-tungstic acid.
Tannic acid.	[Not cadmium-potassic iodide.]
Picric acid.	Mercuric-potassic iodide.
[Not platinum chloride.]	[Not Millon's reagent.]

Colour tests:

Concentrated sulphuric acid, violet, becoming cherry-red.

Hydrochloric acid, light red on warming, an odour of isobutyric acid being perceptible.

§ 227. **VERNONIA** nigritiana; *Compositæ—Tubulifloræ* ('Batiatior' or 'Batjentjor'); W. Africa. Investigators: Heckel and Schlagdenhauffen.

VERNONIN G.; amorphous white powder giving a resinous substance, $C_4H_{10}O_3$, and sugar on hydrolysis. Cardiac poison $\frac{1}{50}$th power of Digitalin.

Soluble in alcohol, slightly in ether and chloroform.

Concentrated sulphuric acid, brown, becoming purple (lasting for several hours).

§ 228. **VIBURNUM** prunifolium.

VIBURNIN B. (possibly G.); brown, resinous, bitter.

Soluble in alcohol.

§ 229. **VICIA** sativa (Tares); *Leguminosæ—Papilionaceæ*. The seeds. Investigators: Ritthausen, *Berichte d. ch. Ges.*, 1876; E. Schultze, and others. In addition to (*a*), (*b*), and (*c*) the following have been found by E. Schultze: ASPARAGINE; VERNINE, $C_{16}H_{20}N_8O_8+3H_2O$, which gives Guanine with boiling hydrochloric acid; BETAINE (0·006 per cent.); CHOLINE (0·0015 per cent.); and AMYGDALIN (see Prunus).

(*a*) *VICINE* A., $C_{28}H_{51}N_{11}O_{21}$? (Ritthausen). Crystallizes in needles, loses water on heating, then melts at 180°. Forms crystalline sulphate and hydrochloride.

Soluble in 108 parts water at 22° C.; **insoluble** in absolute alcohol, but dissolves in hot commercial alcohol.

Reactions:

Alkaline hydrates, Barium hydrate, Calcium hydrate } dissolve.

Ammonia hydrate dissolves slowly; Vicine becomes bluish-gray in an ammoniacal atmosphere, or purple in dry ammonia gas.

Concentrated nitric acid causes it to swell up; soluble after evaporation, the residue from which becomes violet around the edges.

(*b*) *CONVICINE* A., $C_{10}H_{14}N_3O_7+H_2O$.

Soluble with difficulty in water or alcohol.

Precipitated completely by mercuric nitrate.

(*c*) *DIVICINE* A., $C_{31}H_{50}N_{30}O_{16}$? from (*a*) by boiling with 20 per cent. sulphuric acid.

§ 230. **VISCUM** album (Phoradendron; Mistletoe), *Loranthaceæ*; Atractylis gummifera, *Compositæ*; Gentiana lutea, *Gentianaceæ*; Lychnis viscosa, *Caryophyllaceæ*; Robinia viscosa, *Leguminosæ*; Saxifraga tridactylites, *Saxifragaceæ*.

VISCIN; $C_{20}H_{48}O_8$ (Rheinsch, *N. Jahrb. Pharm.*, **14**, 129); thick viscid substance; on heating to 100° becomes more fluid; acid reaction, nearly tasteless; sp. gr. 1·0.

Soluble in ether.

§ 231. **VITEX** agnus castus (*Verbenaceæ*).

CASTINE A.; crystalline.

Soluble in alcohol and ether, not in water. Hydrochloride crystalline.

§ 232. **XANTHIUM** strumaria; *Compositæ*. The seeds (A. Zauder, *Pharm. Z. Russl.*, 1881).

XANTHOSTRUMARIN G.; poisonous; succinic odour on warming.

Soluble in alcohol and ether.

Precipitants:

[Alkaline hydrates, dissolve dark yellow.]
Neutral lead acetate, yellow pp.
[Not tannic acid.]
Ferric chloride, dark green pp.
Silver nitrate, slight reduction on warming; at once with ammonia.
Cupric acetate, greenish-blue pp.
Fehling's solution, reduced after hydrolysis of the glucoside.
Stannous chloride, pp.
[Not tartar emetic.]

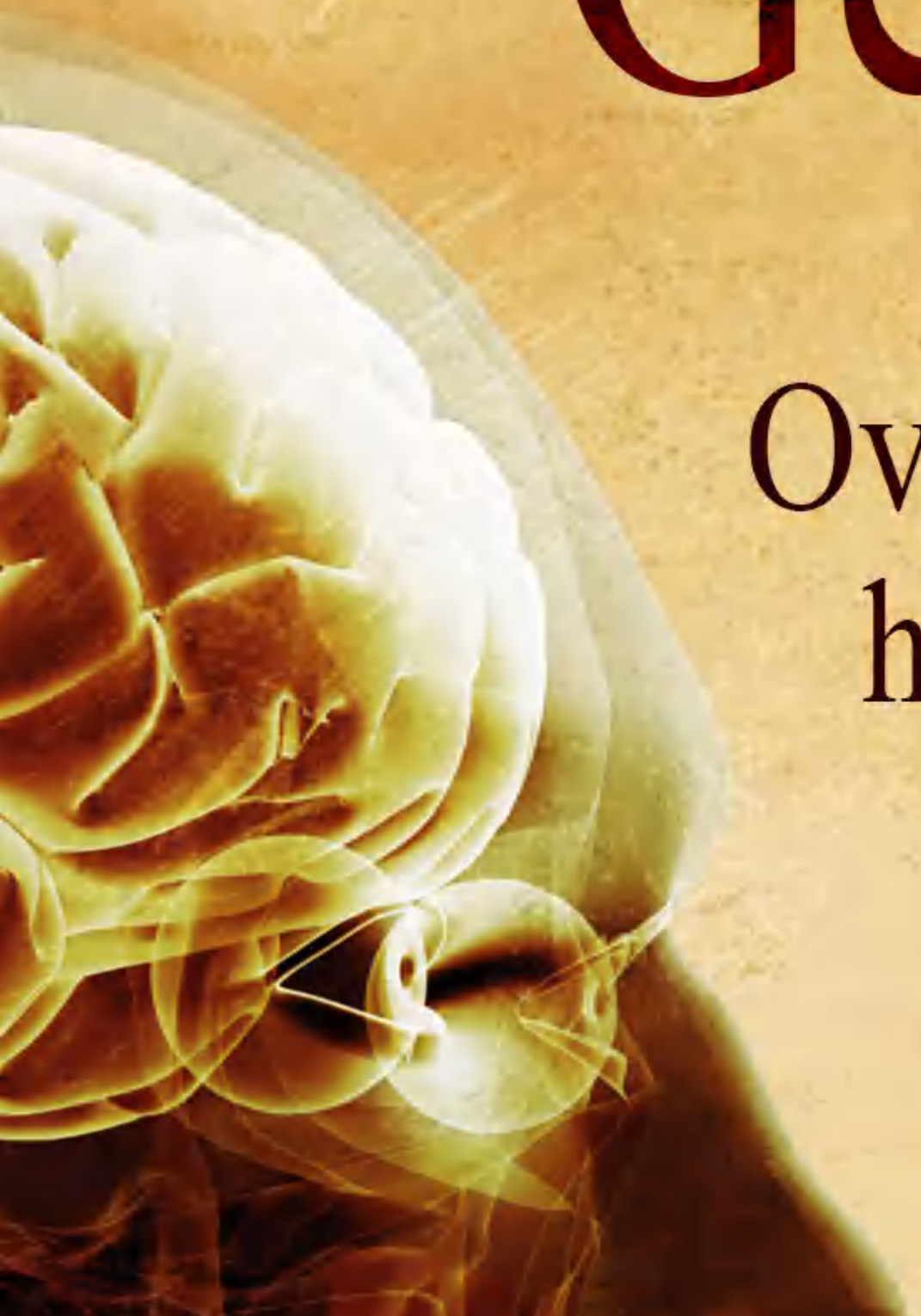
§ 233. **XANTHOXYLON** Senegalense, *D.C.* (Artar root; variety of Prickly Ash); *Xanthoxylaceæ*. Substance (*a*) 0·4 per cent. of the root,

also an alkaloid crystallizing in blood-red needles; soluble in water and forming yellow salts; further, a crystalline substance, $C_{10}H_{10}O_3$, resembling Cubebin. Some varieties of Xanthoxylon contain Berberine (see Berberis).

(*a*) *ARTARINE* A., $C_{21}H_{23}NO_4$ (P. Giacosa); rose-gray amorphous powder; becomes brown at 210°, then melts at 240° with decomposition; bitter taste. Salts golden yellow. Hydrochloride precipitated by ether from alcoholic solution. Free acids diminish the solubility of the salts.

Soluble in ether, boiling amyl alcohol, and warm acetone; with difficulty in warm chloroform, sparingly in boiling 98 per cent. alcohol; warm methyl alcohol dissolves it if the base be freshly precipitated. **Insoluble** in water or benzene.

Precipitants: [Phospho-antimonic acid, slight pp., soluble in excess.]
Bismuth-potassic iodide, red flocculent pp, insoluble in excess.
Cadmium-potassic iodide } yellow flocculent pp., insoluble in excess.
Mercuric-potassic iodide }
Chlorine water (to hydrochloride), yellow tinge; colourless on addition of ammonia.
Concentrated sulphuric acid, colourless.
Nitre added to the solution in sulphuric acid, blood-red.



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§ 1 ACHILLEA.] The substances are inserted here in the order of the paragraphs in Part I.

	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents.					
							Water.	Alcohol.	Ether.	Chloro-form.	Other Solvents.	
ACHILLEA	Achillein G	$C_{20}H_{38}N_2O_{15}$	red-bn., amorphs.	...	alkaline	bitter	easily	difficultly	insoluble	...	(solutions yellow)	1
	Ivain B	$C_8H_{14}O$?	yellow, amorphs.	...	...	bitter	insoluble	soluble	...	...	*idem*	2
	Moschatine A	$C_{21}H_{27}NO_7$	red-bn., amorphs.	below 100	...	bitter	scarcely	difficultly	...	...	...	3
ACHRAS	Sapotin G	$C_{29}H_{52}O_{20}$	micros. crystals	240	...	burning	easily	sparingly	insoluble	insoluble	hot alcohol	4
	Sapotine A	...	...	...	...	salts b.	insoluble	soluble	soluble	soluble	...	5
ACONITUM	Aconitine A	$C_{33}H_{45}NO_{12}$	rhombic prisms	188·5	alkaline	scarce b.*	4,431 cold	20 boiling	100 blng.	230	(not CS_2; petr. ether)	6
	Aconine A	$C_{26}H_{41}NO_{11}$	amorphous	about 130	alkaline	bitter	soluble	soluble	sparingly	soluble	(not benz.: petr. ether)	7
	Pseudaconitine A	$C_{36}H_{49}NO_{11}$	ndls. or varnish	104—105	...	...	scarcely	soluble	diff.	...	...	8
	Pseudaconine A	$C_{27}H_{41}NO_8$	amorphous	...	alkaline	bitter	soluble	soluble	soluble	...	...	9
	Picraconitine A	$C_{31}H_{45}NO_{10}$	...	...	...	...	insoluble	...	soluble	soluble	...	10
	Lycoctonine A	...	needles or prisms	about 100	alkaline	bitter	soluble	soluble	soluble	soluble	CS_2; petroleum ether	11
	Acolyctine A	...	amorphous	...	alkaline	bitter	readily	readily	insoluble	readily	...	12
	Myoctonine A	$C_{27}H_{30}N_2O_8 \cdot 5H_2O$	amorphous	143—144	...	bitter	difficultly	easily	scarcely	easily	CS_2, benz., amyl alcohol	13
	Aticine A	$C_{22}H_{31}NO_2$	amorphous	85	...	bitter	scarcely	soluble	soluble	...	benzene	14
ACORUS	Acorin G	$C_{30}H_{60}O_6$	resinous	...	neutral	hitter aromatic	insoluble	difficultly	soluble	soluble	benzene	15
	Calamine A	...	...	...	alkaline	...	insoluble	soluble	insoluble	soluble	acetone, dilute acids	16
ADANSONIA	Adansonine A ?	...	needles	...	...	bitter	slightly	soluble	6 cold	...	...	17
ADONIS	Adonin G	$C_{24}H_{40}O_9$	...	...	neutral	very b.	sol. cold, not hot	soluble	insoluble	soluble	...	18
ÆSCULUS	Æsculin G	$C_{15}H_{16}O_9$	needles or prisms	160	acid	slightly b	672 cold, 12½ boiling	90 cold, 24 boiling	insoluble	soluble	...	19
	Æsculetin G-deriv.	$C_9H_6O_4 \cdot H_2O$	silky needles or plates	270	neutral	bitter	difficultly	difficultly	insoluble	insoluble	...	20
	Argyræscin G	$C_{27}H_{42}O_{12}$	silvery crys.	...	...	...	difficultly	soluble	insoluble	...	...	21
	Propæscinic Acid G	...	...	...	...	...	soluble	soluble	...	...	...	22

* Causes tingling.

SUMMARY.

For full details of tests for any one substance, see Part I. ; for particulars of OTHER substances giving similar reactions, see lists Part III. § 7 ÆSCULUS.

	Caustic Alkalies.	Lead Acetate.	Tannic Acid.	Platinum Chloride.	Gold Chloride.	Sulphuric Acid, Concentrated.	Nitric Acid, Concentrated.	Sundry Precipitants, etc.
1	no pp.	no pp.	no pp.	...	...	...	...	no pp. ferrous sulphate
2	...	no pp. neut.	...	...	...	...	...	...
3	...	...	...	...	...	...	...	...
4	insoluble	pp. basic, sol. in excs.	...	...	...	garnet red	...	Fehling's solution not reduced
5	insoluble	..	...	pp. yellow	...	...	...	Mayer's solution, mercuric chloride, etc.
6	pp. and carbonates	no pp.	pp.	not if dilute	yel. amorphs.	gradually violet ; red with sugar	red-brown	Mayer's solution, hydriodic acid, etc.
7	(Ammonia - gelatinous pp.)	pp. sol. in excess	pp.	...	pp. and reduction	no colour	...	Fehling's solution reduced
8	...	...	pp.	not if dilute	pp.	alcoholic potash to residue from HNO_3, purple.	...	Mayer's solution, mercuric chloride, etc.
9	...	...	...	...	...	...	...	silver nitrate (Fehling not reduced)
10	pp. on heating	...	pp.	no pp.	pale yellow, amorphous	no colours	no colours	Mayer's solution
11	...	...	pp.	no pp.	...	no colours	no colours	(not phospho-molyb.). Mayer's soln., etc.
12	(pp. carbonates)	pp.	pp.	...	yellow	...	...	phospho-molybdic acid, white
13	...	...	*	*	*	(Vitali's test : reddish-brown)	..	pp. most alkaloid reagents
14	(pp. ammonia)	...	yellow-bn.	...	...	yellow⁓purple (permanent)	colourless	Mayer's solution, white pp.
15	...	...	pp.	(reduced)	(reduced)	...	...	Mayer's solution. Fehling reduced
16	...	...	pp.	(reduced)	(reduced)	...	...	Mayer's solution. Fehling not reduced
17	soluble yellow	...	...	no pp.	no pp.	(no pp. metallic salts)	...	green tinge, ferric chloride
18	(no decompositn.)	...	...	...	pp.	deep red	deep blue	picric acid, mercuric chloride, etc.
19	soluble	pp basic.	...	...	...	(no pp. metallic salts)	yellow ; red with AmHo	fluorescent 1 in 1,500,000
20	soluble yellow	pp. neutral or basic	...	...	...	...	...	silver nitrate reduced
21	decomp. gradually	...	...	...	...	yellow ; red with water	...	...
22	*idem.*	pp. basic	...	...	...	...	...	acids pp. from aqueous solution

* See last column.

	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents.				
							Water.	Alcohol.	Ether.	Chloro-form.	Other Solvents.
Æsculus	Æscinic Acid G	$C_{24}H_{40}O_{42}$	crystalline	...	...	...	diff. cold	soluble	insoluble	...	...
	Æscigenin G-dv.	...	indistinctly crys.	...	...	...	insoluble	soluble	...	...	...
	Aphrodæscin G	$C_{52}H_{82}O_{23}$	amorphous	...	...	...	soluble	soluble	...	...	...
Æthusa	Cynapine A	...	rhombic crys.	...	alkaline	...	soluble	soluble	insoluble	...	...
Agaricus	Muscarine A	...	crys., deliques.	100	alkaline	tasteless	∝	∝	insoluble	scarcely	...
Aloe	Aloin B.	$C_{16}H_{18}O_7$	yellow needles or granules	soft at 100	neutral	bitter	6,000 cold, 10 boiling	soluble	sparingly	insoluble	amyl alcohol; not benz. or petroleom ether
Alstonia	Alstonine A	$C_{24}H_{20}N_2O_4$	brown, amorphs.	195 auby-drous	alkaline	...	difficultly	soluble	...	soluble	amyl alcohol
	Ditamine A	$C_{16}H_{19}NO_3$	amorphous	75	alkaline	some-what b.	...	soluble	soluble	soluble	benzene
	Echitamine A	$C_{22}H_{28}N_2O_4$	prisms	296 auby-drous	alkaline	bitter	soluble	soluble	soluble	soluble	acetone. Diff. benzene
	Echitenine A	$C_{20}H_{27}NO_4$	brown, amorphs.	120	alkaline	bitter	difficultly	soluble	soluble*	soluble*	not petroleum ether
	Porphyrine A	$C_{21}H_{25}N_3O_3$	amorphs. or crys.	97	alkaline	bitter	...	soluble	soluble	soluble	petroleum ether
	Alstonidine A	...	needles	181	...	bitter	...	...	...	...	acetone
Amaryl-lis	Amarylline A	...	needles	196	...	...	difficultly	soluble	soluble	soluble	...
	Bellamarine A	...	needles	181	....	..	....	soluble	soluble	soluble	...
Ammi	Kellin G	...	silky needles	...	neutral	very b.	sol. hot	sol. hot	readily	sol. hot	..
Anagyris	Anagyrine A	$C_{14}H_{18}N_2O_2$	amorphous	...	alkaline	bitter	soluble	soluble	soluble	...	petroleum ether
Anchieta	Anchietine A	...	yellow needles	...	feebly alk.	pungent	insoluble	soluble	insoluble	...	...
Andira	Andirin B	...	yellow-brown	...	...	...	soluble	soluble	soluble	...	...
Anemone	Anemonin B	$C_{10}H_8O_4$	prisms	152—156 volatile	...	...	sol. hot	...	scarcely	soluble	oils fixed and volatile
Angelica	Angelin	$C_{18}H_{30}O$	pearly plates	...	neutral	tasteless	soluble	diff. cold	soluble	soluble	benzene, turpentine
Angos-tura	Angosturin B	$C_{18}H_{24}O_{10}$	4-sided needles	45	neutral	...	scarcely	soluble	insoluble	...	...
	Galipine A	$C_{20}H_{21}NO_3$	silky needles	115·5	...	...	...	soluble	soluble	soluble	petroleum ether
	Galipinine A	$C_{19}H_{19}NO_3$	silky leaflets	111	...	...	...	...	...	...	petroleum ether
	Cusparine A	$C_{20}H_{19}NO_3$	broad needles	89	...	...	...	...	...	...	(less sol. than above)
	Cusparidine A	$C_{19}H_{17}NO_3$	needles (rosettes)	78	...	...	...	...	...	...	petroleum ether

* If freshly precipitated.

	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents.					
							Water.	Alcohol.	Ether.	Chloro-form.	Other Solvents.	
Anthemis	Anthemine A	...	crys. regular	...	alkaline	tasteless	diff. cold	...	insoluble	soluble	acetic ether	1
Antho-cercis	Anthocercine	...	yellow oily	liquid	alkaline	...	difficultly	soluble	soluble	...	...	2
Antiaris	Antiarin G	$C_{14}H_{20}O_5$	glittering plates	220	neutral	...	254 at 22°	70 at 22°	2,792 at 22°	...	...	3
Apocynum	Apocynin G	...	...	...	...	...	scarcely*	soluble	soluble	...	...	4
Apocyneæ	Ophioxylin	$C_{16}H_{18}O_6$	orange crystals	71·8	...	burning	sparingly	soluble	...	very sol.	petroleum ether, CS_2	5
Aralia	Araliin G	...	yellow powder	...	neutral	acrid	easily	scarcely absolute	insoluble	insoluble	(insoluble benzene)	6
Arariba	Aribine A	$C_{23}H_{20}N_4$	rhomb. octahed.	229	neutral	bitter	7,762 cold	soluble	diff.	...	difficultly amyl alcohol	7
Arbutos	Arbutin G	$C_{25}H_{34}O_{14}+2H_2O$	silky needles	144—146	neutral	bitter	soluble hot	sol. hot	scarcely	...	...	8
group	Ericolin G	$C_{24}H_{56}O_{21}$	amorphous brown	partly 100	...	...	...	...	...	...	benzene	9
	Andromedo-toxin B	$C_{31}H_{51}O_{10}$	needles	229	neutral	bitter	35	about 9	about 1,400	about 385	scarcely benzene	10
	Vaccinin B	(N. free)	prisms	fusible	...	...	soluble	soluble	scarcely	...	...	11
	Urson	$C_{20}H_{17}O_2$	crystalline	198—200	...	tasteless	insoluble	difficultly	diff.	...	...	12
	Asebofuscin G	$C_{18}H_{18}O_8$	reddish-brown	...	...	...	scarcely	soluble	insoluble	insoluble	acetic acid	13
	Asebotin G	$C_{24}H_{28}O_{12}$	needles	145·5	...	bitter	soluble hot	soluble	diff.	diff.	*idem*	14
	Asebogenin	$C_{18}H_{18}O_7$?	needles	...	...	...	difficultly	soluble	soluble	diff.	...	15
Areca	Arecoline A	$C_8H_{13}NO_2$	oil (volatile)	(B.P. 209)	strongly alk.	...	∞	∞	∞	∞	...	16
	Arecaine A	$C_7H_{11}NO_2+2H_2O$	crystalline	213	neutral	...	soluble	scarcely cold	insoluble	scarcely	...	17
	Arecaïdine A	$C_7H_{11}NO_2$	. crystalline	223	...	...	soluble	scarcely absolute	scarcely	scarcely	...	18
	Guvacine A	$C_6H_9NO_2$	crystalline	271—272	nentral	...	soluble	insoluble strong	insoluble	insoluble	...	19
Aristo-lochia	Aristolochin B	$C_{32}H_{22}N_2O_{13}$?	...	...	...	bitter	soluble	soluble	soluble	...	...	20
	Aristolochine A	...	resinous	...	...	...	...	soluble	soluble	soluble	benzene	21
	Aristine A	...	gold laminæ	decomp. 260	neutral	...	...	scarcely	soluble	soluble	benzene	22
Arnica	Arnicin B	$C_{20}H_{30}O_4$?	amorphous	...	...	bitter	scarcely	soluble	soluble	...	...	23
Artemisia	Absynthiin B	$C_{15}H_{20}O_4$	yellow amorphous	65	neutral	bitter	difficultly	soluble	soluble	soluble	benzene	24
	Abrotine A	$C_{21}H_{22}N_2O_2$	needles or powder	...	...	...	diff. hot	...	...	...	...	25

* *Apocynein* is soluble.

	CAUSTIC ALKALIES.	LEAD ACETATE.	TANNIC ACID.	PLATINUM CHLORIDE.	GOLD CHLORIDE.	SULPHURIC ACID, CONCENTRATED.	NITRIC ACID, CONCENTRATED.	SUNDRY PRECIPITANTS, ETC.
1	...	...	...	...	...	...	...	...
2	pp. and carbonates	...	pp.	...	...	yellowish-brown	pale yellow	phospho-molybdic acid, Mayer's solution, etc.
3	soluble	...	no pp.	...	...	*idem*	...	(silver nitrate reduced)
4	...	...	...	...	...	...	...	
5	...	...	...	...	...	...	...	
6	no action cold	pp. basic	pp. hot (not cold)	no pp.	no pp.	decolourized	no action	...
7	pp. and carbs.	no pp.	no pp.	...	...	strong acids pp. from solution	...	chlorine or bromine water
8	...	no pp. neut.	...	...	...	...	...	(ferric chloride, blue colour)
9	...	no pp.	no pp.	...	no pp.	yellowish-brown	...	charac. odour with warm dilute acid
10	soluble	no pp.	...	...	...	magnificent red on slow evaporation	...	no pp. usual reagents
11	...	...	...	...	...	...	...	...
12	insoluble	...	...	...	...	yellow	...	insoluble dilute acids
13	soluble	...	...	...	...	...	...	(violet with alcoholic HCl)
14	soluble	pp. basic	...	...	...	(acids pp. from alkaline solution)	...	...
15	...	*idem*	...	...	...	...	...	...
16	no pp.	...	no pp.*	no pp.	pp. difficultly soluble	...	...	phospho-molybdic acid, Mayer's solution, etc.
17	...	...	cloud	...	...	...	...	Mayer's solution, potassic iodide, etc.
18	...	...	...	...	...	...	...	(ferric chloride, red colour)
19	...	...	...	...	...	...	...	(ferric chloride, deep red colour)
20	pp.	pp. neutral	...	...	...	dark green	...	(Fehling's solution not reduced)
21	pp.	...	...	...	...	...	...	...
22	give red compound	...	...	...	...	blue becoming greenish	...	
23	soluble	pp. basic†	pp.	pp.	...	...	...	silver nitrate, mercuric chloride
24	soluble	no pp.	...	...	pp.	brown~green~blue	oxidizes	(Fröhde, brown~violet-blue)
25	...	...	...	salt diff. sol.	...	(sulphate crystalline)	...	(blue fluorescence)

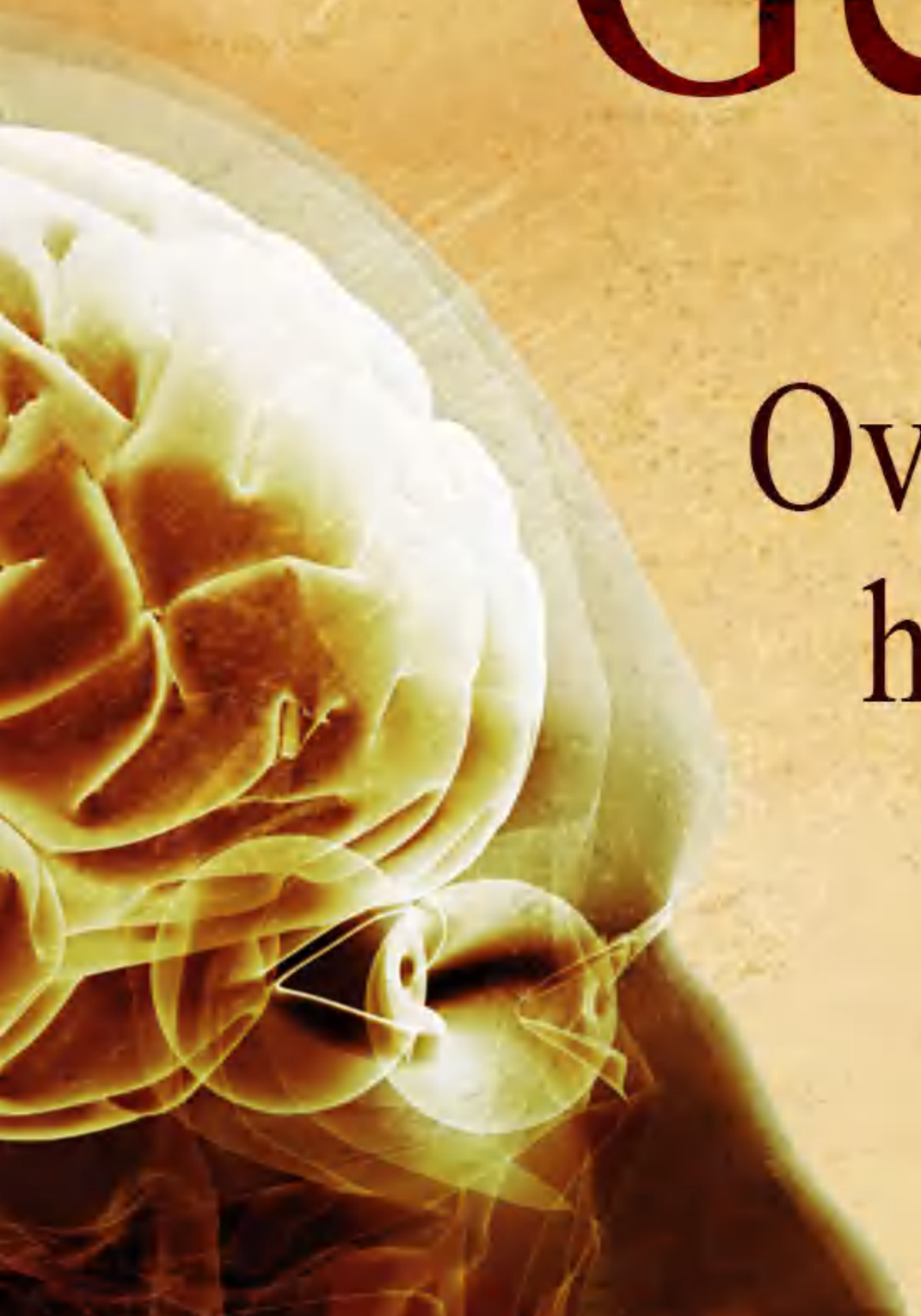
* The free alkaloid gives pp.

† Alcoholic solution.

	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents. Water.	Alcohol.	Ether.	Chloroform.	Other Solvents.	
ARTEMISIA	Santonin B	$C_{15}H_{18}O_3$	crystalline	169—170	feebly acid	bitter in solution	4,000 cold	72 cold	42 hot	4·35 cold	petroleum ether, acetic acid	1
	Hyposantonin	$C_{15}H_{18}O_2$	shining plates	152	...	...	insoluble	sol. warm	soluble	...	benzene, acetic acid	2
ASCLEPIADACEÆ	Asclepiadin G	...	yellow amorphous	...	...	bitter	soluble hot	soluble	soluble	...	...	3
ASIMINIA	Asiminine A	...	amorphous	...	alkaline	tasteless	scarcely	soluble	soluble	diff.	...	4
ATHAMANTHA	Athamanthin B	$C_{24}H_{30}O_7$	...	...	...	b. rancid	insoluble	soluble	soluble	...	...	5
	Oreoselon	$C_{14}H_{10}O_3$	...	...	...	not b.	insoluble	difficultly	diff.	...	...	6
	Ostruthiin	$C_{18}H_{20}O_3$	...	...	...	...	scarcely	soluble	soluble	...	...	7
	Peucedanin B	$C_{15}H_{14}O_{14}$	crystalline	76—82	...	burning	insoluble	sol. hot	soluble	soluble	benzene, CS_2	8
ATHEROSPERMA	Atherospermine A	$C_{30}H_{40}N_2$	powder	128	feebly alk.	bitter	6,000 cold	32 cold	diff.	soluble	CS_2, turpentine	9
ATROPA, ETC.	Atropine A	$C_{17}H_{23}NO_3$	pillars and needles	*	alkaline	bitter metallic	300 cold	30 cold	very soluble	3 cold	40 benzene	10
	Tropine A	$C_8H_{15}NO$	rhombic leaflets	61†	...	...	readily	readily	...	...	...	11
	Homatropine A	$C_{16}H_{21}NO_3$	prisms	95—98	...	...	difficultly	...	readily	readily	...	12
	Hyoscyamine A	$C_{17}H_{23}NO_3$	needles, or plates if pure	108	alkaline	...	soluble hot	soluble	soluble	soluble	benzene, amyl alcohol	13
	Hyoscine A	$C_{17}H_{21}NO_4$*	amorphous	55—59	...	...	difficultly	soluble	soluble	...	benzene	14
	Atropamine A	$C_{17}H_{21}NO_2$	varnish	under 60	slightly alk.	bitter	sparingly	soluble	soluble	soluble	benzene	15
	Belladonnine A	$C_{17}H_{21}NO_2$	amorphous	(volatile)	alkaline	burning	difficultly	soluble	soluble	soluble	benzene	16
	Stramonine A	...	...	150	neutral	not bit.	...	...	...	...	oils	17
	Piturine A	$C_{12}H_{16}N_2$	heavy liquid	B.P. 243	alkaline	sharp	∞	∞	∞	...	...	18
	Scopolin G ?	...	amorphous	...	...	...	difficultly	...	...	...	...	19
	Scopoletin	$C_{10}H_8O_4$	needles	198—199	...	...	difficultly	soluble	soluble	soluble	...	20
	Mandragorine A	$C_{17}H_{23}NO_3$	amorphous	77—79	...	...	soluble	soluble	soluble	...	...	21
	Nicotine A	$C_{10}H_{14}N_2$	liquid	B.P.240—250	alkaline	burning	∞	∞	∞	soluble	benzene, petroleum ether, etc.	22
	Fabianin G	...	...	...	...	...	soluble	soluble	soluble	soluble	...	23
	Fabiin	...	crystalline	non-fusible	neutral	...	insoluble	soluble hot	soluble	soluble	...	24

* Authorities differ.

† Boils 229°.



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	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents. Water.	Alcohol.	Ether.	Chloroform.	Other Solvents.	
AURANTIACEÆ.	Hesperidin G	$C_{22}H_{26}O_{12}$	crystalline	250—251	...	...	5,000 hot	soluble	insoluble	insoluble	acetic acid	1
	Hesperetin	$C_{10}H_{16}O_4$	crystalline plates	224—226	...	sweet	scarcely	soluble	diff.	...	...	2
	Aurantiamaric Acid	$C_{10}H_{12}O_4$	resinous	...	...	very bitter	soluble hot	...	...	...	...	3
	Naringin G	$C_{23}H_{26}O_{12}{\cdot}4H_2O$	yellow crys.	171	...	bitter	300 cold	...	insoluble	insoluble	...	4
	Limonin G	$C_{22}H_{26}O_7$	micros. crys.	224	neutral	very b.	scarcely	soluble	scarcely	...	acetic acid	5
	Murrayin G	$C_{18}H_{22}O_{10}$	needles	170	...	feebly bitter	diff. cold, easily hot	easily	scarcely	...	...	6
	Murrayetin	$C_{12}H_{12}O_5$	...	...	...	tasteless	soluble hot	soluble	diff.	...	...	7
BACHARIS	Bacharine A	...	needles	...	...	...	diff. cold ; sol. hot	difficultly →	diff. →	...	amyl alcl. ; easier sol. alcl. and ether than in water	8
BAPHIA	Baphiin B	$(C_{12}H_{10}O_4)_n$	lustrous needles or plates*	below 100 in part.	neutral	...	insoluble	soluble	soluble	...	difficultly benzene, and carbon bisulphide	9
BAROSMA	Barosmin G	...	microscopic needles	243	...	...	...	scarcely cold	soluble	...	volatile oils; dilute acids	10
BEBEERU	Bebeerine A	$C_{18}H_{21}NO_3$	crys. or amorphs., salts amorphs.	198	alkaline	bitter	6,000 cold ; 1,500 blg.	5 (absolute)	13	soluble	amyl alcohol, acetone	11
	Nectandrine A	$C_{20}H_{23}NO_4$	amorphous powder	below 100	...	bitter	insoluble	...	250	soluble	...	12
BERBERIS	Berberine A	$C_{20}H_{17}NO_4$	yellow prisms or needles	120†	neutral	bitter	300 cold ; soluble hot	difficultly cold	very diff.	diff.	difficultly benzene; insoluble petroleum ether	13
	Oxyacanthine A	$C_{19}H_{21}NO_3$	crystalline	138—150‡	alkaline	bitter	difficultly	30 cold ; 4 boiling	soluble	∞	benzene (petroleum ether slightly)	14
	Hydrastine A	$C_{21}H_{21}NO_6$	crystalline, colourless if pure	135	alkaline	bitter in solution	scarcely	soluble	about 80	very soluble	benzene (insoluble petroleum ether)	15
	Xanthopuccine A	...	crys., orange-yl. powder	...	...	...	...	sol. hot	insoluble	insoluble	...	16
BOLDOA	Boldine A	...	...	...	alkaline	bitter	very diff.	soluble	soluble	soluble	difficultly benzene	17
BRAYERA	Kosin B	$C_{31}H_{38}O_{10}$	crys., yl., rhombic	142 (194 ?)	acid	bitter & sharp	scarcely	soluble	soluble	soluble	benzene, CS_2, acetic (glacial)	18
BRYONIA	Bryonin G	...	...	...	...	very bitter	soluble	soluble	insoluble	...	...	19

* Baphiin colourless only if pure, otherwise red.

† Fleitmann, or brown at 110°, black at 160° (Perkin).

‡ The dried pp. by ammonia; or 208—214° the crys. from alcohol.

	CAUSTIC ALKALIES	LEAD ACETATE.	TANNIC ACID.	PLATINUM CHLORIDE.	GOLD CHLORIDE.	SULPHURIC ACID, CONCENTRATED.	NITRIC ACID, CONCENTRATED.	SUNDRY PRECIPITANTS, ETC.
1	soluble	no pp.	...	...	...	intense red on warming	...	(ferric chloride, reddish colour)
2	...	...	...	...	...	...	...	ferric chloride, brown
3	...	...	...	...	...	...	...	(lævo-rotatory)
4	...	...	...	...	...	..	...	(ferric chloride, reddish-brown colour)
5	soluble	...	pp. (alcl.)	...	...	blood-red solution	...	picric acid
6	sol. fluorescent	...	...	...	...	...	...	...
7	...	pp. basic	...	...	...	...	...	(ferric chloride, blue-green coloration)
8	...	...	→	→	→	...	...	pp. most alkaloid precipitants (Dupuy)
9	gives Baphinitin on boiling	pp. white (alcoholic)	...	...	...	...	...	dry HCl ; red~violet~green
10	soluble, yellow	...	...	...	...	soluble yellow	...	Fehling, not reduced
11	pp., difficultly sol. excess	no pp. (neutral)	pp. yl. (sol. HCl warm)	pp. yellow, insol. HCl	pp. yellow-white	dirty olive-green ~ lighter in 15 hours	brown (para-buxine gn.)	picric acid, Mayer's reagent, etc., etc. See § 41 and lists
12	...	...	...	...	...	with MnO_2 magnificent green~violet (like strychnine)	...	...
13	becomes brown-resinous	no pp. neutral	cloud at 3,000	pp.	pp. orange, part soluble cold HCl.	yellow or olive-green ; MnO_2 colours like strychnine	dark brown-red	red ring on pouring Cl solution on the acid Berberine solution. For other tests see § 42 and lists
14	pp., soluble excess	no pp. nl. or basic	pp. white	pp. yellow, soluble HCl	pp.	red.-brown (colourless ~ yellow, O. Hesse)	brown-yellow	pp. phospho-molybdic, $AgNO_3$, $HgCl_2$, etc. See § 42 and lists
15	pp.	...	pp.	pp. yellow-red	pp. yellow-red	faint yellow with pure H_2SO_4	orange (blue, fluorescent on dilution)	K, ferrocy. picric acid, KI+I, etc. See § 42 and lists. Becomes fluorescent with $K_2Mn_2O_8$, not in excess
16	(pp. ammonia)	...	...	...	...	red-brown	red	...
17	pp., soluble excess	...	...	...	...	red	red	KI+I, KI+HgI_2, iodine tincture
18	soluble (and in carbonates)	pp. (neutral)	...	...	...	gives red amorphous substance	...	red coloration ferric chloride
19	...	no pp. (basic.)	pp.	pp.	...	...	...	

	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents.					
							Water.	Alcohol.	Ether.	Chloroform.	Other Solvents.	
CAILCEDRA	Cailcedrin B	...	brittle, resinous	70—80	neutral	bitter	difficultly	soluble	soluble	soluble	...	1
CALABAR BEAN	Physostigmine A	$C_{15}H_{21}N_3O_2$	crystalline, or varnish	69	alkaline	feebly bitter	difficultly	soluble	soluble	soluble	benzene, CS_2 (insoluble petroleum ether)	2
	Calabarine A	...	known only in solution	...	...	...	soluble	soluble	insoluble	...	benzene, CS_2? (insol. petroleum ether?)	3
CALENDULA	Calendulin	...	amorphs., yellow, transparent	...	neutral	tasteless	swells up	soluble	insoluble	...	acetic (glacial)	4
CALOTROPIS	Mudarin B	...	amorphous, yellow	...	...	bitter	sol. cold; jelly hot	soluble	insoluble	...	(insoluble turpentine)	5
CALUMBA ROOT	Calumbin B	$C_{21}H_{22}O_7$	amorphous	182	neutral	bitter	difficultly	diff. cold; sol. hot	diff. cold; sol. hot	soluble	acetic (glacial)	6
CALYCANTHUS	Calycanthine A	...	crystalline	...	...	...	insoluble	insoluble	soluble	...	...	7
CANTHARIDES	[Cantharidin] B	$C_{10}H_{12}O_4$	...	subl. 180	(acid nature)	...	very diff. (NaCl assists)	800 (salts difficultly)	900	80	amyl alcohol, 500 benzene, 2,000 CS_2	8
CAPSICUM	Capsicol A	$C_9H_{14}NO_2$ (T.)*	[liquid red-brown (B.), crys. (T.)]	76; volat. 115 (T.)	...	...	difficultly	soluble (B. and T.)	sol. (B. and T.)	soluble (B.)	(T.) benz., CS_2, acetic (glacial), acetic ether, oils; diff. petr. ether.	9
	Capsicin B	...	amorphous, soft yellow-brown	58	...	burning	difficultly	soluble	soluble	...	benz., petr. ether, acetic ether, oils; diff. CS_2, turpentine	10
CARAPA	Carapin B	C55, H6·5, O38·4 per cent.	resinous	...	...	...	not easily soluble	soluble	not easily soluble	soluble	...	11
	Tulucunnin B	...	amorphous, pale yellow	...	acid	very bitter	150 cold	soluble	insoluble	soluble	...	12
CAROBA LEAVES	Carobin B	...	needles	...	...	...	soluble hot	sol. hot (not cold)	insoluble	...	...	13
	Sparattospermin B	$C_{19}H_{24}O_{10}$	microscopic needles	24	...	...	insoluble	soluble	scarcely	insoluble	(insol. amyl alcohol, petroleum ether)	14
CASSIA	Cathartic acid G	(conts. N. and S.?)	amorphs., brown-black	...	acid	...	...	soluble	insoluble	...	...	15
	Sennapicrin G	$C_{34}H_{58}O_{17}$	amorphous	...	...	bitter	difficultly	soluble	insoluble	...	...	16

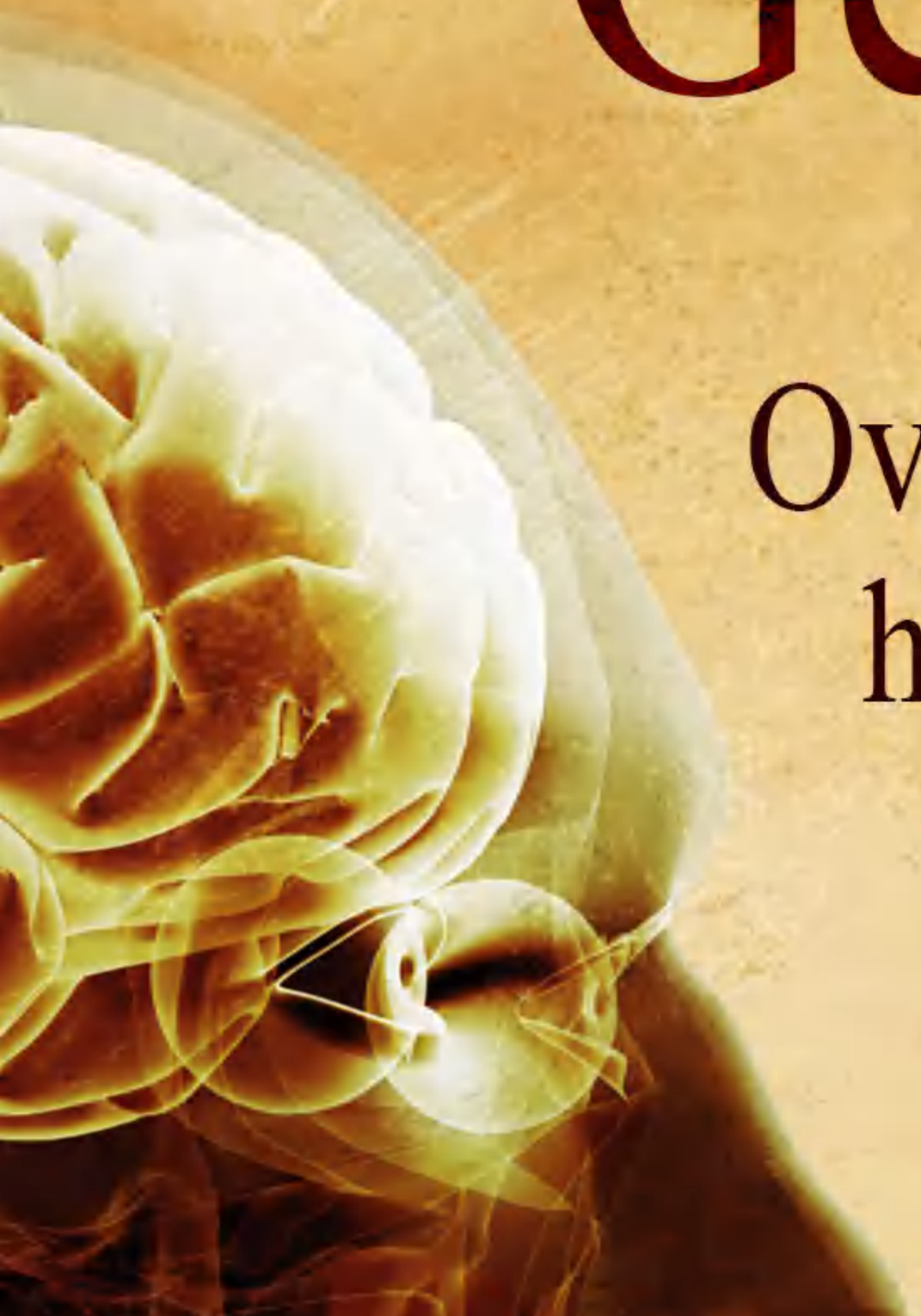
* T refers to Thresh's statements on Capsaicin, and B to Buckheim's on Capsicol.

	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents.				
							Water.	Alcohol.	Ether.	Chloroform.	Other Solvents.
CATECHU	Catechin B	$C_{18}H_8O_8$?	needles	217	slightly acid	slightly bitter	1,133 cold, 3 boiling	5 to 6 cold, 2 or 3 blng.	soluble	...	acetic (glacial); (insoluble turpentine)
CEANOTHUS	Ceanothine A	(conts. N.)	crystalline	190	neutral	bitter	...	...	→	soluble	soluble mixture ether +chloroform
CEPHALANTHUS	Cephalanthin G	$C_{22}H_{34}O_6$	amorphous	...	...	extreme. bitter	difficultly (frothy)	soluble	diff.	diff.	amyl alcohol
	Cephalin G	...	crystalline	...	...	...	very diff.	soluble	soluble	soluble	...
CERBERA	Thevetin G	$C_{54}H_{84}O_{24}$?	micros. plates	170	...	bitter	122 cold	soluble	insoluble	...	acetic (glacial)
	Theveresin G-drv.	...	...	140	neutral	...	difficultly	soluble	scarcely	insoluble	...
	Thevetosin G	...	four-sided prisms	...	...	...	insoluble	...	slightly	...	...
	Odollin	...	...	...	...	...	soluble	soluble	...	insoluble	...
	Cerberin B	(N. free)	crystalline	...	neutral	bitter & burning	insoluble	soluble	sol. 80 % ether	soluble	acetic (glacial)
CHÆROPHYLLUM	Chærophylline A	...	volatile	...	alkaline	...	(sulphate soluble)	(sulphate soluble)	(sulphate soluble)	...	...
CHAMÆLIRIUM	Chamælirin G	...	yellow-red, amorphous	...	...	bitter	soluble frothy	soluble	diff.	insoluble	acetic acid
CHARA	Characin	...	fatty	volatile	...	...	...	soluble	soluble	...	...
CHELIDONIUM	Chelerythrine A	$C_{21}H_{17}NO_4$	needles	resinous at 65	alkaline	bitter	...	soluble	soluble	soluble	benzene, petroleum ether, etc.
	Chelidonine A	$C_{19}H_{17}NO_3 \cdot 2H_2O$	crystalline plates	130 volat.	alkaline	bitter	insoluble	scarcely	scarcely	soluble	amyl alcohol, oils
	Beta-homo-chelidonine A	$C_{21}H_{21}NO_5$	monoclinic crys.	159	...	...	...	...	...	...	...
	Alpha-homo-chelidonine A	$C_{21}H_{21}NO_5$	crystalline	182	...	...	...	...	...	...	...
	Glaucopicrine A	...	crys. or amorphs.	...	alkaline	bitter	soluble	soluble	diff.	...	...
	Glaucine A	$C_{18}H_{19}NO_4$	pearly crystals	below 100	alkaline	bitter & sharp	scarce. cold, sol. hot	soluble	soluble	...	petroleum ether
CHENOPODIUM	Chenopodine A	$C_6H_{13}NO_2$?	microscopic needles	180	neutral	tasteless	11 cold	202 cold	...	...	...
CHIOCOCCA	Caincin G	$C_{40}H_{64}O_{18}$	needles	...	acid	grad. b.	difficultly	sol. hot	diff.	...	...
CHIRETTA	Chirettin B	$C_{26}H_{48}O_{15}$	resinous	...	neutral	very b.	...	soluble	soluble	...	...
	Ophelia Acid	$C_{13}H_{20}O_{10}$	yellow syrup	...	acid	sour and bitter	soluble	soluble	...	...	...

CAUSTIC ALKALIES.	LEAD ACETATE.	TANNIC ACID.	PLATINUM CHLORIDE.	GOLD CHLORIDE.	SULPHURIC ACID, CONCENTRATED.	NITRIC ACID, CONCENTRATED.	SUNDRY PRECIPITANTS, ETC.
...	pp. white (nl. or bc.)	...	reduced	pp. red-brown	soluble warm, purple-red	soluble red fumes	Green-brown, $FerCl_6$. Fehling, $CuSO_4$ and $AgNO_3$ reduced.
...	...	...	...	...	...	...	Fehling reduced after boiling
...	...	...	...	...	...	...	*idem*
soluble	...	...	...	...	...	...	...
...	→	...	→	→	red-brown⁓cherry⁓violet	...	no pp. metallic salts
soluble yellow	...	...	...	...	*idem*	...	...
...	...	...	...	...	...	...	yields sugar and a resin
...	no pp. nl.	...	...	...	violet	...	...
...	...	...	...	...	*idem*	...	decomposed by acids, but no sugar formed
...	...	pp.	...	...	...	...	picric acid
...	...	...	...	...	...	...	...
...	...	...	...	...	...	...	...
pp. gray	...	yellow-red	...	dark red	yellowish-red	...	mercuric chloride, K. chromate, etc.
...	...	...	...	...	...	...	...
...	...	...	...	...	violet	...	phospho-molybdic, bromine water, etc.
...	...	pp. soluble excess	...	...	yellow	...	*idem*
pp.	no pp. nl.	pp.	...	...	dark green on warming	...	...
pp.	*idem*	pp.	...	...	blue⁓violet⁓red	...	...
...	...	...	pp.	...	...	...	...
soluble	pp. basic	...	...	...	...	...	(no pp. ferric chloride)
...	no pp.	pp.	...	...	...	...	(Fehling's solution not reduced)
soluble yellow	pp. neutral	...	...	...	...	...	Fehling's solution reduced

	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents.					
							Water.	Alcohol.	Ether.	Chloro-form.	Other Solvents.	
Chrysan-themum	Tanacetin B	...	amorphous	...	...	bitter	soluble	soluble	insoluble	...	...	
	Chrysanthemine A	$C_4H_{28}N_2N_8$	needles (or syrup)	above 100	alkaline	...	soluble	soluble	insoluble	insoluble	methyl alcohol	
Cichorium	Chicorin G	$C_{38}H_{34}O_{19}\cdot 4\frac{1}{2}H_2O$	crystalline	215—220	...	bitter	soluble hot	soluble	insoluble	...	...	
Cinchona	Quinine A	$C_{20}H_{24}N_2O_2+Aq$	ndls. or amorphs.	171—172*	alkaline	very b.	1,667 cold†	2 parts	23 or 60‡	1·8 cold	benz., petr. ether, etc.	
	Quinicine A	$C_{20}H_{24}N_2O_2$	amorphous	about 60	alkaline	bitter	difficultly	soluble	soluble	soluble	...	
	Apoquinine A	$C_{19}H_{22}N_2O_2\cdot 2H_2O$	crystalline	160	alkaline	bitter	soluble hot	soluble	soluble	soluble	...	
	Hydroquinine A	$C_{20}H_{26}N_2O_2$	crystalline	168	alkaline	bitter	...	soluble	soluble	soluble	benzene, CS_2, etc.	
	Cinchonine A	$C_{19}H_{22}N_2O$	needles or prisms	165§	alkaline	bitter	3,800 cold, 2,500 blng.	140 cold	sc. cold, 371 hot	40 cold	amyl alcohol, hot benzene	
	Isocinchonine A	$C_{19}H_{22}N_2O$	...	125 volat.	alkaline	...	insoluble	easily	easily	easily	benzene, acetone	
	Apocinchonine A	$C_{19}H_{22}N_2O$	prisms	209	alkaline	bitter	insoluble	soluble	diff.	diff.	...	
	Apocinchonicine A	...	...	...	...	...	...	soluble	soluble	soluble	...	
	Diapocinchonine A	$C_{38}H_{44}N_4O_2$	yellow amorphous	...	...	...	...	soluble	soluble	soluble	...	
	Hydrocinchonine A	$C_{19}H_{24}N_2O$	amorphous	256	alkaline	...	...	...	...	...	...	
	Cinchonicine A‖	$C_{19}H_{22}N_2O$	yellowish amphs.	about 50	alkaline	bitter	...	soluble	soluble	soluble	benzene, acetone	
	Cinchotine A	$C_{19}H_{24}N_2O$	prisms and leaflets	268—270	...	...	1,360 cold	soluble	534 cold	...	...	
	Dicinchonine A	$C_{40}H_{46}N_4O_3$	amorphous	...	alkaline	...	insoluble	soluble	soluble	soluble	benzene, acetone	
	Quinidine A	$C_{20}H_{24}N_2O_2+2\frac{1}{2}H_2O$	prisms or needles	168*	feebly alk.	bitter	2,000 cold	26 cold 80 per cent.	30 cold	diff.	amyl alcohol, benzene	
	Apoquinidine A	$C_{19}H_{22}N_2O_2$	amorphous	137*	alkaline	...	...	soluble	soluble	...	...	
	Hydroquinidine A	$C_{20}H_{26}N_2O_2\cdot 2\frac{1}{2}H_2O$	needles or plates	166—167	...	...	...	readily	diff.	readily	...	
	Cinchonidine A	$C_{19}H_{22}NO_2$	prisms	175—206§	alkaline	bitter	1,680 at 10°	19 of 80 %	76	soluble	...	
	Apocinchonidine A	$C_{19}H_{22}NO_2$	crystalline	...	...	very bitter	...	...	...	...	...	
	Cinchamidine A	$C_{19}H_{24}NO_2$?	needles or plates	230	...	...	...	soluble	diff.	soluble	...	
	Aricine¶	$C_{20}H_{26}N_2O_4\cdot 2H_2O$	long needles	188	alkaline	not bitter	scarcely	100 cold, 11 boiling	33	...	...	
	Cusconine A	*idem*	lustrous prisms or plates	110*	feebly alk.	not bitter	scarcely	soluble	soluble	very soluble	acetone	
	Cusconidine A	...	brown, amorphous	...	...	...	...	...	...	...	...	

* The anhydrous base. † Sestini's figures. ‡ 23 = Vandenburg, 60 = Merck § Authorities differ. ‖ For other derivatives of Cinchonine see Part I.
¶ Other Cinchona bases are described in Part I.



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	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents.				
							Water.	Alcohol.	Ether.	Chloroform.	Other Solvents.
CINCHONA	Concusconine A	$C_{23}H_{26}N_2O_4 \cdot H_2O$	monoclinic prisms	144—206*	neutral	tasteless	insoluble	sparingly	readily	readily	benzene
	Cuscamine A	...	crystalline	218	...	rather b.	...	soluble	readily	readily	...
	Chairamine A	$C_{22}H_{26}N_2O_4 \cdot H_2O$	needles or prisms	233†	neutral	...	difficultly	540 of 97 %	readily	readily	(insol. hydrochloric)
	Conchairamine A	$C_{22}H_{26}N_2O_4$	prisms	120†	nearly neutral	...	...	sol. hot	readily	readily	(scarcely hydrochloric)
	Chairamidine A	*idem*	amorphous	127	neutral	...	insoluble	soluble	soluble	soluble	benzene
	Conchairamidine A	*idem*	prisms	114—115	neutral	...	insoluble	soluble	soluble	soluble	benzene, acetone
	Quinamine A	$C_{19}H_{24}N_2O_2$	prisms	172	...	tasteless‡	1,516 at 16° C.	105 at 20° C.	48 at 16° C.	soluble	petroleum ether, benzene boiling
	Quinamicine A	*idem*	crystalline	109	...	...	...	soluble	soluble	soluble	...
	Conquinamine A	*idem*	triclinic prisms	121	...	...	difficultly	soluble	soluble	soluble	benzene, CS_2
	Cupreine A	$C_{19}H_{22}N_2O_2$	prisms	198†	alkaline	...	...	soluble	diff.	diff.	...
	Homoquinine A	$C_{39}H_{46}N_4O_4$	radiating prisms	177	alkaline	...	...	soluble	diff.§	soluble	...
	Cinchonamine A	$C_{19}H_{24}N_2O$	hexagonal prisms	195	...	bitter	difficultly	soluble	soluble	soluble	benzene, CS_2
	Paricine A	$C_{16}H_{18}N_2O$	yellow powder	...	feebly alk.	bitter	scarcely	soluble	soluble	...	...
	Quinovin G	$C_{30}H_{48}O_8$	resinous powder	...	neutral	gradual. bitter	scarcely cold	soluble	soluble	soluble	oils
	Quinovic Acid	$C_{24}H_{38}O_4$	needles or scales	...	...	tasteless	...	slightly	slightly	slightly	...
	Cascarillin B	...	resinous	...	neutral	bitter	scarcely	soluble	soluble	...	benzene
	Copalchin B	...	resinous	...	...	bitter	scarcely	soluble	'partly'‖	soluble	...
	Lignoin B	$C_{20}H_{23}NO_8$	brown, amorphous	...	...	...	...	soluble	...	...	...
	Californin B	...	yellow, amorphs.	...	neutral	bitter	soluble	soluble	insoluble	...	...
	Moradeine A	...	opaque prisms	199—200	...	...	slightly	readily	readily	readily	...
	Javanine A	...	rhombic plates	...	...	...	soluble	...	soluble	...	...
CNICUS	Cnicin B	$C_{42}H_{56}O_{15}$	silky needles	fusible	neutral	bitter	scarcely cold	∞	scarcely	soluble	benzene
COCA	Cocaine A	$C_{17}H_{21}NO_4$	4-6-sided prisms	98	alkaline	slightly bitter¶	704 cold	soluble	readily	sol. hot	benzene
	Benzoyl-cegonine A	$C_{16}H_{19}NO_4 \cdot 4H_2O$	prisms or needles	variable 87—140	...	...	soluble hot	soluble	scarcely	sol. hot	acetone, wood spirit

* Melts, resolidifies, then remelts at the higher temperature.
§ Easier soluble when freshly precipitated.
† The anhydrous base.
‖ Presumably a mixture.
‡ Salts bitter.
¶ Sense of taste temporarily destroyed.

Caustic Alkalies.	Lead Acetate.	Tannic Acid.	Platinum Chloride.	Gold Chloride.	Sulphuric Acid, Concentrated.	Nitric Acid, Concentrated.	Sundry Precipitants, etc.
pp. resinous	no pp. nl.	...	yellow flocks	dirty yellow	blue-green	dark green	...
pp.	...	...	...	...	...	...	not fluorescent. No colour ferric chloride
pp.	...	...	yellow needles	...	colourless	gradual green	Fröhde's solution, colourless
pp. insoluble excess	...	...	dark yellow	...	brown⁓intense green	to HCl solution, green	potass. sulphocyanide
idem	...	...	yellow flocks	...	yellow⁓dark green	*idem* →	insoluble nitric acid alone
pp. oily⁓crystalline	...	...	...	...	intense green	*idem*	*idem*
...	...	...	...	yellow to red, Au reduced	...	...	(non-fluorescent)
pp. insol. excess	...	...	...	yellow if pure	(dilute acids in excess, pp.)	←	potass. iodide, sodium chloride, etc.
...	...	...	...	yl.⁓purple	...	...	...
pp. soluble excess	...	...	pp.	...	...	...	gives thalleioquin reaction
decomposes	...	...	...	...	fluorescent	not decompd.	*idem*
...	...	...	...	...	...	...	no thalleioquin reaction
...	...	...	...	...	yellow-green	...	...
soluble	no pp.	...	→	→	dark red	soluble warm; decomposed	pp. metallic salts. Fehling not reduced
soluble frothy	...	...	...	...	solution unchanged	...	copper sulphate, green pp.
...	no pp.	no pp.	...	...	blood-red	red-violet	(hydrochloric acid, violet colour)
...	no pp. nl.	pp.	...	...	red	...	...
soluble brown	pp. flesh-colour	...	...	...	...	...	
...	...	no pp.	no pp.	...	brown-red	...	no pp. mercuric chloride
...	...	...	...	...	...	...	...
...	...	...	...	...	intense yellow	...	...
...	partial pp. basic	...	no pp.	...	red ; benzoic odour on warming	...	(silver nitrate not reduced)
pp. amorphous ⁓ crystalline	no pp.	pp. white	whitish 12,500	light yellow	benzoic acid vapour on warming	...	phospho-molybdic, Mayer's solution, etc.
soluble	...	...	...	yellow crystals	...	...	

	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents. Water.	Alcohol.	Ether.	Chloroform.	Other Solvents.	
Coca	Ecgonine A	$C_9H_{15}NO_3$	rhombic crystals	205*	neutral	bitter & sweet	readily	difficultly	insoluble	diff.	...	1
	Cocamine A	$C_{19}H_{23}NO_4+\frac{1}{2}H_2O$	amorphous	80	neutral	bitter	diff. cold	soluble	soluble	soluble	benzene, acetone	2
Cocculus Ind.	Picrotoxin B	$C_{30}H_{34}O_{13}$	silky needles	199—201	neutral	very bitter	150 cold, 25 boiling	3	250	soluble	amyl alcohol, acetic acid	3
	Cocculin	$C_{19}H_{26}O_{10}$	needles	...	...	...	scarcely cold	scarcely	scarcely	...	...	4
	Menispermine A	$C_{18}H_{24}N_2O_2$	prisms	120	alkaline	not b.	insoluble	soluble	soluble	...	.	5
Colchicum	Colchicine A	$C_{22}H_{25}NO_6$	amorphous usually	140	feebly alk.	bitter	∞	readily	scarcely	soluble	benzene, amyl alcohol	6
	Colchiceine A	$C_{21}H_{23}NO_6$	needles	161—172	neutral	...	difficultly cold	readily	scarcely	soluble	*idem*	7
Colocynthis	Colocynthin G	$C_{56}H_{84}O_{23}$	usually amorphous	...	neutral	bitter	8 cold	soluble	diff.	diff.	...	8
Conium	Coniine A	$C_8H_{17}N$	liquid	†	alkaline	burning	100 cold	∞	6 cold	soluble	benz., petr. ether, etc.	9
	Conhydrine A	$C_8H_{17}NO$	pearly plates	126·6	alkaline	...	soluble	readily	readily	...	...	10
	Cicutoxin B	...	resinous	...	acid	disagreeable	soluble hot	soluble	soluble	soluble	not petroleum ether	11
Convallaria	Convallamarin G	$C_{46}H_{44}O_{24}$?	crystalline	...	neutral	b.-sweet	soluble	soluble	insoluble	soluble	amyl alcohol	12
	Convallarin G	$C_{34}H_{31}O_{11}$?	rectangular plates	...	neutral	sharp	scarcely	soluble	insoluble	...	...	13
Convolvulaceæ	Convolvulin G	$C_{31}H_{50}O_{16}$	amorphous	150	feebly acid	tasteless	scarcely	soluble	insoluble	scarcely	...	14
	Jalapin G	$C_{34}H_{56}O_{16}$	resinous	150	*idem*	*idem*	difficultly	soluble	soluble ?	soluble	amyl alcohol	15
	Turpethin G	*idem*	brown-yellow, amorphous	183	...	sharp & bitter	insoluble	soluble	insoluble	...	...	16
	Tampicin G	*idem*	resinous	130	...	...	...	soluble	soluble	...	...	17
Coriaria	Coriamyrtin G	$C_{30}H_{36}O_{10}$	clinorhombic prisms	220	neutral	tasteless	70	50	soluble ?	soluble	benzene	18
Cornus	Cornin B	...	silky needles	...	neutral	bitter	soluble	soluble	soluble	...	...	19
Corydalis	Corydaline A	$C_{22}H_{29}NO_4$	prisms or needles	134	alkaline	bitter in solution	insoluble	difficultly	soluble	soluble	benzene, amyl alcohol, etc.	20
	Fumarine A	$C_{20}H_{19}NO_2$	clinorhom. prisms	...	alkaline	bitter	difficultly	soluble	sparingly	11	78 benz., amyl alcl., etc.	21

* Anhydrous base.

† Authorities differ as to boiling point.

	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents.				
							Water.	Alcohol.	Ether.	Chloroform.	Other Solvents.
Corynocarpus	Karakin G ?	...	stellate needles	100	feebly acid	bitter	soluble hot	soluble	insoluble	insoluble	...
Cratægus	Cratægin B	...	grayish crystals	...	...	bitter	soluble	difficultly	insoluble	...	...
Crepis	Crepin	...	...	...	...	bitter	...	...	soluble	...	...
Cucumis	Melonemetin B	...	brown amorphs.	...	...	...	soluble	soluble	insoluble	...	...
	Prophetin G	$C_{20}H_{36}O_7$	resinous	...	...	bitter	difficultly	very sol.	soluble	...	...
Cyclamen	Cyclamin G	$C_{20}H_{34}O_{10}$	crys. or amorphs.	236	neutral	sharp	sol. frothy	71 of 96 %	insoluble	insoluble	...
	Cyclamiretin	$C_{14}H_{16}O_6$	resinous	198	...	tasteless	insoluble	soluble	*	...	...
Cytisus	Cytisine A	$C_{11}H_{14}N_2O$	crystalline	154	alkaline	rather b.	very sol.	very sol.	scarcely	scarcely	amyl alcohol
Damiana	Damianin B	(N. free)	brownish amorphous	...	...	bitter	soluble	soluble	insoluble	insoluble	(insoluble benzene, petroleum ether)
Daphne	Daphnin G	$C_{15}H_{16}O_9 \cdot 2H_2O$	needles or rectangular plates	200	neutral	b. & astringent	soluble hot	soluble hot	insoluble	soluble	warm amyl alcohol
	Daphnetin	$2(C_9H_6O_4) \cdot H_2O$	prisms	253	feebly acid	astringent	soluble boiling	soluble hot	diff.	insoluble	(solutions yellow)
	Coccognin B	$C_{20}H_{22}O_8$	crystalline	...	...	...	difficultly	soluble	insoluble	...	...
Datisca	Datiscin G	$C_{21}H_{22}O_{11}$	silky needles	186	neutral	bitter	soluble hot	soluble	diff.	...	...
Delphinium	Delphinine A	$C_{22}H_{35}NO_6$	rhombic crystals	*	alkaline	bitter & sharp	1,594	44	53	readily	benzene
	Staphisagrine A	$C_{22}H_{35}NO_5$	amorphous	90	alkaline	bitter	200	soluble	855	soluble	...
	Delphinoidine A	$C_{42}H_{68}N_2O_7$?	amorphous	110—120	alkaline	...	scarcely	18 of 90 %	37 absolute	soluble	benzene, petr. ether
	Delphisine A	$C_{27}H_{46}N_2O_4$	crystalline	189	...	...	slightly	370	71 absol.	readily	*idem*
	Calcitrapine A	...	...	...	...	...	...	...	soluble	soluble	...
Digitalis	Digitalin G	...	needles	243	neutral	grad. b.	slightly	readily	diff.	soluble	...
	Digitogenin	$C_{15}H_{24}O_3$	...	...	...	...	...	100 of 93 %	...	soluble	acetic acid
	Digitoxin	...	...	...	...	...	insoluble	sol. hot	diff.	sol. hot	...
	Digitonin G	$C_{31}H_{52}O_{17}$	crystalline or amorphous	235	...	...	soluble	sparingly	insoluble	insoluble	(insoluble benzene)
	Digitalein G	(N. free)	gummy	fusible	...	bitter	soluble	soluble	sparingly	...	...
Ecbalium	Elaterin B	$C_{20}H_{28}O_5$	crystalline	200	neutral	bitter	diff. cold	soluble	125	soluble	amyl alcohol
	Ecbalin B	$C_{20}H_{34}O_4$	resinous	...	...	bitter	20	soluble	soluble	...	...
	Elaterid B	$C_{20}H_{32}O_{12}$	...	...	...	bitter	insoluble	sol. dilute	insoluble	...	...

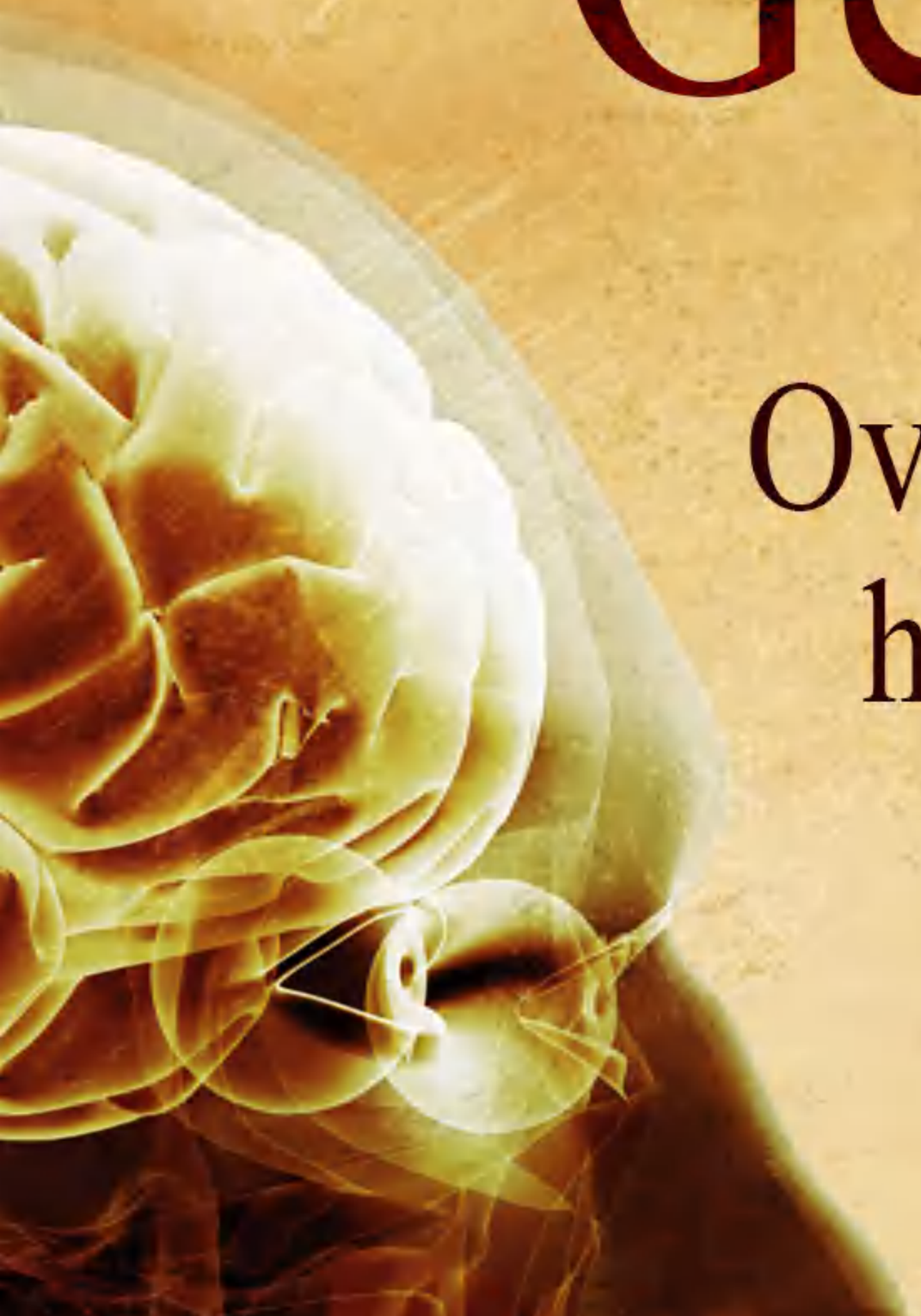
* Authorities differ.

CAUSTIC ALKALIES.	LEAD ACETATE.	TANNIC ACID.	PLATINUM CHLORIDE.	GOLD CHLORIDE.	SULPHURIC ACID, CONCENTRATED.	NITRIC ACID, CONCENTRATED.	SUNDRY PRECIPITANTS, ETC.
soluble	...	no pp.	...	...	dark rose on warming	...	Fehling, green pp.
...	no pp.	...	...	...	...	...	
...	no pp.	...	...	...	...	...	
...	...	...	...	...	...	...	...
...	no pp.	pp.	...	...	red-brown	...	...
...	pp. (nl.)	...	...	...	red	...	Fehling, white pp.
...	...	...	...	...	...	...	Fehling, pp.
soluble	no pp. (bc.)	pp. 1 in 300	no pp. dilute	pp.	colourless	...	phos.-molybdic, 1 in 10,000 ; picric, etc.
...	no pp.	⟶	⟶	⟶	...	...	no pp. usual reagents
soluble yellow	pp. (basic)	...	...	...	red	red	
soluble red	pp. (neutral)	...	...	...	yellow on warming	gradual intense red	Fehling reduced
...	...	...	...	...	...	...	...
soluble yellow	pp. yellow	no pp.	...	...	...	...	ferric chloride, brown-green pp.
pp. insoluble excess	...	pp.	greenish	lemon	no effect, or light brown	no effect or yellow	phospho-molybdic, picric, etc.
...	...	...	...	...	colours uncertain	no effect	...
...	...	...	...	...	red-brown	no marked colour	(Fröhde's solution, blood-red colour)
...	...	...	...	...	...	...	..
...	...	...	cloud	yellow	dark-brown⇝violet⇝gray-brown	...	phospho-molybdic, yellow⇝blue
...	no pp.	pp.	...	crystalline	green or brown	colourless*	no pp. with most reagents
...	...	...	...	...	...	...	...
	...	...	...	...	...	...	
(pp. ammonia)	pp. nl. or basic	pp.	...	...	red ; violet on dilution	...	..
...	...	pp.	...	...	colourless or reddish-brown	...	(hydrochloric, brown-green colour)
soluble	no pp.	...	...	...	yellow to dark red	...	(Fröhde's solution, yellow colour)
...	...	...	...	...	...	red with decomposition	...
soluble	no pp.	pp.	...	...	...	...	...

* Authorities differ.

	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Water.	Alcohol.
Ecbalium	Hydroelaterin	$C_{20}H_{30}O_6$	yellow amorphs.	...	...	not b.	soluble	soluble
Ergot	Ergotine A	$C_{50}H_{52}N_2O_8$	amorphous	...	alkaline	feebly b.	soluble	soluble
	Ecboline A	...	amorphous	...	alkaline	slightly bitter	soluble	soluble
	Ergotinine A	$C_{35}H_{40}N_4O_6$	crystalline	...	...	bitter	insoluble	soluble
Erythrea	Erythrocentaurin G	$C_{27}H_{24}O_8$?	crystalline or amorphous	136	neutral	bitter ?	1,360 cold	48 of 86 %
Erythrophlœum	Erythrophlœine A	...	crystalline	...	alkaline	...	soluble	soluble
Euonymus	Euonymin G	...	amorphous	...	...	bitter	slightly	soluble
Eupatorium	Eupatorine A	$C_{20}H_{25}O_{36}\cdot HNO_3$?	micros. needles	*	alkaline	bitter	insoluble	soluble
	Eupatorin G	...	...	...	...	acid	soluble hot	soluble
Euphorbia	Euphorbon	$C_{13}H_{22}O$?	needles or prisms	106—116	neutral	burning	insoluble	soluble hot
Fraxinus	Fraxin G	$C_{16}H_{18}O_{10}$	crystalline	320	neutral	feebly bitter	1,000 cold	difficultly cold
	Fraxetin	$C_{10}H_8O_5$	yellow needles	227	acid	not bitter	10,000 cold, 33 boiling	difficultly
Fritillaria	Imperialine A	$C_{35}H_{60}NO_4$	needles	254	...	very bitter	slightly	soluble hot
Garcinia	Mangostin B	$C_{20}H_{22}O_5$	yellow plates	190	neutral	tasteless	insoluble	soluble
Gardenia	Gardenin B	$C_{14}H_{12}O_6$	dark yellow crystals	163—164	...	...	scarcely	soluble
Gentiana	Gentiopicrin G	$C_{20}H_{30}O_{12}\cdot H_2O$	yellow amorphs., or crystalline	120—125	neutral	bitter	soluble	soluble
	Gentiogenin	...	yl.-bn., amorphs.	...	...	bitter	difficultly	soluble
	Gentianin	$C_{14}H_{10}O_5$	yellow crystalline	...	neutral	not b.	scarcely	difficultly
Geum	Geum bitter	...	yellow amorphs.	...	neutral	bitter	difficultly	soluble
Globularia	Globularin	$C_{15}H_{20}O_8$	amorphous	...	neutral	bitter	soluble	soluble
Glycyrrhiza	Glycyrrhizin B	$C_{44}H_{68}NO_{18}$	needles†	200	acid	bitter-sweet	soluble hot	soluble‡

* Not fusible without decomposition. † From glacial acetic acid. ‡ Spa



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	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents.					
							Water.	Alcohol.	Ether.	Chloro-form.	Other Solvents.	
Gono-lobus	Gonolobin G	$C_{40}H_{74}O_6$	yellow powder	112	...	...	slightly	slightly	insoluble	...	insoluble petroleum ether	1
Gratiola	Gratiolin G	$C_{42}H_{36}O_{14}$?	needles	200	...	bitter	893 cold	soluble	1000 cold	soluble	...	2
Harmala	Harmaline A	$C_{13}H_{14}N_2O$	octahedra	238	alkaline	slightly bitter	6,000 cold	sol. hot	diff. cold	...	slightly petroleum ether	3
	Harmine A	*idem*	prisms	256—257	alkaline	b. in soln.	difficultly	diff. cold	diff.	...	...	4
Hedera	Helixin G	$C_{82}H_{54}O_{11}$	needles	233	acid	b.-sweet	insoluble	soluble	diff.	insoluble	hot benzene	5
Helio-tropum	Heliotropine A	...	rhombic crystals	...	alkaline	bitter	soluble	...	soluble	...	...	6
Helle-borus	Helleborein G	$C_{26}H_{44}O_{15}$	fine needles	partly 280	slightly acid	sweetish	soluble	soluble	insoluble	soluble	...	7
	Helleboretin	$C_{14}H_{20}O_3$	bluish-gray	...	neutral	tasteless	insoluble	sol. (violt.)	scarcely	...	..	8
	Helleborin G	$C_{36}H_{42}O_6$?	needles	above 250	neutral	burning	diff. cold	soluble	diff.	soluble	difficultly in oils	9
Humulus	Lupulinic Acid	(N free)	crystalline	56	acid	bitter	insoluble	soluble	soluble	...	...	10
	Alpha-resin	*idem*	soft red-brown	...	feebly acid or neutral	very bitter	slightly	soluble	soluble	soluble	petroleum ether	11
	Beta-resin	*idem*	softer than A	...	*idem*	bitter	slightly	soluble	soluble	soluble	petroleum ether	12
	Gamma-resin	*idem*	brown, brittle	...	*idem*	not b.	slightly	soluble	soluble	soluble	insoluble petr. ether	13
Hura	Hurin B	...	oily or crystalline	above 100*	neutral	burning	insoluble	soluble	soluble	...	oils	14
Hyæn-anche	Hyænanchin B	...	crystalline or amorphous	...	...	bitter	soluble	soluble	soluble	...	benzene, oils	15
Hymeno-dictyon	Hymenodictyonine A	$C_{24}H_{40}N_3$?	needles or amorphous	66	alkaline	bitter	difficultly	soluble	soluble	soluble	benzene, petroleum ether	16
Ilex	Ilexanthin	$C_{17}H_{22}O_{11}$	yellow, crystalline	198	...	...	scarce. cold, soluble hot	soluble	insoluble	...	...	17
Ipecacu-anha	Emetine A	$C_{30}H_{40}N_2O_5$?	needles, plates, or amorphous	50—70†	alkaline	bitter	slightly	readily	diff.	readily	benzene, turpentine, etc.	18
Isatis	Indican G	$C_{26}H_{31}NO_{17}$	yellow, syrupy	...	acid	bitter	soluble	soluble	diff.	...	...	19
Juniperus	Juniperin B	...	yellow, amorphs.	...	...	...	60	sol. hot	soluble	...	...	20
Lactuca	Lactucin B	$C_{22}H_{18}O_7$?	pearly crystals	fusible	neutral	bitter	soluble hot	soluble	insoluble	...	..	21
	Lactucon	$C_{15}H_{24}O$?	crystalline stars	185	...	tasteless	scarcely	soluble	soluble	...	petroleum ether, oils	22

* Volatile at lower temperatures.

† Authorities differ.

CAUSTIC ALKALIES.	LEAD ACETATE.	TANNIC ACID.	PLATINUM CHLORIDE.	GOLD CHLORIDE.	SULPHURIC ACID, CONCENTRATED.	NITRIC ACID, CONCENTRATED.	SUNDRY PRECIPITANTS, ETC.
...	...	...	...	...	...	...	(no pp. Mayer's solution)
(soluble ammonia)	no pp.	pp.	...	...	orange⁓brown (red at edges)	yellow	...
pp. and carbonates	...	...	light yellow crystals	...	(acids pp. from solution)	...	potass. ferrocyanide, etc.
idem	...	...	yl. amorphs.	...	...	...	*idem*
soluble green hot	pp. neutral	...	...	...	bright red	no colour	Fehling not reduced
oily pp.	...	→	→	→	yellow, gradual red	...	pp. most alkaloid reagents
(no effect)	no pp. basic	pp.	no pp.	...	brown-red (shade of violet)	...	phospho-molybdic, etc.
...	...	...	...	...	brown-red	...	...
...	...	...	...	...	blood-red	...	...
...	...	...	...	...	...	...	(converted to resins on exposure)
soluble brownish	pp. neutral*	...	...	...	...	...	ferric chloride, brown pp.
idem	no pp. alcoholic	...	...	...	...	...	(strong hop odour)
idem	*idem*	...	...	...	...	...	no cupric reaction, see Part I.
unchanged	...	...	...	...	...	resinified	vapours cause inflammation
...	no pp. neutral	...	...	...	...	...	...
pp. gelatinous	...	pp. brown	pp.	...	...	...	phospho-molybdic, Mayer's solution, etc.
orange pp.	yellow pp.	...	...	...	...	...	(Fehling not reduced)
p. and carbonates	...	pp. (acid solution)	yellowish	pale yellow	dirty brown	yellow-brown	phospho-molybdic, picric, etc.
decompose	pp. neutral or basic	...	...	...	(acids produce indigo)	...	Fehling reduced
soluble ammonia)	...	...	...	...	light yellow	...	..
soluble red	no pp.	...	...	...	colourless⁓red	...	Fehling.redneed
...	...	...	...	...	...	...	no pp. reagents soluble alcohol

* Soluble in excess, if alcoholic solution.

	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents.					
							Water.	Alcohol.	Ether.	Chloroform.	Other Solvents.	
Laserpitium	Laserpitin B	$C_{16}H_{22}O_4$	prismatic crystals	114—118	neutral	bitter in solution	insoluble	9 absolute	3·6	soluble	CS_2, benz., petr. ether	1
Laurineæ	Laurotetanine A	...	crystalline	...	...	...	...	...	scarcely	diff.	...	2
Leguminosæ (Indian)	Derrid B	(N free)	...	...	slightly acid	sharp	slightly	soluble	soluble	soluble	amyl alcohol	3
	Erythrine A	...	(sulphate crys.)	...	...	...	...	...	soluble	...	...	4
	Pithecolobine A	...	amorphous	...	...	burning	milky soln.	...	...	...	...	5
Linaria	Linarin B	...	crystalline	...	...	bitter	soluble	soluble	...	...	...	6
Linum	Linin B	(N free)	silky needles	...	neutral	b. insoln.	insol. cold	very sol.	soluble	soluble	glacial acetic acid	7
	Linamarin G	N 5·55 per cent.	*idem*	134	neutral	bitter	readily	soluble	insoluble	...	...	8
Liriodendron	Liriodendrin B	...	needles and leaflets	82	neutral	bitter	scarcely cold	soluble	soluble	...	...	9
Lobelia	Lobeline A-G	...	oily, yellowish	...	alkaline	acrid	soluble yellow	soluble	soluble	soluble	amyl alcohol, benzene, etc.	10
Loliam	Loliin G	...	amorphous, dirty white	...	...	bitter	soluble	soluble	insoluble	...	...	11
Lonicera	Xylostein G	...	needles	100	neutral	slightly bitter	scarcely cold	soluble	soluble	...	...	12
Lotus bark	Loturine A	...	prisms	234	alkaline	...	scarcely	soluble	soluble	soluble	acetone	13
	Colloturine A	...	prisms	subl. 234	alkaline	...	...	soluble	soluble	...	...	14
	Loturidine A	...	amorphous	...	...	...	...	...	...	...	...	15
Lupinus	Lupinin G	$C_{29}H_{32}O_{16}{\cdot}7H_2O$	yellowish needles	...	...	bitter	difficultly	difficultly	...	...	...	16
	Lupinidine A	$C_8H_{15}N$	yellow, oily	...	alkaline	pungent, bitter	soluble	soluble	diff.	...	...	17
	Lupinine A	$C_{21}H_{40}N_2O_2$	rhombic crystals	67—80	alkaline	bitter	sol. cold*	soluble	soluble	soluble	benzene, petrol. ether	18
	Lupanine A	$C_{15}H_{24}N_2O$	yellow, syrupy	liquid	alkaline	very b.	sol. cold*	difficultly	soluble	soluble	...	19
Lycopodium	Lycopodine A	$C_{32}H_{52}N_2O_3$	clino-rhom. prms.	114—115	alkaline	bitter	soluble	soluble	soluble	soluble	benzene, amyl alcohol	20
	Piliganine A	$C_{15}H_{24}N_2O$	yellow, amorphs.	...	alkaline	...	soluble	soluble	diff.	soluble	...	21
Marrubium	Marrubiin B	...	needles or amorphous	148	neutral	bitter	slightly	soluble	soluble	soluble	(insoluble petroleum ether)	22
Menyanthes	Menyanthin G	$C_{30}H_{46}O_{14}$	amorphous	100—115	neutral	bitter	difficultly cold	soluble	insoluble	soluble	benzene	23
Mikania	Guacin B	...	brown, amorphous	100	neutral	bitter	diff. cold	soluble	soluble	...	...	24

* *Less* soluble warm.

	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents.					
							Water.	Alcohol.	Ether.	Chloro-form.	Other Solvents.	
MORRENIA	Morrenine A	...	dark bn., amorphs.	106	...	bitter	soluble	soluble	...	soluble	amyl alcohol	1
NANDINIA	Nandinine A	$C_{19}H_{19}NO_4$	amorphous	...	...	...	insoluble	soluble	soluble	soluble	benzene	2
NARTHECIUM	Narthecin B	...	crystalline	35	acid	biting	scarcely	soluble	...	...	...	3
NERIUM	Conessine A	$C_{12}H_{20}N$	whitish powder	122	...	bitter	difficultly	difficultly	scarcely	...	carb. bisulphide	4
	Oleandrine A	...	yellow, resinous	above 56	...	bitter	soluble*	soluble	scarce.?*	soluble	oils (not benzene)	5
	Rosaginin G	(N free)	microscopic crys.	171	...	...	insoluble	soluble	insoluble	insoluble	...	6
NIGELLA	Melanthin G	$C_{20}H_{33}O_7$	crystalline	205	...	slightly bitter	scarcely	soluble hot	scarcely	slightly	scarcely benzene	7
	Damascenine A	$C_{20}H_{17}NO_6$	crys. below 27° C.	27	alkaline	...	insoluble	easily	easily	...	CS_2, benzene, etc.	8
NUPHAR	Nupharine A	$C_{14}H_{24}N_2O_2$	amorphous	65	...	not b.†	...	soluble	soluble	soluble	amyl alcohol, acetone	9
ONONIS	Ononin G	$C_{30}H_{34}O_{13}$	micros. prisms, etc.	235	...	not b.	insol. cold	difficultly	insoluble	...	...	10
	Ononid B	$C_{18}H_{22}O_8$	dark yellow, amorphous	...	acid	slightly bitter	soluble	soluble	...	...	...	11
OPIUM	Morphine A	$C_{17}H_{19}NO_3{\cdot}H_2O$	rhombic prisms	about 230	alkaline	bitter	slightly	40 cold	scarcely	scarcely, pure‡	warm amyl alcohol	12
	Apomorphine A	$C_{17}H_{17}NO_2$	Greenish crystals	...	alkaline	...	difficultly	soluble	soluble purple	soluble violet	benzene	13
	Pseudomorphine A	$C_{17}H_{18}NO_3$	needles or leaflets	not fusible	neutral	not bitter	insoluble	insoluble	insoluble	insoluble	hot amyl alcohol	14
	Narcotine A	$C_{22}H_{23}NO_7$	rhombic prisms or needles	115—176*	neutral	tasteless†	scarcely	soluble	166 cold	readily	benzene, etc.	15
	Hydrocotarnine A	$C_{12}H_{15}NO_3{\cdot}1\frac{1}{2}H_2O$	monoclinic prisms	50—55	alkaline	...	...	soluble	soluble	soluble	benzene	16
	Cotarnine A	$C_{12}H_{13}NO_3{\cdot}H_2O$	radiating needles	132	...	...	diff. cold	soluble	readily	...	...	17
	Nartinic Acid	$C_{20}H_{16}N_2O_6$	orange-coloured	...	...	...	...	...	...	...	...	18
	Cupronine A	$C_{20}H_{18}N_2O_6$	black powder	...	...	...	....	...	...	...	HBr salt copper colour	19
	Tarnine A	$C_{11}H_9NO_4+1\frac{1}{2}H_2O$	orange needles	220	...	...	sol. hot	sol. hot	insoluble	...	...	20
	Cuprine A	$C_{11}H_7NO_3$	copper colour	...	...	...	sol. green	sol. green	insoluble	...	...	21
	Papaverine A	$C_{20}H_{21}NO_4$	crystalline prisms	98*	feebly alk.	...	scarcely	50 cold	258	soluble	benzene, acetone, etc.	22
	Codeine A	$C_{18}H_{21}NO_3$	octahedra or prisms	115	alkaline	slightly bitter†	80 cold	readily	readily	10	amyl alcohol, benzene, etc.	23
	Thebaine A	$C_{19}H_{21}NO_3$	needles or prisms	193	alkaline	sharp	insoluble	10 cold	140 cold	18 cold	amyl alcohol, benzene	24

* Authorities differ. † Salts are bitter. ‡ Pure chloroform; easier if containing alcohol.

CAUSTIC ALKALIES	LEAD ACETATE.	TANNIC ACID.	PLATINUM CHLORIDE.	GOLD CHLORIDE.	SULPHURIC ACID, CONCENTRATED.	NITRIC ACID, CONCENTRATED.	SUNDRY PRECIPITANTS, ETC.
...	...	→	→	→	...	...	pp. by most alkaloid reagents
soluble	...	...	yellowish pp.	...	blue if containing HNO_3	...	...
...	no pp. neut.	...	...	...	...	...	...
pp. ammonia	...	pp.	pp.	pp.	...	...	mercuric chloride
...	pp. basic	pp.	pp.	pp.	...	...	not mercuric chloride
(no pp. ammonia)	no pp.	no pp.	no pp.	...	red-brown	...	*idem*
soluble	pp. basic*	...	...	...	rose colour⁓violet	(orange with nitric)	mercuric chloride, cloud
oily pp.	...	...	pp.	pp.	with bichromate, blood-red to violet	gradual red-violet	phospho-molybdic, Mayer's solution, etc.
...	no pp. neut.	→	→	→	on warming: brown, then greenish	...	pp. most alkaloid reagents
...	no pp. neut.	...	...	...	yellow-red⁓carmine	yellow	(Fröhde's solution, red colour)
...	pp.	...	...	...	(acids pp. from water)	...	pp. by Ag, Cu, and Hg salts
p. soluble excess	no pp.	no pp.	if not dilute	pp.	colourless ; violet with nitric	orange	Fröhde, magnificent violet
pp. greenish, soluble excess	no pp.	yellow-green	yellow	purple	(Fröhde's solution, violet)	violet-red	zinc chloride, ferric chloride, etc.
soluble	no pp.	...	...	...	pure acid, colourless ; otherwise blue	blood-red	(ferric chloride, blue or green colour)
pp. insoluble excess	no pp.	cloud	(pp. concentrated)	(pp. concentrated)	colourless, gradually yellow	yellow	phospho-molybdic, Mayer's solution, etc.
...	...	...	...	...	yellow ; red warm	...	...
pp. insoluble	...	pp.	yellow crys.	...	...	red	copper sulphate, ferrous salts
soluble greenish	...	...	...	...	...	...	(ferric chloride, red-brown colour)
oluble red-brown	...	...	...	...	fuchsine red ; bl.-violt. with water	...	(HBr salt difficultly soluble water)
...	...	...	...	...	...	...	(HBr salt soluble)
...	...	...	...	...	deep blue	deep blue	...
)p. and carbonates	no pp.	yellow	nearly white	dirty yellow	if pure, colourless ; otherwise blue-violet	yellow	not phospho-molybdic (characteristic)
pp. slightly soluble	no pp.	white	yellow†	brown	...	...	Mayer's solution, phospho-molybdic, etc.
pp. insol. excess	...	yellow	pp.	pp.	blood-red	yellow	bromine water, phospho-molybdic, etc.

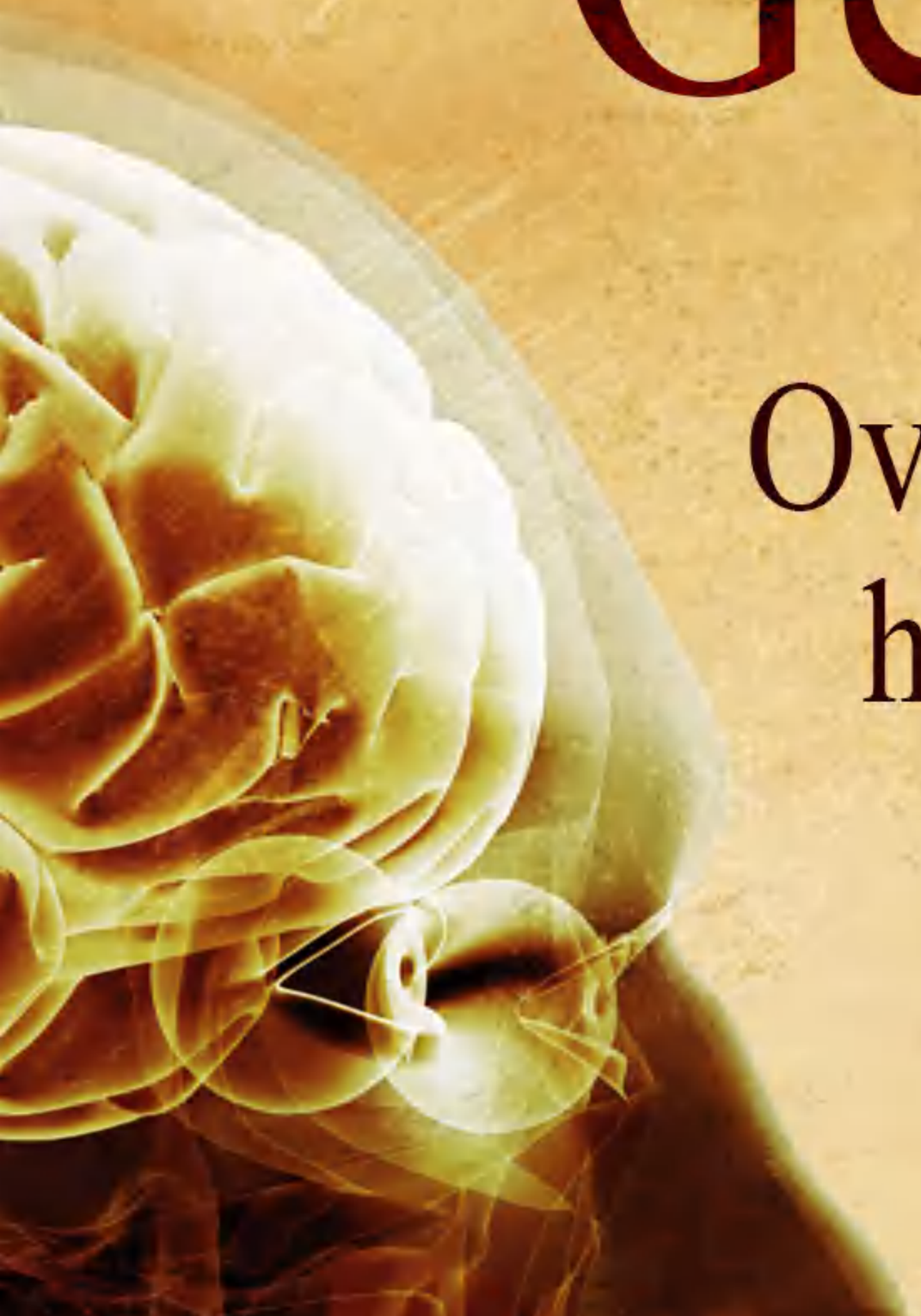
* Soluble in excess.

† Not 1 in 250.

	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt °C.	Reaction.	Taste.	Solvents. Water.	Alcohol.	Ether.	Chloroform.	Other Solvents.	
OPIUM	Thebaicine A	$C_{19}H_{21}NO_3$	amorphous	...	...	...	insoluble	diff. hot	insoluble	...	insoluble benzene	1
	Thebenine A	*idem*	amorphous	...	...	...	insoluble	*idem*	insoluble	...	*idem*	2
	Narceine A	$C_{23}H_{29}NO_9$	rhombic prismatic needles	92	neutral	feebly bitter	slightly	difficultly cold	scarcely	scarcely	insoluble benzene	3
	Rheadine A	$C_{21}H_{21}NO_6$	radiate prisms	232	neutral	tasteless	scarcely cold	scarcely	1,280	scarcely	scarcely benzene	4
	Oxynarcotine A	$C_{22}H_{23}NO_8$	minute crystals	blackens	...	...	difficultly	difficultly	insoluble	insoluble	insoluble benzene	5
	Gnoscopine A	$C_{34}H_{36}N_2O_{11}$	needles	233	neutral	...	insoluble	1,500 cold	insoluble	soluble	CS_2, difficultly benzene	6
	Codamine A	$C_{20}H_{25}NO_4$	6-sided prisms	121	alkaline	bitter	diff. cold	soluble	soluble	readily	benzene	7
	Meconidine A	$C_{21}H_{23}NO_4$	brown, amorphous	58	alkaline	tasteless	insoluble	soluble	soluble	soluble	benzene, acetone	8
	Lanthopine A	$C_{23}H_{25}NO_4$	microscopic prisms	browns	alkaline	tasteless	insoluble	difficultly	diff.	soluble	scarcely benzene	9
	Cryptopine A	$C_{21}H_{23}NO_5$	6-sided prisms	217	alkaline	bit. hot	slightly	1,200 cold	insol. *	soluble	*idem*	10
	Laudanine A	$C_{20}H_{25}NO_4$	radiate prisms	166	alkaline	bitter	...	diff. cold	647	readily	benzene	11
	Protopine A	$C_{20}H_{19}NO_5$	crystalline	204	alkaline	...	insoluble	difficultly	diff.	soluble	scarcely benzene	12
	Tritopine A	$C_{42}H_{54}N_2O_7$	prismatic crystals	182	...	...	...	...	slightly	easily	...	13
	Laudanosine A	$C_{21}H_{27}N_4O$	*idem*	89	alkaline	bitter	insoluble	readily	19	readily	benzene, petrol. ether	14
	Meconin B	$C_{10}H_{10}O_4$	hexagonal prisms	110	neutral	bitter	700 cold, 22 boiling	soluble	soluble	soluble	benzene, amyl alcohol. etc.	15
	Meconic Acid	$C_7H_4O_7$	prisms or scales	decomp. 120	...	...	difficultly	soluble	soluble	scarcely	amyl alcohol	16
	Meconoisin B	$C_8H_{10}O_2$	crystalline	88	...	...	scarcely	easily	soluble	...	...	17
	Opionin B	(N free)	needles	227	neutral	...	...	...	...	...	...	18
PANAX	Panaquilon B	$C_{24}H_{25}O_{18}$	yellow, amorphs.	fusible	...	b.-sweet	soluble	soluble	insoluble	...	...	19
PAPAYA	Carpaine A	...	crystalline	115	...	very bit.	scarcely	soluble	soluble	soluble	...	20
PAREIRA	Geissospermine A	$C_{19}H_{24}N_2O_2 \cdot H_2O$	small prisms	160	...	...	insoluble	sol. hot	insoluble	soluble	benzene	21
	Pareirine A	$C_{19}H_{24}N_2O$	gray, amorphous	124	...	...	difficultly	soluble	soluble	soluble	...	22
PARIS	Paridin G	$C_{16}H_{28}O_7 + Aq$	silky needles	...	neutral	pungent	soluble†	50 of 94 %	scarcely	...	...	23
	Paristyphnin G	$C_{38}H_{64}O_{18}$	yellowish amrphs.	...	...	bitter	soluble	soluble	insoluble	...	...	24
PARMELIA	Ceratophyllin	...	prismatic crystals	147	neutral	burning	soluble hot	soluble	soluble	...	...	25
	Physodin	...	micros. pillars	125	neutral	...	...	sol. hot	soluble	...	...	26

* Soluble if freshly precipitated, but redeposits.

† Difficultly; see Part I.



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	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents.				
							Water.	Alcohol.	Ether.	Chloroform.	Other Solvents.
Payta bark	Paytine A	$C_{21}H_{24}N_2O \cdot H_2O$	prismatic crystals	156	...	...	difficultly	soluble	soluble	soluble	benzene, petrol. ether
	Paytamine A	*idem*	amorphous	...	...	...	...	...	soluble	...	...
Petroselinum	Apiin G	$C_{27}H_{22}O_{13}$	fine needles	228	neutral	tasteless	soluble hot, jelly cold	390 cold	insoluble	...	...
Phillyrea	Phillyrin G	$C_{27}H_{34}O_{11} \cdot 1\frac{1}{2}H_2O$	silvery scales	160	...	grad. b.	1,300 cold	40 cold	scarcely	...	warm glacial acetic
	Philligenin	$C_{21}H_{24}O_6$	pearly crystals	fusible	...	...	scarcely	soluble	soluble	...	...
Phyllanthus	Phyllanthin B	$C_{30}H_{37}O_8$	needles	volat. 200	...	very bitter	slightly	soluble	soluble	soluble	benzene, petrol. ether
Physalis	Physalin B	$C_{14}H_{16}O_5$	amorphous	partly 190	...	bitter	difficultly	soluble	diff.	soluble	...
Pilocarpus	Pilocarpine A	$C_{11}H_{16}N_2O_2$	amorphs. or crys.	...	...	...	easily	soluble	soluble	soluble	benzene
	Pilocarpidine A	$C_{10}H_{14}N_2O_2$	deliquescent	...	...	...	soluble	soluble	diff.	readily	amyl alcohol
	Jaborine A	$C_{22}H_{32}N_4O_4$	amorphous	non-volat.	alkaline	...	soluble	soluble	soluble	...	..
Pimpinella	Pimpinellin B	(N free)	needles	37	...	...	insoluble	soluble	diff.	...	scarcely petroleum ether
Pinus, etc.	Coniferin G	$C_{16}H_{22}O_8 \cdot 2H_2O$	silky needles	185	...	...	196 cold	difficultly	insoluble	...	...
	Pinipicrin G	$C_{22}H_{36}O_{11}$	yellow crystals	100	...	bitter	soluble	soluble	insoluble absolute	...	...
	Thuyin G	$C_{20}H_{22}O_{12}$	yl. 4-sided plates	...	...	astringt.	soluble hot	soluble	...	...	...
	Thuyigenin	$C_{14}H_{12}O_7$	micros. needles	...	...	...	difficultly	soluble	...	...	...
	Thuyetin	$C_{14}H_{14}O_8$	yellow crystals	...	...	...	difficultly	soluble	soluble	...	...
Piper	Piperine A	$C_{17}H_{19}NO_3$	monoclinic prisms	100	neutral	almost tasteless	difficultly	30 cold	60 cold	soluble	benzene
	Piperidine A	$C_5H_{11}N$	liquid	B.P. 106	alkaline	burning	∞	∞	...	...	...
Piscidia	Piscidin B	$C_{29}H_{24}O_8$	prismatic crystals	...	...	...	insoluble	sol. hot	diff.	soluble	benzene
Plumeria	Agoniadin G	$C_{10}H_{14}O_6$	silky needles	155	...	very b.	scarce. cold	diff. cold	...	insoluble	carbon bisulphide
Podophyllum	Podophyllotoxin G	$C_{28}H_{24}O_9$	crystalline	93—95	slightly acid	...	7,000 cold	readily	diff.	soluble	acetone
Populus	Populin G	$C_{20}H_{22}O_8 + 2H_2O$	needles	180	...	sweet	2,000 cold, 79 boiling	100 cold	scarcely	soluble	amyl alcohol, acetic acid
	Chrysin	$C_{15}H_{10}O_4$	yellow crystals	275	...	...	...	180 cold	diff.	scarcely	aniline
Prunus, etc.	Amygdalin G	$C_{20}H_{27}NO_{11} + Aq.$	pearly crystals	120	neutral	slightly bitter	12 cold	148 cold *	insoluble	...	insoluble petroleum ether

* Alcohol of sp. gr. 0·939.

USTIC ALKALIES.	LEAD ACETATE.	TANNIC ACID.	PLATINUM CHLORIDE.	GOLD CHLORIDE.	SULPHURIC ACID, CONCENTRATED	NITRIC ACID, CONCENTRATED	SUNDRY PRECIPITANTS, ETC.
luble gradually	...	...	dark yellow	(reduced)	...	colourless~~red	potassic iodide
...	...	...	pp.	*idem*	...	...	(perchloric acid, fuchsine red)
soluble	...	...	...	...	(acids pp. from alkaline solution)	...	(ferric chloride, dark red colour)
insoluble	→	...	→	→	red-violet	decomposes	no pp. metallic salts
soluble	...	...	...	...	red	active decomposition	(soluble ammonia)
...	...	...	...	...	...	...	
luble ammonia)	pp. basic	...	...	...	...	...	...
...	...	...	→	pp.	colourless	...	pp. most alkaloid precipitants
	...	...	...	no pp.	...	...	phospho-molybdic acid
(pp. ammonia)	...	...	...	...	...	...	...
...	...	...	...	...	red	...	
...	no pp.	...	...	...	violet-blue	...	special test, see Part I.
...	no pp.	...	...	...	...		...
)l. yellow to red	yellow pp.	...	...	...	...	...	(ferric chloride, dark green colour)
...	pp. basic	...	...	...	...	...	(ammonia, blue-green colour)
...	red pp.	...	...	...	...	...	*idem*
...	...	...	(pp. concntd.)	...	yellow ; green in 20 hours	orange~~greenish	phospho-molybdic, ferric chloride, etc.
no pp.	...	...	(salt soluble)	...	...	...	reacts like ammonium hydrate
...	...	...	...	...	...	...	...
soluble brown	no pp.	pp. gradual	no pp.	no pp.	light yellow~~green	golden-yl.	(soluble in ammonia)
...	...	...	no pp.	...	cherry-red~~green-blue~~violet	...	(hydrochloric acid, red colour)
pp.	no pp.	...	...	...	red	soluble	no pp. other metallic salts
soluble yellow	partial pp.*	...	...	...	acids pp. from alkaline solution	...	calcium chloride, yellow crystalline pp.
...	...	...	...	...	light violet red	...	...

* Alcoholic solution.

	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents.					
							Water.	Alcohol.	Ether.	Chloro-form.	Other Solvents.	
PUNICA	Pelletierine A	$C_8H_{15}NO$	liquid	B.P. 195	alkaline	...	20 cold	∞	∝	∞	...	1
	Methyl-pelletierine A	$C_9H_{17}NO$	liquid	B.P. 215	...	...	25 cold	soluble	soluble	soluble	...	2
	Pseudo-pelletierine A	$C_9H_{15}NO{\cdot}2H_2O$	prismatic crystals	46	alkaline	...	soluble	soluble	9 cold	soluble	...	3
	Iso-pelletierine A	$C_8H_{13}NO$	liquid	B.P. 195	...	...	soluble	very sol.	very sol.	very sol.	...	4
PYRUS, ETC.	Phlorrhizin G	$C_{21}H_{24}O_{10}{\cdot}2H_2O$	needles	106—109	neutral	slight. b., then swt.	2,000 cold readily wm.	soluble	scarcely	...	methyl alcohol, acetic acid	5
	Phloretin	$C_{15}H_{14}O_5$	...	180	...	sweet	scarcely	soluble	scarcely	...	methyl alcl., hot acetic	6
QUASSIA	Quassiin B	$C_{10}H_{12}O_3$	pillars or plates	215—217	neutral	very b.	222 cold	soluble	diff.	soluble	benzene	7
QUEBRACHO COLORADO	Loxopterigine A	$C_{26}H_{34}N_2O_2$	amorphous	81	alkaline	bitter	difficultly	soluble	soluble	soluble	benzene, acetone	8
QUEBRACHO BARK	Aspidospermine A	$C_{22}H_{30}N_2O_2$	needles or prisms	205—206	feebly alk.	bitter	6,000 cold	48 of 99 %	106 absolute	soluble	benzene, amyl alcohol	9
	Aspidospermatine A	$C_{22}H_{28}N_2O_2$	needles	162	...	...	soluble*	soluble	soluble	soluble	hot petroleum ether	10
	Aspidosamine A	*idem*	crystalline	100	...	...	scarcely	soluble	soluble	soluble	benzene	11
	Quebrachine A	$C_{21}H_{26}N_2O_2$	crystalline	214—216	alkaline	bitter	scarcely	sol. hot	sol. hot	soluble	amyl alcohol	12
	Hypoquebrachine A	*idem*	yellow amorphous	...	alkaline	bitter	...	soluble	soluble	soluble	...	13
	Quebrachamine A	...	needles or leaflets	142	...	bitter	scarcely	difficultly	diff.	diff.	difficultly benzene	14
QUERCUS	Quercitrin G	$C_{36}H_{38}O_{20}$	yl. micros. plates	160—200	neutral	bitter	2,435 cold	23 cold	scarcely	insoluble	amyl alcohol	15
	Quercetin	$C_{24}H_{16}O_{11}$ (?)	yellow needles	250	neutral	bitter	scarcely	229 cold	soluble	...	warm acetic	16
RATANHIA	Ratanhine	$C_{10}H_{13}NO_3$	needles or leaflets	40—45	neutral	tasteless	9,480 cold	scarcely	scarcely	scarcely	benzene	17
RHAMNUS	Rhamnin G	$C_{23}H_{28}O_{14}$	microscopic needles or plates	...	neutral	nearly tasteless	soluble yellow	soluble	scarcely	scarcely	acetic acid	18
	Rhamnetin	$C_{11}H_{10}O_5$	yl. prismatic plts.	...	...	tasteless	scarcely	sol. hot	soluble	...	phenol	19
	Frangulin G	$C_{20}H_{20}O_{10}$	yellow needles	226	...	...	insoluble	diff. sol.	scarcely	diff.	hot acetic	20
	Emodin	$C_{15}H_{10}O_5$	micr. orange ndls.	above 250	...	...	insoluble	soluble	diff.	...	amyl alcohol	21
RHINANTHUS	Rhinanthin G	$C_{29}H_{52}O_{20}$ (?)	microscopic needles	...	neutral	bitter-sweet	soluble	soluble	insoluble	...	...	22

* If freshly precipitated.

	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents.					
							Water.	Alcohol.	Ether.	Chloro-form.	Other Solvents.	
Rhus	Fustin G	$C_{58}H_{46}O_{32}$	needles	218—219	...	...	soluble hot	easily	diff.	...	...	1
	Fisetin	$C_{15}H_{10}O_6$	yellow needles or prisms	browns at 270	...	...	scarcely	soluble	diff.	diff.	difficultly benzene	2
Robinia	Robinin G	$C_{25}H_{30}O_{16}\cdot5\frac{1}{2}H_2O$	pale yellow ndls.	195	neutral	astringt.	soluble hot	sol. hot	insoluble	...	...	3
Rottleria	Rottlerin B	$C_{22}H_{20}O_6$	yellow crystals	fusible	...	...	soluble	sol. hot	solubie	...	...	4
Rubus	Villosin G	...	...	173—175	...	bitter in solution	difficultly	soluble	scarcely	insoluble	...	5
Ruta	Rutin G	$C_{25}H_{28}O_{15}\cdot2\frac{1}{2}H_2O$	yellow needles	chars	...	b. insoln.	...	...	...	...	...	6
Saffron	Crocin G	$C_{44}H_{70}O_{28}$	red	...	...	sweet	soluble	difficultly	scarcely	...	(solutions red)	7
	Crocetin	$C_{34}H_{46}O_9$	red	...	...	...	scarcely	easily	easily	...	...	8
	Picrocrocin G	$C_{38}H_{66}O_{17}$	prisms	75	...	bitter*	easily	easily	diff.	soluble	...	9
Salix	Salicin G	$C_{13}H_{18}O_7$	plates, needles, etc.	above 100	neutral	bitter	29 cold	soluble	insoluble	...	amyl alcohol	10
	Saligenin	$C_7H_8O_2$	plates or rhombhd.	82	...	...	15 cold	readily	readily	...	...	11
	Helicin G	$C_{13}H_{16}O_7$	needles	175	neutral	slightly bitter	64 cold	soluble	insoluble	...	...	12
Sophora	Sophorine A	...	amorphous	...	alkaline	...	soluble	soluble	soluble	soluble	...	13
Saponaria	Saponin G	$C_{32}H_{54}O_{18}$	amorphous	...	neutral	sharp	sol. frothy	sol. hot	insoluble	soluble	...	14
	Sapogenin	$C_{14}H_{22}O_2$	needles	...	...	...	...	soluble	soluble	...	...	15
Sarsa-parilla	Smilacin G	$C_{16}H_{30}O_6+Aq$	needles	...	neutral	bitter in solution	sol. hot (frothy)	sol. hot	insoluble	soluble	amyl alcohol	16
	Glycyphyllin G	$C_{21}H_{24}O_9+3H_2O$	crystalline	175—180†	...	...	soluble hot	sol. hot	slightly	insoluble	(insoluble benzene)	17
Scilla	Scillain G	(N free)	amorphous	...	...	bitter	difficultly	soluble	diff.	diff.	...	18
	Scillin B	...	light yellow crys.	...	...	...	difficultly	soluble	sol. hot	...	...	19
	Scillipicrin B	...	yellow, amorphs.	...	...	bitter	soluble	...	...	...	...	20
	Scillitoxin	...	brown, amorphs.	...	...	...	insoluble	soluble	insoluble	...	...	21
Simaba	Cedrine A	...	needles	...	neutral	bitter	soluble	soluble	soluble	...	...	22
	Valdivin G ?	$C_{18}H_{24}O_{10}\cdot5H_2O$	hexagonal prisms	230	...	bitter	600 cold	190 absol.	insoluble	easily	...	23
Sinapis	Sinalbin G	$C_{30}H_{44}N_2S_2O_{16}$	pale yellow needles	fusible	...	...	soluble	insol. cold absol.‡	insoluble	...	(insoluble CS_2)	24
	Sinapine-thio-cyanide A	$C_{16}H_{23}NO_5\cdot HSCN$	prisms	130	neut.§	bitter	soluble yellow	soluble yellow	insoluble	...	*idem*	25
	Potass. Myronate G	$C_{10}H_{18}KNS_2O_{10}$	prisms	fusible	neutral	bitter	soluble	difficultly	insoluble	insoluble	(insoluble benzene)	26

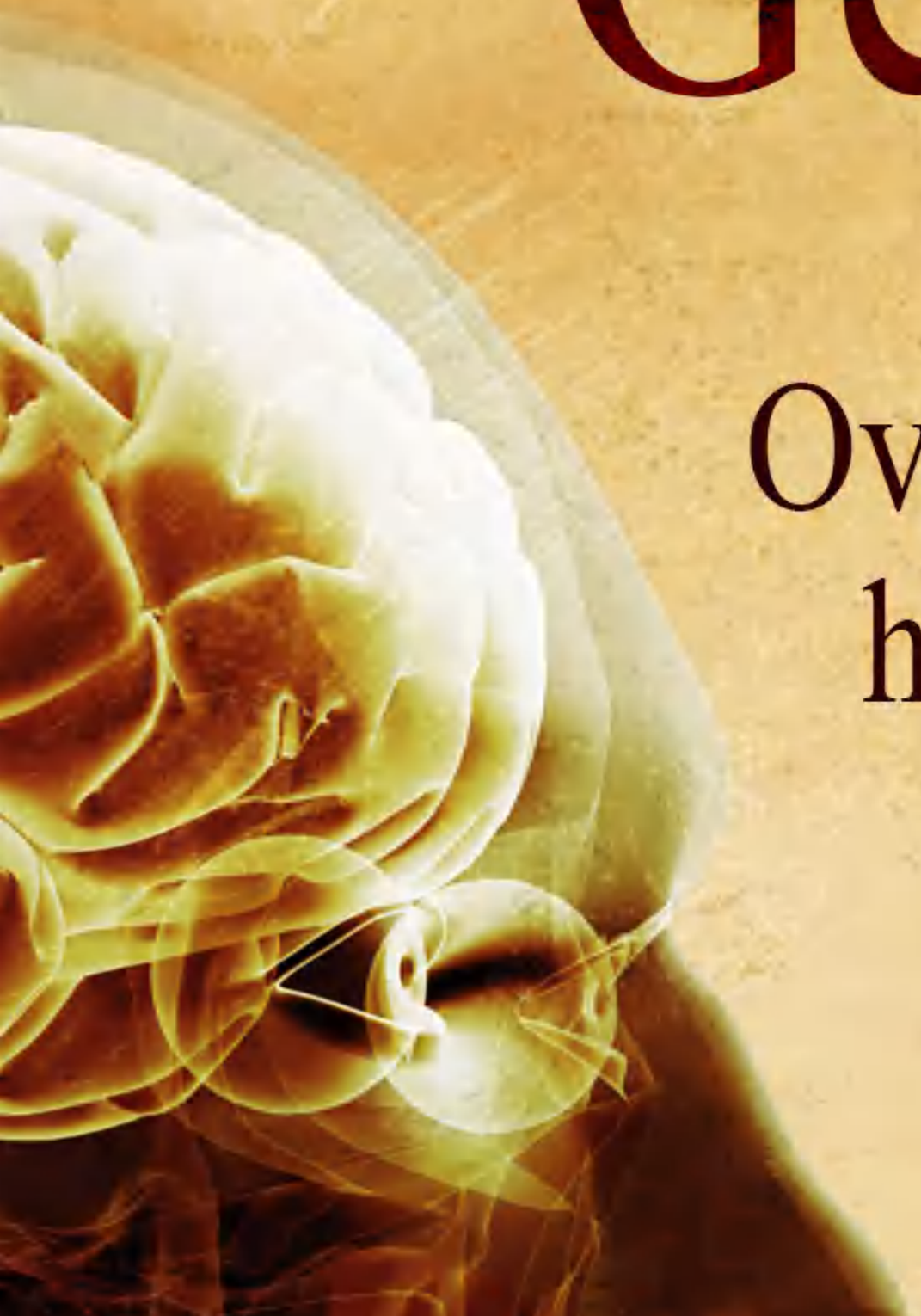
* Taste characteristic. † Decomposed at 110°-115°. ‡ Readily in hot 86 per cent. alcohol. § Free Sinapine is alkaline.

CAUSTIC ALKALIES.	LEAD ACETATE.	TANNIC ACID.	PLATINUM CHLORIDE.	GOLD CHLORIDE.	SULPHURIC ACID, CONCENTRATED.	NITRIC ACID, CONCENTRATED.	SUNDRY PRECIPITANTS, ETC.
soluble	pp. neutral	...	...	...	...	...	...
...	yl. pp.*	...	...	...	...	...	ferric chloride, green black
soluble, yellow	pp. basic	...	...	(reduced)	...	pic.ac.formed	Fehling reduced
deep red solution	no pp.*	...	...	...	...	...	..
gradual yellow colour	pp. basic	...	...	...	brown ; with water, deep violet	...	ferric chloride, no change
..	....	...	...	...	...	...	....
l. (decomposed)	red pp. warm†	...	...	...	deep blue~violet~red~brown	blue~brown	copper sulphate, green pp.
soluble red	pp. basic	...	...	...	*idem*	...	...
l.(decomposed)‡	...	...	...	...	(dilute acids give volatile oil, etc.)	...	Fehling reduced on warming
soluble	no pp.§	...	...	...	red (water precipitates Rutilin)	...	(Fröhde, magnificent violet)
soluble	no pp.	...	...	...	intense red	...	(ferric chloride, indigo-blue)
soluble	...	...	...	...	yellow	...	(no violet with Fröhde)
...	...	...	...	...	...	...	(ferric chloride, blood-red)
oluble frothy	pp.	pp.	...	no pp.	yellow-red to violet	...	Fehling reduced slowly
soluble	...	...	...	...	gradual violet	soluble	Fehling, slight reduction
soluble	pp. (alcoholic)	...	...	...	dark red ; violet warm	soluble	Fehling reduced
soluble	pp. basic	...	...	...	...	...	...
...	no pp.	pp.	...	...	brown (fluorescent green)	...	concentrated HCl, gradual green pp.
...	...	...	...	...	red-brown	yellow~gn.	...
...	...	...	...	...	...	...	(hygroscopic)
...	...	...	...	...	red~brown	yl. ; gn. warm	...
...	...	pp.	→	→	...	...	pp. by most alkaloid reagents
...	no pp.	pp.	...	...	colourless	colourless	Fehling reduced after decomposition
oluble yellow	...	...	...	...	...	transient red	Fehling reduced, some CuS formed
idem	...	...	...	...	...	...	(ferric chloride, blood-red warm)
...	pp. neutral	...	...	...	...	...	

pp. if concentrated.

	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents.				
							Water.	Alcohol.	Ether.	Chloro-form.	Other Solvents.
Solanum	Solanine A-G	$C_{42}H_{75}NO_{19}$	4-sided prisms	235	feebly alk.	bitter & hot	scarcely	soluble hot	4,000	...	amyl alcohol
	Solanidine A	$C_{25}H_{41}NO_2$	needles	208	alkaline	bitter & sharp	scarcely	soluble	soluble	soluble	benzene
	Jurubebine A	...	amorphous	...	...	bitter	difficultly	soluble	...	soluble	...
	Dulcamarin B	...	...	...	...	bitter	...	...	insoluble	insoluble	acetic ether
	Grandiflorine A	(Molec. Wt. 236·4)	amorphous	...	...	bitter	insoluble	soluble	...	...	...
Spartium	Sparteine A	$C_{15}H_{16}N_2$	liquid	B.P. 276*	alkaline	bitter	scarcely	soluble	soluble	soluble	petroleum ether
	Scoparin G	$C_{21}H_{22}O_{10}$	pale yellow crys.	...	...	...	soluble hot	soluble	...	...	...
Spigelia	Spigeline A	...	crystalline	volatile	alkaline	bitter	soluble	soluble	...	...	...
Stro-phanthus	Strophanthin G	$C_{29}H_{44}O_{14}$	crystalline or amorphous	172·5	faintly acid	very bitter	soluble	55 cold	insoluble	insoluble	acetone
Strych-nos, etc.	Strychnine A	$C_{21}H_{22}N_2O_2$	4-sided prisms	*	alkaline	*idem*	6,667 cold	120 cold	1,250	5 to 7	160 benzene
	Brucine A	$C_{23}H_{26}N_2O_4$	prisms, plates, etc.	178	alkaline	bitter	850 cold	readily hot	not absol.	about 2	64 benzene
	Curarine A	$C_{10}H_{15}N$	crystalline (from $CHCl_3$)	...	feebly alk.	bitter	∞	∞	insoluble	diff.	...
	Akazgine A	...	amorphous	fusible	alkaline	bitter	13,000 cold	60 absol.	120 abso.	soluble	benzene, CS_2
	Gelsemine A	$C_{22}H_{38}N_2O_4$	crystalline	45	...	bitter	difficultly	difficultly	soluble	soluble	*idem*
Syringa	Syringin G	$C_{17}H_{24}O_9+Aq$	needles	191—192	neutral	tasteless	soluble hot	sol. hot	insoluble	soluble	...
	Syringopicrin B	$C_{26}H_{24}O_{17}$	yellow, amorphs.	under 100	slightly acid	bitter	soluble	soluble	insoluble	...	...
Taraxa-cum	Taraxacin B	...	crystalline	fusible	neutral	bitter	soluble hot	...	soluble	...	...
Taxus	Taxine A	$C_{37}H_{52}NO_{10}$	amorphs. (crys. ?)	80—82	...	bitter	scarcely	readily	readily	soluble	carbon bisulphide
Thalic-trum	Macrocarpine A	C 58·36 %	yellow needles	blackens	neutral	...	insoluble	soluble	insoluble	...	amyl alcohol
	Thalictrine A	...	radiate needles	...	alkaline	...	scarcely	soluble	soluble	soluble	...
Thea	Theine A†	$C_8H_{10}N_4O_2+H_2O$	silky needles	*	feebly alk.	bitter	98 cold	97 cold	194 cold	soluble	benzene, amyl alcohol
	Theophylline A	$C_7H_8N_4O_2+H_2O$	crystalline	...	...	...	soluble	soluble	...	...	...
	Theobromine A	$C_7H_8N_4O_2$	micros. needles	not fus.	neutral	bitter	1,600 at 0°	1,400 cold	sparingly	diff.	...
	Camellin G	$C_{53}H_{84}O_{19}$	...	...	...	...	...	...	...	...	...
	Assamic Acid G	$C_{18}H_{29}O_{10}$	...	...	...	...	soluble	insol. (?) ‡	...	...	...

* Authorities differ. † Caffeine. ‡ Precipitates the aqueous solution as jelly.



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	Substance.	Formula.	Appearance or Crystalline Form.	Melting Pt. °C.	Reaction.	Taste.	Solvents. Water.	Alcohol.	Ether.	Chloroform.	Other Solvents.	
TRIANOSPERMA	Trianospermine A	...	needles	...	alkaline	sharp	soluble	soluble	insoluble	...	...	1
	Trianospermitine A	...	crystalline	...	...	tasteless	insoluble	difficultly	soluble	...	...	2
TRIGONELLA	Trigonelline A	...	flat prisms	decomp.	neutral	...	very soluble	soluble	insoluble	insoluble	insoluble benzene	3
TYLOPHORA	Tylophorine A	...	*idem*	...	alkaline	...	sparingly	...	soluble	...	...	4
URECHITIS	Urechitin G	$C_{28}H_{42}O_8+Aq$	needles or prisms	...	...	very bitter	...	soluble hot	soluble	easily	benzene, amyl alcohol	5
	Urechitoxin G	$C_{13}H_{20}O_5$	not readily crys.	...	...	...	difficultly	...	soluble	soluble	*idem*	6
USTILAGO	Ustilagine A	...	...	...	...	bitter	soluble	soluble	soluble	...	...	7
VALLARIA	Vellarin B	...	yellow oily	...	neutral	bitter	emulsion	soluble	soluble	...	oils	8
VARIOLARIA	Variolarin B	...	needles	...	...	...	insoluble	soluble	soluble	...	...	9
	Picrolichenin B	$C_{12}H_{20}O_6$	rhombohedra	110	acid	bitter	diff. hot	soluble	soluble	...	oils	10
VERATRUM	Veratrine A	$C_{32}H_{49}NO_{11}$	crys. or varnish	205	alkaline	burning	insol. cold	soluble	soluble	soluble	amyl alcohol, benzene	11
	Jervine A	$C_{26}H_{37}NO_3+2H_2O$	ndls. (from alcl.)	237—39	...	...	...slightly	soluble	diff.	soluble	1,625 benzene	12
	Sabadilline A	$C_{20}H_{26}N_2O_5$	crys. or varnish	...	alkaline	sharp	soluble	soluble	...	...	benzene	13
	Sabadine A	$C_{29}H_{51}NO_8$	crystalline	238—240	...	...	...	readily	readily	...	...	14
	Sabadinine A	$C_{27}H_{45}NO_8$	hair-like crystals	indefinite	...	...	soluble	soluble	diff.	soluble	...	15
	Veratroidine A	$C_{51}H_{78}N_2O_{16}$	crystalline	about 149	...	...	...	very soluble	readily	readily	benzene	16
	Pseudojervine A	$C_{29}H_{49}NO_{12}$	rhombic crys.	blackens	...	...	...	184 absol.	...	4	372 benzene	17
	Protoveratrine A	$C_{32}H_{51}NO_{11}$	microscopic 4-sided plates	245—250	...	...	...	difficultly hot	insoluble	diff. hot	insoluble benzene	18
	Protoveratridine A	$C_{26}H_{45}NO_8$	*idem*	265	...	bitter	...	scarcely	insoluble	scarcely	*idem*	19
VERNONIA	Vernonin G	...	amorphous	...	...	...	...	soluble	slightly	slightly	...	20
VICIA	Vicine A	$C_{28}H_{51}N_{11}O_{21}$?	needles	180	...	...	108	insoluble absolute	...	...	...	21
XANTHIUM	Xanthostrumarin G	...	...	...	...	...	...	soluble	soluble	...	...	22
XANTHOXYLON	Artarine A	$C_{21}H_{23}NO_4$	amorphous	240	...	bitter	insoluble	scarcely	soluble	diff. warm	(insoluble benzene)	23

	Caustic Alkalies.	Lead Acetate.	Tannic Acid.	Platinum Chloride.	Gold Chloride.	Sulphuric Acid, Concentrated.	Nitric Acid, Concentrated.	Sundry Precipitants, etc.
1	...	...	...	pp.	...	...	...	...
2	...	...	...	...	...	...	...	
3	...	...	→	→	→	red-brown⁓green⁓blue	purple-red	pp. usual reagents
4	...	...	...	no pp.	...	...	...	(ferric chloride, red colour)
5	...	...	...	...	...	yellow⁓red⁓mauve⁓purple	...	(oxidizing agents hasten colours with sulphuric acid)
6	...	pp. basic	...	...	...	*idem*	...	...
7	...	...	...	...	...	brown⁓intense green	...	(ferric chloride, dark yellow colour)
8	insoluble	...	...	...	...	(acids pp. from ammoniacal soln.)	...	(soluble in ammonia)
9	(no colour)	...	...	...	...	colourless	colourless	...
10	sol. gradual red	...	...	..	...	colourless	...	(soluble in ammonia, red)
11	pp. incomplete	...	pp.	not dilute	(salt M.P.182°)	yellow⁓carmine	yellow	iodo-potassic iodide, etc.
12	...	...	...	...	...	yellow to greenish-brown	no effect	(H_2SO_4 with sugar, blue)
13	no pp.	...	...	not dilute	not dilute	red	yellow	phospho-molybdic, Mayer's solution
14	soluble (pp. hot)	...	...	...	...	yellow⁓blood-red⁓violet	no colour	(soluble in ammonia, pp. on heating)
15	...	...	...	...	...	persistent blood-red	...	...
16	...	...	→	→	→	yellow⁓orange-red (fluorescent)	transient rose ⁓yellow	pp. most alkaloid reagents
17	...	...	...	...	...	no colours if pure	no colour	phospho-molybdic, Mayer's solution
18	pp.	...	no pp.	no pp.	pp.	gradual green⁓blue⁓violet	...	*idem*
19	(pp. ammonia)	...	pp.	no pp.	...	violet⁓cherry-red	...	picric acid, Mayer's solution
20	...	...	...	...	...	brown⁓purple	...	...
21	soluble	...	...	...	...	...	(violet after evaporation)	(dry NH_3, purple)
22	soluble yellow	pp. yellow	no pp.	...	...	...	...	ferric chloride, dark green pp.
23	...	...	...	...	...	colourless (blood-red with KNO_3)	...	Mayer's solution, etc.

PART III.

CLASSIFICATION OF REACTIONS, ETC.

1. REACTION WITH ALKALINITY TESTS.

The behaviour towards litmus is so easily seen on reference to column 6, in Part II., that another list here appears unnecessary.

Towards Phenolphthalein, it is interesting to note that nearly all alkaloids are neutral; solutions of their salts being indicated as if the acids were free. The following are exceptions, *i.e.*, *react alkaline to phenolphthalein:*

Atropine.
Coniine.
Homatropine.
Hyoscine.
Hyoscyamine.
Nicotine.

See Plugge, *Arch. Pharm.*, [3], **25**, 45.

2a. SOLUBILITY IN WATER.

The following substances are soluble in water, and may be regarded as exceptions, for the greater number of the A., G. and B. are insoluble. Those in italics are very soluble; ∞=miscible in all proportions; those in brackets are soluble with difficulty; df.c.=difficultly cold.

Alkaloids.

Acolyctine.
Aconine.
Anagyrine.
Arecaidine.
Arecaine.
Arecoline.
Aspidospermatine.
Atropine, 1 in 200.
[Bacharine, df.c.]
Berberine, 1 in 300.
Caffeine.
Calabarine.
Cedrine.
Chrysanthemine.
Codeine, 1 in 80.
[Colchiceine, df.c., sol. hot.]
Colchicine, very.
Conhydrine.
Coniine, 1 in 100.
Cuprine.
Curarine.
Cynapine.
Cytisine ∞.
Dibromapophylline.
Ecboline.
Ecgonine, very.
Echitamine.
[Ephedrine.*]
Ergotine.
Erythrophlœine.
Glaucopicrine.
Guachamanine.
Guvacine.
Heliotropine.
Hyoscyamine.
Isopelletierine.
Jaborine.
Laburnine.
Lobeliine.
[Lupanine.*]
Lupinidine.
Lycoctonine.
Macrocarpine (?).
Mandragorine.
Methylpelletierine.
[Morrenine.*]
Muscarine ∞.
Nicotine.
Pelletierine.
Piliganine.
Pilocarpidine.
Pilocarpine.
Piperidine ∞.
Piturine.
Pseudaconine.
Pseudopelletierine.
Sabadilline, very.
Sabadinine.
Sabatrine.
Sinapine thiocyanide (yellow).
Sophorine.
Spigeline.
[Staphisagrine, 1 in 200.]
Theine, 1 in 98.
Theophylline.
Trianospermine.
Trigonelline, very.
Ulexine, very.
[Valdivine, hot.]
[Violine, df.]

Glucosides.

Achillein.
Adonin.
Amygdalin.
Aphrodæscin.
Apocynein.
[Arbutin, df.c., sol. hot.]
[Asclepiadin, df.]
[Asebotin, df.]
Bergenin.
Bryonin.
Cathartic acid.
Chamælirin.
Colocynthin.
Coniferin, 1 in 196.
Convallamarin.
Convolvulin.
Coriamyrtin.
Crocin, reddish-yellow.
[Cyclamin.]
[Datiscin, df.c., sol. hot.]
Digitalein.
[Digitalin.*]
Digitonin ∞.
Echugin.
Fabianin.
[Fustin, sol. hot.]
Gentiopicrin.
Globularin.
Gratiosolin.
Helicin, 1 in 64.
Indican.
Loliin.
[Menyanthin, df.c, sol. hot.]
[Murrayin, df.c., sol. hot.]
Naringin, 1 in 300.
Paristyphnin.
Pinipicrin.
Propæscinic acid.
Rhamnin (yellow).
Rhinanthin.
[Robinin, df.c., sol. hot.]
Salicin.
Saponin.
Scoparin, greenish-yellow.
Sinalbin.
Sophorin.
Strophanthin.
Thevetin, 1 in 122.

Amaroids, etc.

Agoniapicrin.
Andirin.

* Authorities differ.

Surinamine.
Tarnine.
Thebaicine.
Thebenine.
Theobromine, 1 in 17,000.
Trianospermine.
Trigonelline.
Tritopine, sp.
Ulexine.
Valdivine.
Xanthopuccine.

Amaroids, etc., insoluble or difficultly soluble in Ether.

Aloin, very df.
Andromedotoxin.
Angosturin.
Anemonin, sc.
Calumbin, df.
Cantharidin, 1 in 900.
[Carapin, not easily.]
Carobin.
Chrysin.
Cnicin, sc.
Coccognin.
Dulcamarin.
Elaterid.
Elaterin, 1 in 290.
Gentiopicrin.
Glycyphyllin.
Hymenodictyon, amaroid from.
Ilixanthin.
Karakin.
Lactucin.
Lygustron.
Melonemetin.
Mudarin.
[Oreoselon, df.]
Panaquilon.
[Physalin, df.]
Picrocrocin.
Pimpinellin.
Piscidin.
[Podophyllotoxin, df.]
[Quassiin, df.]
Scillotoxin.
Sennapicrin.
Syringopicrin.
Tanacetin.
Tulucunnin.
Vaccinin.

Glucosides soluble in Ether.

Acorin.
[Chinovin, df.]
[Coriamyrtin ?]
[Digitalin.*]
Diosmin.
Fabianin.
Fabiin.
[Indican.*]
Jalapin.
Sophorin.
Tampicin.
Urechitin.

3. **ALKALINE HYDRATES** (Caustic Soda and Potash).—Alkaloids are usually precipitated, many being soluble in excess of the reagent; the majority of the Glucosides and Amaroids are soluble.

Substances soluble in water are not precipitated by caustic alkalies, except perhaps from concentrated solutions.

cb. = pp. also by carbonates, sol. xs. = soluble in excess of caustic alkalies, sc. = scarcely, df. = difficultly, crys. = crystalline.

I. Alkaloids precipitated.

Aconitine, cb., sc. sol. xs.
Akazgine, cb.
Alpha-oxy-cinchonine.
Alstonidine.
Alstonine.
Anthocercine, cb.
Apocinchonicine.
Apocinchonidine.
Apocinchonine.
Apoquinidine, df. sol. xs.
Apoquinine, sol. xs.
Aribine, cb.
Aricine, cb.
Aristolochine.
Aspidospermine.
Atherospermine, cb.
[Atropine, if concd., sol. xs.]
Bebeerine, cb., df. sol. xs.
Beta-oxy-cinchonine.
Boldine, sol. xs.
Brucine, cb.
Cannabinine.
Carpaine.
Chairamidine, cb.
Chairamine, insol. xs.
Chelerythrine.
Cinchonibine.
Cinchonidine, cb.
Cinchonifine.
Cinchonigine.
Cinchoniline.
Cinchonine.
Cocaine, insol. xs.
Cocamine.
Codamine, sol. xs.
Codeine, partly; not cb.
[Colchicine.]
Conchairamidine, cb.
Conchairamine, insol. xs.
Concusconine.
Corydaline, sol. xs.
Cotarnine, sc. sol. xs.
Cryptopine, insol. xs.
Cupreine, sol. xs.
Cuscamine.
Damascenine, cb., oily.
Delphinine, cb., insol. xs.
Diapocinchonine.
[Echitenine, if concd.]
Emetine, cb.
Erythrophlœine.
Fumarine.
Geissospermine, cb.
Gelsemine, sol. xs.
Glaucine.
Glaucopicrine.
Gnoscopine, insol. xs.
Harmaline, cb.
Harmine.
Heliotropine, oily
Hydrastine.
Hydroquinine.
Hymenodictyonine, gelat.
Hyoscyamine, as Atropine.
Hypoquebrachine.
Igasurine, sol. xs.
Imperialine, cb.
Isocinchonine.
Isopelletierine.
Laudanine, sol. xs.
Laudanosine (?).
Laurotetanine, sol. xs.
Lupinine.
[Lycopodine, if coned.]
Menispermine.
Methylpelletierine
Morphine, sol. xs.
Narceine.
Narcotine, insol. xs.
Oxynarcotine.
Papaveramine, sc. sol. xs.
Papaverine.
Paytine, df. sol. xs.
Pelletierine.
[Physostigmine, if concd.]
Piliganine.
Pithecolobine, oily.
Protopine, insol. xs.
Protoveratrine.
Pseudopelletierine.
Quebrachine, cb.
Quinamicine.
Quinicine, partly.
Quinine, cb.
Ratanhine, sol. xs.
Rheadine, df. sol. xs.
Solanidine.
Solanine.
Sparteine, sol. xs.
Strychnine, crys.
Taxine.
Thebaicine, sc. sol. xs.
Thebaine, insol. xs.
Theobromine, sol. xs.
Tritopine, insol. xs.
Ulexine, sol. xs.
Veratralbine.
Veratrine, cb.

II. Glucosides, Amaroids, etc., insoluble in Caustic Alkalies.

Aristolochin.
Derrid.
Euphorbon.
Gardenin.
Hurin.
Ilixanthin.
Laserpitin.
Liriodendrin.

* Authorities differ.

Phillyrin.
Populin.
Sapotiu.
Sikimin.
Vellarin.

III. Alkaloids not precipitated or soluble.

(See also I. for some alkaloids that are soluble in excess.)

Aconine.
Adansonine.
Apomorphine.
Arecoline.
[Atropine, see I.]
Benzoyl-ecgonine.
Chrysanthemine.
Colchiceine.
[Colchicine.]
Conhydrine.
Coniine.
Cupronine.
Cytisine.
Ecgonine.
Eupatorine.
[Gelsemine, see I.]
[Hyoscyamine, see I.]
Lanthopine.
[Lycopodine, see I.]
Meconidine.
[Morphine, see I.]
Nandinine.
Nartinic Acid.
Nicotine.
Physostigmine.
Picraconitine, pp. hot
Pseudomorphine.
Sabadilline.
Sabadine.
Sabadinine.
Sabatrine.
Sinapine thiocyanide.
[Sparteine, see I.]
Surinamine.
[Thebenine, see I.]
Theine.
Theobromine.
Ulexine.
Vicine.

IV. Glucosides, Amaroids, etc., soluble in Caustic Alkalies.

(y. = yellow, r. = red, pl. = purple, bn. = brown, gn. = green.)

Absynthiin.
Æsculetin, y.
Æsculin.
[Agoniadin, sol. hot.]
Aloin, y.
Anemonin.
Angosturin.
Antiarin.
Apiin.
Araliin.
Arnicin.
Baphiin (?).
Brucamarin.
Caincin.
Calendulin.
Calumbin.
Cantharidin.
Capsicin.
Cathartic Acid.
Cephalanthin.
[Ceratophyllin, sol. hot.]
Chicorin, y.
Chrysin, y.
Cicutoxin.
Convolvulin.
Crocin, y.
Daphnetin, r.
Daphnin, y.
Datiscin, deep y.
Diosmin, y.
Dulcamarin, r.-bn.
Ecbalin.
Elaterid.
Elaterin.
Emodin, r.
Erythrocentaurin.
Fabianin, deep y.
Fabiin.
Fisetin.
Frangulin, pl.
Fraxin, y.
Fustin.
Gentiopicrin, y.
Geum bitter.
Glycyphyllin.
Glycyrrhizin, r.-y.
Helixin, gn.
Helleborein.
Hesperidin.
Hop resins.
Indican, decomp.
Jalapin.
Karakin.
Kosin.
Lactucin, r.
Laserol, y.
Lignoin, bn.
Limonin.
Linin, y.
Lupinin.
Mangostin.
Meconic Acid.
Meconin.
Melanthin.
Murrayin.
Ophelic Acid.
Opionin.
Panaquilon, bn.
Paristyphnin, y.
Phillygenin.
Phloretin.
Phlorrhizin.
Physodin, y. to r.
Picrocrocin.
Picrolichenin, r.
Quercetin.
Quercitrin.
Quinovaic Acid.
Quinovin.
Rhamnetin, y.
Rhinanthin.
Robinin, y.
Rottlerin, deep r.
Rutin, y.-bn.
Salicin.

Santonin.
Sapogenin.
Saponin.
Scoparin, y.-bn.
Sennapicrin.
Sinalbin, y.
Smilacin.
Strophanthin, y.
Thuyin, y. to bn.-r.
Variolarin.
Villosin, y.
Xanthostrumarin, deep y.

4. LEAD ACETATE gives usually no precipitates with alkaloids, bitter principles, or glucosides; the following are exceptions, *i.e.*, *are precipitated.*

nl. = neutral, bc. = basic, aml. = ammoniacal, alcc. = alcoholic lead acetate solution.

I. Alkaloids.

Acolyctine.
Aconine, sol. in excess.
Nicotine. The free alkaloid gives pp. of $Pb(OH)_2$.
Sparteine.
Theobromine, bc.

II. Glucosides or their derivatives.

Æscinic acid, nl.
Æsculetin.
Æsculetin hydrate, pp., nl. or bc.
Aphrodæscin, bc.
Assamic acid, nl. or bc.
Cephalanthin.
Crocin, red, bc.
Daphnin.
Datiscin, yellow.
Digitonin.
Dulcamarin, bc.
Fraxin, yellow, aml.
Indican, bc.
Lupinin, bc.
Myronic acid, yellow.
Phloretin.
Phlorrhizin.
Rutin.
Saponin.
Sapotin, sol. in excess.
Smilacin.
Thuyigenin, bc.
Thuyin.
Villosin, bc.

III. Bitter Principles, etc.

Araliin (Glucoside ?).
Baphiin.
Calendulin, alcc.
Calumbic acid.
Cantharidin.
Catechin.
[Cnicin — authorities differ.]
Gentiopicrin, aml.
Glycyphyllin, bc.
Glycyrrhizin, aml.
Hop resins (entirely or nearly so).
Ilexanthin, yellow.
Kosin.
Laserpitin.
Lignoin.
Mangostin.
Ononid.
Ophelic acid.
Scoparin.

TANNIC ACID (5*a*).]

5a. TANNIC ACID gives precipitates with the great majority of alkaloids, bitter principles and glucosides; the following, however, are *not precipitated*.

I. Alkaloids.

Arginine.
Aribine.
Arecaine, turbidity.
Arecoline.
Chrysanthemine.
Grandiflorine, cloud.
Theine, cloud.
Theobromine, cloud.

II. Glucosides.

Antiarin.
Chamælirin.
Chirettogenin.
Datiscin.
Ericolin (or slightly).
Frangulin.
Strophanthin, pp. if concentrated.

III. Bitter Principles.

Angelin.
Calumbin.
Cascarillin.
Capsicin.
Cornin.
[Erythrocentaurin, authorities differ.]
Marrubiin.
Ononid.
Picrotoxin.

5b. TANNIC ACID (Tannin), second list. The following give a pp.:

I. Alkaloids.

Aconine.
Acolyctine.
Aconitine.*
Alpha-homo-chelidonine.
Amarylline (?).
Anagyrine (?).
Anthocercine.
Apomorphine.
Aricine.
Asiminine (?).
Aspidospermine.
Atherospermine.
Atisine.
Atropine, sol. HCl.
Bebeerine, sol. warm HCl.
Belladonnine.
[Berberine, cloud 1 in 1,000.*]
Brucine, sol. HCl.
Calamine.
Cannabine.
Cannabinine (?).
Carpaine.
Ceanothine.
Cedrine.
Chelerythrine.
Chelidonine.
Cinchonine, sol. warm HCl.*
Cocaine.
[Codeine, cloud 1 in 3,000, sol. HCl.]
Colchicine.*
Conessine.
Coniine, sol. warm HCl.
Corydaline.
Cotarnine.
Curarine, sol. HCl.
[Cytisine, 1 in 300.]
Damascenine.
Delphinine.
Ecboline.
Emetine.
Ergotine.
Ergotinine.
Erythrine.
Glaucine.
Glaucopicrine.
Guachamanine.
Hyoscyamine.
Igasurine.
Imperialine.
[Ineine (?).]
Laurotetanine.
Lobeline.
Lupanine.
Lycoctonine.
Morphine.
Morrenine.
Muscarine.
[Narceine, cloud 1 in 3,000.]
Narcotine.*
[Nicotine, cloud 1 in 500, sol. warm HCl.]
Nupharine.
Oleandrine.
Oxyacanthine.
Papaïn.
Papaverine, sol. *warm* HCl.*
Pelletierine.
Physostigmine, red pp., sol. HCl.
Piliganine.
Piturine, 1 in 100.
Protoveratridine.
Pseudaconitine.
Quinidine, sol. warm HCl.
Quinine, sol. wm. HCl.
Rhæadine.
Sapotine.
Scopoline.
Solanine.*
Strychnine, sol. HCl.
Thalictrine.
Thebaine, sol. *warm* HCl.*
Trigonelline.
Tylophorine.
Veratrine, not *at once*.*
Veratroidine (?).

II. Glucosides.

Acorin.
Agoniadin, gradually.
Asclepiadin.
Bryonin.
Colocynthin.
Digitalein.
Digitalin.
Digitonin.
Euonymin, slight.
Helleborein.
Limonin.

[PICRIC ACID (6).

Paristyphnin.
Rosaginin.
Saponin, cloud.
Sapotin.
Solanine, sol. warm not cold HCl.

III. Bitter Principles, etc.

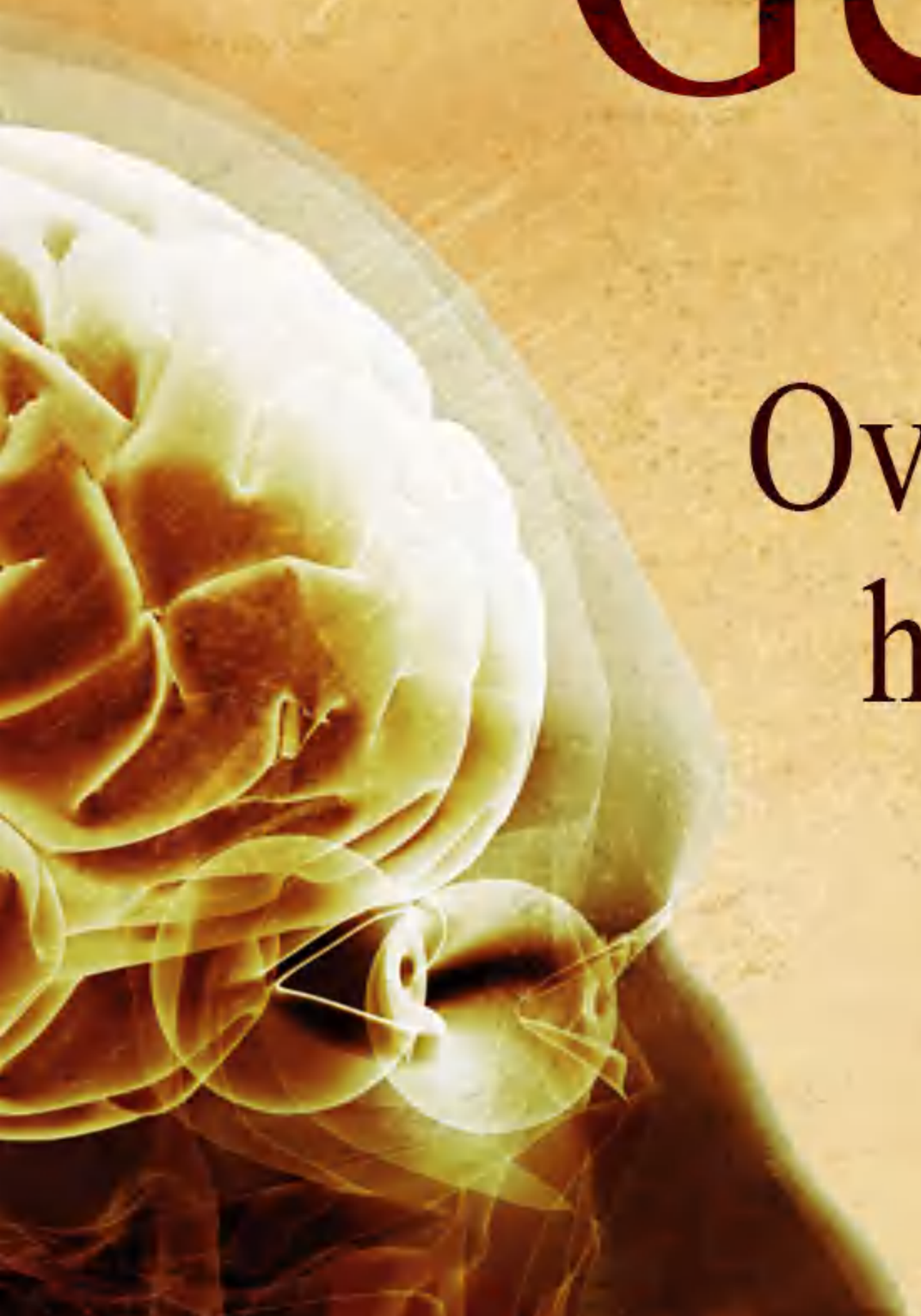
Absynthiin.
Agoniapicrin.
Angosturin.
Arnicin.
Cailcedrin.
Chirettin.
Copalchin.
Dulcamarin.
[Erythrocentaurin, authorities differ.]
Menyanthin.
Panaquilon.
Quassiin.
Santonin.
Scrophularin.
Syrincopicrin.

6. PICRIC ACID.

I. The following Alkaloids give precipitates.

Amarylline.
Akazgine.
Anthocercine.
[Antypyrine.]
Apomorphine.
Arecoline.
Arganine.
Aspidospermine.
Atherospermine.
Bebeerine.
Bellamarine.
Berberine, crys.
Brucine, 1 in 10,000 crys.
Calcitrapine.
Cannabinine.
Carpaine.
Chærophylline.
Cinchonine, 1 in 100,000 crys.
Cocaine.
Cocamine.
[Codeine, 1 in 250.]

* Insoluble or difficultly soluble in cold dilute hydrochloric acid.



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Rhinanthin, no colour.
Sinalbin, no colour till boiled with alkali.
Villosin.

Amaroids and Glucoside derivatives.
Cailcedrin.
Caincin.
Capsicin, see III.
Cornin.
Erythrocentaurin.
Opionin.
Picrotoxin.
Syringopicrin.

8. PLATINUM CHLORIDE forms with hydrochlorides of the alkaloids double salts, which are mostly soluble with difficulty in water; others, however, are readily soluble, and are consequently not precipitated, or are only deposited from concentrated solutions. The formulæ of these double salts are usually $(NR{\cdot}HCl)_2\,PtCl_4$, or $N_2R{\cdot}(HCl)_2{\cdot}PtCl_4$, so that the ratio of nitrogen to platinum is N_2 : Pt, but there are a few exceptions to this rule.

Platinum chloride is reduced by some substances, and occasionally gives colour effects.

I. Alkaloids precipitated or causing reduction (pp. unless otherwise stated).

Abrotine, df. sol.
[Aconitine, concentrated only.]
Akazgine.
Apomorphine, y.
Aricine.
Aspidosamine.
[Aspidospermine, blue solution, or y. pp.]
Atherospermine, y.
[Atropine, concd. only]
Bebeerine, y., insol. HCl.
Bellamarine, y.
Berbamine, white crys.
Berberine.
Brucine, y.
Calamine, reduction.
Cannabine.
Ceanothine.
Chairamidine, y. flocct.
Chairamine, needles.
Chenopodine.
Cinchonidine, sc. sol.
Cinchonine, 1 in 500.
Cocaine, peculiar crys.
Cocamine.
Codeine, y.
Conchairamine, dark y.
Concusconine, y.
Conessine.
Corydaline, y. crys.
Cotarnine, y. crys.
Cryptopine, sol. warm.
Cupreine, df. sol. (dark gn. on heating the dry salt).
Curarine, y.
[Cytisine, strong soln.]
Damascenine.
Delphinine.
Ditamine, crys.
Ecboline, y.
Echitenine, y.
Emetine, whitish.
Ergotine.
Ergotinine.
Erythrine.
Geissospermine, y.
Gelsemine, y.
Gnoscopine, pale y.
Grandiflorine, y.
Harmaline, crys.
Harmine, y., grad. crys.
Hydrastine, y.-r.
Hymenodictyonine.
[Hyoscyamine, concd. only.]
Hypoquebrachine, y. then r.
Igasurine, y.
Laudanosine, y.
Laurotetanine.
Lobeliine.
Mandragorine.
Meconidine, y.~red.
Morphine, 1 in 100, or after 24 hours 1 in 3,000.
Narceine, y. crys.
[Narcotine, strong solution.]
[Nicotine, sol. HCl.]
Oleandrine.
[Oxyacanthine, y., sol. HCl.]
Papaverine, nearly white.
Paytamine.
Paytine, dark y.
[Physostigmine, not 1 in 250.]
[Piperine, coned. only.]
[Piturine, not 1 in 100.]
Porphyrine.
[Pseudaconitine, concentrated only.]
Quinidine, nearly white.
Quinine, nearly white.
Rheadine, y.
Sapotine, y.
[Scopoline, concd., y.]
Sparteine, y.
Strychnine, crys.
Taxine.
Thebaine.
[Theine, concd. only.]
Theobromine, slowly.
Trianospermine.
Ulexine, y.
[Veratrine, not very dilute.]

II. Non-basic substances precipitated or causing reduction.

Acorin, reduction.
Agoniapicrin.
Aloin, colour effects, see Part I.
Arnicin, pp. (alcoholic).
Bryonin, pp.
Catechin, reduction.
Mangostin.
Papaïn (ferment).

III. Alkaloids not precipitated, or only in concentrated solutions.

[Aconitine, pp. concd.]
Amarylline.
Angeline.
Arecoline.
[Atropine, pp. coned.]
Calcitrapine, cloud.
Chrysanthemine.
[Coniine, pp. concd.]
[Cytisine, pp. concd.]
Ecgonine.
Ergotine (pp. in ether + alcohol).
[Hyoscyamine, pp. concd.]
Jurubebine.
Lycoctonine.
[Morphine, see list I.]
[Narcotine, „ „]
[Nicotine, „ „]
Pelletierine.
[Physostigmine, see I.]
Picraconitine.
Piliganine.
[Piperine, pp. coned.]
[Piturine, not 1 in 100.]
Protoveratridine.
Protoveratrine.
[Pseudaconitine, pp. concd.]
Ratanhine (Angeline).
[Sabadilline, see I.]
[Sabatrine, see I.]
[Theine. pp. concd.]
Trigonelline.

IV. Non-basic substances giving no pp.

Araliin.
Cailcedrin.
Californin.
Coriamyrtin.
Erythrocentaurin.
Helleborein.
Panaquilon.
Podophyllotoxin.
Picrotoxin.
Rosaginin.
Strophanthin.

9. GOLD CHLORIDE forms with hydrochlorides of alkaloids compounds which are mostly insoluble or difficultly soluble in water; the double salt has usually the formula $R{\cdot}HCl{\cdot}AuCl_3$. Reduction of the gold salt very

frequently occurs on standing, or even immediately. Glucosides and other substances also cause reduction.

I. Alkaloids giving pp. or causing reduction.

(*Pp. understood unless otherwise mentioned.*)

Acolyctine.
Aconine, y. and reduction.
Aconitine, y. (M.P. 35°).
Akazgine.
Apomorphine, purple.
Arecoline, y., oily, df. sol.
Aricine, y., amorphous.
Aspidosamine.
Atherospermine, y.
Atropine, y.
Bebeerine, light y.
Benzoyl-ecgonine, sol. hot alcohol.
Berberine, orange.
Brucine, pp. sol. HCl; dirty y., amorphous.
Calcitrapine, y.
Carpaine, limit 25,000.
Chelerythrine, dark r. (M.P. 233°).
Chelidonine, dark r.
Chrysanthemine, y. crys.; sol. hot.
Cinchonidine, y. (M.P. about 100°).
Cinchonine, limit 200,000.
Cocaine, light y. fern-like crys.
Cocamine.
Codeine, bn. (not 1 in 1,000).
Colchiceine.
Colchicine.
Colloturine, y., flocculent.
Concusconine, dirty y. and reduction.
Conessine.
Coniine.
Conquinamine, y. to purple.
Corydaline, y. crys.
Cytisine.
Damascenine.
Daturine, y.
Delphinine, y.
Ditamine, sol. hot; crys. cold.
Ecboline.
Ecgonine, y., amorphous.
Emetine, y., amorphous.
Ergotine.
Ergotinine.
Erythrine.
[Eupatorine, coloration only.]
Geissospermine, deep red colour.
Gelsemine, y., sol. hot.
Hydrastine, reddish y.
Hyoscine.
Hyoseyamino, whitish-yellow.
Hypoquebrachine, y., then violet.
Laurotetanine.
Lobeliine, insol. HCl.
Lupanine, y., sol. warm, crys. cold.
Mandragorine, sol. hot (M.P. 153°-155°.)
Meconidine, dirty y., amorphous.
Morphine, at once, 1 in 5,000.
Narceine, 1 in 5,000.
[Narcotine, only strong solutions.]
Nicotine, cloud 1 in 10,000.
Oleandrine.
Papaverine, dirty y., distinct at 5,000.
[Paytamine, reduction.]
[Paytine, reduction.]
Pelletierine, pp.; reduction on warming.
Physostigmine, pp. and reduction, blue coloration.
Picraconitine, very insol.
Piliganine, y., altered by light.
Pilocarpine.
Piturine, pale red; 1 in 100.
Porphyrine.
Protoveratrine.
Pseudaconitine.
Pseudopelletierine.
Quebrachine, y., amorphous.
Quinamicine.
Quinamidine, purple.
Quinamine, y. to r. and reduction.
Quinicine, y., amorph.
Quinidine, y.
Quinine, y., amorphous.
Rheadine, y., flocculent.
Sabadilline, not very dilute.
Sabatrine, not very dilute.
Scopoline, y.
Solanine.
Strychnine, pp. sol. HCl.
Taxine (?).
Thebaine, y.-bn.; limit 10,000.

[Theine, slowly crys.]
[Theobromine, slowly crys.]
[Veratrine, not very dilute; M.P. 182°.]

II. Non-basic substances causing reduction.

(*Pp. only where stated.*)

Absynthiin, pp. cold; reduction on warming.
Acorin.
Adonin.
Aloin, red, then violet.
Catechin, red-brown.
Digitalin, crys. pp.
[Ericolin, reduction warm.]
Mangostin.
[Menyanthin, reduction warm.]
Quercetin.
Robinin.
Sophorin, y. crys. pp.

III. No reaction with following:

(A. = Alkaloid.
G. = Glucoside.
B. = Bitter Principle.)

Capsicin, B., no pp.
Cnicin, B.
[Menyanthin, G., cold; see II.]
[Narcotine, A., see I.]
[Picrotoxin, B., slight reduction on warming.]
Pilocarpidine, A.
Quassiin, B.
Saponin, G.
Strophanthin, B.
Theine, A. } not at once; see I.
Theobromine }

10. SILVER NITRATE is reduced by the following, or gives pp. (where this is mentioned):

I.

Alkaloids.

Aconine.
Anthocercine, slight pp.
Capsaicin, purple on heating.
Capsicin, pp. in strong alcoholic solution.
Capsicol, pp. soluble in ammonia.
Cupronine.
Dibromapophylline.
Lobeliine, pp. soluble in ammonia and HNO_3.
Macrocarpine, pp. = pure Macroc.
Morphine, red colour.
Oxyacanthine, white pp.
Pseudaconine (ammoncl. $AgNO_3$).
Quinidine, pp.
Theobromine, pp. gradually crystallizes.

SILVER NITRATE (10.)]

Glucosides.

Antiarin (ammoniacal $AgNO_3$).
Chicorin.
Convolvulin, pp. (?).
Cyclamin.
[Daphnin, slight reduction on warming.]
Gentiopicrin (ammoniacal $AgNO_3$).
Menyanthin „ „
Quercetrin.
Rhamnin.
Robinin, slow reduction.
[Saponine, slow reduction on warming.]
Sinalbin, white pp., then black pp. of Ag_2S.
Solanine.
Strophanthin.
Xanthostrumarin, slight reduction on warming, or at once with ammoniacal $AgNO_3$.

Bitter principles and Glucoside derivatives.

Æsculetin.
Arnicin (alcoholic).
Cantharidin, white crys. pp.
Catechin, gn. pp. ~ violet-black (ammoniacal $AgNO_3$).
Cornin, white crys.
Cyclamiretin.
Daphnetin.
Lactucin (ammoniacal $AgNO_3$).
Lygustron „ „
Ononid, pp.
Ophelic Acid.
Quercetin.
Rhamnetin.
Sapogenin (ammoniacal $AgNO_3$).

II. No reaction with following :

Aloin.
Ceratophyllin.
Cnicin.
Ericolin.
Erythrocentaurin.
Picrotoxin.
Quassiin.
Scoparin.
Viliosiu.

11. FEHLING'S SOLUTION (alkaline copper tartrate) is reduced by the following :

Alkaloids.

Aconine, on boiling.
Aspidospermine.
Ceanothine, on boiling.
[Echitamine, after treatment with acid].
[Lobeliine, after treatment with acid.]
[Valdivine, after treatment with alkali.]

Glucosides.

Acorin.
Bergenin.
Camellin.
Cephalanthin, after boiling.
Chamælirin, after boiling.
Chicorin.
Colocynthin.
Coriamyrtin.
Cyclamin, white precipitate.
[Helixin, after inversion.]
Indican.
[Melanthin, after inversion.]
Menyanthin.
[Phillyrin, after inversion.]
Picrocrocin, on warming.
Quercitrin, on boiling.
Robinin.
Saponin, slowly.
Sinalbin (copper sulphide also formed).
Smilacin.
[Villosin, after inversion.]

[POTASSIUM FERROCYANIDE (12).

[Xanthostrumarin, after inversion.]

Bitter Principles and Glucoside derivatives.

Æsculetin.
Cyclamiretin, pp.
Daphnetin.
Fisetin (on warming).
Karakin, gn. pp. or reduction.
Kellin.
Lactucin.
Ophelic Acid.
Picrotoxin.
Quercetin, on boiling.
Rhamnetin, cold.
Sapogenin, slightly.

Not by the following (amongst others) :

Alkaloids.

Calamine.
Lobeliine (not till inverted).
Macrocarpine.
Pseudaconine.
Theine, if pure.

Glucosides.

Arbutin.
Diosmin.
Gentiopicrin.
Helixin, not till inverted.
Linamarin.
Melanthin, not till inverted.
Phillyrin, not till inverted.
Rhamnin, no pp.
Sapotin.
Solanine.
Strophanthin.
Villosin, not till inverted.
Xanthostrumarin, not till inverted.

Bitter Principles and Glucoside derivatives.

Aristolochin.
Chirettin.
Ilixanthin.
Marrubiin.
Opionin.
Sikimin.
Syringopicrin.

12. POTASSIUM FERROCYANIDE.

I. Gives precipitates with the following alkaloids :

Akazgine.
Apomorphine, reddish-y. (gn. warm).
Atherospermine.
Bebeerine, y.
Brucine, y. crys. (blue on exposure), not 1 in 500.
Cinchonidine, reddish-y. crys.
Cinchonine, yellowish-white, not 1 in 500, sol. warm.
Cocaine, sol. in excess.
Codeine, white (alcoholic), slight pp. 1 in 1,000.
[Colchicine, concd.]
Curarine.
Emetine, 1 in 1,000.
Harmaline, red.
Hydrastine, white.
Hymenodictyonine.

POTASSIUM BICHROMATE (16).]

Berberine, pp. amorphous.
Brucine, gradually crys.
Cinchonine, grad. crys.
Codeine, grad. pp. 1 in 3,000.
Corydaline, pp.
Curarine, pp. amorphous.
Damascenine, pp. y.
Emetine, grad. y. pp.
Erythrine, pp.
Erythrophlœine, y. pp.
Fumarine, pp.
Harmine, pp.
Hydrastine, y. pp.
Hymenodictyonine, y. pp.
Imperialine, y. crys.
[Narceine, not neutral soln., grad. pp. if acid.]
Narcotine, slight pp. 1 in 400.
Papaverine, pp.
Piliganine, y. pp.
Porphyrine, y. pp and red coloration.
Quinidine, y. pp.
Quinine, y. pp.
Strychnine, y. crys. pp.
Thebaine, pp.
Theobromine, cloud, then grad. pp. at 1 in 3,000.
Veratrine, grad. pp. 1 in 3,000.

No pp.

Aconitine, cloud.
Amarylline.
[Cocaine, not dilute.]
Delphinine.
Morphine, scarcely cloud 1 in 100.
[Narceine, not neutral ; see I.]
Sabadilline, not 1 in 150.
Sabatrine, „ „
Theine, not 1 in 3,000.

17. PHOSPHO-MOLYBDIC ACID (Sonnenschein's Reagent) precipitates alkaloids with very few exceptions ; those liable to oxidation give bluish colours on addition of ammonia, this coloration being in other cases observed even in acid solution. The following list includes all the more important bases, together with some others of lesser significance.

*=blue with ammonia.

I. Alkaloids precipitated :

Acolyctine.
Aconine, gray-bluish.
Aconitine, light y. flocculent.
Anagyrine.
Angeline, partly.
Anthocercine.
Alpha-homo-chelidonine.
Arecoline.
Aspidospermine, white.
Atherospermine, dirty y.
*Atropine, y. flocct.
*Bebeerine.
Beta-homo-chelidonine.
Brucine, orange flocculent.
Calamine.
Calcitrapine, y. ; bl.-gn. after 24 hrs.
Cannabine.
Cannabinine, white.
Carpaine. limit 75,000.
Ceanothine.
Cinchonine, cloud up to 200,000.
Cocaine, cloud up to 50,000.
Codeine, limit 50,000.
Colchiceine.
Colchicine.
*Coniine.
Curarine.
Cytisine, 1 in 10,000.
Damascenine, white.
Delphinine, gray.-y.
Ecboline.
Ecgonine, y.
Emetine, up to 1 in 25,000.
Erythrophlœine, dirty gn.
Ergotine.
Ergotinine.
Hymenodictyonine, y.
Hyoscyamine, y. flocculent.
Laurotetanine.
*Lobeliine, yellowish-white.
Morphine.
Narceine, limit 50,000.
Narcotine, bn.-y. ; limit 4,000.
Nicotine (cloud at 40,000).
Oxyacanthine, y.-white.
Pelletierine.
Physostigmine, 1 in 25,000.
Piliganine, y.
Piperine, bn.
Protoveratrine.
Pseudojervine, cloud 1 in 10,000.
Pseudopelletierine, pale y.
Quinine, white ; y. with ammonia.
Sabadilline, cloud 1 in 5,000.
Sabatrine, „ „
Scopoline, white.

[PHOSPHO-TUNGSTIC ACID (18).
Solanine, light y.
Sparteine, white.
Strychnine, y.-white.
Taxine.
Thebaine, extreme limit 50,000.
Theine.
Theobromine, y.
Veratralbine, cloud 1 in 3,500.
Veratrine, cloud 1 in 5,000.

II. Non-basic substances precipitated.

Acorin, reduction ; blue.
Chamælirin, y.-white.
Euonymin, gn.-y.
Helleborein.
Menyanthin, y.
[Solanine, A.-G., see I.]
[Strophanthin, if concd.]

III. No precipitate with following :

Arecaine, A., cloud.
Chrysanthemine, A.
Coriamyrtin, G.
Digitalin, G.
Eupatorine, A., green colour.
Lycoctonine, A.
Papaverine, A.
Picrotoxin, B.
Salicin, G.
[Strophanthin, emerald solution ; pp. if concd.]

18. PHOSPHO-TUNGSTIC ACID (Scheibler's Reagent).

Pp.

Aconitine.
Alpha-homo-chelidonine.
Anthocercine.
Aricine (cloud 1 in 50,000).
Arganine.

PHOSPHO-TUNGSTIC ACID (18).]

Atropine.
Beta-homo-chelidonine.
Calabarine.
Cannabine.
Cocaine, gelatinous.
Corydaline.
Cytisine, 1 in 30,000.
Helleborein, G., and by Meta-tungstic acid.
Hyoscine.
Hyoscyamine.
Laurotetanine.
[Lobeliine, pp. by Meta-tungstic acid.]
Mandragorine.
[Morphine, cloud 1 in 10,000.]
Narcotine, limit 1 in 8,000.
Pilocarpine, white.
[Spigeliue, pp. by Meta-tungstic acid.]
Strychnine, 1 in 200,000.

No pp.

[Chamælirin, G., Meta-tungstic acid.]
Strophanthin, G.

19. PHOSPHO-ANTIMONIC ACID (Schultz's Reagent).

Pp.

Artarine, slight pp., sol. in excess.
Atropine, up to 1 in 5,000; sol. warm.
Brucine.
Cinchonine, bluish-white pp. 1 in 1,000, cloud at 1 in 5,000.
Narcotine, y. flocculent 1 in 1,000, cloud 1 in 2,500.
Quinine.
Strychnine, white flocculent 1 in 5,000, cloud 1 in 25,000.
Veratrine.

No pp.

[Codeine, dirty white, cloud 1 in 1,000.]
[Digitalin, G., cloud 1 in 1,000.]
[Morphine, not 1 in 1,000.]
[Nicotine, cloud 1 in 250.]
[Piperine, yellow coloration in dilute solutions.]
Theine.
[Theobromine, cloud 1 in 1,000.]

20. IODO-POTASSIC IODIDE (Wagner's Solution) precipitates almost every alkaloid without exception. The following list includes all the more important bases, besides many of the rarer ones. The strength of the reagent may vary without greatly affecting the result. Precipitates are generally brown or reddish-brown, and amorphous.

k = kermes-coloured.

Aconitine.
Akazgine.
Anthocercine.
Apomorphine, blood-red, sol. warm.
Aricine, pp., cloud 50,000.
Aspidospermine, y.
Atherospermine, bn.-yellow.
Atropine, r.-bn.
Bebeerine, k
Berberine, k.
Brucine, k.
Calamine.
Calcitrapine.
Cannabinine.
Ceanothine.
Cinchonine, k., 1 in 500,000.
Cocaine, bn.; cloud 200,000.
Codeine, k., crys.
Colchicine, 1 in 2,500.
Conhydrine.
Coniine, k., 1 in 8,000.
Corydaline, bn.
Cryptopine.
Cytisine, dark bn.
Damascenine, bn.-purple.
Delphinine, k., 3,000.
Ecgonine, r.-bn.
Emetine, 1 in 25,000.
Erythrophlœine, y.-r.
Ergotine.
Grandiflorine, y.
Hydrastine, bn.
Hymenodictyonine, bn.
Hyoscine, oily.
Hyoscyamine, orange.
Igasurine, bn.
Imperialine, dark y.
Lobeliine, bn.
Lycoctonine.
Mandragorine.
Morphine, k., 5,000.
Narceine, crys.
Narcotine, k., 8,000.
Nicotine, r.-bn.
Papaverine, 1 in 50,000.
Pelletierine.
Physostigmine, k., 25,000.
Piliganine, light bn.
Protopine.
Pseudojervine.
Pseudopelletierine, pale bn. needles.
Quinidine, K.
Quinine, r.-bn.
Sabadilline, k.
Sabatrine, k.
Scopoline, bn.
Spigeline, bn.-r.
Strychnine, k. crys, limit 50,000.
Taxine, y.
Thebaine, 1 in 5,000; extreme limit 50,000.
[Theine, if concentrated.]
[Theobromine, if concentrated.]
Tritopine.
Veratrine, k.

Glucosides and bitter principles are not precipitated as a rule. An exception to former class is:

Euonymin, r.-bn. pp.

[BISMUTH-POTASSIC IODIDE (21).

21. BISMUTH-POTASSIC IODIDE (Dragendorff's Reagent).

Pp.

Aconitine, orange, 1 in 40,000.
Amarylline, y.
Anagyrine.
Anthocercine.
Apomorphine, limit 10,000.
Arecaine, red, becoming crys.
Arecoline, „ „
Artarine, r. flocculent.
Atropine, orange, 1 in 10,000 flocculent.
Bebeerine, orange-red.
Bellamarine, white.
Berberine, orange-r.; limit 5,000.
Brucine, orange-r.; 1 in 10,000.
Calcitrapine, orange.
Ceanothine.
Chelidonine, orange-red.
Chrysanthemine, orange, grad. crys.
Cinchonine, orange-r.; cloud 200,000.
Codeine, orange; slight at 50,000.
Colchicine.

BISMUTH-POTASSIC IODIDE (21).]

Coniine, y., limit 6,000.
Corydaline.
Cryptopine.
Curarine.
Delphinine, orange-r.
Emetine, 1 in 25,000.
Erythrophlœine, y.
Hymenodictyonine, r.
Hyoscyamine, r. amorphous.
Imperialine, orange.
Lycoctonine, 1 in 40,000.
Morphine (faint at 5,000.
Narceine.
Narcotine, orange-r.; limit 4,000.
Nicotine, pp. cloud 1 in 40,000.
Papaverine, orange-r.; slight at 10,000.
Piperidine, the Hydriodide.
Quinine, orange-r.; cloud 50,000.
Strychnine, orange-red; limit 250,000.
Taxine, y.
Thebaine, orange; slight at 10,000.
Theine, 1 in 3,000.
[Theobromine, cloud if dilute.]
[Veratrine, not dilute.]

Of the Glucosides, Menyanthin gives y. pp., and Digitalin pp. in concentrated solutions.

22. CADMIUM-POTASSIC IODIDE (Marmé's Reagent).

Pp.

Ancoihtine. wite 1 in 1,000, cloud 1 in 2,500.
Alpha-homo-chelidonine.
Amarylline, y.
Artarine, y. flocculent.
Atropine.
Bebeerine.
Bellamarine, white.
Beta-homo-chelidonine.
Berberine.
Brucine, crys.
Calcitrapine, orange-y.
Ceanothine.
Chelidonine.
Cinchonine, hair-like crys.; 50,000.
Codeine, grad. crys. 1 in 500.
[Colchicine, only concd.]
Coniine.
Curarine.
Cytisine.
Damascenine, white.
Delphinine.
Emetine, y. amorphous.
Erythrophlœine, white.
Hymenodictyonine.
Hyoscyamine, 1 in 10,000.
Imperialine, white flocculent.
Lycoctonine.
Morphine, needles 1 in 1,000 after a time.
Narceine, 1 in 1,000.
Narcotine, 1 in 3,000; limit 8,000.
Nicotine.
Papaverine, 1 in 1,000.
Pelletierine.
Physostigmine, pale y. 1 in 1,000.
Piperine.
Protoveratrine.
Pseudopelletierine.
Quinidine.
Sparteine, white.
Strychnine, flocculent.
Taxine.
Thebaine, up to 1 in 10,000.
Veratrine.

No pp.

Æsculin, G.
Amygdalin, G.
Chamælirin, G.
[Colchicine, only concd.]
Colocynthin, G.
Digitalin, G.
Glycyrrhizin.
Helleborein, G.
Helleborin, G.
Ononin, G.
Protoveratridine.
Salicin, G.
Saponin, G.
Solanine (not dilute).
Strophanthin, G.
Theine.
[Theobromine, not dilute.]

[MERCURIC-POTASSIC IODIDE (24).

23. ZINCIC-POTASSIC IODIDE.

Pp.

Berberine, amorphous (1 in 3,000 to 6,000).
Codeine, long hair-like crys. (not 1 in 1,000).
Narceine, crys. become blue on exposure (feeble at 1 in 5,000).
Papaverine, 1 in 10,000.
Quinidine.
Quinine.
Strychnine.
Thebaine (not 1 in 500).
[Veratrine, see opposite.]

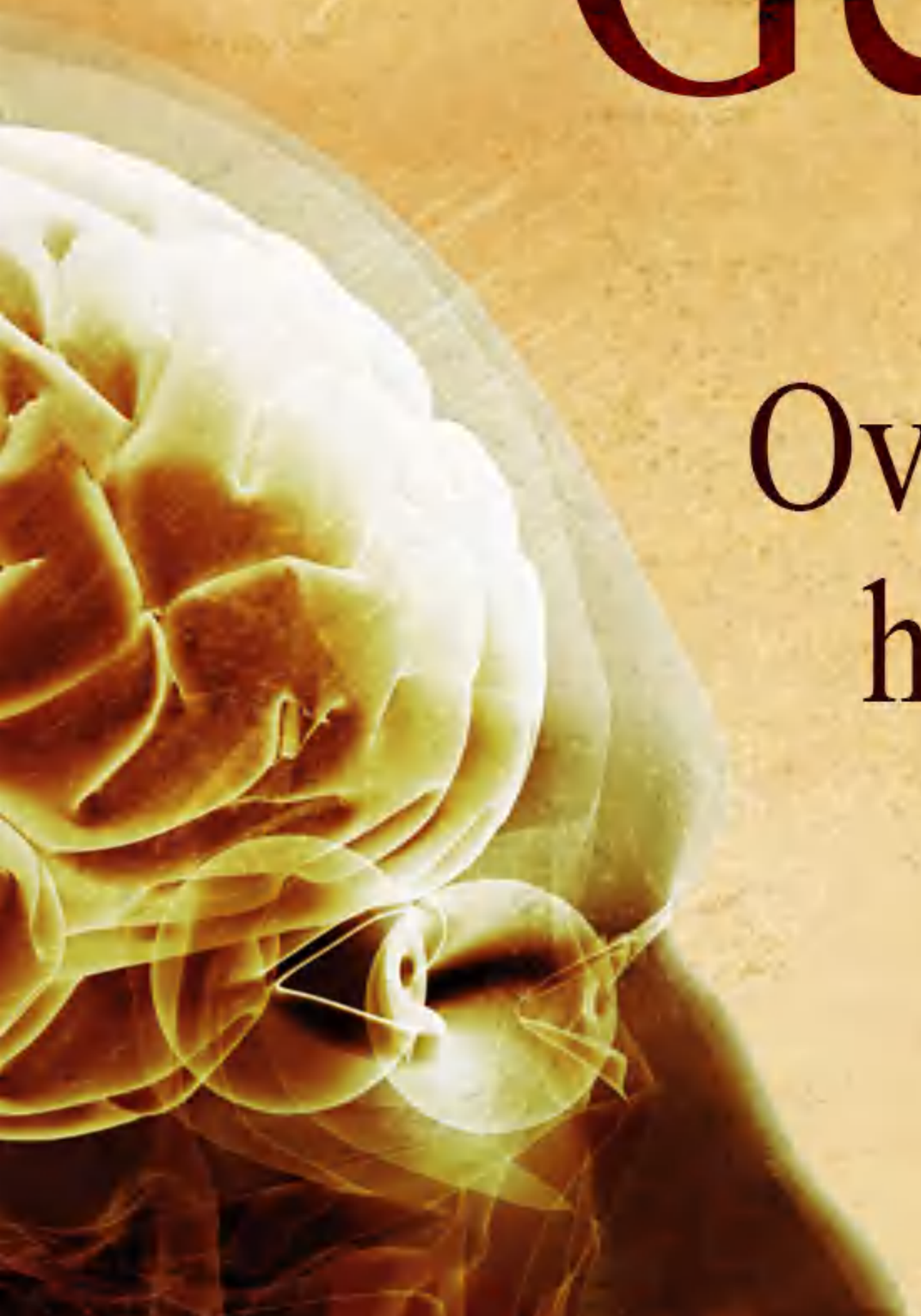
No pp. or scarcely.

Atropine.
Cinchonidine.
Cinchonine.
Coniine.
Morphine.
Narcotine.
Nicotine.
Theine.
[Veratrine, very slight pp.]

24. MERCURIC-POTASSIC IODIDE (Mayer's Solution), a very general precipitant of alkaloids, much used for quantitative determinations.

I. Alkaloids precipitated.

Aconitine, white.
Akazgine.
Amarylline, y.-green.
Anagyrine.
Angeline, partially.
Anthocercine.
Arecaine, y.-crys.
Arecoline, y. oily, grad. crys.
Aricine, pp.; cloud 50,000.
Artarine, y. flocculent.
Aspidospermine, y. flocculent.
Atisine, white.
Atropine, up to 1 in 7,000.
Bebeerine, white.
Bellamarine, y.-green.
Berberine.



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MERCURIC CHLORIDE (25).]
Theine, 1 in 1,000 grad. crys.
Theobromine, pp.; cloud at 3,000.
Veratrine, M.P. 172° (no pp. 1 in 500).

II. Glucosides, etc., precipitated:

Adonin, G.
[Aloin, B., pp. with mercurous nitrate.]
Arnicin, B. (alcoholic solution).
Calendulin.
Cantharidin, B., w. crys.
Catechin, dirty w. (also pp. by mercurous nitrate).
[Convallamarin, pp. mercurous nitrate.]
Emodin.
[Melanthin, cloud.]
Neriodorin.
Ononid, B.
[Santonin, B., pp. mercurous nitrate.]
Sinalbin, G.

III. Following substances *not* precipitated:

Aloin, B.
Angelin.
Angeline.
Araliin, G.
Arecoline, A.
Californin, B.
Chrysanthemine, A.
[Colchicine, pp. concd.]
Cornin, B.
Cytisine, see I.
Ergotinine, A.
Erythrocentaurin, G.
Helleborein, G.
Lobeliine, A.
Lycoctonine, A.
Ratanhine (free).
Rosaginin (alcoholic).
Panaquilon, B.
Picrotoxin, B.
Protoveratrine, A.
Sabadilline, not 1 in 150.
Sabatrine, not 1 in 150.
Scoparin, B.
Solanine, G.
Taxine, A.
Trigonelline, A.

26. SOLUTIONS OF THE HALOGENS.

Cl. = Chlorine water, Br. = Bromine water, I = Iodine tincture, Th = *Thalleioquin test* (green coloration on addition of ammonia after treatment with chlorine water; compare Quinine.)

I. Following alkaloids give reactions:

Akazgine, pp.
Alpha-homo-chelidonine, Br. pp.
Anagyrine, I. bn. pp.
Apoquinidine, Th. faint.
Apoquinine, Th. faint.
Arecoline, I. bn. pp.
Aribine, Cl., Br. and I. pp.; sol. warm.
Atropine, Br., y. crys.; see Part I.
Berberine, Cl. red ring; see Part I.
„ Br. y. pp.
„ I. gn. y.-bn. crys.
Beta-homo-chelidonine, Br. pp.
Boldine, I. bn. pp.
Brucine, Cl., colours y. then r., then white pp.
Brucine, Br., violet coloration.
„ I., bn. crys.
Carpaine, I. pp. (limit 1 in 200,000).
Chelerythrine, I. pp.
Chelidonine, I. pp.
Cinchonidine.
Cinchonine.
Cocaine, I. bn. pp.
Codeine, r.-bn. with Th. test; Br. pp.
Corydaline, I. pp.
Cupreine, Th.
Delphinine, I. pp.
Dicinchonine.
Ergotinine, Br. pp.
Erythrine, I. pp.
Homoquinine, Th.
Hydroquinidine, Th.
Hydroquinine, Th.
Hyoscyamine, I. pp.
Laurotetanine, I. pp.
Lupanine, I. bn.-r. pp.
Lycoctonine, Br. pp. crys.
Lycopodine, I. bn., cloud.
Muscarine.
Narceine, I. crys.
Narcotine, Cl. greenish; y. with ammonia.
Narcotine, Br. y. pp.
Nicotine, I. long crys. (ethereal solution).
Oxyacanthine, Br. pp.
Oxyacanthine, I. r.-bn. pp.
Papaverine, Cl. greenish; bn. with ammonia.
Papaverine, I. pp. crys.
Pelletierine, Br. pp.
Physostigmine, Br. r. coloration
[SOLUTIONS OF HALOGENS (26).
Piliganine, Br., I., y. pp.
Quinicine, Th.
Quinidine, Th.
Quinine, Th.
Sabadilline, Br. cloud 1 in 5,000.
Sabatrine, Br. cloud 1 in 5,000.
Spigeline, I., bn.-r. pp.
Strychnine, Cl. white pp. (sensitive).
Thebaine, r.-bn. with Th. test.
„ Br. pp.
Veratrine, golden y. with Th. test.
„ Br. cloud 1 in 5,000.

II. Following non-basic substances give reactions:

Acorin, I. pp.
Aloin, Br. pp.
Menyanthin, I. y. pp.
Picrolichenin, Cl. y. coloration.
Santonin, Cl. grad. crys.
Xanthostrumarin, Cl. y. coloration.

III. No reaction or pp. with following:

Atherospermine, Th.
Cinchonamine, Th.
Cinchonicine, Th.
Cinchonidine, Th.
Cinchonine, Th.
Cornin, I.
Digitalin, I.
Erythrocentaurin, Cl., Br., or I.
Fumarine, Cl. no coloration.
Hydrastine, Cl.
Macrocarpine, I.
Picrotoxin, I. no pp.
Sabadilline, Cl.
Sabatrine, Cl.

27. CONCENTRATED SULPHURIC ACID produces colour changes with an extremely large number of alkaloids, glucosides, and amaroids, when a fragment of the substance is brought in contact with it; in many instances a succession of colours is given.

I. Alkaloids giving red, purple, and violet shades.*

[Aconitine, if impure r. or grad. v.]
Apomorphine, r.
Asiminine, gn.~r.~colourless.
Atisine, y.~p.
Beta-homo-chelidonine, v.
Boldine, r.
[Brucine, r. if acid, not pure.]
Calcitrapine, see II.
Chelerythrine, y.-r.
Cinchonamine, r.-y.
[Cocaine, r. if not pure; see IV.]
Corydaline, y.-r. (?); see IV.
" (Adermann's), y.~v.
Cupronine, r. (bl.-v. with water).
Curarine, r. or bl.~r.
Delphinoidine, r.-v.
Ditamine, r.
Ecboline, dark r.
Echitamine, p.
Echitenine, r.-v.
Eupatorine, r.-bn.
Fumarine, v.
Grandiflorine, y.~r.
Heliotropine, y.~r
[Hydrastine, r. if acid, not pure.]
Hypoquebrachine, v.
Laudanine, pink.
Laudanosine, r.-bn.
Laurotetanine, pale rose.
[Narceine, if impure blood-r. or bl.; see II.]
[Papaverine, if impure bl.-v.~v.; see IV.]
Pareirine, v.
Physostigmine, grad. r.; see II., III.
Porphyrine, p.
Protoveratridine, v.~cherry-r.
Rheadine, r., or III.
Rubijervine, r.
Sabadilline, r.
Sabadine, y.~blood-r.~v.
Sabadinine, blood-r.
Sabatrine, grad. r.
Solanine, r.-y.
Staphisagrine, r.~v., or IV.
Taxine, p.~v.
Thebaine, blood-r.
Veratralbine, y.~r.
Veratrine, y.~r.
Veratroidine, see II.

Ia. Non-alkaloidal substances giving red, purple, and violet shades.

Adonin, G., deep r.
Angelin, B., r.
Amygdalin, G., light v.-r.
Assamic acid, G., y.~r.~bl.~v.
Brucamarin, v.
Californin, B., bn.-r.
Calumbin, r.
Camellin, r.
Cascarillin, B., blood-r.
Cerberin, v.
Chamælirin, B., r.~black.
Cnicin, B., r.~v.
[Convallamarin, v., acid to aqueous solution.]
Convolvulin, G., grad. r.
Copalchin, B., r.
Cyclamin, G., r.
Digitonin, G., v.
Dulcamarin, B., r.~rose.
Elaterin, B., dark r.
Frangulin, G., gn.~p.~dark r.
Gratiosolin, bn.-r.
Helixin, G., bright r.
Helleborein, G., bn.-r.
Helleboretin, bn.-r.
Helleborin, G., blood-r.
Hesperidin, r. (intense on warming).
Hydrocarotin, r.
Jalapin, G., grad. r.
Kosin, B., r.
Laserpitin, B., r.
Limonin, G., blood-r.
Linin, B., dark v.
Mangostin, B., y.-r.
Melanthin, G., rose~v.-r.
Menyanthin, G., y.-bn.~v.
Ononin, G., r.-y.~r.
Panaquilon, p.-r.
Phillygenin, r.
Phillyrin, G., r.-v.
Phloretin, r.
Phlorrhizin, G., y.~r.
Physodin, v.
Pimpinellin, B., r.
Podophyllotoxin, r.~gn.~bl.~v.
Populin, G., r.
Quinovin, B., grad. r.
Rosaginin, G., r.-bn.
Salicin, G., r.
Saligenin, intense r.
Sapogenin, grad. v.
Saponin, G., light r.
Sapotin, G., r.
Scillin, r.-bn.
Scillitoxin, r.~bn.
Smilacin, dark r., or II.
Sophorin, G., flesh-colour.
Sparattospermin, gn.~r.
Tampicin, G., grad. r.
Theveresin, G.-d., as Thevetin.
Thevetin, G., r.-bn.~r.~v.
Turpethin, G., grad. r.
Urechitin, Urechitoxin, } y.~r.~p.
Vernonin, G., bn.~p.

II. Alkaloids giving yellow to brown colours:

Alpha-homo-chelidonine, y.
Amarylline, r.-bn.~bn.
Anthocercine, y.-bn.
Atisine, see I.
Berberine, y.
Calcitrapine, dark bn.~v.~gray-bn.
Calycanthine, pale y.
Ceanothine, r.-bn.
Chelerythrine, y.-r.
Cinchonamine, r.-y.
Colchiceine, Colchicine, } y.
Conchairamidine, y.~dark gn.
Conchairamine, bn.~intense gn.
Corydaline (Adermann's), y.~v.
Delphinine, light bn., or IV.
Delphinoidine, r.-bn.
Emetine, dirty bn. to gn.
Gelsemine, y.~r.-bn., or IV.
Gnoscopine, y.
Grandiflorine, y.-r.
Heliotropine, y.~r.

* It must not be overlooked that impure alkaloids frequently give colours, although the pure substances do not.

SULPHURIC ACID (27).]

Hydrocotarnine, y. (r. warm).
Imperialine, pale y.
Javanine, intense y.
Jervine, y.
Laudanosine, r.-bn.
Lobeliine, r.-bn.
Narceine, if pure, bn.; see I.
Oxyacanthine, r.-bn. (?); see II.
Paricine, y.-gn.
Physostigmine, y.~gn.(?); see also I., III.
Piperine, y.~gn. slowly.
Sabadine, see I.
Solanine.
Ustilagine, see III.
Veratralbine, y.~r.
Veratrine, y.~r.
Veratroidine, orange~r.

IIa. Non-alkaloidal substances giving yellow to brown colours.

Absynthiin, see III.
Agoniadin, G., light y.~gn.
Andromedotoxin, B., bn.
Antiarin, B., y.-bn.
Assamic Acid, G., see I.
Capsicin, bn.
Digitalein, G., r.-bn., or IV.
Diosmin, G., y.
Ericolin, G., y.-bn.
Euonymin, G., y.~r.-bn.
Eupatorin, G., r.-bn.
Euphorbon, B., y.-bn.
Fraxetin, G.-d., y.
Gentiopicrin, G., y.
Gratiolin, y.-bn.
Gratiosolin, bn.-r.
Guacin, B., r.-bn.
Helixin, G., y.
Helleborein, bn.-r.
Helleborin, bn.-r.
Junipicrin, B., light y.
Liriodendrin, orange.
Mangostin, B., y.-r.
Menyanthin, G., y.-bn.~v.
Ononin, G., r.-y.~r.
Phlorrhizin, G., y.~r.
Picrotoxin, B., y.
Prophetin, r.-bn.
Rosaginin, G., r.-bn.
Scillain, G., bn.
Scillin, r.-bn.
Scillitoxin, r.~bn.
Smilacin, G., y.~r., or I.
Solanine, A.-G.
Strophanthin, G., pale y.-bn.
Syringopicrin, gn.-bn.
Urechitin, Urechitoxin } see I.
Xylostein, B., bn.
Vernonin, see I.
Villosin, G., light bn.

III. Alkaloids giving blue and green colours.

Aricine, dark gn.
Aristine, bl.~gn.-bl.
Asiminine, see I.
Bebeerine, dirty olive-gn.
Chairamidine, y.~dark gn.
Codamine, gn.
Conchairamidine, intense gn.
Conchairamine, bn.~intense gn.
Concusconine, bl.-gn.~olive-gn.
Cryptopine, intense bl.
Cuprine, deep bl.
Curarine, r. or bl.~r.
Cusconine, y.-gn.
Emetine, see II.
Geissospermine, colourless~bl.~colourless.
Glaucine, bl.
[Papaverine, if impure; see I. and IV.]
Paricine, y.-gn.
Physostigmine, y.~olive-gn. (?); see I.
Piperine, see II.
Protoveratrine, grad. gn.~bl.
Pseudomorphine, olive-gn., or IV.
Quebrachamine, bl.
Rheadine, olive-gn., or I.
Thebaicine, bl.
Thebenine, bl.
Tylophorine, r.-bn.~gn.~bl.
Ustilagine, bn.~intense gn.

IIIa. Non-alkaloidal substances giving blue and green colours.

Absynthiin, B., bn.~gn.~bl.
Agoniadin, G., light y.~gn.
Aristolochin, B., dark gn.
Asclepiadin, G., y.-gn.~deep gn.
Assamic Acid, G., see I.
Coniferin, G., v.-bl.
Crocetin, G.-d., bl.
Crocin, G., deep-bl.
Digitalin, G., gn. (authorities differ).
Frangulin, G., gn.~p.~dark r.
Sparattospermin, gn.~r.
Strophanthin, G., pale y.~bn.~gn.
Syringin, G., dark bl.
Syringopicrin, B., gn.~bn.
Tulucunnin, bn.~v.

IV. Alkaloids giving no colour, or very pale yellow merely.

Aconine.
Aconitine, see I.

[SULPHURIC ACID (27).

Akazgine.
Alstonidine.
Anagyrine.
Artarine.
Atherospermine.
Atropine.
Brucine (red with trace HNO_3).
Carpaine.
Chairamine.
Chelidonine.
Chrysanthemine.
Cinchonidine.
Cinchonine.
Cinchovaic Acid.
Cocaine.
Codeine.
Coniine.
Coptine.
Corydaline, see I.
Cytisine.
Delphinine, see II.
Geissospermine, colourless at first; see III.
Gelsemine, or III.
Hydrastine.
Hyoscine.
Hyoscyamine.
Lanthopine.
Lycoctonine.
Morphine.
Narcotine.
Nicotine.
Oxyacanthine, see II.
Papaverine, if pure; see I.
Picraconitine.
Pilocarpine.
Pseudojervine, if pure.
Pseudomorphine, if pure acid; see III.
Quinidine.

NITRIC ACID (28).]

Cusconine, gn.
Ditamine, see II.
Echitamine, see I.
Echitenine, see I.
Parabuxine, gn.
Solanine, A.-G., colourless~bl.~colourless.

III*a*. Non-alkaloidal substances giving green or blue colours.

Adonin, G., indigo bl.
Crocin, G., momentary bl.~bn.
[Digitalin, gn., or II*a*., IV*a*.]
Scillain, G., y.~gn.
Solanine, A.-G., see III.

IV. Alkaloids giving no colours or but faint yellow.

Aconitine, if pure; see also II.
Atisine.
Atropine.
Carpaine.
Cinchonidine.
Cinchonine.
Conhydrine.
Coniine.
Delphinine.
Delphinoidine.
Fumarine.
Hyoscyamine.
Lycoctonine.
Nicotine.
Paytamine, at first colourless; see I.
Piturine.
Quinidine.
Quinine.
Ratanhine, if pure; see I.
Scopoline.
Solanine, A.-G., at first colourless; see III.
Sparteine.
Strychnine, or II.
Surinamine (Ratanhine), if pure; see I.
Taxine, or II.
Theine.
Theobromine.
Valdivine.
Veratrine, or II.
Vicine, see I.

IV*a*. Non-alkaloidal substances giving no colours or but faint yellow.

Araliin, G.
Ceratophyllin, faint y.
[Digitalin, G., authorities differ; see II., III.]
Erythrocentaurin.
Helixin, G.
Liriodendrin, B.
Solanine, A.-G.; see IV.
Variolarin, B.

29. FRÖHDE'S SOLUTION (concentrated sulphuric acid containing molybdic acid).

I. Alkaloids giving red, purple, or violet shades.

Apomorphine, v.
Beta-homo-chelidonine, y.~v.~gn.
Brucine, r.~y.
Cocaine, blood-r.
Corydaline (Adermann's), v. or gn. streaked with v.
Delphinoidine, blood-r.
Emetine, r.~gn.
Ergotinine, see III.
Fumarine, v.~dark gn.
Hypoquebrachine, v.
Lobeliine, r.-bn.
Loxopterigine, v.~bl.
Morphine, magnificent v.~gn.~bn.-gn.~y.; after 24 hours bl.-v.
Narceine, see II.
Oxyacanthine, v.
Papaverine, see III.
Porphyrine, p.
Sabadilline, grad. r.-v.
Sabatrine, grad. r.-v.
[Solanine, A.-G., cherry~bn.~r.~y.]
Staphysagrine, bn.-v.
Taxine, r.-v.
Veratrine, y.~cherry.
Veratroidine, grad. cherry.

I*a*. Non-basic substances giving above colours.

Absynthiin, see II.
Colocynthin, grad. r.~bn.
Digitalin, see II.
Ononin, pure r.
Populin, v.
Salicin, magnificent v.~dark cherry.
Solanine, see I.
Syringin, blood-r.~v.-r.

II. Alkaloids giving yellow to brown colours.

Aconitine, y.-bn.
Amarylline, bn.-gn.
Bellamarine, bn.
Beta-homo-chelidonine, see I.
Bebeerine, bn.-gn.

[FRÖHDE'S SOLUTION (29).

Colchiceine, y.
Colchicine, y.~y.-gn.
Conchairamine, brownish.
Delphinine, r.-bn.~dirty gn.; or IV.
Hydrastine, gn.~bn.
Imperialine, gn.-y.
Lobeliine, r.-bn.
Narceine, bn.~gn.~r.~v.
Piperine, y.~bn.
Thebaine, orange.
Veratrine.

II*a*. Non-basic substances giving yellow to brown colours.

Absynthiin, bn.~v.-bl.
Chamælirin, y.-bn.
Colocynthin, see I.
Digitalin, orange~r.~bn.~black.
Elaterin, y.
Quassiin, bn.
Villosin, dark bn.

III. Alkaloids giving green or blue colours.

Alpha-homo-chelidonine, dirty bn.-gn.
Amarylline, bn.-gn.
Aricine, bl.~gn., or IV.
Aspidosamine, bl.
Bebeerine, bn.-gn.
Berberine, bn.-gn.~gn.
Beta-homo-chelidonine, y.~v.-gn.
Calcitrapine, olive-gn.
Ceanothine, bl.
Codeine, dirty gn.~indigo-bl. after 24 hours.
Conchairamidine, intense gn.
Corydaline (Adermann's), see I.

Emetine, r.~gn.
Ergotinine, v.~bl.
Fumarine, v.~dark gn.
Hydrastine, gn.~bn.
Imperialine, see II.
Laurotetanine, indigo-bl.
Loxopterigine, v.~bl.
Morphine, see I.
Narceine, see II.
Narcotine, gn.
Papaverine, gn.~bl.~v.~cherry.
Pareirine, bl.
Quebrachine, bl.
Quinidine, greenish.
Quinine, greenish.
Tylophorine, bright gn.

IIIa. Non-basic substances giving green or blue colours.

Absynthiin, see II*a*.

30. VITALI'S TEST. — Evaporation with fuming nitric acid at 100° C., then touching the residue with a drop of *alcoholic* potash solution freshly prepared.

Atropine, magnificent violet, changing gradually to cherry-red.
Brucine, greenish.
Homatropine, yellow.
Strychnine, red.
Veratrine, similar to Atropine.*

Other alkaloids that were similarly tested gave negative results.

31. GERRARD TEST.—To $\frac{1}{10}$ grain of *free* alkaloid about 20 drops of a 2 per cent. solution of mercuric chloride in proof-spirit are added gradually, avoiding excess.

* If the test be modified by substituting nitrous acid or a nitrite for the nitric acid, and aqueous for alcoholic potash; atropine still gives a reddish-violet whilst veratrine produces a yellow coloration (E. Beckmann).

Phlorrhizin, bl.
Phloretin, bl.

IV. The following alkaloids give no colours.

Alstonidine.
Aricine.
Atropine.
Chairamine.
Cinchonine.
Coniine.
Delphinine, see II.
Hyoscyamine.
Nicotine.
Theine (Caffeine).
Theobromine.
Strychnine.

Pp. red or yellow.

Atropine, immediate red coloration, then pp.
Homatropine, red pp.
Hyoscyamine, at first yellow coloration, then red pp. (no pp. if excess of reagent).
Scoparin, yellow pp.

No red or yellow pp. (a white pp. is frequently obtained).

Aconitine.
Arbutin, G., no pp.
Brucine.
Cinchonidine.
Cinchonine.
Cocaine (white pp.).
Codeine.
Condurangin, no pp.
Coniine.
Gelsemine.
Hyoscine, no coloration.
Morphine.
Quinidine.
Quinine.
Sparteine, no pp.
Strychnine, no pp.
Theine, no pp.

32. PERMANGANATE OF POTASH, added to a solution of the hydrochloride salt of the alkaloids, gives the following results (Beckurt and List, *Pharm. J. Trans.*, and *Jahresb.*, 1886, II.).

I. Immediate oxidation and pp. of Manganese Peroxide.

Aconitine.
Brucine.
Cinchonamine.
Cinchonidine.
Cinchonine.
Codeine.
Colchicine.
Coniine.
Nicotine.
Physostigmine.
Quinine.
Thebaine.
Veratrine.

II. Red solution, oxidation being slower than with alkaloids under I.

Atropine.
Berberine.
Hyoscyamine.
Pilocarpine.
Piperine.
Strychnine.

III. Alkaloids behaving in an exceptional manner.

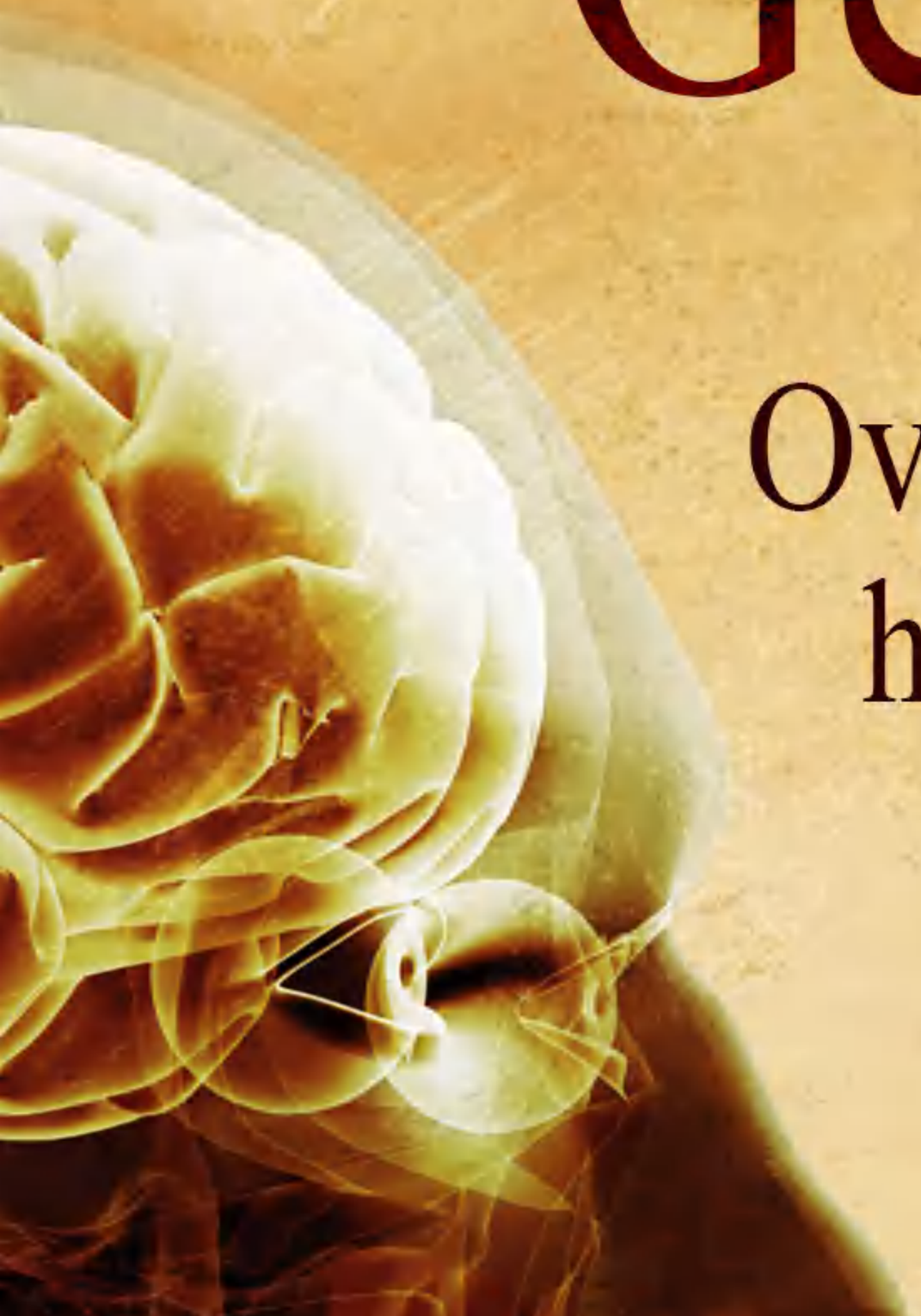
Apomorphine, intense green.
Morphine, white crys. pp. (=Oxidimorphine).
Cocaine, *stable light violet pp. of Cocaine Permanganate* (characteristic).
Narceine, Narcotine, Papaverine } also give permanganate salts, but which are not stable.

BOTANICAL INDEX.

THE sign ⊣ opposite the name of a plant indicates that it forms a group-heading, and may consequently be directly found in its alphabetical position in Parts I. and II.

Any other plant is to be found under the group-heading specified in this index.

The first column gives the plant-name with botanical order, and the second the group-heading to which to refer. The names in these two columns, although frequently synonymous, must not be regarded as necessarily so.



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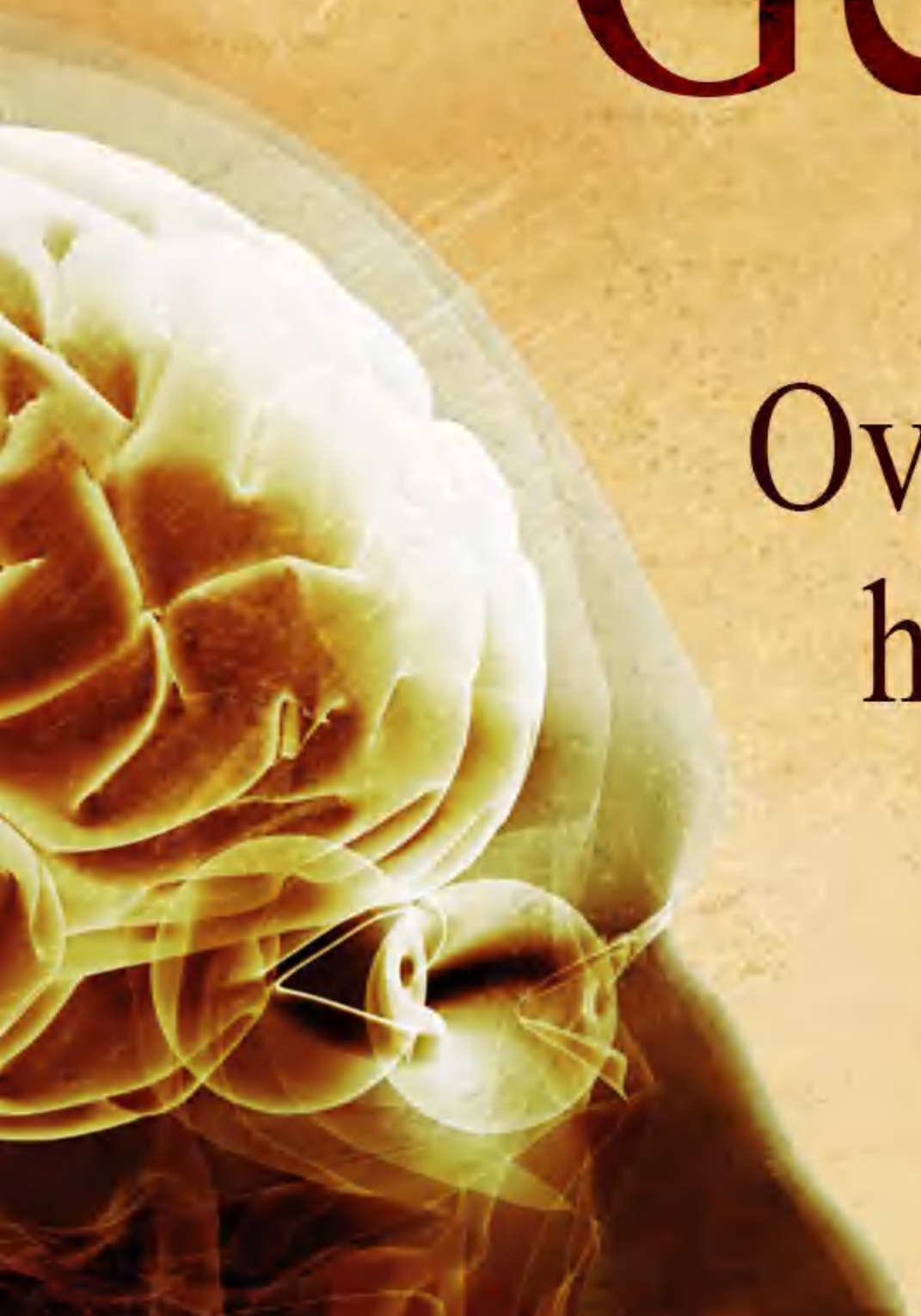

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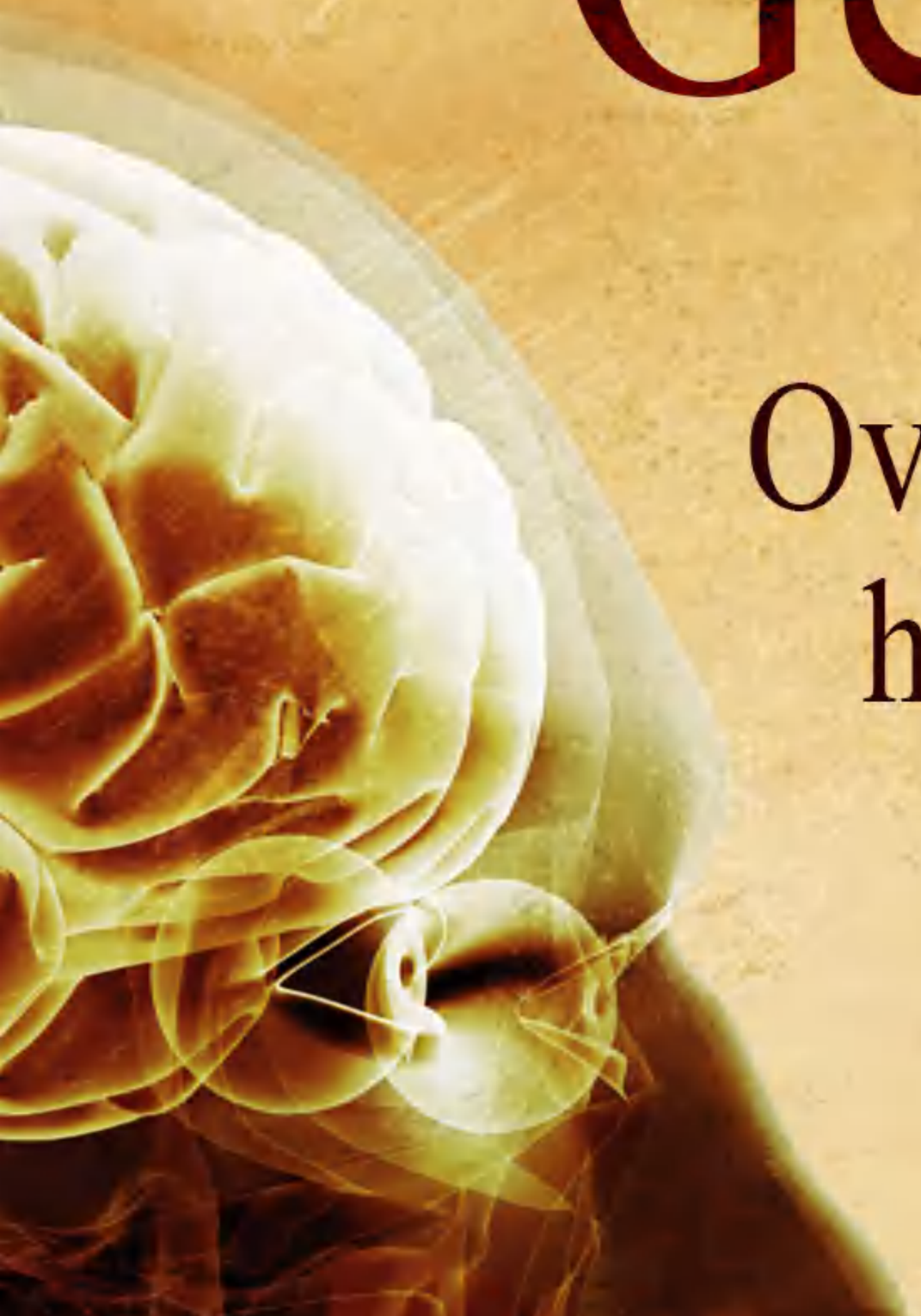
INDEX OF SUBSTANCES, REAGENTS, ETC.

THIS index gives the pages on which the substance looked for will be found in Parts I. and II.; to Part III. no references here are needed, each list in that section being alphabetically arranged.



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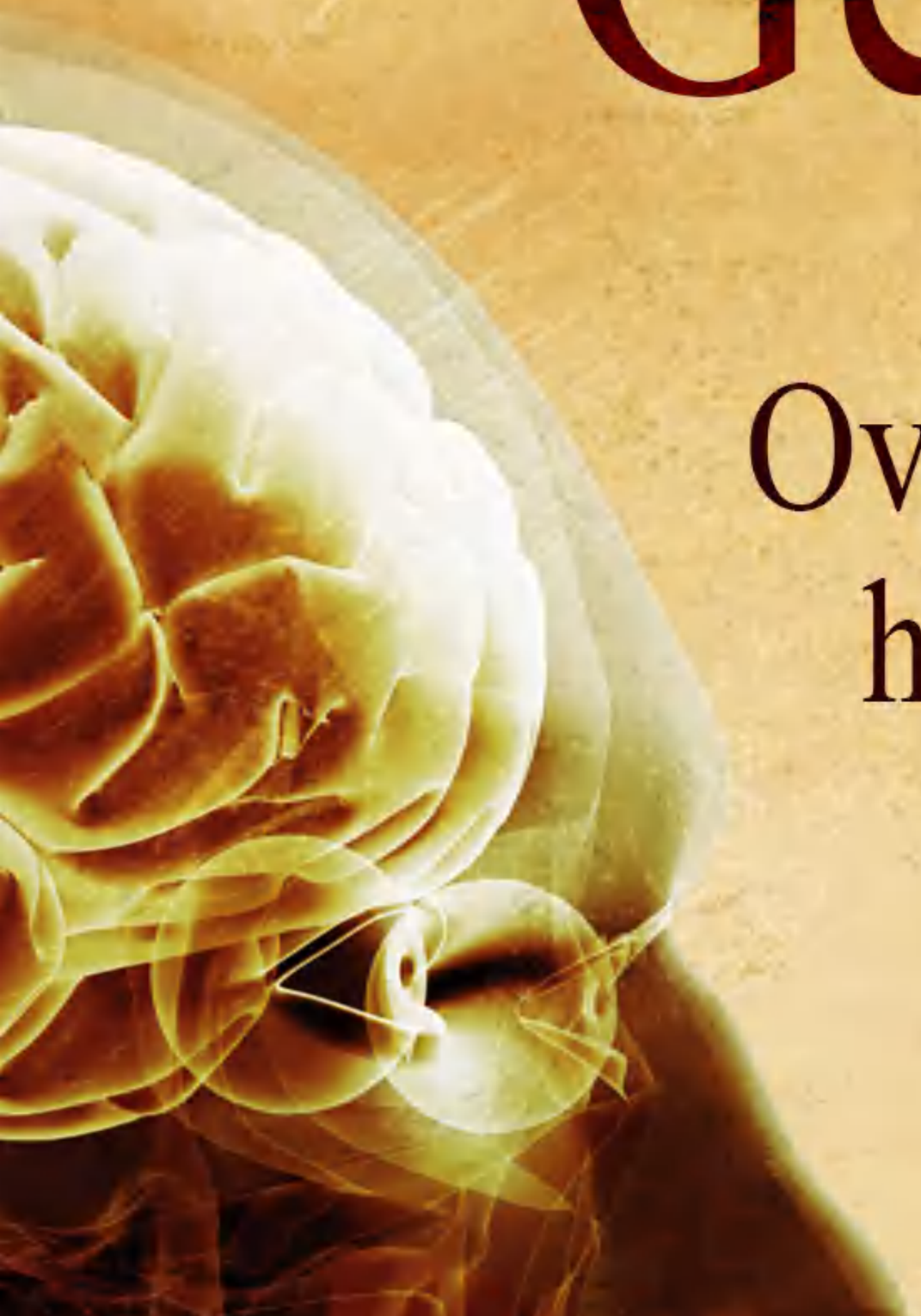

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Soluble in boiling concentrated acetic acid (separating in yellow needles on cooling), less readily dissolved by ether and alcohol, scarcely by water.

Reactions:

Potassic hydrate, sol. red on boiling; the solution, on cooling, becomes a crimson jelly, admixed with crystalline needles.
Potassic carbonate
Calcic hydrate solution } scarcely dissolve.
Baric hydrate solution, dissolves traces, becoming coloured dark red.
Lead acetate (alcoholic), no pp.
Cupric acetate (alcoholic), brick-red pp.
Concentrated sulphuric acid, red solution, decomposing on standing, or on warming, into glucose and Rubiadin.

ROTTLERIA. A. G. Perkin (*Chem. Soc. Jnl.*, Aug., 1893) finds for *ROTTLERIN* (compare p. 87) M.P. 191°-191·5°; pale flesh-coloured, slender transparent needles (not yellow, as Anderson had found).

Soluble in ether readily, in chloroform, benzene, and toluene; sparingly in carbon bisulphide and glacial acetic acid.

Reactions:

Alkaline hydrates and carbonates, dissolve.
Ferric chloride, brown coloration.

ISOROTTLERIN was also isolated, in the form of a glittering salmon-coloured mass; M.P. 198°-199°.

Soluble in a mixture of ether and chloroform, sparingly in ether alone. **Insoluble** in hot benzene, carbon bisulphide, chloroform.

Reactions:

Alkaline hydrates, sol. cold, orange-red.
„ carbonates, sol. boiling, orange-red.
Ferric chloride (alcoholic), brownish-black coloration.

STACHYS tuberifera. A. v. Planta and E. Schulze (*Berichte d. d. Chem. Ges.*, 26, 939) have separated *STACHYDRINE*, $C_7H_{18}NO_2$, resembling *Betaine;* transparent deliquescent crystals; M.P. 210° (after losing water at 100°); neutral reaction.

THE END.

Baillière, Tindall and Cox, King William Street, Strand, London